DRAFT
TOXICOLOGICAL PROFILE FOR SELENIUM

U.S. DEPARTMENT OF HEALTH AND HUMAN SERVICES
Public Health Service
Agency for Toxic Substances and Disease Registry

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UPDATE STATEMENT

Toxicological profiles are revised and republished as necessary, but no less than once every three years. For information regarding the update status of previously released profiles, contact ATSDR at:

Agency for Toxic Substances and Disease Registry
Division of Toxicology/Toxicology Information Branch
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FOREWORD

This toxicological profile is prepared in accordance with guidelines developed by the Agency for Toxic Substances and Disease Registry (ATSDR) and the Environmental Protection Agency (EPA). The original guidelines were published in the *Federal Register* on April 17, 1987. Each profile will be revised and republished as necessary.

The ATSDR toxicological profile succinctly characterizes the toxicologic and adverse health effects information for the hazardous substance described therein. Each peer-reviewed profile identifies and reviews the key literature that describes a hazardous substance's toxicologic properties. Other pertinent literature is also presented, but is described in less detail than the key studies. The profile is not intended to be an exhaustive document; however, more comprehensive sources of specialty information are referenced.

The focus of the profiles is on health and toxicologic information; therefore, each toxicological profile begins with a public health statement that describes, in nontechnical language, a substance's relevant toxicological properties. Following the public health statement is information concerning levels of significant human exposure and, where known, significant health effects. The adequacy of information to determine a substance's health effects is described in a health effects summary. Data needs that are of significance to protection of public health are identified by ATSDR and EPA.

Each profile includes the following:

(A) The examination, summary, and interpretation of available toxicologic information and epidemiologic evaluations on a hazardous substance to ascertain the levels of significant human exposure for the substance and the associated acute, subacute, and chronic health effects;

(B) A determination of whether adequate information on the health effects of each substance is available or in the process of development to determine levels of exposure that present a significant risk to human health of acute, subacute, and chronic health effects; and

(C) Where appropriate, identification of toxicologic testing needed to identify the types or levels of exposure that may present significant risk of adverse health effects in humans.

The principal audiences for the toxicological profiles are health professionals at the Federal, State, and local levels; interested private sector organizations and groups; and members of the public. We plan to revise these documents in response to public comments and as additional data become available. Therefore, we encourage comments that will make the toxicological profile series of the greatest use.

Comments should be sent to:

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The toxicological profiles are developed in response to the Superfund Amendments and Reauthorization Act (SARA) of 1986 (Public Law 99-499) which amended the Comprehensive Environmental Response, Compensation, and Liability Act of 1980 (CERCLA or Superfund). This public law directed ATSDR to prepare toxicological profiles for hazardous substances most commonly found at facilities on the CERCLA National Priorities List and that pose the most significant potential threat to human health, as determined by ATSDR and the EPA. The availability of the revised priority list of 275 hazardous substances was announced in the *Federal Register* on October 21, 1999 (64 FR 56792). For prior versions of the list of substances, see *Federal Register* notices dated April 17, 1987 (52 FR 12866); October 20, 1988 (53 FR 41280); October 26, 1989 (54 FR 43619); October 17, 1990 (55 FR 42067); October 17, 1991 (56 FR 52166); October 28, 1992 (57 FR 48801); February 28, 1994 (59 FR 9486); April 29, 1996 (61 FR 18744); and November 17, 1997 (62 FR 61332). Section 104(i)(3) of CERCLA, as amended, directs the Administrator of ATSDR to prepare a toxicological profile for each substance on the list.

This profile reflects ATSDR's assessment of all relevant toxicologic testing and information that has been peer-reviewed. Staff of the Centers for Disease Control and Prevention and other Federal scientists have also reviewed the profile. In addition, this profile has been peer-reviewed by a nongovernmental panel and is being made available for public review. Final responsibility for the contents and views expressed in this toxicological profile resides with ATSDR.

Jeffrey P. Koplan, M.D., M.P.H.
Administrator
Agency for Toxic Substances and
Disease Registry

QUICK REFERENCE FOR HEALTH CARE PROVIDERS

Toxicological Profiles are a unique compilation of toxicological information on a given hazardous substance. Each profile reflects a comprehensive and extensive evaluation, summary, and interpretation of available toxicologic and epidemiologic information on a substance. Health care providers treating patients potentially exposed to hazardous substances will find the following information helpful for fast answers to often-asked questions.

Primary Chapters/Sections of Interest

Chapter 1: Public Health Statement: The Public Health Statement can be a useful tool for educating patients about possible exposure to a hazardous substance. It explains a substance's relevant toxicologic properties in a nontechnical, question-and-answer format, and it includes a review of the general health effects observed following exposure.

Chapter 2: Relevance to Public Health: The Relevance to Public Health Section evaluates, interprets, and assesses the significance of toxicity data to human health.

Chapter 3: Health Effects: Specific health effects of a given hazardous compound are reported by *type of health effect* (death, systemic, immunologic, reproductive), by *route of exposure*, and by *length of exposure* (acute, intermediate, and chronic). In addition, both human and animal studies are reported in this section.

> ***NOTE:*** Not all health effects reported in this section are necessarily observed in the clinical setting. Please refer to the Public Health Statement to identify general health effects observed following exposure.

Pediatrics: Four new sections have been added to each Toxicological Profile to address child health issues:

Section 1.6 How Can (Chemical X) Affect Children?
Section 1.7 How Can Families Reduce the Risk of Exposure to (Chemical X)?
Section 3.7 Children's Susceptibility
Section 6.6 Exposures of Children

Other Sections of Interest:

Section 3.8 Biomarkers of Exposure and Effect
Section 3.11 Methods for Reducing Toxic Effects

ATSDR Information Center

Phone: 1-888-42-ATSDR or (404) 498-0110 ***Fax:*** (404) 498-0057
E-mail: atsdric@cdc.gov ***Internet:*** http://www.atsdr.cdc.gov

The following additional material can be ordered through the ATSDR Information Center:

Case Studies in Environmental Medicine: Taking an Exposure History—The importance of taking an exposure history and how to conduct one are described, and an example of a thorough exposure history is provided. Other case studies of interest include *Reproductive and Developmental Hazards; Skin Lesions and Environmental Exposures; Cholinesterase-Inhibiting Pesticide Toxicity;* and numerous chemical-specific case studies.

Managing Hazardous Materials Incidents is a three-volume set of recommendations for on-scene (prehospital) and hospital medical management of patients exposed during a hazardous materials incident. Volumes I and II are planning guides to assist first responders and hospital emergency department personnel in planning for incidents that involve hazardous materials. Volume III—*Medical Management Guidelines for Acute Chemical Exposures*—is a guide for health care professionals treating patients exposed to hazardous materials.

Fact Sheets (ToxFAQs) provide answers to frequently asked questions about toxic substances.

Other Agencies and Organizations

The National Center for Environmental Health (NCEH) focuses on preventing or controlling disease, injury, and disability related to the interactions between people and their environment outside the workplace. *Contact:* NCEH, Mailstop F-29, 4770 Buford Highway, NE, Atlanta, GA 30341-3724 • Phone: 770-488-7000 • FAX: 770-488-7015.

The National Institute for Occupational Safety and Health (NIOSH) conducts research on occupational diseases and injuries, responds to requests for assistance by investigating problems of health and safety in the workplace, recommends standards to the Occupational Safety and Health Administration (OSHA) and the Mine Safety and Health Administration (MSHA), and trains professionals in occupational safety and health. *Contact:* NIOSH, 200 Independence Avenue, SW, Washington, DC 20201 • Phone: 800-356-4674 or NIOSH Technical Information Branch, Robert A. Taft Laboratory, Mailstop C-19, 4676 Columbia Parkway, Cincinnati, OH 45226-1998 • Phone: 800-35-NIOSH.

*The National Institute of Environmental Health Sciences (*NIEHS) is the principal federal agency for biomedical research on the effects of chemical, physical, and biologic environmental agents on human health and well-being. *Contact:* NIEHS, PO Box 12233, 104 T.W. Alexander Drive, Research Triangle Park, NC 27709 • Phone: 919-541-3212.

Referrals

The Association of Occupational and Environmental Clinics (AOEC) has developed a network of clinics in the United States to provide expertise in occupational and environmental issues. *Contact:* AOEC, 1010 Vermont Avenue, NW, #513, Washington, DC 20005 • Phone: 202-347-4976 • FAX: 202-347-4950 • e-mail: AOEC@AOEC.ORG • Web Page: http://www.aoec.org/.

The American College of Occupational and Environmental Medicine (ACOEM) is an association of physicians and other health care providers specializing in the field of occupational and environmental medicine. *Contact:* ACOEM, 55 West Seegers Road, Arlington Heights, IL 60005 • Phone: 847-818-1800 • FAX: 847-818-9266.

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THE PROFILE HAS UNDERGONE THE FOLLOWING ATSDR INTERNAL REVIEWS:

1. Health Effects Review. The Health Effects Review Committee examines the health effects chapter of each profile for consistency and accuracy in interpreting health effects and classifying end points.

2. Minimal Risk Level Review. The Minimal Risk Level Workgroup considers issues relevant to substance-specific minimal risk levels (MRLs), reviews the health effects database of each profile, and makes recommendations for derivation of MRLs.

3. Data Needs Review. The Research Implementation Branch reviews data needs sections to assure consistency across profiles and adherence to instructions in the Guidance.

PEER REVIEW

A peer review panel was assembled for selenium. The panel consisted of the following members:

1. Maryce M. Jacobs, Ph.D., Private Consultant, McLean, Virginia

2. Donald Lisk, Ph.D., Cornell University, Ithaca, New York

3. Craig Schnell, Ph.D., Private Consultant, Fargo, North Dakota

4. Calvin Willnite, Ph.D., Staff Toxicologist, Freemont, California

These experts collectively have knowledge of selenium's physical and chemical properties, toxicokinetics, key health end points, mechanisms of action, human and animal exposure, and quantification of risk to humans. All reviewers were selected in conformity with the conditions for peer review specified in Section 104(I)(13) of the Comprehensive Environmental Response, Compensation, and Liability Act, as amended.

Scientists from the Agency for Toxic Substances and Disease Registry (ATSDR) have reviewed the peer reviewers' comments and determined which comments will be included in the profile. A listing of the peer reviewers' comments not incorporated in the profile, with a brief explanation of the rationale for their exclusion, exists as part of the administrative record for this compound. A list of databases reviewed and a list of unpublished documents cited are also included in the administrative record.

The citation of the peer review panel should not be understood to imply its approval of the profile's final content. The responsibility for the content of this profile lies with the ATSDR.

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1. PUBLIC HEALTH STATEMENT

This public health statement tells you about selenium and the effects of exposure.

The Environmental Protection Agency (EPA) identifies the most serious hazardous waste sites in the nation. These sites make up the National Priorities List (NPL) and are the sites targeted for long-term federal cleanup activities. Selenium has been found in at least 494 of the 1,585 current or former NPL sites. However, the total number of NPL sites evaluated for this substance is not known. As more sites are evaluated, the sites at which selenium is found may increase. This information is important because exposure to this substance may harm you and because these sites may be sources of exposure.

When a substance is released from a large area, such as an industrial plant, or from a container, such as a drum or bottle, it enters the environment. This release does not always lead to exposure. You are exposed to a substance only when you come in contact with it. You may be exposed by breathing, eating, or drinking the substance, or by skin contact.

If you are exposed to selenium, many factors determine whether you'll be harmed. These factors include the dose (how much), the duration (how long), and how you come in contact with it. You must also consider the other chemicals you're exposed to and your age, sex, diet, family traits, lifestyle, and state of health.

1.1 WHAT IS SELENIUM?

Selenium is a naturally occurring, solid substance. It is widely but unevenly distributed in the earth's crust. It is commonly found in rocks and soil. Selenium, in its pure form of metallic gray to black hexagonal crystals, is often referred to as elemental selenium or selenium dust. However, in the environment, selenium is not often found in this pure form. It is usually combined with other substances. Much of the selenium in rocks is combined with sulfide minerals or with silver, copper, lead, and nickel minerals.

In the environment, selenium combines with oxygen to form several substances. The most common are sodium selenite and sodium selenate. Pure sodium selenite and selenate are white or colorless crystals. Selenium sulfide is an insoluble, bright red-yellow powder that is used in anti-dandruff shampoos. When used in anti-dandruff shampoos, it is known by the common trade name Selsun Blue. Some selenium compounds are gases. Hydrogen selenide (also called selenium hydride) is a colorless gas with a disagreeable odor. Although other selenium forms may be in dust in the air, hydrogen selenide is probably the only gaseous selenium compound that might pose a health concern in occupational settings. Selenium dioxide is an industrially

produced compound that easily dissolves in water to form selenious acid. Selenious acid can be found in gun blueing (a liquid solution used to clean the metal parts of a gun). Levels at which you begin to smell or taste selenium compounds in air or water are not available. The only exception is for selenium dioxide. It has a sharp, sour odor that you can smell when its concentration in the air goes above about 200 nanograms per cubic meter (ng/m^3; a nanogram is $1/1,000,000,000^{th}$ of a gram).

Plants easily take up selenium compounds from water and change them to organic selenium compounds such as selenomethionine (a transparent solid in pure form). Selenium is an essential nutrient for humans and animals, and both can use inorganic as well as organic selenium compounds. The human recommended dietary allowance (RDA) of selenium for maintenance of good health is 55 µg/day (a µg is $1/1,000,000^{th}$ of a gram) for adults. Although it is an essential nutrient, selenium can harm people and animals when consistently eaten in amounts higher than those needed for good nutrition.

Burning coal—especially the fly ash that comes from burning coal—is the major man-made source of selenium in the environment. Elemental selenium is also commercially produced, primarily as a by-product of copper refining. In 1996, U.S. production of selenium was 379 metric tons. Examples of uses for selenium and its compounds are some photographic devices (because of its semiconductor and photoelectric properties), plastics, paints, anti-dandruff shampoos, vitamin and mineral supplements, fungicides, and certain types of glass. Selenium is also used to prepare drugs and as a nutritional feed for poultry and livestock. For instance, in some European countries, sodium selenate is found in drenches (large doses of liquid medicine) used to supplement the diets of cattle and sheep grazing on selenium-deficient soils. When selenium is found in vitamin and mineral supplements, it is usually listed on the content label as "selenium."

Selenium can exist at hazardous waste sites in many forms. However, specific forms have not been analyzed at sites where selenium is reported. At hazardous waste sites, you could be exposed to selenium by swallowing the soil or water, or by breathing the dust. Eating plant products that have taken up selenium from soil and water in areas of high concentration could also result in indirect exposure at regulated waste sites. The form selenium takes at any given time depends on the environmental conditions. This is further discussed in Chapter 6.

Chapter 4 has more information on the physical and chemical properties of selenium and its compounds. Chapter 5 provides more information on the sources and uses of selenium.

1.2 WHAT HAPPENS TO SELENIUM WHEN IT ENTERS THE ENVIRONMENT?

Although selenium occurs naturally in the environment, it also can be released by both natural and manufacturing processes. As an element (its simplest form), selenium cannot be created or destroyed. However, selenium can change forms in the environment. Weathering of rocks to soil may cause low levels of selenium in water or may cause selenium to be taken up by plants or released on fine dust-like particles into the air. Volcanic eruptions are suspected of contributing to selenium in air, and soils in the areas around volcanos contain more selenium.

More commonly, selenium enters the air from burning coal or oil. Much of the selenium in air is attached to fly ash and to suspended particles. The elemental selenium that may be present in fossil fuels forms selenium dioxide when burned. Selenium dioxide can then form selenious acid with water or sweat. Selenium anhydride is released during the heating of copper, lead, and zinc ores when there is selenium in them. Hydrogen selenide breaks up rapidly in air to form elemental selenium and water, thus eliminating the danger from this compound for most people, except those who are exposed to it in their workplace.

Airborne particles of selenium, such as in ash, can settle on soil or surface water Disposal of selenium in commercial products and waste could also contribute to selenium levels in soil. However, the amount of selenium released to soil from ash and hazardous waste sites has not been measured. The forms and fate of selenium in soil depend largely on the acidity of the surroundings and its interaction with oxygen. In theory, with no oxygen present, deep soil selenium may be present as elemental selenium. In the absence of oxygen when the soil is acidic, the amount of biologically available selenium should be low. Elemental selenium that cannot dissolve in water and other insoluble forms of selenium (such as selenium sulfide and heavy metal selenides) are less mobile and will usually remain in the soil, posing less of a risk for exposure. Active agricultural or industrial processes may increase the amount of biologically available selenium by decreasing the acidity of the soil and increasing the oxygen and the soluble selenium compounds. Selenium compounds that can dissolve in water are very mobile. For example, selenates and selenites are water soluble, and thus mobile, so there is an increased chance of exposure to them. Selenium may enter surface water in irrigation drainage waters. Factors that may affect how fast selenium moves through soil are temperature, moisture, season of year, concentration of water-soluble selenium, organic matter content, and microbiological activity.

Some evidence indicates that selenium can be taken up in tissues of organisms and possibly increase in concentration in aquatic organisms as the selenium is passed up through the food chain. Selenium concentrations in organisms living in water have been a problem as a result of irrigation runoff in some dry areas of the United States. It is important to remember that

selenium's behavior in the environment is largely affected by the surrounding conditions and by how it interacts with other compounds. Chapter 6 contains more information on what happens to selenium in the environment.

1.3 HOW MIGHT I BE EXPOSED TO SELENIUM?

People can be exposed to selenium in several ways and environments. They could be exposed to higher-than-normal levels of selenium at hazardous waste sites. Exposure may occur through locally grown food, water, soil, and/or air that contains selenium particularly at higher-than-normal levels. Research shows that the most significant exposure to selenium comes from eating food, and to a lesser extent, from water intake. For example, in certain areas of China, villagers used coal ash high in selenium to fertilize their crops. After many years, the vegetables from these areas, their main source of food, had large amounts of selenium. Selenium can also be washed from the soil at hazardous waste disposal sites into surface water or can flow into groundwater. Occasionally, water containing selenium may seep from abandoned uranium or coal mining areas into groundwater, or into rivers or streams. This can eventually enter into drinking water systems.

People may also be exposed to selenium from industrial sources. Humans are normally not exposed to large amounts of selenium in the air, unless selenium dust or volatile selenium compounds are formed in their workplace. Occupations in which humans may be exposed to selenium in the air are the metal industries, selenium-recovery processes, paint manufacturing, and special trades. Selenium dioxide and elemental selenium can be released into the air during the burning of coal and oil. In addition, irrigation runoff may cause high levels of selenium in fresh water and can lead to high concentrations in some of the fish and birds that live there.

Because selenium is naturally occurring and widespread, people are also exposed to low levels of selenium daily through food, water, and air. Estimates of the average intake of selenium from food for the U.S. population range from 71 to 152 micrograms of selenium per person per day (μg selenium/day). Generally, the levels in food in the United States are enough to protect against diseases that may result from too little selenium. Most of the daily intake of selenium comes from eating grains, cereals, and meat. Selenium in food is primarily in the form of selenoamino acids (organic selenium compounds). In some parts of the United States, especially in the western states, the soils naturally have high levels of selenium compounds. Some plants can build up selenium to levels that harm livestock feeding on them. In these areas, people could be exposed to too much selenium if they eat a lot of locally grown grains and vegetables that have built up high levels of selenium. However, tests of people in areas of the western United States with high soil selenium levels show none of the negative health effects associated with

selenium that were seen in Chinese villagers who had eaten much larger amounts of selenium for longer periods.

Low levels of selenium can also be found in drinking water. Most of the water sources tested in the United States have very low levels of selenium compared with the levels found in food. Selenium levels were less than 10 ppb (10 parts of selenium in a billion parts of water) in 99.5% of drinking water sources tested. The 10-ppb concentration is lower than the 50-ppb Maximum Contaminant Level (MCL), which the EPA believes will protect against adverse health effects. Less than 1% of the daily intake of selenium is estimated to come from drinking water.

Chapter 6 contains more information on how people can be exposed to selenium.

1.4 HOW CAN SELENIUM ENTER AND LEAVE MY BODY?

Selenium from the environment mainly enters the body when people eat food containing selenium. The human body easily absorbs the organic selenium compounds (for example, selenoamino acids) when eaten and makes them available where needed in the body. The selenium in drinking water is usually in the form of inorganic sodium selenate and sodium selenite; these are also easily absorbed from the digestive tract. The human body can change these inorganic selenium compounds into forms it can use. Selenium in the air may also enter your body when you breathe it.

Hazardous waste sites at which selenium is present could present a major source of exposure to selenium. The way that it can enter the body from a particular site depends on such things as whether vegetables are grown on soil in which selenium from the site has been deposited, whether water at the site contains selenium and is able to flow into drinking water supplies, and whether selenium dust blows into the air. As mentioned earlier, specific conditions at a site can greatly influence which selenium compounds form and whether they can move in the environment to places where people may be exposed. Therefore, it is important to know that the presence of selenium at a site does not necessarily mean that people are being exposed to it. Specific tests of locally grown food, drinking water, and air must be done to find out whether exposure is occurring. You should also be aware that selenium compounds, including those used in some medicated dandruff shampoos, are not easily absorbed through the skin.

Most of the selenium that enters the body quickly leaves the body, usually within 24 hours. Beyond what the body needs, selenium leaves mainly in the urine, but also in feces and breath. Selenium in the urine increases as the amount of the exposure goes up. Selenium can build up in the human body, however, if exposure levels are very high or if exposure occurs over a long time. It builds up mostly in the liver and kidneys but also in the blood, lungs, heart, and testes.

Selenium can build up in the nails and in hair, depending on time and amount of exposure. Chapter 3 contains more information on how selenium enters and leaves the human body.

1.5 HOW CAN SELENIUM AFFECT MY HEALTH?

To protect the public from the harmful effects of toxic chemicals and to find ways to treat people who have been harmed, scientists use many tests.

One way to see if a chemical will hurt people is to learn how the chemical is absorbed, used, and released by the body; for some chemicals, animal testing may be necessary. Animal testing may also be used to identify health effects such as cancer or birth defects. Without laboratory animals, scientists would lose a basic method to get information needed to make wise decisions to protect public health. Scientists have the responsibility to treat research animals with care and compassion. Laws today protect the welfare of research animals, and scientists must comply with strict animal care guidelines.

The general public rarely breathes high levels of selenium, although some people may be exposed to selenium dust and selenium compounds in workplace air. Dizziness, fatigue, and irritation of mucous membranes have been reported in people exposed to selenium in workplace air at concentrations higher than legal levels. In extreme cases, collection of fluid in the lungs (pulmonary edema) and severe bronchitis have been reported. The exact exposure levels at which these effects may occur are not known, but they become more likely with increasing amounts of selenium and with increasing frequency of exposure.

The normal intake of selenium by eating food is enough to meet the RDA of 55 μg/day for adults for this essential nutrient. However, as discussed in Chapter 2 of this profile, selenium compounds can be harmful at daily dietary levels that are higher than needed. The seriousness of the effects of excess selenium depends on how much selenium is eaten and how often. Intentional or accidental swallowing of a large amount of sodium selenate or sodium selenite (for example, a very large quantity of selenium supplement pills) could be life-threatening without immediate medical treatment. Even if mildly excessive amounts of selenium are eaten over long periods, brittle hair and deformed nails can develop. In extreme cases, people may lose feeling and control in arms and legs. These health effects, called selenosis, were seen in several villages in China where people were exposed to foods high in selenium for months to years. No human populations in the United States have been reported with long-term selenium poisoning, including populations in the western part of the country where selenium levels are naturally high in the soil. Because most people in the United States eat foods produced in many different areas, overexposure to selenium in food is unlikely to occur.

In some regions of China where soil levels of selenium are very low, not eating enough selenium has resulted in health effects. Selenium is used by the body in enzymes that protect against damage to tissues done by oxygen, and in an enzyme that affects growth and metabolism. Not eating enough selenium can cause heart problems and muscle pain. Muscle pain has also been noted in people fed intravenously for a long time with solutions that did not contain selenium. Babies born early may be more sensitive to not having enough selenium, and this may contribute to lung effects. In the United States, selenium in food is sufficient to meet the RDA and prevent harmful effects from not enough selenium.

Upon contact with human skin, selenium compounds have been reported to cause rashes, redness, heat, swelling, and pain. Brief, acute exposure of the eyes to selenium dioxide as a dust or fume in workplace air may result in burning, irritation, and tearing. However, only people who work in industries that process or use selenium or selenium compounds are likely to come into contact with levels high enough to cause eye irritation.

Studies of laboratory animals and people show that most selenium compounds probably do not cause cancer. In fact, studies of cancer in humans suggest that lower-than-normal selenium levels in the diet might increase the risk of cancer. But levels of selenium in the diet that are much higher than normal have not been shown to reduce the risk of cancer in humans. Taking selenium so that your daily amount is greater than that required may just increase your risk of selenium poisoning.

Based on studies up until 1987, the International Agency for Research on Cancer (IARC) determined that selenium and selenium compounds could not be classified as to their ability to cause cancer in humans. However, since then the EPA has determined that one specific form of selenium, called selenium sulfide, is a probable human carcinogen. Selenium sulfide is the only selenium compound shown to cause cancer in animals. Rats and mice that were fed selenium sulfide daily at very high levels developed cancer. Selenium sulfide is not present in foods, and it is a very different chemical from the organic and inorganic selenium compounds found in foods and in the environment. Also, if introduced into the environment, selenium sulfide does not dissolve readily in water and would probably bind tightly to the soil, further reducing any chance of exposure. Because selenium sulfide is not absorbed through the skin, the use of anti-dandruff shampoos containing selenium sulfide is generally considered safe.

Very high amounts of selenium have resulted in decreased sperm counts, increased abnormal sperm, changes in the female reproductive cycle in rats, and changes in the menstrual cycle in monkeys. The relevance of the reproductive effects of selenium exposure in animals studied to potential reproductive effects in humans is not known. Selenium compounds have not been shown to cause birth defects in humans or in other mammals.

Chapter 3 contains more information on the health effects of selenium and selenium compounds in humans and animals.

1.6 HOW CAN SELENIUM AFFECT CHILDREN?

This section discusses potential health effects from exposures during the period from conception to maturity at 18 years of age in humans.

Children living near selenium waste sites or coal burning plants are likely to be exposed to higher environmental levels of selenium through breathing, touching soil, and eating contaminated soil. Children living in areas of China with high selenium in the soil had higher levels of selenium in the blood than adults from that area. Very few studies have looked at how selenium can affect the health of children. Children need small amounts of selenium for normal growth and development. Children will probably show the same sort of health effects from selenium exposure as adults, but they may be less susceptible to health effects of selenium than adults.

We do not know if exposure to selenium will result in birth defects in people. Selenium compounds have not been shown to cause birth defects in humans or in other mammals. We have no information to suggest that there are any differences between children and adults in where selenium is found in the body or in how fast it enters or leaves the body. Studies in laboratory animals have shown that selenium crosses the placenta and enters the fetus. Studies in humans show that infants are supplied with selenium through breast milk, and therefore, women who were exposed to selenium by living near a waste site might transfer selenium to their babies. However, babies in areas of China with high selenium in the soil did not show any signs of health effects due to selenium, even though some of their parents did.

1.7 HOW CAN FAMILIES REDUCE THE RISK OF EXPOSURE TO SELENIUM?

If your doctor finds that you have been exposed to significant amounts of selenium, ask whether your children might also be exposed. Your doctor might need to ask your state health department to investigate.

Since selenium occurs naturally in the environment, we cannot avoid exposure to it. Certain dietary supplements and shampoos contain selenium. You should not exceed the recommended dosages when using these products.

Children living near selenium waste sites or coal burning plants are likely to be exposed to higher environmental levels of selenium through breathing, touching soil, and eating

contaminated soil. Some children eat a lot of dirt. You should discourage your children from eating dirt. Make sure they wash their hands frequently and before eating. Discourage your children from putting their hands in their mouths or from other hand-to-mouth activity.

The primary route of human exposure to selenium is through eating food. People who irrigate their home gardens with groundwater containing high levels of selenium may grow and eat plants that contain high levels of selenium because this element accumulates in some plants. Fishermen and hunters of waterfowl who regularly eat fish and game from waterways with high selenium content may also swallow above average levels of selenium. To reduce your family's exposure to selenium, obey any wildlife advisories issued by your state. Information on fish and wildlife advisories in your state is available from your state public health or natural resources department.

1.8 IS THERE A MEDICAL TEST TO DETERMINE WHETHER I HAVE BEEN EXPOSED TO SELENIUM?

Selenium can be measured in the blood, urine, and fingernails or toenails of exposed individuals. However, since selenium is an essential nutrient normally present in foods, low levels of selenium are normally found in body tissues and urine. Tests for selenium are most useful for people who have recently been exposed to high levels. Samples of blood, urine, or nails can be properly collected in a physician's office and sent to a laboratory that has the special equipment needed to measure selenium. Urine can be used to determine short-term exposure. Because red blood cells last about 120 days before they are replaced by newly made red blood cells, the presence of selenium in red blood cells can show whether a person was exposed to selenium during the 120 days before testing, but not if exposed more than 120 days before testing. Toenail clippings can be used to determine longer term exposure.

Many methods are available to measure selenium levels in human tissue and the environment. However, none of the methods that are routinely available can measure or detect each selenium compound in one test, and better tests that measure lower levels of different selenium compounds are needed. Also, these tests cannot determine the exact levels of selenium you may have been exposed to or predict whether health effects will occur.

In the United States, normal selenium amounts have been reported in blood as 80–300 ppb and in urine as 7–200 ppb. The results of other studies of average populations and those that have been exposed to higher levels may serve as a guideline. In China, the average selenium amounts in the general population were 95 ppb in whole blood and 26 ppb in urine. In 350 healthy Italian subjects living in the same region of northern Italy, the mean selenium value was 107.5 ppb in blood and 22.1 ppb in urine. Among ranchers who live in western South Dakota and eastern

Wyoming where the soil is naturally high in selenium, the amount of selenium in blood was 317 ppb, and selenium in urine ranged from 10 to 462 ppb. The latter study is considered by some to suggest that, when people are exposed over a long period to higher-than-normal amounts of selenium, their bodies adjust to the higher amounts. Unfortunately, the differences between studies on populations in such areas make it difficult to compare the results.

People in China who had eaten food that was very high in selenium over a long time developed signs of selenosis. These people had an average blood level of 3,200 ppb, about 7 times the highest amount found in the study of the western U.S. population living in an area with high selenium levels in the soil. The amount of selenium in blood in the U.S. study was also compared with an average amount of 206 ppb found in subjects from 19 other U.S. cities. The researchers reported that in the ranchers mentioned above, there was no association between how often physical symptoms occurred and the amount of selenium in blood (up to 206 ppm). However, very high amounts of selenium in blood are clearly related to selenosis.

Chapter 3 contains more information on studies that have measured selenium amounts in human tissues.

The length of time that selenium stays in the body after exposure stops depends on the form of selenium to which the person was exposed. Thus, it is difficult to predict how useful a test will be if some time has gone by since exposure stopped. Chapter 7 contains more information on the methods available to measure selenium in human tissues and in the environment.

1.9 WHAT RECOMMENDATIONS HAS THE FEDERAL GOVERNMENT MADE TO PROTECT HUMAN HEALTH?

The federal government develops regulations and recommendations to protect public health. Regulations can be enforced by law. Federal agencies that develop regulations for toxic substances include the Environmental Protection Agency (EPA), the Occupational Safety and Health Administration (OSHA), and the Food and Drug Administration (FDA). Recommendations provide valuable guidelines to protect public health but cannot be enforced by law. Federal organizations that develop recommendations for toxic substances include the Agency for Toxic Substances and Disease Registry (ATSDR) and the National Institute for Occupational Safety and Health (NIOSH).

Regulations and recommendations can be expressed in not-to-exceed levels in air, water, soil, or food that are usually based on levels that affect animals; then they are adjusted to help protect people. Sometimes these not-to-exceed levels differ among federal organizations because of

different exposure times (an 8-hour workday or a 24-hour day), the use of different animal studies, or other factors.

Recommendations and regulations are also periodically updated as more information becomes available. For the most current information, check with the federal agency or organization that provides it. Some regulations and recommendations for selenium include the following:

The EPA Office of Drinking Water regulates the amount of selenium allowed in drinking water. Public water supplies are not allowed to exceed 50 ppb total selenium.

The FDA regulations allow a level of 50 ppb of selenium in bottled water. OSHA is responsible for setting regulations on selenium levels allowable in the workplace. The exposure limit for selenium compounds in the air for an 8-hour period is 0.2 mg selenium/m^3. Table 8-1 contains other regulations and guidelines for selenium.

1.10 WHERE CAN I GET MORE INFORMATION?

If you have any more questions or concerns, please contact your community or state health or environmental quality department or

Agency for Toxic Substances and Disease Registry
Division of Toxicology
1600 Clifton Road NE, Mailstop E-29
Atlanta, GA 30333

* Information line and technical assistance
Phone: 1-888-42-ATSDR (1-888-422-8737)
Fax: 1-404-498-0057

ATSDR can also tell you the location of occupational and environmental health clinics. These clinics specialize in recognizing, evaluating, and treating illnesses resulting from exposure to hazardous substances.

* To order toxicological profiles, contact

National Technical Information Service
5285 Port Royal Road
Springfield, VA 22161
Phone: 1-800-553-6847 or 1-703-605-6000

1. PUBLIC HEALTH STATEMENT

2. RELEVANCE TO PUBLIC HEALTH

2.1 BACKGROUND AND ENVIRONMENTAL EXPOSURES TO SELENIUM IN THE UNITED STATES

Since the publication of the previous version of ATSDR's Toxicological Profile for Selenium in 1996, several events have occurred that focused attention on the mineral selenium, its role in maintaining optimal human health, and any risk it may present to those exposed to excessive amounts of this metallic chemical element.

Late in the decade of the 90's, selenium was found to have entered the environment from old mining operations some northwestern U.S. locations. This resulted in public concern about the potential effects on livestock grazing in near-by areas, and ultimately the effects on humans consuming food products from plants and animals raised in those areas. At the same time, the Food and Nutrition Board of the Institute of Medicine, National Academy of Sciences was in the process of reevaluating the dietary requirements for many of the essential nutrients, including selenium. The result of this latter effort was the establishment a new Dietary Reference Intake (DRI) of 55 ug/day for selenium for both male and female adults (NAS 2000). This new number represented a decrease from the previous Recommended Dietary Allowance of 70 ug/day for male; 55 ug/day was already the RDA for females (NRC 1989). The combination of the increased concern regarding selenium toxicity and the reduction in the dietary selenium recommendation suggested that a reevaluation of selenium from a toxicological perspective might also be in order.

While a health guidance value equivalent to 5 ug/kg/day for selenium had been independently derived by both ATSDR and the U.S. EPA, this value was based upon data from two 1989 studies. To ensure that chronic oral MRL, which is used by a number of agencies at the federal, state, and local levels, is based upon the state-of-the-science knowledge about selenium, an update of the Toxicological Profile for Selenium was undertaken. Althugh the additional studies examined during this update effort did not result in a change in the MRL, they provide

strong support for the existing chronic oral MRL for selenium, thereby increasing our confidence in our health guidance value.

Selenium is a naturally occurring mineral element in the earth's crust. It is distributed widely in nature and is found in most rocks and soils at concentrations between 0.1 and 2.0 ppm. However, selenium is seldom found in its elemental form in the environment, but is obtained primarily as a byproduct of copper refining. Selenium exists in several allotropic forms. The primary factor determining the fate of selenium in the environment is its oxidation state. Selenium is stable in four valence states (-2, 0, +4, and +6) and forms chemical compounds similar to those of sulfur. The heavy metal selenide compounds (-2) are insoluble in water, as is elemental selenium. The inorganic alkali selenites (+4) and selenates (+6) are soluble in water and are, therefore, more bioavailable.

Conditions such as pH, oxidation-reduction potential, and the presence of metal oxides affect the partitioning of the various compounds of selenium in the environment. In general, elemental selenium is stable in soils and is found at low levels in water because of its ability to coprecipitate with sediments. The soluble selenates are readily taken up by plants and converted to organic compounds such as selenomethionine, selenocysteine, dimethyl selenide, and dimethyl diselenide. Selenium is bioaccumulated by aquatic organisms. Very low levels of selenium are found in ambient air.

Most processed selenium is used in the electronics industry. Selenium's semiconductor and photoelectric properties make it useful in "electric eyes," photographic exposure meters, and rectifiers for home entertainment equipment, and it is used to coat the metal cylinders from which a photographic image is transferred in xerography. Selenium is also used in the glass industry to counter coloration that results from iron impurities and in the production of both red and black glasses. Other uses include: as a component of pigments used in plastics, paints, enamels, inks, and rubber; as a catalyst in the preparation of pharmaceuticals, including niacin and cortisone; as a nutritional feed additive for poultry and livestock; in pesticide formulations; as an accelerator and vulcanizing agent in rubber production; as an ingredient in antidandruff shampoos (selenium sulfide); and as a constituent of fungicides (selenium sulfide). Radioactive selenium is used in diagnostic medicine and aids in the visualization of difficult-to-study malignant tumors.

The general public is exposed to selenium by ingestion of both organic and inorganic forms of selenium, which occur naturally in the diet. Most forms of selenium can interconvert either within organisms or within the environment, depending on ambient conditions. Some persons living in areas with high soil concentrations of selenium (as in areas of the western United States) might have higher exposure because

of the natural selenium levels found locally, particularly if they consume primarily crops grown in that area.

Various estimates of the selenium intake for Americans have ranged from 0.071 to 0.152 mg selenium/day. The greatest portion of dietary selenium intake occurs from the ingestion of grains and cereals. Selenium is excreted in breast milk, and levels vary with maternal nutritional status.

2.2 SUMMARY OF HEALTH EFFECTS

Selenium is known to be an essential micronutrient for humans and other animals; both inadequate and excessive selenium intake can cause adverse health effects. However, most people in the United States are unlikely to suffer from selenium deficiency. The current recommended dietary allowance (RDA) for selenium established by the Food and Nutrition Board of the National Research Council is 55 μg/day for adults. Adverse health effects due to selenium are generally observed at doses at least 5 times greater than the RDA.

Selenium deficiency has been associated with two endemic diseases found in selenium-poor regions of China: a cardiovascular condition known as Keshan Disease and an osteoarthropathy called Kashin-Beck Disease. Keshan Disease is characterized by cardiac enlargement, abnormal ECG patterns, cardiogenic shock, and congestive heart failure, with multifocal necrosis of the myocardium. The disease is reported to occur primarily in children and women of child-bearing age and has been successfully treated by selenium supplementation; however, a low incidence of cases persisting after selenium supplementation suggests that there may be other contributing factors. Kashin-Beck Disease is characterized by atrophy, degeneration, and necrosis of cartilage tissue, and occurs primarily in children between the ages of 5 and 13 years; it also has been successfully treated with selenium supplements. Chronically ill people and older people have been shown to have lower organ concentrations of selenium than healthy individuals, but it is not clear if this is a cause or consequence of aging or illness.

The primary target organ in humans and laboratory animals in cases of acute, high-level inhalation exposure to selenium is the lung, with cardiovascular, hepatic, nervous, and renal involvement as well. Lesser effects are observed in other organs/organ systems. Workers exposed to high concentrations of elemental selenium dust reported stomach pain and headaches, while workers briefly exposed to high levels of selenium dioxide dust reported respiratory symptoms such as pulmonary edema, bronchial spasms, symptoms of asphyxiation and persistent bronchitis, elevated pulse rates, lowered blood pressure,

vomiting, nausea, and irritability. No studies were located regarding effects in humans or animals after intermediate-duration inhalation exposure to selenium or to selenium compounds. For chronic inhalation exposure, there are several occupational studies that describe respiratory effects such as irritation of the nose, respiratory tract, and lungs, bronchial spasms, and coughing following exposure to selenium dioxide or elemental selenium dust. Respiratory symptoms similar to those reported for humans have been seen in animals inhaling high doses of elemental selenium fumes or dust, and studies of animals with acute inhalation exposure to hydrogen selenide or elemental selenium fumes or dust have reported hepatocellular degeneration and atrophy of the liver.

Acute oral exposure to high concentrations of selenium produces nausea, vomiting, and diarrhea in both humans and animals. Acute exposure of humans to selenium has occasionally produced cardiovascular symptoms, such as tachycardia, but no electrocardiographic abnormalities were found in 20 individuals from a human population chronically exposed to selenium. In contrast, acute and intermediate oral exposure of laboratory animals to selenium has produced myocardial degeneration.

Chronic oral exposure to high concentrations of selenium compounds produces selenosis, the major effects of which are dermal and neurological. As evidenced by populations in China, chronic exposure to high selenium levels in the diet can cause diseased nails and skin and hair loss. Higher levels can cause neurological problems including unsteady gait and paralysis. In contrast, however, studies of contemporary human populations living in areas of naturally occurring high selenium concentrations in the United States have not revealed adverse health effects in those populations. This difference may result from the lower exposure observed in the U.S. population compared to the Chinese population, as well as a better balanced, higher protein diet in the United States, which may increase tolerance to selenium.

Intermediate and chronic oral exposure of livestock to high levels of dietary selenium compounds also produces dermal and neurological effects. Studies in animals with high selenium concentrations demonstrate that many organ systems retain selenium and are affected. The primary effects in laboratory animals exposed to inorganic selenium salts or to selenium-containing amino acids are cardiovascular, gastrointestinal, hematological, hepatic, dermal, immunological, neurological, and reproductive. Blind staggers has been repeatedly observed in cattle feeding off vegetation in areas with high selenium content in the soil. However, the neurological effects have not been replicated in experimentally-exposed cattle receiving doses of selenium sufficient to induce hoof lesions, and thus, the neurological symptoms associated with blind staggers may be due to other compounds found within this vegetation.

Some evidence for adverse effects on the endocrine system has also been found following intermediate and chronic oral exposure to elevated levels of dietary selenium in humans and animals. Human studies have demonstrated a decrease in triiodothyronine levels in response to increased dietary selenium, although the hormone levels remained within the normal range. Intermediate-duration studies of rats have shown reductions in type-I-deiodinase activity in response to selenium. However, the levels of thyroid hormones in these animals did not show a consistent pattern.

There is no evidence to support a causal association between selenium compounds and cancer in humans. In fact, some epidemiological and experimental evidence suggests that selenium exposure under certain conditions may contribute to a reduction in cancer risk. Currently, the chemopreventive potential of selenium is under research. Selenium sulfide and ethyl selenac are the only selenium compounds that have been shown to be carcinogenic upon oral administration in rodents; however, significant exposure to these forms of selenium is extremely unlikely.

Studies of Chinese populations and laboratory animals exposed to high levels of organic and/or inorganic selenium compounds have not found evidence of selective teratogenic effects in mammals.

Selenosis. Following chronic oral exposure to excessive amounts of the organic selenium compounds found in food, the two principal clinical conditions observed in humans are dermal and neurological effects, as described most completely in the epidemiological study of endemic selenosis in the People's Republic of China. The dermal manifestations of selenosis include loss of hair, deformation and loss of nails, and discoloration and excessive decay of teeth, while neurological effects include numbness, paralysis, and occasional hemiplegia.

Similar clinical manifestations occur in livestock following intermediate and chronic exposure to excessive amounts (more than 30 times the normal dietary amount of selenium) of the organic selenium compounds found in seleniferous plants, including loss of hair and malformation of hooves in pigs, horses, and cattle and poliomyelomalacia in pigs. Histologically, swine with selenium-induced neurological signs exhibit bilateral macroscopic lesions of the ventral horn of the spinal cord. The selenium in the selenium-accumulating plant *Astragalus bisulcatus* appears to be a more potent neurotoxicant than D,L-selenomethionine or selenate. Following a similar dose of selenium, *A. bisulcatus* resulted in complete paralysis in four of five pigs after 5 days of treatment, and in the last pig after 3 weeks of treatment, while in pigs fed selenate, three of five developed complete paralysis, and one pig developed posterior paralysis after 4–21 days of treatment. Although D,L-selenomethionine resulted in

the greatest incidence of selenosis, it was the least potent neurotoxin, resulting in posterior paralysis in two of five pigs after 9 and 24 days of treatment. The pigs that did not develop paralysis were fed D,L-selenomethionine for approximately 31 days. The form of selenium in *A. bisulcatus* is unknown, although Panter et al. indicate that it is nonprotein.

The neurological signs and histopathology observed in livestock following oral exposure to excess selenium compounds have not been recorded in laboratory animals. This suggests that (1) small laboratory mammals may not be appropriate models for selenium toxicity in humans (e.g., laboratory animals absorb selenium compounds to a lesser extent, or metabolize and/or excrete selenium compounds more quickly), (2) some as yet unidentified organic form of selenium contributes to the neurological manifestations of chronic selenosis in humans and in livestock, or (3) unrecognized confounding factors, such as other plant toxins, have contributed to the neurological syndrome associated with chronic selenosis in field studies of humans and livestock.

Endocrine Effects. Selenium is a component of all three members of the deiodinase enzyme family, the enzymes responsible for deiodination of the thyroid hormones, and has a physiological role in the control of thyroid hormone levels. Significant decreases in triiodothyronine levels in response to elevated selenium have been observed in humans. However, the triiodothyronine levels observed in these studies were within the normal human range, so the biological impact of this change is unclear. The effect of increased dietary selenium on other thyroid hormones is also unclear. Intermediate-duration studies in rats show a decrease in type-I-deiodinase activity in response to elevated selenium; however, the levels of thyroid hormones in these animals did not show any consistent changes.

One of the most common effects observed following excess selenium intake in animals is a decrease in growth. It is likely that the selenium-induced reduction in growth has an endocrine component. For example, selenite treatment of young rats decreased somatomedin C levels, and growth hormone secretion in response to the growth hormone releasing factor was also reduced in selenium-treated rats. The primary endocrine target of selenium leading to decreased growth has yet to be elucidated.

Pancreatic toxicity has been observed following excess selenium exposure. Cytoplasmic flocculation was observed in lambs treated with a single oral dose of selenite, and pancreatic damage, which was not further described, was noted in rats following chronic oral treatment with selenate or selenite. Pancreatic toxicity associated with excessive selenium exposure is likely related to the unique ability of that organ to accumulate the element.

Hepatic Effects. Liver effects have not been reported for humans exposed to excessive amounts of selenium. No significant abnormalities were found in blood levels of liver enzymes of individuals living in high selenium areas, or in a "supersonic B" morphological examination of the livers of 20 individuals suffering from severe symptoms of selenosis. In experimental and other animals, however, the liver has been shown to be affected following both inhalation and oral exposure to several different selenium compounds. Hepatocellular degeneration was observed in guinea pigs following inhalation exposure to elemental selenium dust and hydrogen selenide. Cirrhosis, hepatocellular degeneration, and changes in liver enzyme levels in serum have been reported for rats, pigs, and mice orally exposed to selenite, selenate, or organic selenium. The doses in these studies producing adverse effects were approximately 10 times the amount normally found in an adequate diet. The liver appears to be the primary target organ for the oral toxicity of selenium in experimental animals following intermediate and chronic exposure, whereas liver cirrhosis or dysfunction has not been a notable component of the clinical manifestations of chronic selenosis in humans. Selenium sulfide administration has also produced frank hepatotoxicity in rats, but not in mice.

Renal Effects. No reports of renal effects in humans were located. In nonhuman animals, mild kidney effects have been observed following oral exposure to selenium compounds. These effects include hydropic degeneration in sheep exposed to selenite and nephropathy in monkeys exposed to selenomethionine. Rats appear to be more sensitive than mice to the renal effects of selenium compounds. A dose-related increase in renal papilla degeneration, described as mild to minimal, was observed in rats treated with selenate or selenite in the drinking water, while the only kidney effect noted in mice treated with sodium selenate or selenite in the drinking water was increased kidney weight. Selenium sulfide administration to mice has, however, been shown to produce interstitial nephritis.

Body Weight Effects. Two of the most common effects in experimental animals following intermediate or chronic oral administration of excessive inorganic and organic compounds of selenium are reduced growth rate of young animals and loss of weight in older animals. Selenium sulfide administration has been associated with reduced body weight in female mice. As noted under endocrine effects, it is quite possible that the decreased growth rate has an endocrine component.

Reproductive Effects. In samples from more than 200 men, no correlation between seminal fluid selenium and sperm count or mobility was detected. No significant increase in spontaneous abortions was reported among women chronically exposed to drinking water containing excessive amounts of selenium. Oral or injection treatment of rats with sodium selenate or selenite (at doses at least 8 times greater than those normally supplied by an adequate diet) has been shown to increase the number of abnormal sperm, produce testicular hypertrophy, or degeneration and atrophy in males, and to affect the estrous cycle in females. The animals in these studies were not mated, so it is not clear if fertility was affected. A small increase in the number of abnormal length estrous cycles was observed in mice exposed to selenium. Disturbances in the menstrual cycle (anovulation, short luteal and follicular phases) were also observed in monkeys treated orally with L-selenomethionine. Selenium deficiency has also been reported to cause decreased sperm production and motility in rats. The relevance of these reproductive effects of selenium in laboratory animals to potential reproductive effects in humans is not known.

2.3 MINIMAL RISK LEVELS

Inhalation MRLs

No MRLs were derived for inhalation exposure to selenium because of insufficient quantitative data concerning both human and animal exposures. Data on the health effects of inhaled selenium in humans are derived from studies of occupationally exposed workers. These studies suggest that the respiratory system is the most sensitive end point for inhaled selenium dust, but they do not provide quantitative measurements of exposure and are frequently confounded by concurrent exposure to other chemicals. Animal studies support the respiratory system as the target of selenium toxicity, but these are acute studies of exposure to high concentrations of selenium that also produced serious health effects and death.

Oral MRLs

No MRLs were derived for acute or intermediate oral exposure to selenium because of insufficient quantitative data concerning both human and experimental animal exposures. For acute exposure, no quantitative data are available from studies of humans. Some acute animal studies provide less serious LOAELs for organ weight changes, behavioral changes, and reduced body weight, which occur at doses similar to those producing serious LOAELs for paralysis and developmental effects in other studies. For intermediate exposure, one human study provided quantitative data suggesting that a selenium intake of 0.0048 mg/kg/day may affect thyroid hormone levels, but this study was available only as an abstract.

The effects seen in intermediate animal studies include reductions in enzyme activities, changes in liver and body weights, and histological changes in the liver and kidney. However, the relevance of animal models to selenium toxicity in humans is questionable. Humans with selenosis did not display any changes in serum levels of liver enzymes or morphological damage to the liver when examined by "supersonic B" examination, and the liver and kidney effects observed in animal studies occurred at doses ($0.2 mg/kg/day) that were considerably higher than those that produced changes in thyroid hormone levels in humans.

The chronic oral MRL is based on a study by Yang and Zhou. This study was an examination of a group of five individuals recovering from selenosis drawn from a larger population studied by the same authors. Yang et al. examined a population in an area of China where selenosis occurred. The study collected data on selenium levels in the diet, blood, nails, hair, urine, and milk of residents at three sites with low, medium, and high selenium, and compared the incidence of clinical symptoms of selenosis (morphological changes in finger nails) with dietary intake of selenium and selenium levels in blood. They found that selenium levels in blood corresponded to the dietary intake of selenium, and that symptoms of selenosis were found at or above a selenium intake level of 910 µg/day (0.016 mg/kg/day). In 1992, Yang and Zhou reexamined five individuals from the high selenium site who had been suffering from symptoms of selenosis (loss of fingernails and hair), but were recovering (nails were regrowing). Since their earlier report, the living conditions of the population had improved; they had been cautioned against consuming high selenium foods and parts of their locally produced corn had been replaced with rice or cereals. Yang and Zhou found that the concentration of selenium in the blood of these individuals had fallen from 1,346 µg/L (measured in 1986) to 968 µg/L (measured in 1992). Using a regression equation derived from the data in an earlier report, Yang and Zhou calculated that the dietary intake of selenium associated with selenosis in these individuals was 1,270 µg/day, while an intake of 819 µg selenium/day was associated with recovery.

- An MRL of 0.005 mg/kg/day has been derived for chronic oral exposure (>365 days) to selenium. This MRL is based on a NOAEL of 819 µg/day for disappearance of symptoms of selenosis in recovering individuals and uses an uncertainty factor of 3 for human variability. An uncertainty factor of 3 was considered appropriate, because the individuals in this report were sensitive individuals drawn from the larger population in the Yang et al. studies and because of the supporting studies. The MRL is consistent with the NOAELs observed for other human populations.

3. HEALTH EFFECTS

3.1 INTRODUCTION

The primary purpose of this chapter is to provide public health officials, physicians, toxicologists, and other interested individuals and groups with an overall perspective on the toxicology of selenium. It contains descriptions and evaluations of toxicological studies and epidemiological investigations and provides conclusions, where possible, on the relevance of toxicity and toxicokinetic data to public health.

A glossary and list of acronyms, abbreviations, and symbols can be found at the end of this profile.

Selenium is a naturally occurring element that is widely distributed in rocks and soils. Although selenium has been reported at hazardous waste sites where it can occur in many forms, analysis of specific forms present at these sites has not been performed. Selenium has multiple oxidation states (valence states) including -2, 0, +4, and +6. The type of selenium found is a result of its oxidation state which may vary according to ambient conditions, such as pH and microbial activity.

Elemental selenium (selenium[0]) is rarely found naturally, but it is stable in soils. Selenates (selenium[+6]) and selenites (selenium[+4]) are water soluble and can be found in water. Sodium selenate is among the most mobile forms of selenium; because of its high solubility and inability to adsorb to soil particles. More insoluble forms, such as elemental selenium, are less mobile; therefore, there is less risk for exposure. Because of greater bioavailability, water-soluble selenium compounds are probably more lethal than elemental selenium by any route. Selenium is found in nature complexed with multiple compounds, and although various forms are discussed in the profile, many others exist. Plants can contain organic selenium in the form of the amino acids selenomethionine and selenocysteine, along with the dimethyl selenides. Elemental selenium can be oxidized to form selenium dioxide. While the products of oxidation might be expected at the soil surface, elemental selenium would be the expected predominant form in deep soils where anaerobic conditions exist. Selenium sulfides, used in some anti-dandruff shampoos, are not very water soluble and, therefore, like elemental selenium are relatively immobile in the environment.

Much of the selenium released to the environment comes from the burning of coal and other fossil fuels, and from other industrial processes such as the production of rubber. For more information on the

physical and chemical properties of selenium, see Chapter 4. For more information on the potential for human exposure, see Chapter 6.

In humans and animals, selenium is an essential nutrient that plays a role in protecting tissues from oxidative damage as a component of glutathione peroxidase. It is also found in the deiodinases, including type I and II iodothyronine 5'-deiodinase, which convert thyroxine to triiodothyronine and in thioredoxin reductase, which catalyses the NADPH-dependent reduction of the redox protein thioredoxin. The biologically active form of selenium in these enzymes is the modified amino acid, selenocysteine. Humans and animals can be exposed to excessive amounts of selenium through the use of dietary supplements containing selenium. The nutritional role of selenium is further discussed in Section 3.4. Although selenium is an essential nutrient, exposure to high levels via inhalation or ingestion may cause adverse health effects. The mechanism by which selenium exerts toxic effects is unknown, but existing theories are discussed in Section 3.5. Most of the studies available on health effects involve exposure to selenite, selenate, and the organic chemical selenomethionine.

Several factors should be considered when evaluating the toxicity of selenium compounds. The purity and grade of the particular test substance used in the testing are important factors. For example, in studies of selenium sulfide compounds, the amounts of mono- and disulfides are often not specified by the study authors. The solubility and the particle size of selenium compounds also influence their toxicity.

3.2 DISCUSSION OF HEALTH EFFECTS BY ROUTE OF EXPOSURE

To help public health professionals and others address the needs of persons living or working near hazardous waste sites, the information in this section is organized first by route of exposure (inhalation, oral, and dermal) and then by health effect (death, systemic, immunological, neurological, reproductive, developmental, genotoxic, and carcinogenic effects). These data are discussed in terms of three exposure periods: acute (14 days or less), intermediate (15–364 days), and chronic (365 days or more).

Levels of significant exposure for each route and duration are presented in tables and illustrated in figures. The points in the figures showing no-observed-adverse-effect levels (NOAELs) or lowest-observed-adverse-effect levels (LOAELs) reflect the actual doses (levels of exposure) used in the studies. LOAELS have been classified into "less serious" or "serious" effects. "Serious" effects are those that evoke failure in a biological system and can lead to morbidity or mortality (e.g., acute respiratory distress or death). "Less serious" effects are those that are not expected to cause significant dysfunction

or death, or those whose significance to the organism is not entirely clear. ATSDR acknowledges that a considerable amount of judgment may be required in establishing whether an end point should be classified as a NOAEL, "less serious" LOAEL, or "serious" LOAEL, and that in some cases, there will be insufficient data to decide whether the effect is indicative of significant dysfunction. However, the Agency has established guidelines and policies that are used to classify these end points. ATSDR believes that there is sufficient merit in this approach to warrant an attempt at distinguishing between "less serious" and "serious" effects. The distinction between "less serious" effects and "serious" effects is considered to be important because it helps the users of the profiles to identify levels of exposure at which major health effects start to appear. LOAELs or NOAELs should also help in determining whether or not the effects vary with dose and/or duration, and place into perspective the possible significance of these effects to human health.

The significance of the exposure levels shown in the Levels of Significant Exposure (LSE) tables and figures may differ depending on the user's perspective. Public health officials and others concerned with appropriate actions to take at hazardous waste sites may want information on levels of exposure associated with more subtle effects in humans or animals (LOAELs) or exposure levels below which no adverse effects (NOAELs) have been observed. Estimates of levels posing minimal risk to humans (Minimal Risk Levels or MRLs) may be of interest to health professionals and citizens alike.

Estimates of exposure levels posing minimal risk to humans (Minimal Risk Levels or MRLs) have been made for selenium. An MRL is defined as an estimate of daily human exposure to a substance that is likely to be without an appreciable risk of adverse effects (noncarcinogenic) over a specified duration of exposure. MRLs are derived when reliable and sufficient data exist to identify the target organ(s) of effect or the most sensitive health effect(s) for a specific duration within a given route of exposure. MRLs are based on noncancerous health effects only and do not consider carcinogenic effects. MRLs can be derived for acute, intermediate, and chronic duration exposures for inhalation and oral routes. Appropriate methodology does not exist to develop MRLs for dermal exposure.

Although methods have been established to derive these levels (Barnes and Dourson 1988; EPA 1990), uncertainties are associated with these techniques. Furthermore, ATSDR acknowledges additional uncertainties inherent in the application of the procedures to derive less than lifetime MRLs. As an example, acute inhalation MRLs may not be protective for health effects that are delayed in development or are acquired following repeated acute insults, such as hypersensitivity reactions, asthma, or chronic

bronchitis. As these kinds of health effects data become available and methods to assess levels of significant human exposure improve, these MRLs will be revised.

A User's Guide has been provided at the end of this profile (see Appendix B). This guide should aid in the interpretation of the tables and figures for Levels of Significant Exposure and the MRLs.

3.2.1 Inhalation Exposure

Table 3-1 and Figure 3-1 describe the health effects observed in experimental animals that inhaled elemental selenium dust or hydrogen selenide. Studies using other forms of selenium were not used in the LSE tables and figures (Table 3-1 and Figure 3-1) because either the reporting of the studies was incomplete or no studies on other forms were located. All doses are expressed in terms of total selenium.

3.2.1.1 Death

No studies were located regarding death in humans after inhalation of elemental selenium or selenium compounds.

In animals, the acute lethality of hydrogen selenide and elemental selenium dust when inhaled has been investigated. In guinea pigs exposed to hydrogen selenide for 2, 4, or 8 hours, 5/16 died within 10 days of exposure at 12 mg selenium/m^3, 3/16 died at 6 mg selenium/m^3, and 8/16 died at 6 mg selenium/m^3, respectively (Dudley and Miller 1941).

No deaths were observed among rabbits or guinea pigs exposed to elemental selenium dust at levels of 31 mg selenium/m^3 for 4 hours every other day for 8 exposure days (Hall et al. 1951). Higher levels were not tested.

All LOAEL values from each reliable study for death in each species and duration category are recorded in Table 3-1 and plotted in Figure 3-1.

Table 3-1. Levels of Significant Exposure to Selenium - Inhalation

Key to figure[a]	Species (strain)	Exposure/ duration/ frequency	System	NOAEL (mg/m3)	LOAEL Less serious (mg/m3)	LOAEL Serious (mg/m3)	Reference Chemical Form
ACUTE EXPOSURE							
Death							
1	Gn Pig (NS)	4 hr				6 (3/16 died)	Dudley and Miller 1941 hydrogen selenide
2	Gn Pig (NS)	8 hr				1 (8/16 died)	Dudley and Miller 1941 hydrogen selenide
3	Gn Pig (NS)	2 hr				12 (8/16 died)	Dudley and Miller 1941 hydrogen selenide
Systemic							
4	Rat (NS)	8 hr	Resp			33 F (pulmonary hemorrhage, pneumonitis)	Hall et al. 1951 elemental
			Hepatic		33 F (congestion; mild central atrophy)		
			Renal	33 F			
			Endocr	33 F			
			Bd Wt	33 F			
5	Gn Pig (NS)	4 hr	Resp			8 (pneumonitis)	Dudley and Miller 1941 hydrogen selenide
			Cardio	8			
			Hepatic		8 (fatty metamorphosis, increased liver weight)		
			Renal	8			
			Endocr	8			

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Table 3-1. Levels of Significant Exposure to Selenium - Inhalation (continued)

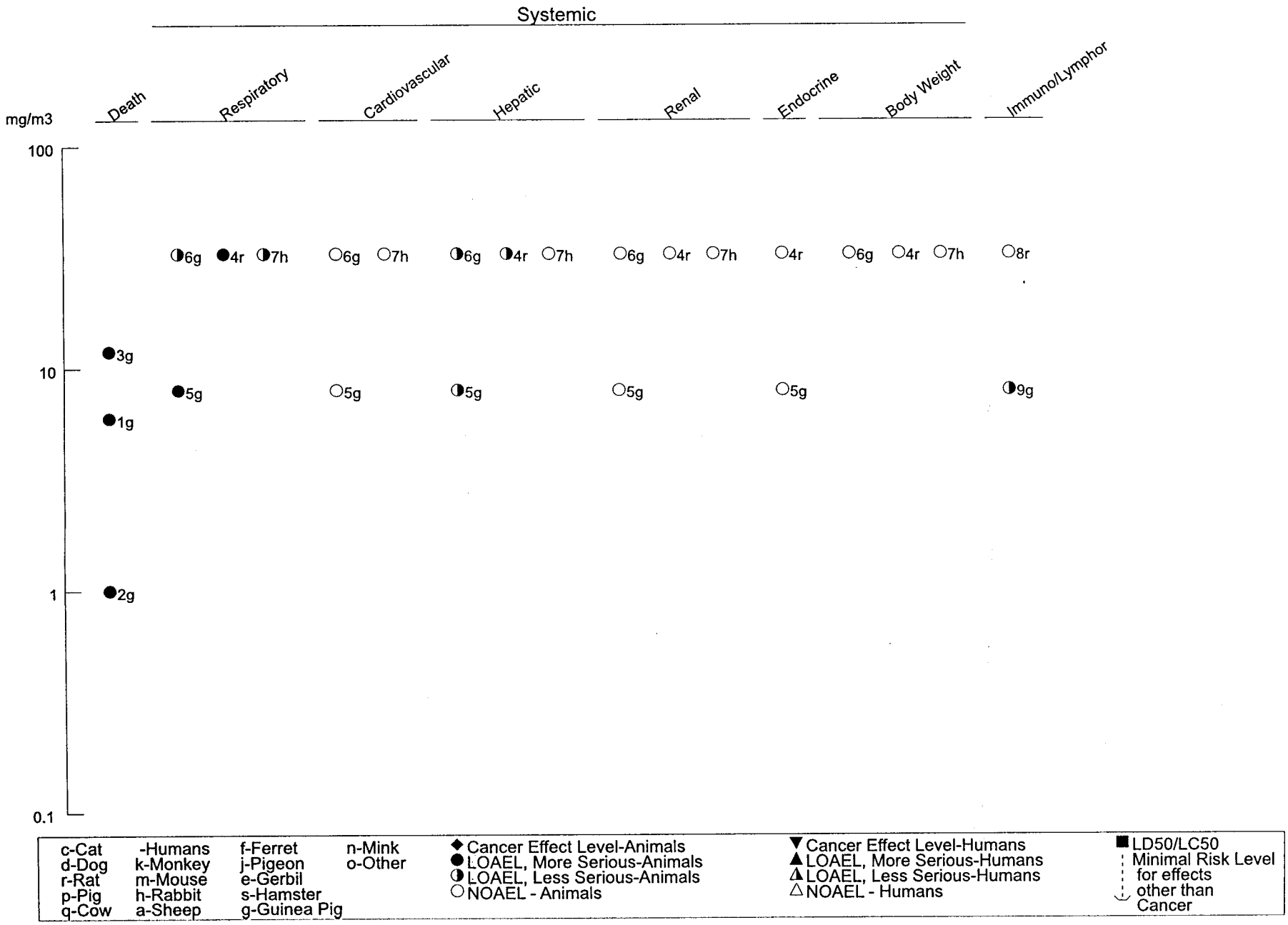
Key to figure[a]	Species (strain)	Exposure/ duration/ frequency	System	NOAEL (mg/m3)	LOAEL Less serious (mg/m3)	LOAEL Serious (mg/m3)	Reference Chemical Form
6	Gn Pig (NS)	8 d 4hr/2d	Resp		33 M (mild congestion; mild to moderate interstitial pneumonitis; slight emphysema)		Hall et al. 1951 elemental
			Cardio	33 M			
			Hepatic		33 M (congestion; central atrophy; fatty metamorphosis)		
			Renal	33 M			
			Bd Wt	33 M			
7	Rabbit (NS)	8 d 4hr/2d	Resp		33 F (congestion, mild pneumonitis)		Hall et al. 1951 elemental
			Cardio	33 F			
			Hepatic	33 F			
			Renal	33 F			
			Bd Wt	33 F			
Immunological/Lymphoreticular							
8	Rat (NS)	8 hr		33			Hall et al. 1951 elemental
9	Gn Pig (NS)	4 hr			8 (splenic hyperplasia)		Dudley and Miller 1941 hydrogen selenide

[a]The number corresponds to entries in Figure 3-1.

Bd Wt = body weight; Cardio = cardiovascular; CEL = cancer effect level; d = day(s); (F) = feed; Endocr = endocrine; F = female; gastro = gastrointestinal; Hemato = hematological; hr = hour(s); LOAEL = lowest-observed-adverse-effect level; M = male; Metab = metabolic; Musc/skel = musculoskeletal; NOAEL = no-observed-adverse-effect level; (NS) = not specified; Resp = respiratory

Figure 3-1. Levels of Significant Exposure to Selenium - Inhalation

Acute (≤14 days)



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3.2.1.2 Systemic Effects

The selenium compounds that are most likely to be encountered in air in occupational settings are dusts of elemental selenium, hydrogen selenide, and selenium dioxide. Other volatile selenium compounds (e.g., dimethyl selenide, dimethyl diselenide) might be encountered in some naturally occurring situations. Because selenium is converted from one form to another, as in plant biosynthesis of selenoamino acids, it is not clear which specific forms may be encountered at hazardous waste sites. If a hazardous waste site specifically contains deposits of compounds of selenium, those compounds could be released off-site in dust or air. Toxicity data for exposures via inhalation are available for elemental selenium, selenium dioxide, selenium oxychloride, hydrogen selenide, and dimethyl selenide. Because there are few studies of inhalation of selenium of any single form, all available studies of inhalation exposures to selenium compounds will be included in this discussion.

In studies of human occupational exposures, it appears that the respiratory tract is the primary site of injury after inhalation of selenium dust or selenium compounds, but gastrointestinal (possibly due to swallowed selenium) and cardiovascular effects, as well as irritation of the skin and eyes, also occur. Little of the available information for humans, however, relates health effects exclusively to measured concentrations of the selenium dust or compounds because of the possibility of concurrent exposures to multiple substances in the workplace. In animals, the respiratory tract is also the primary site of injury following inhalation exposure to selenium dust and hydrogen selenide. Hematological and hepatic effects have also been noted in animals. Inhalation data from laboratory animal studies are available only for acute exposures.

No information was located regarding hematological, musculoskeletal, dermal, or ocular effects in humans or laboratory animals after inhalation exposure to selenium or selenium compounds. The systemic effects that have been observed after inhalation exposure are discussed below. The highest NOAEL values and all LOAEL values for each reliable study for systemic effects in each species and duration category are recorded in Table 3-1 and plotted in Figure 3-1.

Respiratory Effects. In humans, the respiratory system is the primary site of injury after inhalation of elemental selenium or selenium compounds. The largest number of reported human exposures occurred in occupational settings, especially in industries that extract, mine, treat, or process selenium-bearing minerals and in industries that use selenium or selenium compounds in manufacturing. The reports of occupational exposure do not link observed symptoms to specific air concentrations of

elemental selenium or selenium compounds. Several reports, however, have noted common effects associated with inhalation exposure in occupational settings.

Selenium dioxide is formed when selenium is heated in air. Direct exposure to selenium dioxide is, therefore, primarily an occupational hazard and not likely to be a risk at hazardous waste sites. Selenium dioxide forms selenious acid on contact with water, including perspiration, and can cause severe irritation. Acute inhalation of large quantities of selenium dioxide powder can produce pulmonary edema as a result of the local irritant effect on alveoli (Glover 1970). Bronchial spasms, symptoms of asphyxiation, and persistent bronchitis have been noted in workers briefly exposed to high concentrations of selenium dioxide (Wilson 1962). Kinnigkeit (1962) reported that selenium dioxide concentrations of 0.007–0.05 mg selenium/m^3 in a selenium rectifier plant produced slight tracheobronchitis in 9 of 62 exposed workers.

Hydrogen selenide, a highly poisonous selenium compound, is a gas at room temperature, with a density much higher than air. Selenium oxychloride, also highly toxic, is more irritating and corrosive to the human respiratory tract than are other forms of selenium because the compound hydrolyzes to hydrogen chloride (HCl), which can then form hydrochloric acid in humid air and in the respiratory tract (Dudley 1938). Hydrogen selenide and selenium oxychloride are occupational exposure hazards that are not expected to be much of a concern at hazardous waste sites.

Acute inhalation exposure to elemental selenium dust, possibly including some selenium dioxide, in occupational settings has been shown to irritate mucous membranes in the nose and throat and produce coughing, nosebleed, loss of olfaction, and in heavily exposed workers, dyspnea, bronchial spasms, bronchitis, and chemical pneumonia (Clinton 1947; Hamilton 1949). Chronic exposure of 40 workers at a copper refinery produced increased nose irritation and sputum (Holness et al. 1989). The exact concentration of selenium was not given, but the concentration was reported to exceed 0.2 mg selenium/m^3. Confounding variables in this study include concurrent exposure to several other metals including copper, nickel, silver, lead, arsenic, and tellurium.

In experimental animals, the respiratory tract is the primary site of injury following acute inhalation exposure to elemental selenium and selenium compounds. Rats exposed to selenium fumes (selenium concentration and particle size were not reported) for 2–16 minutes experienced moderate to severe respiratory effects, including hemorrhage and edema of the lungs (Hall et al. 1951). Rats exposed to selenium dust (average particle diameter, 1.2 μm) at levels of 33 mg selenium/m^3 for 8 hours experienced

severe respiratory effects, including hemorrhage and edema of the lungs, and several animals died (Hall et al. 1951). Histopathological examinations of surviving animals revealed chronic interstitial pneumonitis. Acute exposure of rabbits and guinea pigs to selenium dust (average particle diameter, 1.2 μm) at a concentration of 33 mg selenium/m^3 resulted in mild interstitial pneumonitis or congestion, and slight emphysema in both species (Hall et al. 1951). Other histological findings included vascular lymphocytic infiltration and intra-alveolar foci of large macrophages.

Acute inhalation exposure of guinea pigs to 8 mg selenium/m^3 as hydrogen selenide for 4 hours, produced diffuse bronchopneumonia and pneumonitis (Dudley and Miller 1941). The investigators do not indicate if any of these guinea pigs died as a result of the exposure. Histologic examination of animals that had died following exposure to higher concentrations revealed thickening of the alveolar walls and congestion of alveolar capillaries (Dudley and Miller 1937). In contrast, 1-hour exposure of rats to 25,958 mg selenium/m^3 as dimethyl selenide produced only minor effects (increased weight of lung and liver) 1 day postexposure. These changes disappeared by 7 days postexposure (Al-Bayati et al. 1992). Enzymatic methylation of selenium compounds is the primary route of detoxification and may explain the low toxicity of dimethyl selenide (Al-Bayati et al. 1992). Although this form of selenium is environmentally relevant since it is formed in soil, plants, and microorganisms, dimethyl selenide appears to be relatively nontoxic in comparison to occupational exposure to hydrogen selenide.

The effects of intratracheal instillation of selenium on pulmonary function may be dependent on the form in which it is supplied (Nonavinakere et al. 1999). Instillation of 0.06 mg selenium/100 g body weight as selenium dioxide produced a significant decrease in respiratory rate and a significant increase in lung resistance compared with controls. Instillation with 0.06 mg selenium/110 g body weight as seleno-L-methionine also produced a decrease in respiratory rate and an increase in lung resistance, but the values were not significantly different from controls.

Intratracheal instillation of 0.3 mg selenium as sodium selenite in male Hartley-guinea pigs decreased dynamic-lung-compliance and increased pulmonary resistance compared with control animals instilled with saline (Bell et al. 1997). Analysis of bronchoalveolar-lavage fluid showed increased activities of lactate dehydrogenase, β-glucuronidase, alkaline phosphatase, and protein, suggesting damage to lung tissue.

Histological analysis of guinea pigs that received single intratracheal instillations of 0.3 mg selenium as sodium selenite found mild acute inflammation in approximately one-third of the lung tissue and a

noticeable amount of sloughed epithelium and mucus within the bronchi (Bell et al. 2000). Lungs of animals treated with 0.06 mg selenium showed neutrophils aggregated in the alveoli and some dilation of the alveoli suggestive of emphysema. Relative lung weights and the ratio of wet/dry lung weight were increased in the selenium-treated animals compared with controls; the increase was only significant for those receiving the higher dose of selenium. Leucocyte counts in bronchoalveolar-lavage fluid were decreased for selenium-treated animals compared with controls, and the difference was significant for the animals receiving 0.3 mg selenium, but not the 0.06 mg dosage.

No studies were located regarding respiratory effects in animals after intermediate or chronic inhalation of selenium or selenium compounds.

Cardiovascular Effects. Several workers experienced symptoms of shock, including lower blood pressure and elevated pulse rates, following an acute exposure (at most 20 minutes) to selenium dioxide fumes resulting from a fire (Wilson 1962). The subjects were treated with oxygen and inhalation of ammonia vapor, and pulse rates were normalized within 3 hours.

Cardiovascular effects were not observed in guinea pigs exposed to hydrogen selenide at 8 mg selenium/m^3 for 4 hours (Dudley and Miller 1941), or in guinea pigs and rabbits exposed to elemental selenium dust (average particle diameter, 1.2 μm) every other day at 33 mg selenium/m^3 for eight 4-hour exposure periods (Hall et al. 1951).

Gastrointestinal Effects. Vomiting and nausea were reported in workers exposed to high concentrations of selenium dioxide for a maximum of 20 minutes during a fire (Wilson 1962). Stomach pain was frequently reported by workers exposed to elemental selenium and selenium dioxide at a selenium rectifier plant (Glover 1967), and by copper refinery workers exposed to an unspecified form of selenium (Holness et al. 1989). Exposure concentrations were not reported for the rectifier plant, but were greater than 0.2 mg selenium/m^3 at the copper refinery.

No studies were located regarding gastrointestinal effects in animals after inhalation of selenium or selenium compounds.

Hepatic Effects. No studies were located regarding hepatic effects in humans after inhalation of selenium or selenium compounds.

Hepatoxicity has been observed in experimental animals following inhalation exposure to elemental selenium dust and to hydrogen selenide. One month after an 8-hour exposure to elemental selenium dust at a level of 33 mg selenium/m^3, most rats exhibited slight liver congestion, and a few exhibited mild centrilobular atrophy (Hall et al. 1951). In contrast, 1 week after exposure to 25,958 mg selenium/m^3 as dimethyl selenide for 1 hour, rats showed no observable changes in the liver (Al-Bayati et al. 1992). Three weeks following acute exposure to elemental selenium dust at a level of 33 mg selenium/m^3 for 4 hours every other day for 8 days, 4/10 guinea pigs exhibited slight hepatic congestion with mild central atrophy, and 2/10 showed some fatty hepatocellular degeneration (Dudley and Miller 1941). In contrast, exposure of guinea pigs to lower concentrations of selenium (8 mg/m^3), as hydrogen selenide, for a single 4-hour period produced mild fatty hepatocellular metamorphosis (Dudley and Miller 1941).

Renal Effects. No studies were located regarding renal effects in humans after inhalation of selenium or selenium compounds.

The kidneys do not appear to be affected in guinea pigs (Dudley and Miller 1941; Hall et al. 1951) after acute inhalation exposure to 33 mg selenium/m^3 as hydrogen selenide for 8 hours or to 8 mg selenium/m^3 as elemental selenium dust for 4 hours. Likewise, the kidneys were not affected in rabbits following acute inhalation exposure to 33 mg selenium/m^3 as hydrogen selenide for 8 hours (Hall et al. 1951) or in rats following acute inhalation exposure to 25,958 mg selenium/m^3 as dimethyl selenide for 1 hour or to 33 mg selenium/m^3 as hydrogen selenide for 8 hours (Al-Bayati et al. 1992; Hall et al. 1951).

Endocrine Effects. No studies were located regarding endocrine effects in humans after inhalation of selenium or selenium compounds.

No histopathological changes in the adrenal gland were observed in guinea pigs exposed to hydrogen selenide at 8 mg selenium/m^3 for 4 hours (Dudley and Miller 1941), or in rats exposed to elemental selenium at 33 mg selenium/m^3 for 8 hours (Hall et al. 1951).

Body Weight Effects. No studies were located regarding effects on body weight in humans following inhalation of selenium or selenium compounds.

No effects on body weight were observed in guinea pigs following a single 8-hour exposure to elemental selenium at 33 mg selenium/m^3, or in guinea pigs and rabbits exposed to elemental selenium dust at 33 mg selenium/m^3 every other day for 4 hours for a total of eight exposures (Hall et al. 1951).

3.2.1.3 Immunological and Lymphoreticular Effects

No studies were located regarding immunological or lymphoreticular effects in humans after inhalation exposure to selenium or selenium compounds.

Lymphoid hyperplasia was noted in the spleen of guinea pigs following a single 4-hour exposure at 8 mg selenium/m^3 as hydrogen selenide (Dudley and Miller 1941). Histopathological changes in the spleen were not observed in guinea pigs exposed to elemental selenium dust (average particle diameter, 1.2 μm) at 33 mg selenium/m^3 for 8 hours (Hall et al. 1951). Injury to the spleen was observed in guinea pigs following exposure for 4 hours, every other day, for 8 days to elemental selenium dust at a level of 33 mg selenium/m^3 (Hall et al. 1951). Specific effects included congestion of the spleen, fissuring red pulp, and increased polymorphonuclear leukocytes (Hall et al. 1951).

3.2.1.4 Neurological Effects

Information concerning possible neurological effects caused by inhalation of selenium or selenium compounds is limited. Severe frontal headaches were reported by workers exposed during an accident to high concentrations of selenium fumes (compound not stated) for approximately 2 minutes (Clinton 1947). Workers at a selenium rectifier plant reported symptoms of malaise and irritability when working with selenium (exposure probably to selenium dioxide and elemental selenium, but form was not stated) (Glover 1967). The symptoms resolved whenever the workers were moved to other work. Urinary concentrations of selenium were about 0.08 mg/L, compared to 0.024–0.034 mg/L in unexposed workers.

No studies were located regarding neurological effects in animals after inhalation of selenium or selenium compounds.

No studies were located regarding the following health effects in humans or animals after inhalation exposure to selenium or selenium compounds:

3.2.1.5 Reproductive Effects

3.2.1.6 Developmental Effects

3.2.1.7 Cancer

There are no epidemiologic data that support a causal association between the inhalation of elemental selenium dusts or selenium compounds and the induction of cancer in humans (Gerhardsson et al. 1986; Wester et al. 1981). In one study, postmortem samples were collected from copper smelter workers who were exposed to several different airborne compounds, including selenium compounds. Samples from lung cancer cases had lower concentrations of selenium in lung tissue than samples from controls or from workers who had died from other causes (Gerhardsson et al. 1986). In another autopsy study of smelter workers, Wester et al. (1981) found that the selenium concentrations in kidney tissues from workers who had died of malignancies were lower than the selenium concentrations in kidney tissues from workers who died of other causes. Further discussions regarding the cancer protective effects of selenium can be found in Section 3.2.2.7.

No studies were located regarding carcinogenic effects in laboratory animals after inhalation exposure to selenium or selenium compounds.

3.2.2 Oral Exposure

Table 3-2 and Figure 3-2 describe the health effects observed in humans and experimental animals associated with dose and duration of oral exposure to selenium and selenium compounds (i.e., elemental selenium dust, selenium dioxide dissolved in water [selenious acid], sodium selenate, sodium selenite, potassium selenate, and dietary selenium compounds, which include selenoamino acids). All doses for these compounds are expressed in terms of total selenium. Table 3-3 and Figure 3-3 describe health effects observed in laboratory animals following oral exposure to selenium sulfides (SeS_2 and SeS) at varying doses and exposure durations. All doses for selenium sulfide compounds are expressed in terms of the compound, because selenium sulfide preparations often exist as a variable mixture of the mono- and disulfide forms, precluding accurate expression of the dose in terms of total selenium.

Table 3-2. Levels of Significant Exposure to Selenium - Oral

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
	ACUTE EXPOSURE						
	Death						
1	Rat (Sprague-Dawley)	once (G)				6700 M (LD_{50})	Cummins and Kimura 1971 elemental
2	Rat (Sprague-Dawley)	once (GW)				7 M (LD_{50})	Cummins and Kimura 1971 selenite
3	Rat (Sprague-Dawley)	14 d ad lib (W)				0.418 F (7/12 died)	NTP 1996 sodium selenate
4	Rat (NS)	once (G)				4.8 F (LD_{50})	Pletnikova 1970 selenite
5	Rat (Wistar)	once (GW)				48 (LD_{50})	Singh and Junnarkar 1991 selenium dioxide
6	Mouse (NS)	once (G)				3.2 M (LD_{50})	Pletnikova 1970 selenite
7	Mouse (ICR)	once (G)				35.9 M (LD_{50})	Sayato et al. 1993 D,L-selenocystine
8	Mouse (Swiss)	once (GW)				16 M (LD_{50})	Singh and Junnarkar 1991 selenium dioxide
9	Gn Pig (NS)	once (G)				2.3 F (LD_{50})	Pletnikova 1970 selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
10	Rabbit (NS)	once (G)				1.0 F (LD_{50})	Pletnikova 1970 selenite
	Systemic						
11	Rat (Sprague-Dawley)	14 d ad lib (W)	Bd Wt	0.251 F		0.418 F (significant (36%) reduction in body weight)	NTP 1996 sodium selenate
12	Mouse (BALB/c)	14 d ad lib (W)	Hemato	0.38 M	0.82 M (significant increase in red blood cell count)		Johnson et al. 2000 Selenite
			Hepatic	0.38 M	0.82 M (significant decrease in relative liver weight)		
			Renal	0.17 M	0.38 M (significant increase in relative kidney weight)		
			Bd Wt	0.38 M	0.82 (significant decrease in body weight gain)		
13	Mouse (BALB/c)	14 d ad lib (W)	Hemato	1.36 M			Johnson et al. 2000 selenomethionine
			Hepatic	1.36 M			
			Renal	1.36 M			
			Bd Wt	1.36 M			
14	Pig (NS)	5 d	Resp	1.25			Panter et al. 1996 organic
			Cardio	1.25			
			Hepatic	1.25			
			Renal	1.25			
			Dermal	1.25			
			Bd Wt			1.25 (5% loss of body weight)	

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
Immunological/Lymphoreticular							
15	Mouse (BALB/c)	14 d ad lib (W)		0.38 M	0.82 M (increased proliferation of splenic lymphocytes and LPS-induced production of TNF alpha and IL-1beta)		Johnson et al. 2000 Selenite
16	Mouse (BALB/c)	14 d ad lib (W)		1.36 M			Johnson et al. 2000 selenomethionine
Neurological							
17	Mouse (Swiss)	once (GW)			1.6 (decreased activity, muscle tone, touch response, respiration; hypothermia)		Singh and Junnarkar 1991 selenium dioxide
18	Mouse (BALB/c)	14 d ad lib (W)		0.24 M	0.58 M (significant increase in the levels of striatal dihydroxyphenylacetic acid and homovanillic acid)		Tsunoda et al. 2000 Selenite
19	Mouse (BALB/c)	14 d ad lib (W)		1.96 M			Tsunoda et al. 2000 Organic selenium
20	Pig (NS)	10 d 1x/d (C)				1.3 (hypoactivity, focal symmetrical poliomalacia, histopathological lesions in brain and spinal cord)	Wilson et al. 1989 selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
Developmental							
21	Hamster (Syrian LKV)	once Gd 8 (GW)		7.1		7.9 (encephalocele; decreased crown-rump length)	Ferm et al. 1990 selenite
22	Hamster (Syrian LKV)	once Gd 8 (GW)				7.1 (encephalocele)	Ferm et al. 1990 selenate
23	Hamster (Syrian LKV)	once Gd 8 (GW)			5.9 (decreased fetal crown-rump length)		Ferm et al. 1990 selenomethionine

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
	INTERMEDIATE EXPOSURE						
	Death						
24	Rat (Sprague-Dawley)	6 wk ad lib (F)				0.48 M (1/8 died)	Halverson et al. 1966 selenite
25	Rat (Sprague-Dawley)	6 wk ad lib (F)				0.4 M (1/8 died)	Halverson et al. 1966 organic
26	Rat (Fischer- 344)	13 wk (W)				2.54 (20/20 died)	NTP 1994 selenate
27	Rat (Fischer- 344)	13 wk (W)				1.67 F (2/10 died)	NTP 1994 selenite
28	Rat (Sprague-Dawley)	4-6 wk ad lib (W)				0.84 M (4/6 died)	Palmer and Olson 1974 selenite
29	Rat (Sprague-Dawley)	4-6 wk ad lib (W)				0.84 M (2/6 died)	Palmer and Olson 1974 selenate
30	Rat (Wistar)	1 yr daily ad lib (W)				1.05 F (3/5 died) 1.05 M (1/3 died)	Rosenfeld and Beath 1954 selenate

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
31	Rat (BLU:[LE])	365 d (W)				0.28 (50% males died at 58 days 50% females died at 160 days)	Schroeder and Mitchener 1971a selenite
32	Mouse (ICR)	30 d 6d/wk (G)				14.2 M (15/15 died)	Sayato et al. 1993 D,L-selenocystine
	Systemic						
33	Human	20 wk (IN)	Endocr	0.001			Duffield et al. 1999
34	Human	120 d (F)	Endocr		0.0048 M (significant increase in thyroid stimulating hormone)		Hawkes and Keim 1995 NS
35	Monkey (Macaca fascicularis)	gd 20-50 1x/d (GW)	Gastro	0.025 F	0.15 F (vomiting)		Tarantal et al. 1991 selenomethionine
			Bd Wt	0.025 F	0.15 F (increased weight loss)		
36	Rat (Wistar)	110 d ad lib (F)	Endocr	0.105 M	0.105 M (significant reduction in type I deiodinase activity)		Behne et al. 1992 sodium selenite
			Bd Wt	0.105 M			
37	Rat (Wistar)	110 d ad lib (F)	Endocr		0.118 M (significant reduction in type I deiodinase activity)		Behne et al. 1992 selenomethionine
			Bd Wt		0.118 M (significant reduction in body weight (15%))		

Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
38	Rat (Sprague-Dawley)	2 mo ad lib (F)	Hepatic	0.1 M	0.2 M (nodular regenerative hyperplasia, increased relative liver weight)		Bioulac-Sage et al. 1992 selenite
			Bd Wt	0.2 M			
39	Rat (Sprague-Dawley)	8 wk (F)	Resp	0.45 M			Chen et al. 1993 selenite
			Cardio	0.45 M			
			Gastro	0.45 M			
			Hepatic		0.45 M (diffuse panlobular vacuolar accumulation of glycogen and lipid)		
			Renal	0.45 M			
			Bd Wt		0.25 M (final body weights about 14% lower than controls)	0.35 M (final body weights about 29% lower than controls)	
40	Rat (Sprague-Dawley)	40 d ad lib (F)	Hemato	0.27 M			Eder et al. 1995 sodium selenite
			Endocr	0.026 M	0.055 M (significant reduction in serum tri-iodothyronine levels)		
			Bd Wt	0.27 M			

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
41	Rat (Sprague-Dawley)	6 wk ad lib (F)	Hemato	0.24 M	0.32 M (23% decrease in hemoglobin)	0.56 M (79% decrease in hemoglobin)	Halverson et al. 1966 organic
			Hepatic		0.4 M (6-fold increase in bilirubin)		
			Endocr	0.32 M	0.4 M (pancreas weight 1.4 times greater than diet restricted controls)		
			Bd Wt	0.32		0.4 M (body weight gain 36% lower than controls)	
42	Rat (Sprague-Dawley)	6 wk ad lib (F)	Endocr		0.09 M (significant increase in serum TSH (~30%))		Hotz et al. 1997 sodium selenate
			Bd Wt	0.09 M			
			Metab		0.09 M (significant increase in GSH-Px in kidney (~30%) and erythrocytes (~100%))		
43	Rat (Wistar)	3 mo 1x/d (F)	Hepatic		0.002 M (sporadic infiltrations of mononuclear cells in portal canals and weak activation of Kupffer cells)	0.005 M (distict swelling of Kupffer cells in dilated sinusoidal vessels and necrotic areas comprising single groups of hepatocytes)	Kolodziejczyk et al. 2000 sodium selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
44	Rat (Fischer- 344)	13 wk (W)	Resp	1.57 M			NTP 1994 selenate
			Cardio	1.57 M			
			Gastro	1.57 M			
			Hemato	0.92 M	1.57 M (increased hematocrit and hemoglobin associated with decreased water intake)		
			Musc/skel	1.57 M			
			Hepatic	0.92 M	1.57 M (increased bile acids indicating cholestasis)		
			Renal	0.31 F	0.47 F (minimal papilla degeneration of the kidneys)		
			Endocr	1.57 M			
			Ocular	1.57 M			
			Bd Wt	0.47 F	0.88 F (body weights 10% less than controls)	1.35 F (body weights 29% less than controls, associated with decreased water intake)	

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
45	Rat (Fischer- 344)	13 wk (W)	Resp	1.67 F			NTP 1994 selenite
			Cardio	1.67 F			
			Gastro	1.67 F			
			Hemato	0.86 F	1.67 F (increased hematocrit associated with decreased water intake)		
			Musc/skel	1.67 F			
			Hepatic	1.67 F			
			Renal	0.28 F	0.5 F (mild papilla degeneration)		
			Endocr	1.67 F			
			Ocular	1.67 F			
			Bd Wt	0.98 M		1.59 M (body weights 34% less than controls; associated with decreased water intake)	
46	Rat (Sprague-Dawley)	23-29 d ad lib (W)	Bd Wt	0.167[b] M 0.209 F	0.293[b] M 0.334 F (significant (11%) reduction in body weight)	0.418 (significant (20% male, 39% female) reduction in body weight)	NTP 1996 sodium selenate
47	Rat (Sprague-Dawley)	4-6 wk ad lib (W)	Hepatic			0.84 M (cirrhosis)	Palmer and Olson 1974 selenate
			Bd Wt		0.42 M (body weight gain 10% lower than controls)		
48	Rat (Sprague-Dawley)	6 wk ad lib (F)	Bd Wt	0.125 M			Salbe and Levander 1990a selenate

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
49	Rat (Sprague-Dawley)	6 wk ad lib (F)	Bd Wt	0.125 M			Salbe and Levander 1990a selenomethionine
50	Rat (Wistar)	3-6 wks ad lib (W)	Endocr		0.64 F (decreased somatomedin C)		Thorlacius-Ussing 1990 selenite
			Bd Wt			0.64 F (body weight gain 30% lower than controls)	
51	Rat (Wistar)	12-14 wk ad lib	Cardio			0.324 (degeneration of heart tissue with disruption of myofibrils and sarcomeres)	Turan et al. 1999a sodium selenite
			Hepatic			0.324 (degeneration of liver tissue with dilation of sinusoidal capillaries)	
			Bd Wt		0.324 (significant decrease in body weight (17%))		
52	Mouse (ICR)	90 d (G)	Hepatic	2.4 M	4.7 M (increased serum aspartate aminotransferase and alanine aminotransferase)		Hasegawa et al. 1994 D,L-selenocystine
			Bd Wt	2.4 M	4.7 M (body weights 16% lower than controls)	7.1 M (body weights 22% lower than controls)	

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
53	Mouse (B6C3F1)	13 wk (W)	Resp	7.17 F			NTP 1994 selenate
			Cardio	7.17 F			
			Gastro	7.17 F			
			Hemato	7.17 F			
			Musc/skel	7.17 F			
			Hepatic	7.17 F			
			Renal	1.07 M	1.87 M (increased kidney weight associated with decreased water intake)		
			Endocr	7.17 F			
			Ocular	7.17 F			
			Bd Wt	1.87	2.95 M (body weights 13% lower than controls; decreased water intake)	5.45 M (body weights 24% lower than controls; decreased water intake)	

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
54	Mouse (B6C3F1)	13 wk (W)	Resp	3.83 F			NTP 1994 selenite
			Cardio	3.83 F			
			Gastro	3.83 F			
			Hemato	3.83 F			
			Musc/skel	3.83 F			
			Hepatic	3.83 F			
			Renal	0.91 M	1.61 M (increased relative kidney weight; decreased water intake)		
			Endocr	3.83 F			
			Ocular	3.83 F			
			Bd Wt	1.61 M		3.31 M (body weights 20% lower than controls; decreased water intake)	
55	Mouse (ICR)	30 d 6d/wk (G)	Hepatic	4.7 M	9.4 M (significant 2-3-fold increases in aspartate aminotransferase and alanine aminotransferase)		Sayato et al. 1993 D,L-selenocystine
			Renal	9.4 M			
			Bd Wt		9.4 M (final body weight about 13% lower than controls)	18.9 M (final body weight about 29% lower than controls)	
56	Mouse Balby	12 wk ad lib (F)	Hepatic		0.2 M (vacuolization of hepatocytes)		Skowerski et al. 1997a sodium selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
57	Mouse Balby	12 wk ad lib (F)	Cardio		0.2 M (cardiocytes have numerous damaged mitochondria, large number of lipid droplets and numerous lysosomes)		Skowerski et al. 1997b sodium selenite
			Bd Wt	0.2 M			
58	Rabbit (New Zealand)	3 mo ad lib (F)	Cardio			0.137 (disruption of myofibrils, irregular sarcomeres, and diosrganization of bands in sarcomeres)	Turan et al. 1999b sodium selenite
			Hemato	0.137			
			Bd Wt	0.137			
59	Pig (mixed breed)	8 wk ad lib (F)	Hepatic		1.1 (vacuolar degeneration, portal fibrosis)		Baker et al. 1989 selenate
			Dermal		1.1 (cracked hoof walls)		
			Bd Wt			1.1 (body weight gain 83% lower than controls, accompanied by decreased food intake)	
60	Pig (NS)	35 d ad lib (F)	Dermal	0.014	0.25 (hoof cracking)		Mahan and Magee 1991 selenite
			Bd Wt	0.25		0.47 (body weight gain 78% lower than controls, accompanied by decreased food intake)	

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
61	Pig (crossbred L x Y)	8 wk ad lib (F)	Hepatic	0.33		0.59 (atrophic cirrhosis)	Mihailovic et al. 1992 selenite
			Dermal	0.33	0.59 (hoof cracking, alopecia, redness of skin, petechiae)		
62	Pig (NS)	31 +/- 14 d	Cardio	1.25			Panter et al. 1996 D,L-selenomethionine
			Hepatic	1.25			
			Renal	1.25			
			Dermal		1.25 (symmetrical hair loss, dry scaling skin, cracked overgrown hooves 3/5 pigs)		
			Bd Wt		1.25 (body weight gain 15% less than controls)		
63	Pig (NS)	16 +/- 16 d	Resp	1.25			Panter et al. 1996 selenate
			Cardio	1.25			
			Hepatic	1.25			
			Renal	1.25			
			Dermal		1.25 (symmetrical hair loss, dry scaling skin, cracked overgrown hooves 1/5 pigs)		
			Bd Wt			1.25 (body weight gain 22% less than controls)	

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
64	Pig (NS)	34 d ad lib (F)	Cardio			0.46 (vacuolation, pyknosis of nuclei)	Stowe et al. 1992 NS
			Musc/skel			0.46 (hyperplasia of sarcolemma nuclei; disintegration of myofibrils)	
65	Pig (Duroc)	NS ad lib (F)	Dermal		0.4 F (2/10 alopecia; 1/10 hoof separation)		Wahlstrom and Olson 1959b selenite
			Bd Wt	0.4 F			
66	Cattle Hereford	120 d 1x/d (F)	Resp	0.808 M			O'Toole and Raisbeck 1995 selenomethionine
			Cardio	0.808 M			
			Gastro	0.808 M			
			Musc/skel	0.808 M			
			Hepatic	0.808 M			
			Renal	0.808 M			
			Endocr	0.808 M			
			Dermal	0.158 M	0.288 M (mild parakeratosis of hoof)	0.808 M (severe parakeratosis and epithelial hyperplasia of hoof)	
			Ocular	0.808 M			

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
67	Cattle Hereford	120 d 1x/d (F)	Resp	0.808 M			O'Toole and Raisbeck 1995 sodium selenite
			Cardio	0.808 M			
			Gastro	0.808 M			
			Musc/skel	0.808 M			
			Hepatic	0.808 M			
			Renal	0.808 M			
			Endocr	0.808 M			
			Dermal	0.288 M	0.808 M (mild parakeratosis of hoof)		
			Ocular	0.808 M			
	Immunological/Lymphoreticular						
68	Rat (Sprague-Dawley)	10 wk ad lib (W)			0.7 F (decreased delayed-type hypersensitivity; increased thymus weight)		Koller et al. 1986 selenite
69	Mouse (BALB/c)	47 d ad lib (W)			0.173 NS (reduced B-cell function and OVA-specific antibody concentration)		Raisbeck et al. 1998 selenocystine
70	Mouse (BALB/c)	47 d ad lib (W)			0.173 NS (reduced B-cell function and OVA-specific antibody concentration)		Raisbeck et al. 1998 selenomethionine
71	Mouse (BALB/c)	47 d ad lib (W)			0.173 NS (reduced OVA-specific antibody concentration)		Raisbeck et al. 1998 sodium selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)		Reference Chemical Form
72	Cattle Hereford	120 d 1x/d (F)		0.808 M				O'Toole and Raisbeck 1995 selenomethionine
73	Cattle Hereford	120 d 1x/d (F)		0.808 M				O'Toole and Raisbeck 1995 sodium selenite
	Neurological							
74	Human	120 d (F)		0.0048 M				Hawkes and Hornbostel 1996 selenomethionine
75	Monkey (Macaca fascicularis)	30 d 1x/d (GW)		0.08	0.12 F (hypothermia)			Cukierski et al. 1989 selenomethionine
76	Pig (mixed breed)	7 wk ad lib (F)				1.3	(tetraplegia, poliomyelomalacia)	Baker et al. 1989 organic
77	Pig (crossbred L x Y)	8 wk ad lib (F)		0.33		0.59	(hind limb paresis, hind limb ataxia, symmetric poliomylomalacia of the ventral horn of the spinal cord)	Mihailovic et al. 1992 selenite
78	Pig (NS)	20-42d ad lib (F)		1		2.1	(poliomylelomalacia, paralysis, difuse gliosis of the spinal cord)	Wilson et al. 1983 selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
79	Cattle Hereford	120 d 1x/d (F)		0.808 M			O'Toole and Raisbeck 1995 selenomethionine
80	Cattle Hereford	120 d 1x/d (F)		0.808 M			O'Toole and Raisbeck 1995 sodium selenite
	Reproductive						
81	Monkey (Macaca fascicularis)	30 d 1x/d (GW)		0.06 F	0.08 F (altered menstrual cycle)		Cukierski et al. 1989 selenomethionine
82	Rat (Wild)	5 wk (F)			0.1 M (3.9% abnormal sperm; decrease in live sperm)	0.2 M (24.6% abnormal sperm; decreased live sperm, and sperm motility; decreased testicular weight)	Kaur and Parshad 1994 selenite
83	Rat (Fischer- 344)	13 wk (W)			0.29[b] M (15% decreased sperm counts) 0.31 F (more time in diestrus and less time in proestrus, estrus, and metestrus than controls)		NTP 1994 selenate
84	Rat (Fischer- 344)	13 wk (W)			0.17[b] M (11% decrease epididymal sperm counts)		NTP 1994 selenite
				0.5 F	0.86 F (more time in diestrus and less time in proestrus and estrus)		

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
85	Rat (Wistar)	12-14 wk ad lib				0.324 (testicular hypertrophy)	Turan et al. 1999a sodium selenite
86	Mouse (IVCS)	48 d ad lib (W)		0.17 F	0.34 F (proportion of mice with longer estrus cycles increased by 11.8%)		Nobunaga et al. 1979 selenite
87	Mouse (B6C3F1)	13 wk (W)		7.17 F 5.45[b] M			NTP 1994 selenate
88	Mouse (B6C3F1)	13 wk (W)		3.83 F 3.31[b] M			NTP 1994 selenite
89	Rabbit (New Zealand)	6 wks 1x/wk (GW)			0.001 M (significant reduction in serum testosterone (49%))		El-Zarkouny et al. 1999 sodium selenite
90	Pig (Duroc)	NS ad lib (F)				0.4 (decreased fertility, maternal toxicity)	Wahlstrom and Olson 1959b selenite
	Developmental						
91	Rat (Wistar)	8 wks ad lib (W)			0.64 (decrease weight gain of pups exposed during lactation)		Thorlacius-Ussing 1990 selenite
92	Mouse (IVCS)	pre-Gd:30 d Gd 0-18 ad lib (W)		0.17	0.34 (decreased fetal body weight, delayed vertebral ossification)		Nobunaga et al. 1979 selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
93	Pig (Duroc)	NS ad lib (F)				0.4 (increased number of deaths between birth and weaning; reduced birth weight and reduced body weight at weaning)	Wahlstrom and Olson 1959b selenite
94	Cattle	3 mo ad lib (F)		0.265			Yaeger et al. 1998 sodium selenite

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
CHRONIC EXPOSURE							
Death							
95	Rat (Wistar)	2 yr ad lib (F)				0.5 (reduced longevity from about 500 days to about 60-100 days)	Harr et al. 1967; Tinsley et al. 1967 selenate, selenite
Systemic							
96	Human	>3 yr (F)	Endocr	0.006 F	0.007 F (significant reduction in serum levels of triiodothyronine)		Bratter and Negretti De Bratter 1996 NS
97	Human	>2 yr (F)	Hemato	0.0098			Longnecker et al. 1991 organic
			Musc/skel	0.0098			
			Hepatic	0.0098			
			Dermal	0.0098			
98	Human	lifetime (F)	Dermal	0.015[c]	0.023 (selenosis: sloughing of nails and brittle hair)		Yang and Zhou 1994 Organic
99	Human	yr (F)	Cardio	0.025			Yang et al. 1989a organic
			Hemato	0.015	0.016 (increased prothrombin time)		
			Hepatic	0.025			
			Dermal	0.015	0.016 (brittle nails)		

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
100	Rat (Wistar)	2 yr ad lib (F)	Musc/skel	0.10	0.20 (soft bones)		Harr et al. 1967; Tinsley et al. 1967 selenite, selenate
			Hepatic	0.025	0.10 (hyperplastic lesions)		
			Renal	0.025	0.1 (nephritis)		
101	Rat (Osborne-Mendel)	24 mo ad lib (F)	Resp	0.50 F			Nelson et al. 1943 organic
			Gastro	0.50 F			
			Musc/skel	0.50 F			
			Hepatic			0.25 F (slight to moderate cirrhosis)	
			Endocr	0.5 F			
			Dermal	0.5 F			
102	Mouse (Swiss)	lifetime ad lib (W)	Resp			0.57 (amyloidosis)	Schroeder and Mitchener 1972 selenate
			Cardio			0.57 (amyloidosis)	
			Hepatic			0.57 (amyloidosis)	
			Renal			0.57 (amyloidosis)	
			Endocr			0.57 (amyloidosis of adrenal gland)	
			Dermal		0.57 (poor coat)		
			Bd Wt	0.57			

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)		LOAEL Serious (mg/kg/day)		Reference Chemical Form
103	Mouse (Swiss)	lifetime ad lib (W)	Resp				0.57	(amyloidosis)	Schroeder and Mitchener 1972 selenite
			Cardio				0.57	(amyloidosis)	
			Hepatic				0.57	(amyloidosis)	
			Renal				0.57	(amyloidosis)	
			Endocr				0.57	(amyloidosis of adrenal gland)	
			Dermal		0.57	(poor coat)			
			Bd Wt	0.57					
Neurological									
104	Human	yr (F)		0.027			0.058	(tendon hyperflexia, peripheral anesthesia, pain in extremities, polyneuritis)	Yang et al. 1983 organic
Reproductive									
105	Rat (Wistar)	1 yr daily ad lib (W)		0.21	0.35	(50% reduction in number of pups reared in second generation)	1.05	(decreased fertility, pup survival, maternal toxicity; second generation failed to reproduce)	Rosenfeld and Beath 1954 selenate
106	Mouse (CD)	3 gen ad lib (W)					0.57	(failure to breed in the third generation)	Schroeder and Mitchener 1971b selenate

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Table 3-2. Levels of Significant Exposure to Selenium - Oral (continued)

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)		Reference Chemical Form
Developmental								
107	Mouse (CD)	3 gen ad lib (W)				0.57	(increased number of runts; postnatal lethality)	Schroeder and Mitchener 1971b selenate

[a]The number corresponds to entries in Figure 3-2.
[b]Differences in levels of health effects and cancer effects between males and females are not indicated in Figure 3-2. Where such differences exist, only the levels of effect for the most sensitive gender are presented.
[c]Used to derive a chronic oral minimal risk level (MRL) of 0.005 mg/kg-day; The NOAEL is divided by an uncertainty factor of 3 (for human variability).

ad lib = ab libitum; Bd Wt = body weight; Cardio = cardiovascular; CEL = cancer effect level; d = day(s); Endocr = endocrine; (F) = feed; F = female; (G) = gavage; gastro = gastrointestinal; gd = gestation day; GHS-Px = selenium-dependent glutathione peroxidase; (GW) = gavage in water; Hemato = hematological; (IN) = ingestion; LD_{50} = lethal dose, 50% kill; LOAEL = lowest-observed-adverse-effect level; LPS = lipopolysaccharide; M = male; metab = metabolic; mg/kg/day = milligram per kilogram per day; mo = month(s); Musc/skel = musculoskeletal; NOAEL = no-observed-adverse-effect level; (NS) = not specified; Resp = respiratory; TNF = tumor necrosis factor; TSH = thyroid-stimulating hormone; (W) = water; wk = week(s); x =time(s); yr = year(s)

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Figure 3-2. Levels of Significant Exposure to Selenium - Oral

Acute (≤14 days)

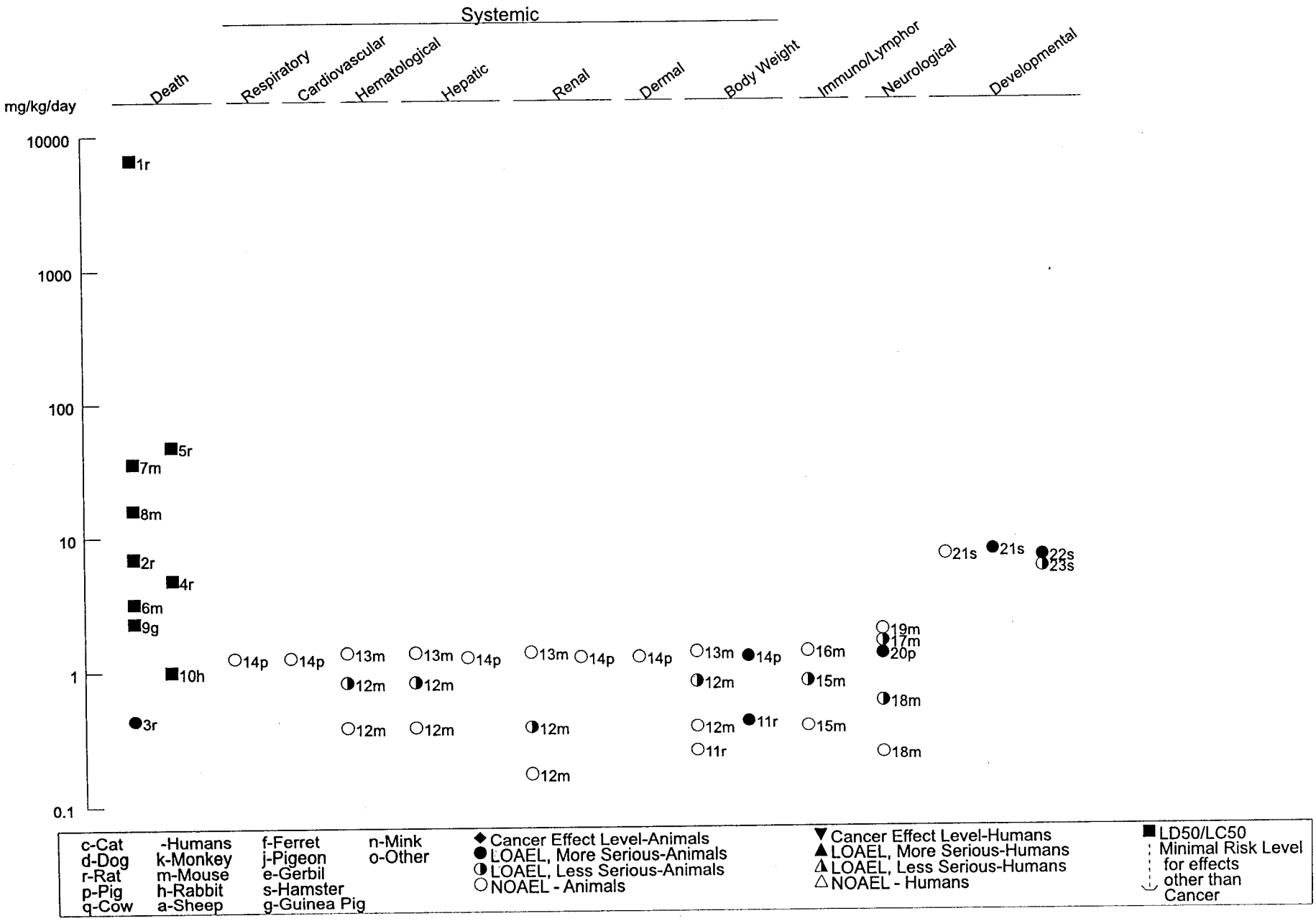

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Figure 3-2. Levels of Significant Exposure to Selenium - Oral (*continued*)

Intermediate (15-364 days)

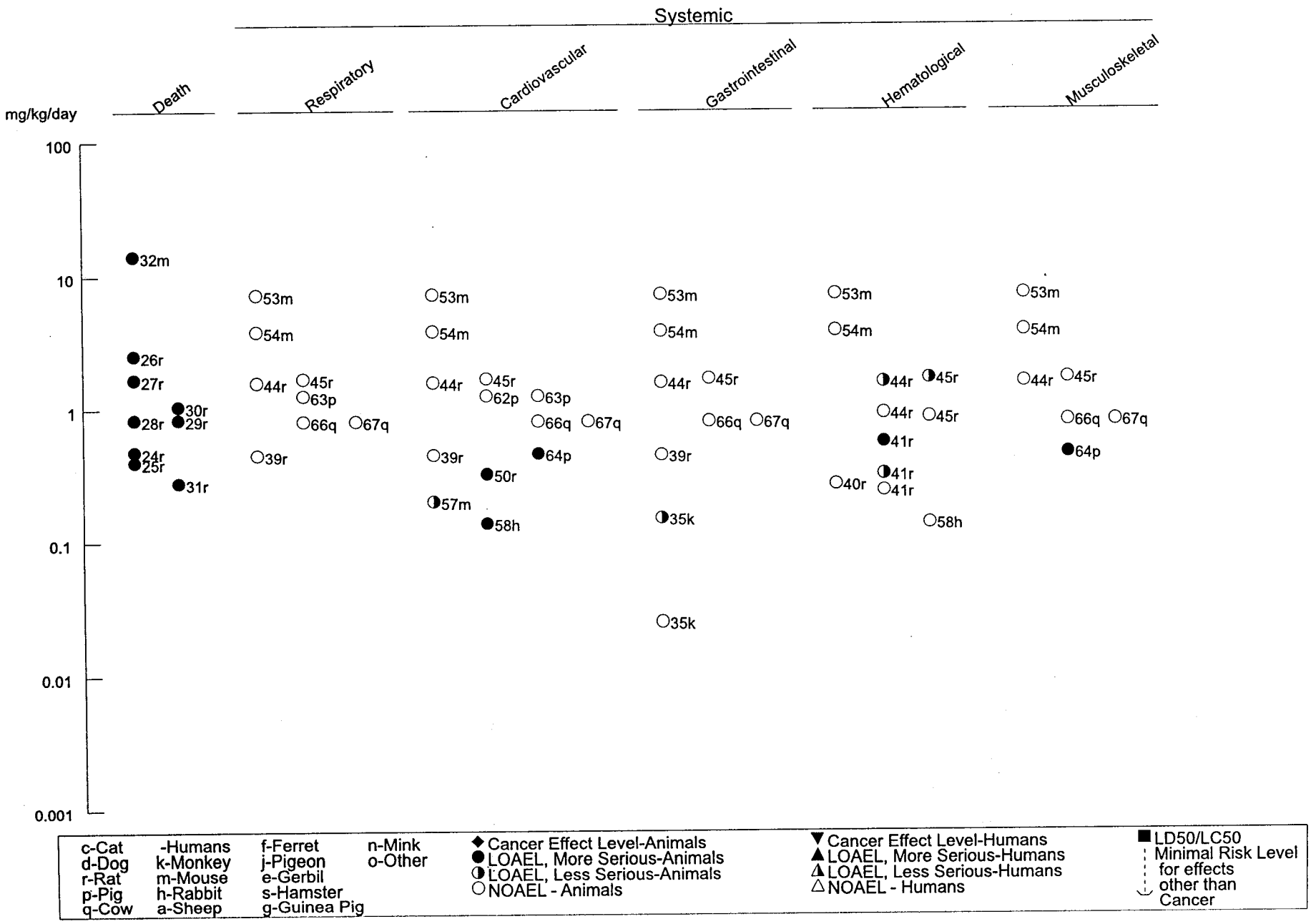

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Figure 3-2. Levels of Significant Exposure to Selenium - Oral (*continued*)

Intermediate (15-364 days)

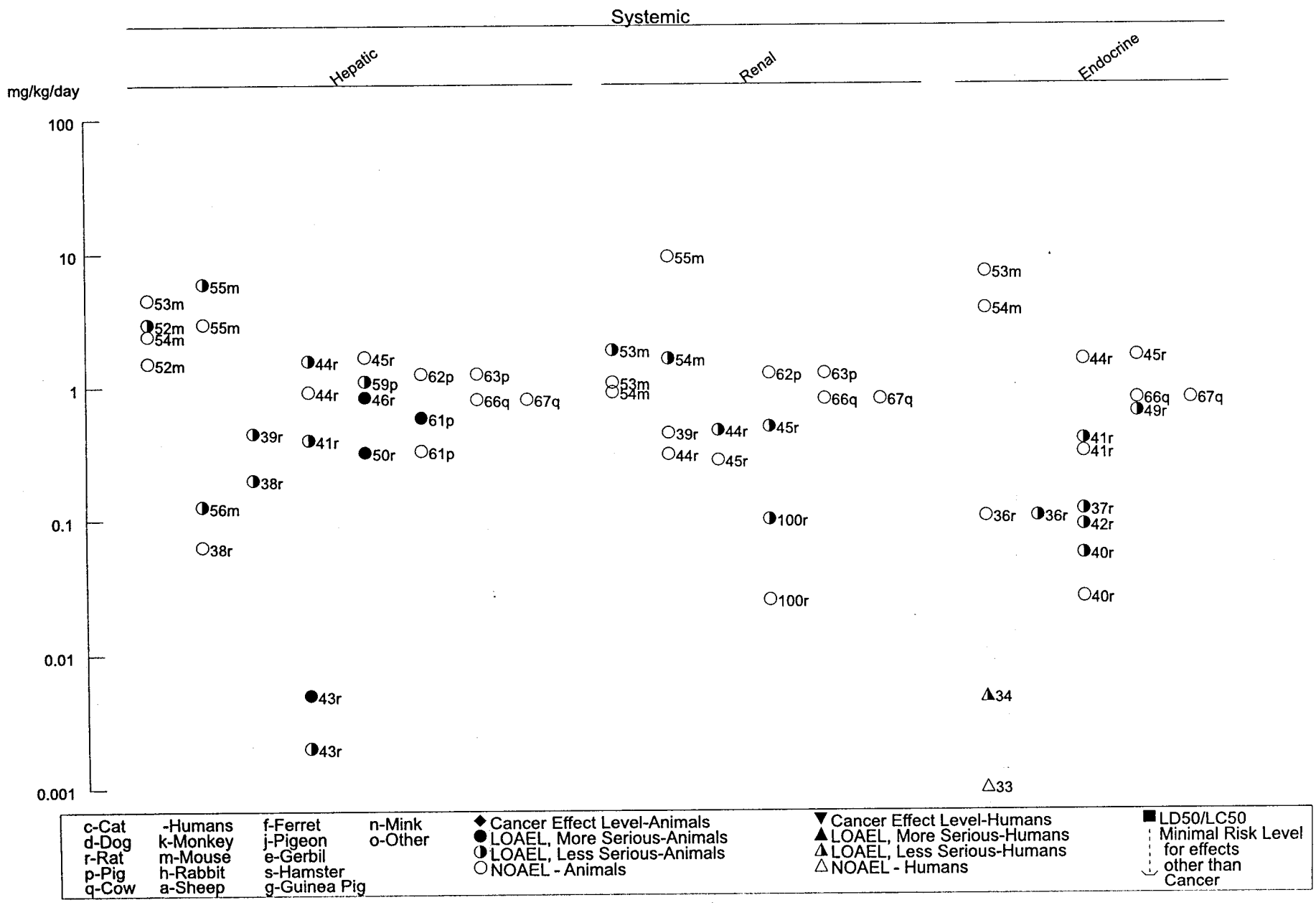

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Figure 3-2. Levels of Significant Exposure to Selenium - Oral (*continued*)

Intermediate (15-364 days)

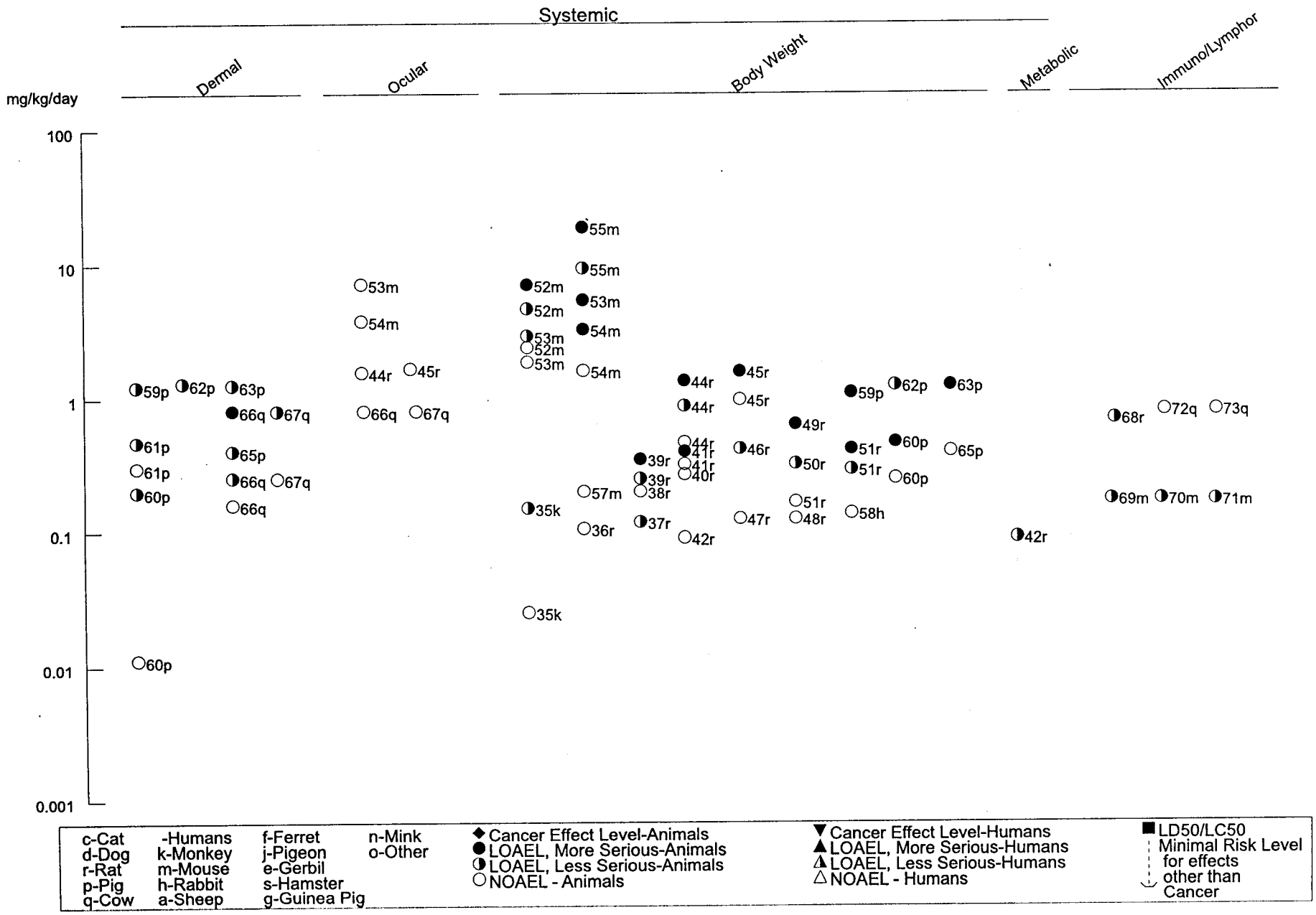

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Figure 3-2. Levels of Significant Exposure to Selenium - Oral (*continued*)

Intermediate (15-364 days)

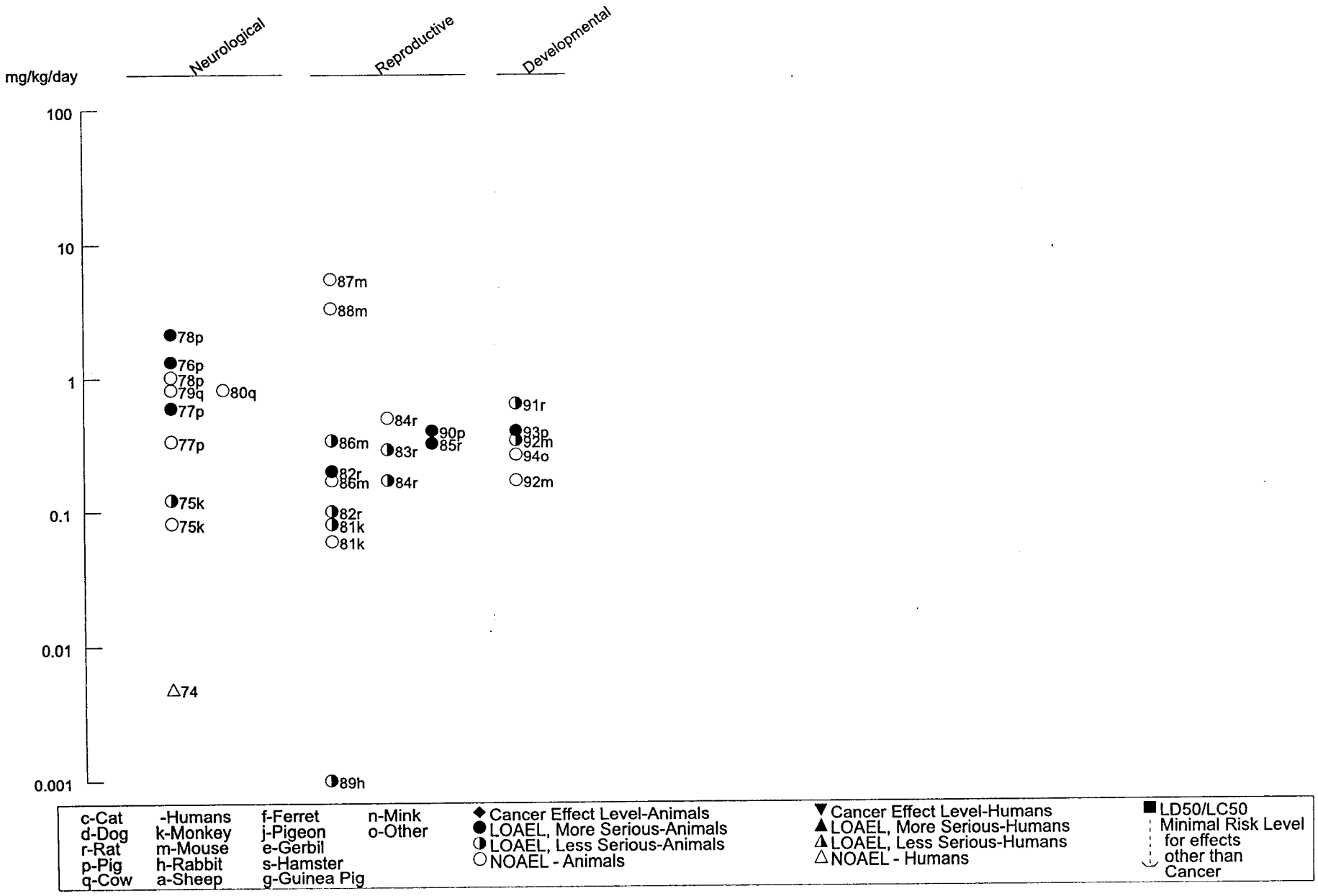

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Figure 3-2. Levels of Significant Exposure to Selenium - Oral (*continued*)

Chronic (≥365 days)

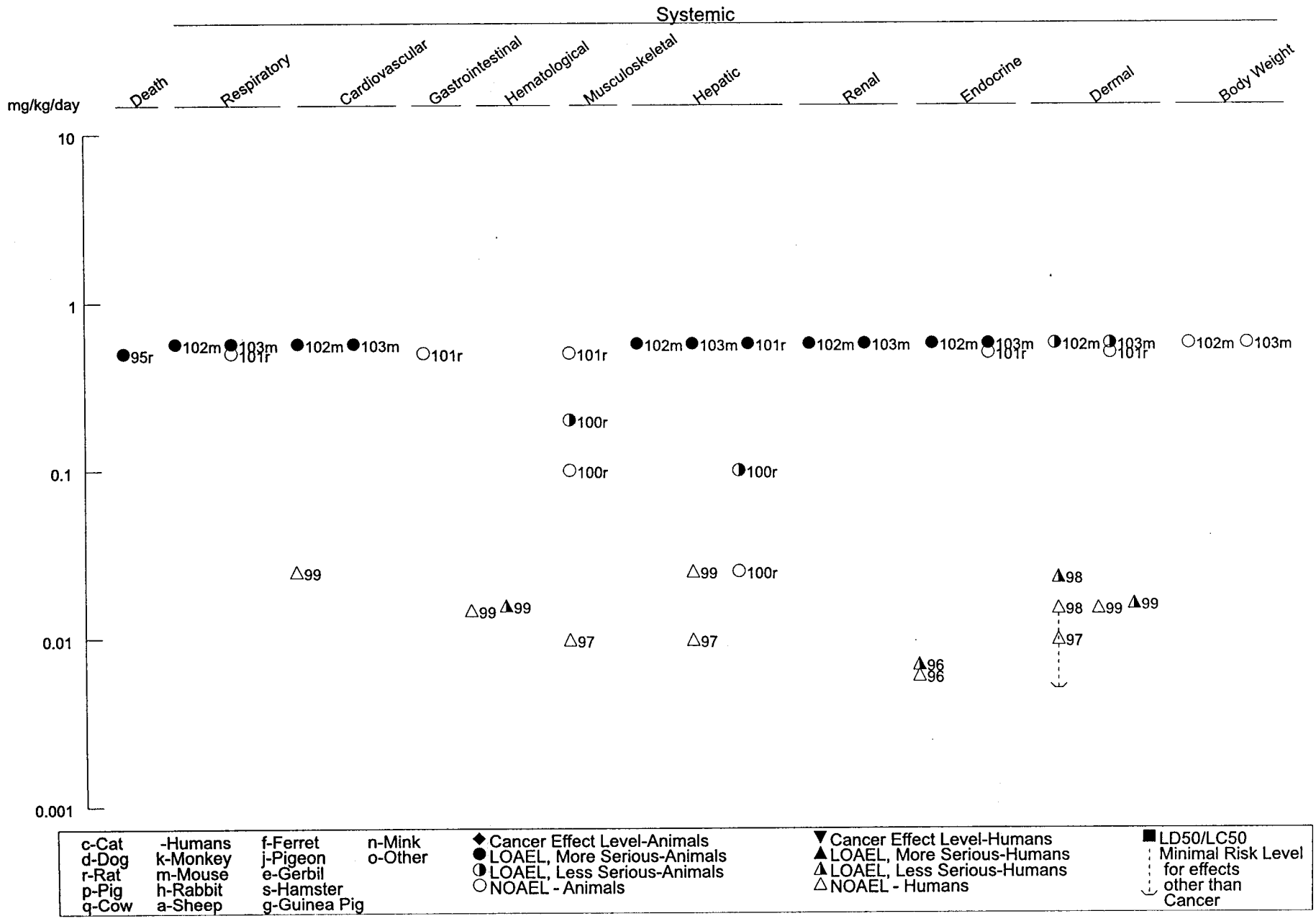

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Figure 3-2. Levels of Significant Exposure to Selenium - Oral (*continued*)

Chronic (≥365 days)

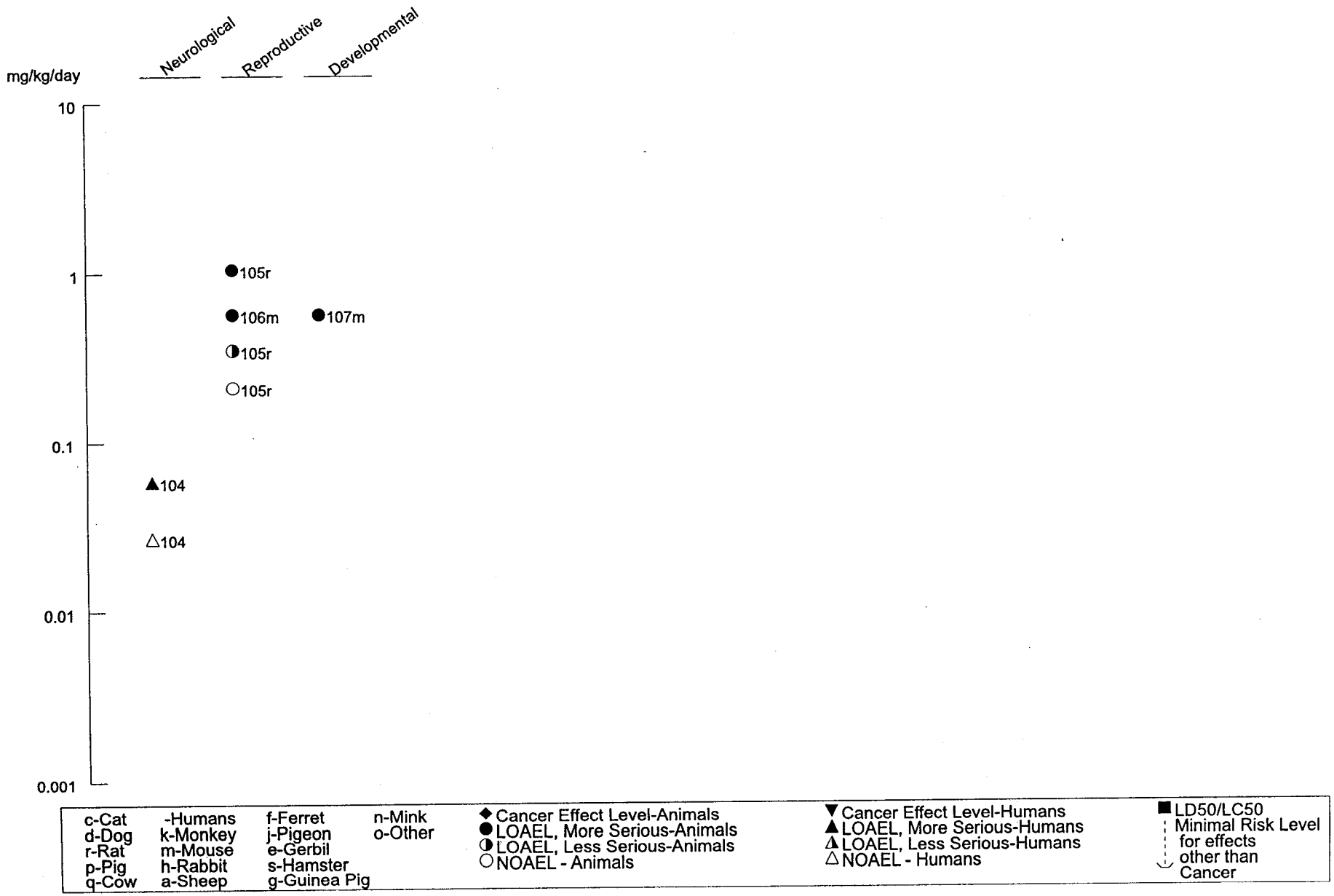

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Table 3-3. Levels of Significant Exposure to Selenium Sulfides - Oral

Key to[a] figure	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
ACUTE EXPOSURE							
Death							
1	Rat (Sprague-Dawley)	once (G)				138 M (LD_{50})	Cummins and Kimura 1971 SeS_2 (aqueous)
2	Rat (Wistar)	once (GO)				75 M (3/6 died)	Moore et al. 1996b SeS
3	Rat (Wistar)	once (GO)				50 (3/15 died)	Moore et al. 1996b SeS
4	Mouse (NMRI)	once (G)				3700 (LD_{50})	Henschler and Kirschner 1969 SeS
Systemic							
5	Rat (Wistar)	once (GO)	Hepatic			75 M (widespread hepatic necrosis)	Moore et al. 1996b SeS

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Table 3-3. Levels of Significant Exposure to Selenium Sulfides - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
	INTERMEDIATE EXPOSURE						
	Death						
6	Rat (Fischer-344)	17 d 1x/d (G)				112 M (LD_{50}) 56[b] F (LD_{50})	NTP 1980c SeS, SeS_2
7	Mouse (B6C3F1)	17 d 1x/d (G)				805 M (LD_{50}) 316[b] F (LD_{50})	NTP 1980c SeS, SeS_2
	Systemic						
8	Rat (Fischer-344)	13 wk 7d/wk 1x/d (G)	Resp	31.6			NTP 1980c SeS, SeS_2
			Cardio	31.6			
			Gastro	31.6			
			Musc/skel	31.6			
			Hepatic	17.6	31.6 (focal necrosis)		
			Renal	31.6			
			Endocr	31.6			
			Dermal	31.6			
			Bd Wt	31.6			

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Table 3-3. Levels of Significant Exposure to Selenium Sulfides - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
9	Mouse (B6C3F1)	13 wk 7d/wk 1x/d (G)	Resp	464			NTP 1980c SeS, SeS_2
			Cardio	464			
			Gastro	464			
			Musc/skel	464			
			Hepatic	464			
			Renal	216	464 (interstitial nephritis)		
			Endocr	464			
			Dermal	464			
			Bd Wt	216 F	464 F (body weight 17% lower than controls)		

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Table 3-3. Levels of Significant Exposure to Selenium Sulfides - Oral (continued)

Key to figure[a]	Species (Strain)	Exposure/ Duration/ Frequency (Specific Route)	System	NOAEL (mg/kg/day)	LOAEL Less Serious (mg/kg/day)	LOAEL Serious (mg/kg/day)	Reference Chemical Form
	CHRONIC EXPOSURE						
	Cancer						
10	Rat (Fischer- 344)	103 wk 7d/wk 1x/d (G)				15 (hepatocellular carcinomas 14/49 males, 21/50 females)	NTP 1980c SeS, SeS_2
11	Mouse (B6C3F1)	103 wk 7d/wk 1x/d (G)				100 F (hepatocellular carcinomas/adenomas 25/49, alveolar/bronchiolar carcinoma/adenomas 12/49)	NTP 1980c SeS, SeS_2

[a]The number corresponds to entries in Figure 3-3.
[b]Differences in levels of health effects and cancer effects between males and females are not indicated in Figure 3-3. Where such differences exist, only the levels of effect for the most sensitive gender are presented.

ad lib = ab libitum; Bd Wt = body weight; Cardio = cardiovascular; CEL = cancer effect level; d = day(s); Endocr - endocrine; F = female; gastro = gastrointestinal; (G) = gavage; gd = gestation day; Hemato = hematological; LOAEL = lowest-observed-adverse-effect level; M = male; Metab = metabolic; Musc/skel = musculoskeletal; NOAEL = no-observed-adverse-effect level; Resp = respiratory; (W) = water; wk = week(s); x = time(s); yr = year(s)

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Figure 3-3. Levels of Significant Exposure to Selenium Sulfides - oral

Acute (≤14 days)

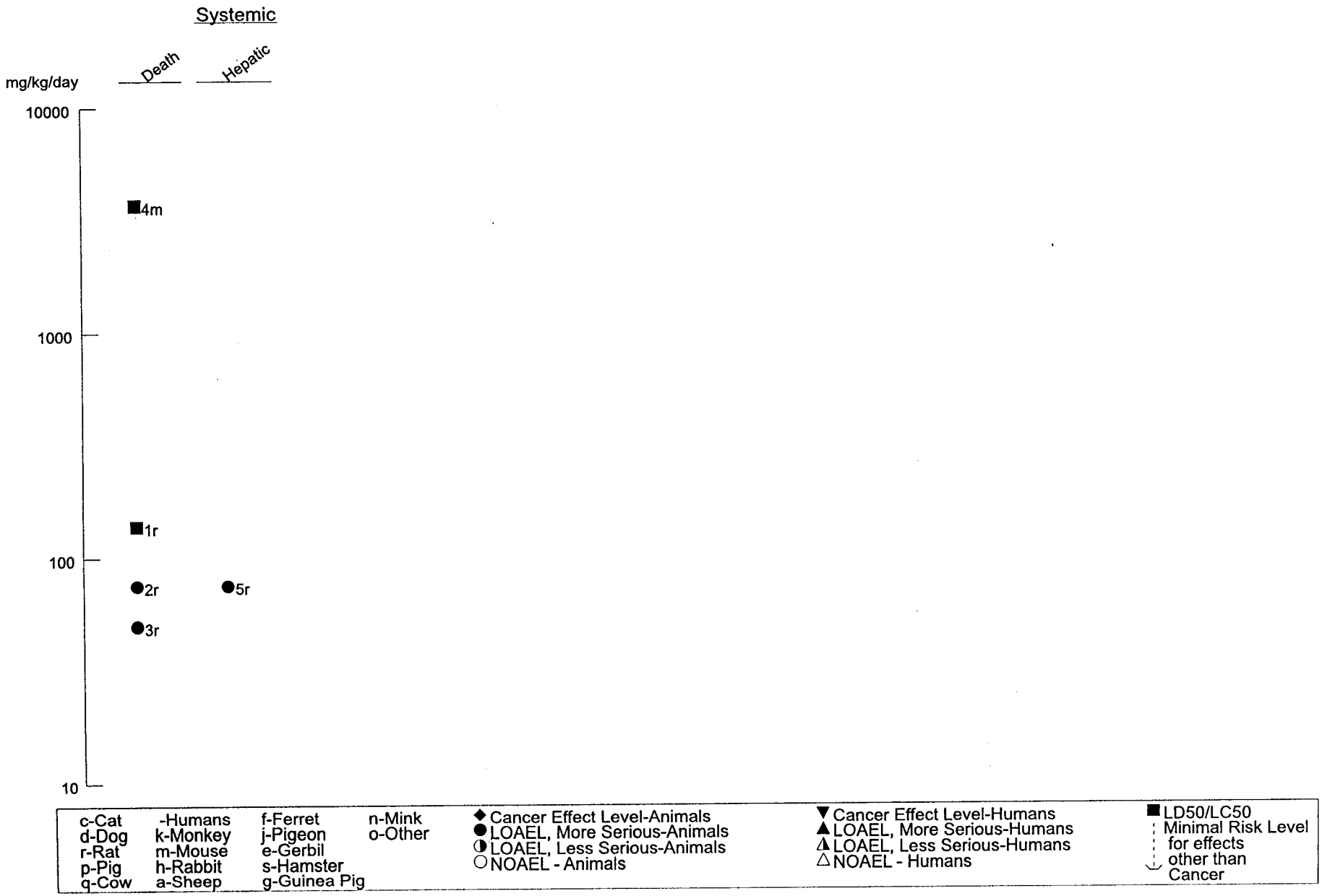

Figure 3-3. Levels of Significant Exposure to Selenium Sulfides - oral (*continued*)

Intermediate (15-364 days)

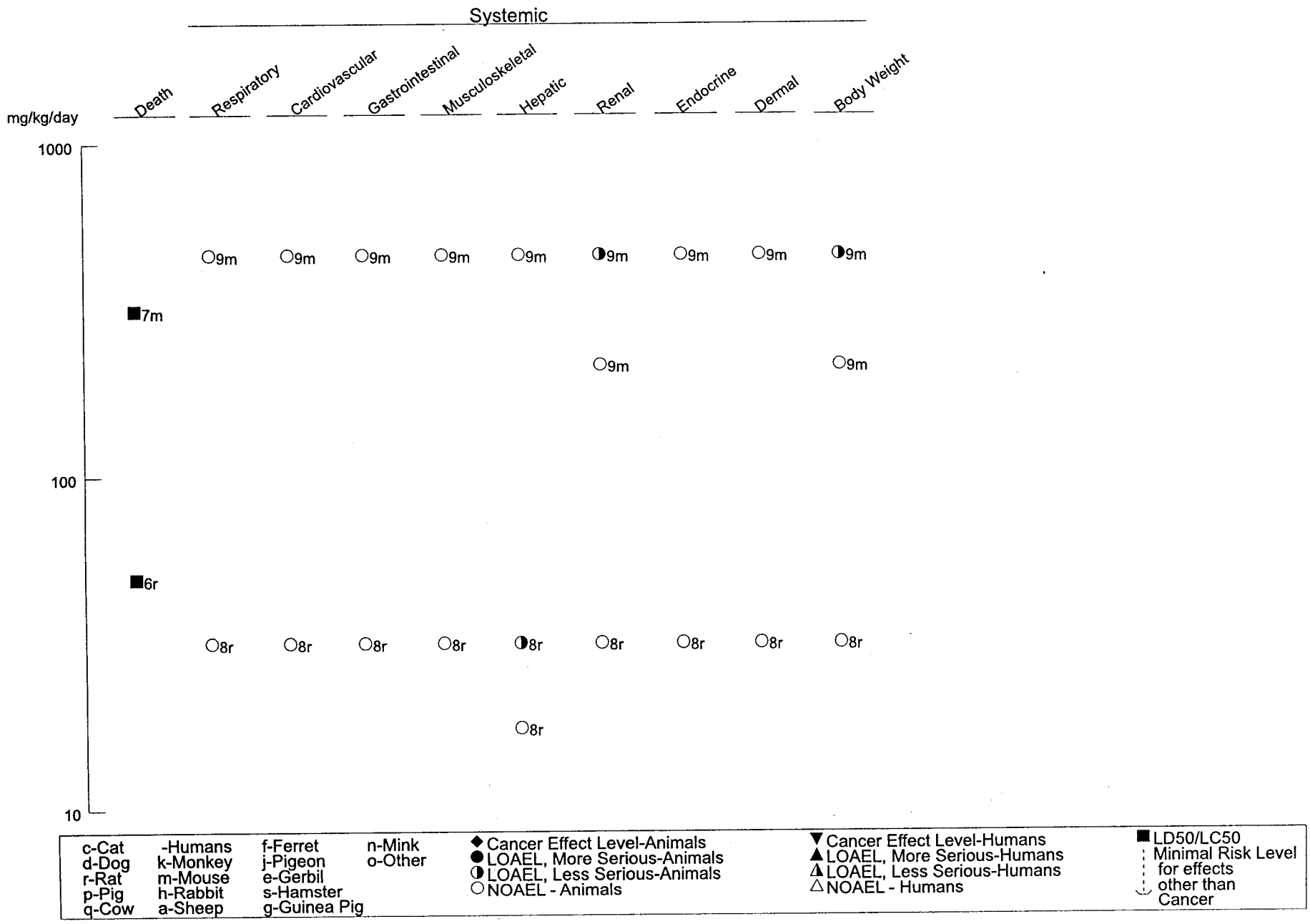

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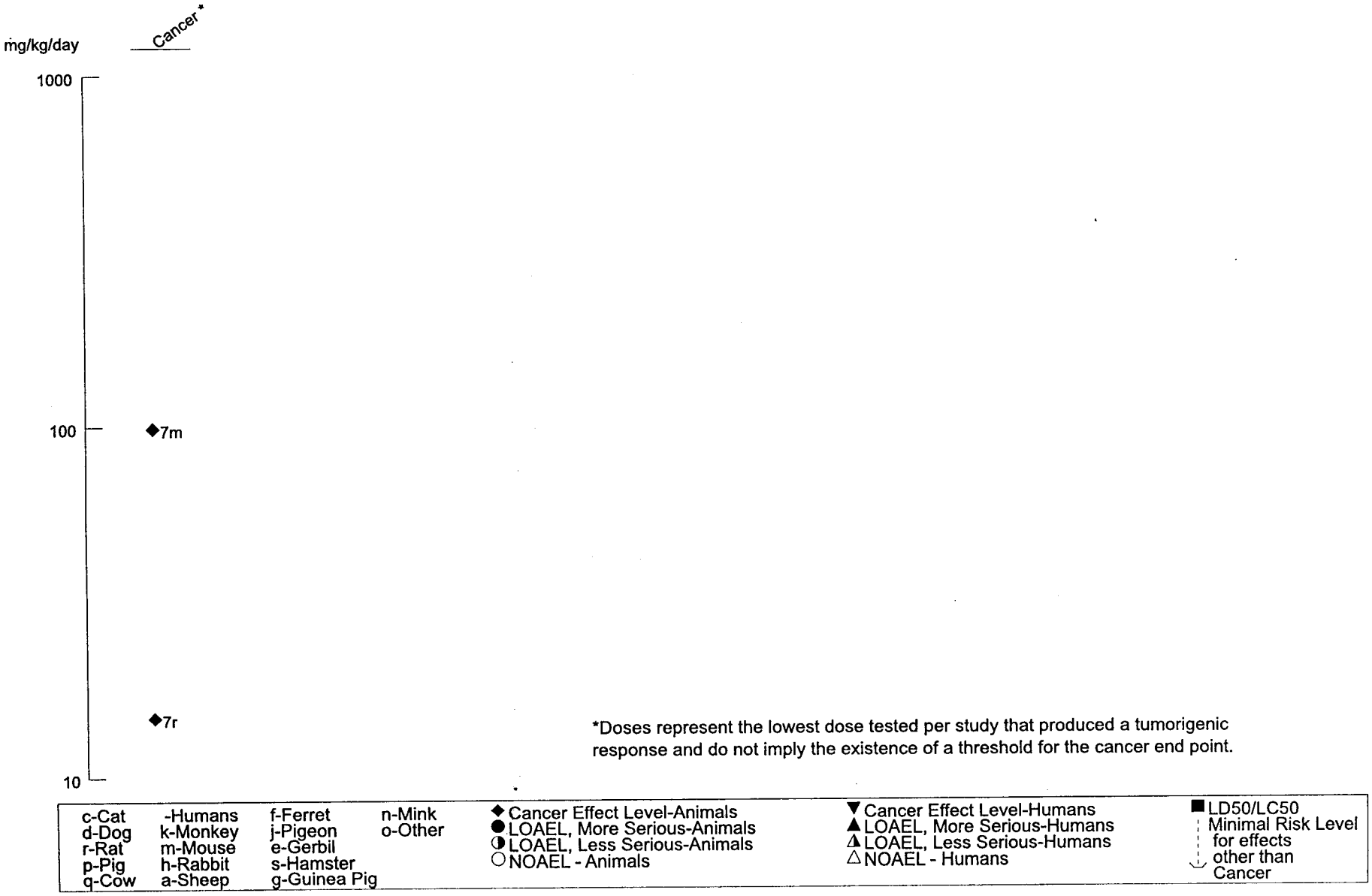
Figure 3-3. Levels of Significant Exposure to Selenium Sulfides - oral (*continued*)

Chronic (≥365 days)



*Doses represent the lowest dose tested per study that produced a tumorigenic response and do not imply the existence of a threshold for the cancer end point.

Most of the available toxicity information for oral exposures to selenium compounds comes from domestic or experimental animal exposures to selenite, selenate, selenium sulfides (mixed), and organic selenium compounds (selenocystine, selenomethionine). Some of the earliest recognized effects of selenium were observed in livestock that grazed on plants in areas of South Dakota, where soil selenium concentrations are naturally high. Selenium-associated effects observed in livestock include blind staggers and alkali disease. Blind staggers is an acute disease in which there is usually a slight impairment of vision, which can result in the animal straying from the herd. As the disease progresses, the blindness becomes more pronounced, and the animal may wander in circles. In the last stage there are various degrees of paralysis, and evidence of abdominal pain; death results from respiratory failure. Alkali disease is a chronic disease in which the animals become emaciated, stiff, and lame; lose long hair from the mane and the tail; and the hooves become deformed. Alkali disease is also associated with atrophy of the heart and liver, while congestion and focal necrosis of the liver are more prominent in blind staggers.

Some epidemiological studies report data from populations exposed to selenium in the food chain in areas with high selenium levels in soil. It is likely that selenite, selenate, and the selenium found in food and in dietary supplements comprise the majority of selenium compounds to which oral, off-site selenium exposures will occur at or near hazardous waste sites. Aside from the variation in effective dose, the health effects from exposure to selenate, selenite, and dietary selenium are not expected to differ greatly. However, oral exposures to many other compounds of selenium could occur (primarily through soil or edible plant ingestion) if those compounds were deposited at the site, or if local environmental conditions greatly favor transformation to those forms. Heavy metal selenides, aluminum selenide, tungsten diselenides, and cadmium selenide are used in industry and may end up in waste sites.

3.2.2.1 Death

Accidental selenium poisonings in humans have occurred, but few fatalities have been reported. The selenium doses associated with the reported deaths are unknown (Carter 1966; Koppel et al. 1986). One 3-year-old boy died 1.5 hours after ingestion of an unknown quantity of selenious acid contained in a gun-blueing preparation (Carter 1966). Clinical signs included excessive salivation, garlic odor on the breath, and shallow breathing. A 15-year-old female survived ingestion of a solution of sodium selenate estimated to have provided 22 mg selenium/kg body weight, probably because she was forced to vomit soon after exposure (Civil and McDonald 1978). Clinical signs included garlic odor of the breath and diarrhea.

No cases of human death in the United States have been attributed to intermediate or chronic oral exposures to selenium or selenium compounds. In the Hubei Province of China, in an area of endemic selenosis, a woman who died was suffering from hemiplegia thought to have been caused by chronic selenosis induced by eating locally grown foods that contained high levels of organic selenium compounds (Yang et al. 1983). However, an autopsy was not performed and no clinical history of previous illness was available.

In nonhuman animals, the most acutely toxic selenium compounds by ingestion appear to be sodium selenite and sodium selenate (Olson 1986). Oral LD_{50} values for sodium selenite, expressed as mg selenium/kg body weight, were reported as 4.8–7.0 in rats, 1.0 in rabbits, 3.2 in mice, and 2.3 in guinea pigs (Cummins and Kimura 1971; Pletnikova 1970). Minimum lethal doses of sodium selenite, expressed as mg selenium/kg body weight, reported for larger animals were 13–18 for pigs and 9.9–11.0 for cows (Miller and Williams 1940); however, these values were estimated on the basis of a small number of animals. Two of four 12-week-old lambs died within 16 hours of administration of 5 mg selenium/kg as sodium selenite (Smyth et al. 1990). Selenium dioxide is reported to have LD_{50} values of 16 mg selenium/kg for mice and 48 mg selenium/kg body weight for rats, but these values are also based on a small number of animals (Singh and Junnarkar 1991). An oral LD_{50} of 35.9 mg selenium/kg has been reported for L-selenocystine given to mice (Sayato et al. 1993). Elemental selenium is less toxic than most selenium compounds, because of its extremely low solubility; an LD_{50} of 6,700 mg selenium/kg body weight has been reported for oral administration of elemental selenium as a suspension (particle size 1–30 m) in 0.5% methylcellulose to rats (Cummins and Kumura 1971).

Lower doses of selenium can cause signs of toxicity if administered over extended periods of time. Eight weaned 5-week-old pigs receiving 1.3 mg selenium/kg/day as sodium selenite in gelatin capsules daily for 10 days died during one study; only one dose level was tested (Wilson et al. 1989). Two long-tailed macaques administered 0.60 mg selenium/kg/day as selenomethionine by nasogastric intubation died of either anorexia or aspirated vomitus secondary to emesis and gastritis after 10 or 15 days of treatment (Cukerski et al. 1989). Seven of 12 female rats receiving diets containing 0.418 mg selenium/kg/day as sodium selenate for 14 days died before the end of the experiment (NTP 1996). Exposure to selenium in drinking water at a level of 0.84 mg selenium/kg/day as selenite or selenate for 4–6 weeks resulted in the death of four of six or two of six male rats, respectively (Palmer and Olson 1974). Feeding male rats diets containing 0.48 mg selenium/kg/day as sodium selenite or 0.4 mg selenium/kg/day as seleniferous wheat for 6 weeks resulted in the death of one of eight rats in each group (Halverson et al. 1966). Administration of sodium selenite in drinking water at a level of 0.28 mg selenium/kg/day for 58 days

resulted in the death of 25 of 50 male rats (Schroeder and Mitchener 1971a). Mortality was observed in rats, but not in mice, receiving either 1.67 mg selenium/kg/day as sodium selenite or 2.54 mg selenium/kg/day as sodium selenate in drinking water for 13 weeks (NTP 1994). Gavage treatment of male mice with selenocystine 6 days per week for 30 days at a dose of 14.2 mg selenium/kg killed all 15 treated animals, while no deaths were noted at 9.4 mg selenium/kg (Sayato et al. 1993). The longevity of hamsters, was not affected by dietary administration of sodium selenite at a dose of 0.42 mg selenium/kg/day for 124–144 weeks (Birt et al. 1986).

Sodium selenate and sodium selenite exhibit similar toxicity in female rats, but male rats appear more susceptible to the toxicity of sodium selenite than selenate (Palmer and Olson 1974; Schroeder and Mitchener 1971a). Sodium selenate in drinking water at 0.28 selenium mg/kg/day for 1 year did not increase mortality of male or female rats compared with control rats (Schroeder and Mitchener 1971a). Ingestion of 0.28 mg selenium/kg/day of sodium selenite in drinking water for 1 year did not increase mortality in female rats, whereas 50% of the males died by day 58 of administration (Schroeder and Mitchener 1971a).

The relative acute toxicities of sodium selenite, potassium selenite, sodium selenate, and potassium selenate in aqueous solution have been examined in mice (Pletnikova 1970). No significant differences among the toxicities of the potassium and sodium salts of selenium were apparent in this study. In another study, rats tolerated a dose of 1.05 mg selenium/kg/day administered in drinking water as potassium selenate for over 8 months with no deaths, but three of five females and one of three males died by the end of 1 year (Rosenfeld and Beath 1954). Decreased survival was reported in rats fed sodium selenate or selenite at 0.5 mg selenium/kg/day in a 2-year cancer study (Harr et al. 1967; Tinsley et al. 1967). No mortality was observed in hamsters fed 0.42 mg selenium/kg/day as sodium selenite in the diet for 82–142 weeks (Birt et al. 1986).

Selenium sulfide (i.e., selenium monosulfide) and selenium disulfide are less water soluble and are of lower acute toxicity than sodium selenate or sodium selenite. There are no reported human deaths due to ingestion of selenium sulfide. The LD_{50} value for the gavage administration of 1–20% selenium disulfide in aqueous 0.5% methylcellulose to rats was 138 mg selenium disulfide/kg (Cummins and Kimura 1971). When 1% selenium disulfide shampoo was administered by gavage, the LD_{50} value was lower (78 mg selenium disulfide/kg) (Cummins and Kimura 1971). The compound administered may have been a mixture of selenium sulfide and selenium disulfide; analysis of the compound was not reported. Henschler and Kirschner (1969) reported an LD_{50} of 3,700 mg selenium sulfide/kg for mice administered

by gavage in aqueous 0.5% carboxymethylcellulose. Administration of single gavage doses of selenium monosulfide to rats produced death in 3/15 animals dosed with 50 mg/kg, 3/6 animals dosed with 75 mg/kg, 1/2 animals dosed with 100 mg/kg, and 2/2 animals dosed with 125 mg/kg (Moore et al. 1996b).

In the case of selenium sulfide, mice are more tolerant than rats, and males of both species appear to be more tolerant than females (NTP 1980c). The daily doses producing 50% mortality for a 17-day gavage administration of a mixture of selenium mono- and disulfides were 112 mg selenium sulfides/kg for male rats, 56 mg selenium sulfides/kg for female rats, and 805 mg selenium sulfides/kg for male mice (NTP 1980c). A 13-week gavage study using the same mixture of selenium mono- and disulfides reported survival as 10/10, 10/10, 10/10, 9/9, 8/9, and 6/10 in female mice and 10/10, 10/10, 10/10, 10/10, 10/10, and 9/10 in male mice receiving 0, 21.6, 46.4, 100, 216, and 464 mg selenium sulfides/kg/day, respectively (NTP 1980c). Although the researchers intended to administer selenium monosulfide to the animals, elemental analysis, melting point, and x-ray diffraction revealed that the compound administered included some selenium disulfide. No other chemical or physical analyses of the selenium compound administered were reported.

The LD_{50} and lethal LOAEL values from each reliable study following oral exposure to elemental selenium dust, selenium dioxide dissolved in water (selenious acid), sodium selenate, sodium selenite, potassium selenate, and dietary selenium for each species and exposure duration are recorded in Table 3-2 and plotted in Figure 3-2. The LOAEL values for death in rats and mice following acute and intermediate oral exposures to selenium sulfide or selenium disulfide are recorded in Table 3-3 and plotted in Figure 3-3.

3.2.2.2 Systemic Effects

The highest NOAEL value and all LOAEL values for each reliable study for systemic effects in each species and duration category are recorded in Tables 3-2 and 3-3 and plotted in Figures 3-2 and 3-3.

Respiratory Effects. Pulmonary edema and lesions of the lung have been noted in case reports of humans (Carter 1966; Koppel et al. 1986) and other animals (Glenn et al. 1964a; Rosenfeld and Beath 1947) after ingestion of lethal doses of selenium compounds. Rabbits orally administered sodium selenite (subroute not specified) at levels approximating the LD_{50} (1–5 mg selenium/kg body weight) developed pulmonary congestion, hemorrhages, and edema; dyspnea; general muscular weakness; and asphyxial

convulsions (Smith and Westfall 1937). Pulmonary edema and hemorrhages were observed in four sheep treated orally (subroute not specified) with a single dose of sodium selenite of 5 mg selenium/kg (Smyth et al. 1990). The lungs may be a target of acute exposure to excess selenium because the metabolite, dimethyl selenide, is exhaled.

The effects of intermediate or chronic exposures to selenium compounds are less clear. Although Harr et al. (1967) stated that absolute lung weights decreased with increasing doses of selenite or selenate chronically administered to rats in the diet in a 2-year cancer study, but they did not report lung weights at specific dose levels. Selenium administration also might have contributed to pneumonic lesions, but again, the authors did not statistically analyze their results or relate the severity of the effect to the doses of selenium administered. Respiratory effects were not observed in rats treated with selenite in the diet for 8 weeks at a dose of 0.45 mg selenium/kg/day (Chen et al. 1993). Effects on the lungs were not observed in pigs fed 1.25 mg selenium/kg as organic selenium found in the plant *Astragalus bisulcatus* for up to 5 days, or D,L-selenomethionine or selenate in the diet for up to 6 weeks (Panter et al. 1996). Treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any signs of respiratory distress or changes in lung weight or histology (O'Toole and Raisbeck 1995). Ingestion of selenium in drinking water for 13 weeks at doses up to 1.67 and 7.17 mg selenium/kg as selenate in rats and mice, respectively, and 1.57 and 3.83 mg selenium/kg as selenite in rats and mice, respectively, did not cause any respiratory effects (NTP 1994). Nelson et al. (1943) reported that no effects on the lungs were apparent in rats administered 0.50 mg selenium/kg/day as seleniferous corn for 2 years.

An increased incidence of amyloidosis of the major organs, including the lungs, was observed in mice following lifetime exposure to sodium selenate or sodium selenite in drinking water at a level of 0.57 mg selenium/kg/day (Schroeder and Mitchener 1972). This effect was noted in 30% of control mice and 58% ($p<0.001$) of selenium-treated mice. Data for individual organs were not provided.

Administration of lethal doses of selenium sulfide particles in carboxymethylcellulose by gavage has been reported to cause irregular breathing in mice (Henschler and Kerschner 1969), but not in rats (Cummins and Kimura 1971). No respiratory effects were seen in mice administered 464 mg selenium sulfides/kg/day or in rats administered 31.6 mg selenium sulfides/kg/day by gavage once daily for 13 weeks (NTP 1980c).

Cardiovascular Effects. Tachycardia has occasionally been reported as a result of a lethal, acute oral exposure to selenium compounds in humans (Carter 1966); however, the dose was not reported in this lethal exposure to a gun-blueing solution containing selenious acid. Although myocardial disorders (cardiogenic shock, congestive heart failure, arrhythmia, multifocal necrosis of the myocardium) have been associated with selenium deficiencies (Yang et al. 1988), none has been reported to be associated with chronic dietary selenosis in humans observed at doses of 0.016 mg/kg/day and greater (Yang et al. 1989a).

In contrast, postmortem studies of sheep that died from acute oral exposure to sodium selenite or sodium selenate, have revealed petechial hemorrhages of the endocardium (Glenn et al. 1964a; Smyth et al. 1990). The sheep were treated with a time-weighted average dose of 0.65 or 0.9 mg selenium/kg/day as selenate over a 171-day period (Glenn et al. 1964a, 1964b), or a single dose of selenite at 5 mg selenium/kg (Smyth et al. 1990). Vacuolation and pyknosis of nuclei were observed in the hearts of pigs fed an unspecified form of selenium at a dose of 0.46 mg selenium/kg/day for 34 days (Stow et al. 1992). In a two year cancer study, Harr et al. (1967) reported the occurrence of myocardial hyperemia, hemorrhage, and degeneration, as well as pericardial edema, in young rats administered sodium selenite or sodium selenate in the feed at doses of 0.5 mg selenium/kg/day, although the authors did not specify the duration of exposure required to produce the effects.

Exposure of pigs to feed containing 54 mg/kg selenium for 1–7 days resulted in severe toxicity and death of several of the animals (Penrith and Robinson 1996). Histological examination of heart tissue from pigs that died revealed myocardial lesions consisting of widespread hypertrophy, atrophy, and disorganization of fibers, occasional fibrosis, and marked medial hypertrophy of the arterioles.

Wistar rats administered 0.324 mg selenium/kg/day as sodium selenite in food for 12–14 weeks showed severe diffuse degenerative changes, including edema in the sub-endocardial connective tissue and the interfibers of prevascular regions, and myofibril swelling with profuse intercellular edema (Turan et al. 1999a). Myocyte borders were irregular, and there was a loss of striations and a degeneration of the sarcolemma and myofibril structure and order. Examination of the mechanical function of the heart *in vitro* using either Langendorff perfusion or papillary muscle recordings showed increased coronary perfusion pressure, increased resting force, and increased heart rate with irregular beating. No difference in contractile force was observed. Chronic heart failure did not occur in any of the animals in the study.

Cardiac damage was also observed in mice exposed to 0.2 mg selenium/kg/day as sodium selenite in food for 12 weeks (Skowerski et al. 1997b). Ultrastructural examination revealed cardiomyocytes that had numerous damaged mitochondria, a large number of lipid droplets, and numerous lysosomes.

Hearts of New Zealand White rabbits administered 0.137 mg selenium/kg/day as sodium selenite in food for 3 months showed distinct, degenerative changes indicating disintegration of the internal structure of the myocytes (Turan et al. 1999b). Muscle fibers were fragmented and separated. Disruption and loss of myofibrils was observed, sarcomeres were irregular, and the I, Z, and H bands were disorganized and discontinuous. Mitochondria were fewer and more variable in size and shape, with disoriented cristae and a loss of matrix substance. Hearts of control animals (0.007 mg selenium/kg/day) had normal histology.

Treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any changes in heart weight or histology (O'Toole and Raisbeck 1995). Histopathological changes in the heart were not observed in pigs fed selenium at 1.25 mg selenium/kg as organic selenium found in the plant *A. bisulcatus* for up to 5 days, or D,L-selenomethionine or selenate for up to 6 weeks (Panter et al. 1996). Histopathological changes were not observed in the hearts of rats treated with selenite in the diet for 8 weeks at a dose of 0.45 mg selenium/kg/day (Chen et al. 1993). Selenium administered to rats and mice in drinking water for 13 weeks at doses up to 1.57 and 7.17 mg selenium/kg/day as selenate, respectively, and up to 1.67 and 3.83 mg selenium/kg/day as selenite, respectively, did not cause any histopathological changes in the heart tissue (NTP 1994). No histopathological changes were noted in mice administered 464 mg selenium sulfides/kg/day or in rats administered 31.6 mg selenium sulfides/kg/day by gavage once daily for 13 weeks (NTP 1980c).

An increased incidence of amyloidosis of the major organs, including the heart, was observed in mice following lifetime exposure to sodium selenate or sodium selenite in drinking water at a level of 0.57 mg selenium/kg/day (Schroeder and Mitchener 1972). This effect was noted in 30% of control mice and 58% ($p<0.001$) of selenium-treated mice. Data for individual organs were not provided.

Gastrointestinal Effects. In humans, gastrointestinal distress, including nausea, vomiting, diarrhea, and abdominal pain, has been reported following ingestion of aqueous sodium selenate (Civil and McDonald 1978; Gasmi et al. 1997; Helzlsouer et al. 1985; Koppel et al. 1986; Sioris et al. 1980). Two studies provided an estimate of dose. In a case report by Civil and McDonald (1978), diarrhea was observed in a 15-year-old girl about 45 minutes after she swallowed sheep drench containing selenate at a dose of about 22 mg selenium/kg. This effect was observed despite the induction of vomiting shortly

after the exposure. In a second case report of a suicide attempt, a 56-year-old man reported that vomiting, diarrhea, and abdominal pain occurred 1 hour after he ingested approximately 11 mg/kg selenium as sodium selenite (Gasmi et al. 1997). Postmortem examinations following two deaths from selenium ingestion revealed dilation of the stomach and small intestine (Carter 1966) and erosive changes of the gastrointestinal tract (Koppel et al. 1986). High (unspecified) levels of dietary selenium compounds have been implicated as causing gastrointestinal disturbances in chronically exposed humans (Smith et al. 1936), but such symptoms are not specific to selenium intoxication.

Exposure of pigs to feed containing 54 mg/kg selenium for 1–7 days resulted in severe toxicity and death of several animals. Clinical signs included anorexia and vomiting, and histological examination (70–79 days after exposure) of three of the exposed animals that died found lesions ranging from small erosions (1–2 mm diameter) to extensive mucosal necrosis (up to 100 mm diameter) near the cardia of the stomach (Penrith and Robinson 1996).

Gross necropsy of steers that died after ingestion of sodium selenite revealed severe gastrointestinal irritation (Baker et al. 1989; Maag et al. 1960). In addition, livestock exhibiting alkali disease, perhaps as a result of long-term consumption of range plants high in selenium, eat and drink less and suffer from ulcers in the upper intestinal tract (Shamberger 1986). A single oral dose of 5 mg selenium/kg as selenite caused edema and congestion of abdominal viscera in lambs (Smyth et al. 1990). However, treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any changes in histology of the gastrointestinal tissues (O'Toole and Raisbeck 1995).

Gastrointestinal effects were not observed in rats treated with selenite in the diet for 8 weeks at a dose of 0.45 mg selenium/kg/day (Chen et al. 1993). Selenium treatment in drinking water for 13 weeks at doses up to 1.57 and 7.17 mg selenium/kg/day as selenate in rats and mice, respectively, and 1.67 and 3.83 mg selenium/kg/day as selenite in rats and mice, respectively, did not cause any gastrointestinal effects (NTP 1994). Gastrointestinal effects were not observed in rats fed organic selenium (seleniferous corn or wheat) at 0.5 mg selenium/kg/day for 24 months (Nelson et al. 1943). Vomiting and anorexia were reported in monkeys receiving 0.15 mg/kg/day selenium as L-selenomethionine by oral intubation during gestation days 20–50 (Tarantal et al. 1991).

Selenium sulfide administration by gavage at lethal levels has been reported to cause diarrhea and anorexia in rats (Cummins and Kimura 1971). No gastrointestinal effects were seen in mice administered

464 mg selenium sulfide/kg/day or in rats administered 31.6 mg selenium sulfide/kg/day by gavage once daily for 13 weeks (NTP 1980c).

Hematological Effects. In one study of exposed humans, increased prothrombin time was reported for individuals chronically exposed to doses estimated as 0.016 mg selenium/kg/day of dietary selenium (Yang et al. 1989a). However, no increase in prothrombin time was found in another study of individuals consuming diets that supplied up to 0.0098 mg/kg/day selenium (Longnecker et al. 1991). A study that compared children from seleniferous and nonseleniferous areas of Venezuela found slightly reduced (no statistical analysis was performed) hemoglobin levels and hematocrit values for the children from the seleniferous area (Jaffe et al. 1972). However, the children from the seleniferous zone had a poorer diet, consumed less milk and meat, and had a greater incidence of intestinal parasites, which may account for the differences observed.

Red blood cell counts were significantly increased in mice which received drinking water containing 9 ppm (0.82 mg selenium/kg/day) selenium as sodium selenite for 14 days (Johnson et al. 2000). However, these mice also had a severe reduction in water consumption (43%) and this may have led to a decrease in blood volume. No significant increase in red blood cell count (or decrease in water consumption) was observed for mice receiving 3 ppm (0.38 mg selenium/kg/day) selenium as sodium selenite, or up to 9 ppm (1.36 mg selenium/kg/day) selenium as selenomethionine for 14 days (Johnson et al. 2000).

No hematological changes (hemoglobin concentration, hematocrit, erythrocyte count, and cell volume) were reported for male Sprague-Dawley rats fed diets providing up to 0.27 mg selenium/kg/day as sodium selenite for 40 days (Eder et al. 1995). Increased hematocrit was observed in rats treated with selenate (1.56 mg selenium/kg/day) or selenite (1.67 mg selenium/kg/day) in the drinking water for 13 weeks, but only at concentrations that decreased water intake (NTP 1994). No effects on hematology end points were observed in mice treated with selenate or selenite in drinking water for 13 weeks at 7.17 mg selenium/kg for selenate and 3.83 mg selenium/kg/day for selenite (NTP 1994).

No difference in blood cell counts or hematological parameters were found in rabbits administered 0.137 mg selenium/kg/day as sodium selenite in the diet for 3 months, compared with control animals receiving a normal laboratory diet (Turan et al. 1999b).

A dose-related decrease in hematocrit was observed in rats fed seleniferous wheat (Halverson et al. 1966). Compared to controls, hemoglobin was decreased 23 and 79% at 0.32 and 0.56 mg selenium/kg/day, respectively. Hemoglobin reductions were most evident in the animals that had died during the experiments. In a 2-year cancer study, Harr et al. (1967) reported that the hemoglobin concentration decreased by 0.5 g/100 mL with each twofold increase of sodium selenate in the diet, but did not specify the lowest dose at which hemoglobin concentrations were significantly reduced compared to the controls (the range of selenium doses used was 0.025–0.40 mg selenium/kg/day). Hematocrit was increased in rats given selenite and selenate in drinking water for 13 weeks at concentrations that also resulted in decreased water intake (NTP 1994). No hematological effects were noted in rats or mice treated with selenate at 0.92 and 7.17 mg selenium/kg/day, respectively, or selenite at 0.86 and 3.83 mg selenium/kg/day, respectively (NTP 1994).

No studies were located regarding hematological effects in humans or other animals after oral exposure to selenium sulfide or selenium disulfide.

Musculoskeletal Effects. No adverse musculoskeletal effects were reported following chronic oral exposure of humans to dietary levels of selenium of up to 0.0098 mg selenium/kg/day (Longnecker et al. 1991).

A single oral (subroute not specified) dose of sodium selenite (5 mg selenium/kg/day) caused edema in skeletal muscles of the diaphragm in sheep (Smyth et al. 1990). Exposure of pigs to feed containing 54 mg/kg selenium for 1–7 days resulted in severe toxicity and death of several animals (Penrith and Robinson 1996). Histological examination of skeletal muscle from animals that died found damage with interstitial oedema and diffuse swelling of fibers. Livestock suffering from chronic alkali disease, a disease once common in the southwestern United States where selenium levels are high, showed lameness due to joint erosion and hoof deformation (Shamberger 1986). However, treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any changes in muscle or bone histology (O'Toole and Raisbeck 1995). Hyperplasia of the sarcolemma nuclei and disintegration of myofibrils were observed in the skeletal muscles of pigs fed an unspecified form of selenium for 34 days (Stow et al. 1992). In a 2-year cancer study, Harr et al. (1967) fed graded doses of selenium in the form of sodium selenate or selenite to rats and reported frank osteotoxicity at doses as low as 0.2 mg selenium/kg/day given for several months (duration specified as less than 100 days). Selenium administered in drinking water for 13 weeks at doses up to 1.57 and 7.17 mg selenium/kg/day as selenate in rats and mice, respectively, and 1.67 and 3.83 mg selenium/kg as

selenite in rats and mice, respectively, failed to cause adverse musculoskeletal effects (NTP 1994). Musculoskeletal effects were not observed in rats fed seleniferous corn or wheat at 0.5 mg selenium/kg/day for 24 months (Nelson et al. 1943). No musculoskeletal effects were seen in mice administered 464 mg selenium sulfide/kg/day or in rats administered 31.6 mg selenium sulfide/kg/day by gavage once daily for 13 weeks (NTP 1980c).

Hepatic Effects. Limited data suggest that hepatotoxicity can occur in humans following acute oral exposure to sodium selenate, but no definitive studies were located regarding hepatic effects in humans after intermediate or chronic oral exposure to selenium compounds. Tests following an acute poisoning of a 15-year-old girl with sodium selenate revealed abnormally elevated serum bilirubin and alkaline phosphatase (Civil and McDonald 1978). Hepatic effects, such as changes in serum liver enzymes or liver morphology (identified by "supersonic B" examination), have not been observed in humans at chronic dietary intakes of 0.0098 mg selenium/kg/day (Longnecker et al. 1991) or 0.025 mg selenium/kg/day (Yang et al. 1989a).

Congestion and/or edema and hemorrhage in the liver have been reported in sheep following the acute oral (subroute not specified) administration of lethal levels of sodium selenate (Hopper et al. 1985) or sodium selenite (Smyth et al. 1990) and in mules and pigs following administration of lethal levels of sodium selenite (Miller and Williams 1940). A significant decrease in relative liver weight was reported for mice exposed to 9 ppm (0.82 mg selenium/kg/day) selenium as sodium selenite in drinking water for 14 days, but not to 3 ppm (0.38 mg selenium/kg/day) (Johnson et al. 2000). No effect on liver weight was observed for mice receiving up to 9 ppm (1.36 mg selenium/kg/day) selenium as selenomethionine in drinking water for 14 days (Johnson et al. 2000).

Administration of single gavage doses of selenium monosulfide to rats produced death and widespread hepatic necrosis in 3/6 animals dosed with 75 mg/kg, 1/2 animals dosed with 100 mg/kg, and 2/2 animals dosed with 125 mg/kg (Moore et al. 1996b).

Hepatic effects have also been reported following intermediate-duration exposure in pigs, but not in cattle. Pigs exposed for 7 weeks to either dietary organic selenium in dried plants (either *A. pruelongus* or *A. bisulcatus*) or sodium selenate (at 1.1 or 1.3 mg selenium/kg/day) exhibited diffuse swelling and vacuolar degeneration of hepatocytes (Baker et al. 1989). The doses used in this study reduced mean survival to only 44 days. Pigs exposed to sodium selenite in feed for 35 days at doses less than half as high as those tested by Baker et al. (1989) (0.47 versus 1.1 or 1.3 mg selenium/kg/day) exhibited no liver

damage (Mahan and Magee 1991). A study of pigs treated with 0.08, 0.33, 0.59, or 1.07 mg selenium/kg/day as sodium selenite in the feed for 8 weeks found hepatic nodules/granules in two pigs treated with 0.59 or 1.07 mg selenium/kg/day (Mihailovic et al. (1992). The lesions were diagnosed as postdystrophic atrophic cirrhosis. However, only these two severely affected pigs (one from each of the highest dose groups of 40 animals each) were selected for histopathological examination. Hepatic effects were not observed in pigs fed selenium at 1.25 mg selenium/kg as organic selenium of the type(s) found in the plant *A. bisulcatus* for up to 5 days, or D,L-selenomethionine or selenate for up to 6 weeks (Panter et al. 1996). Treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any changes in liver weight or histology (O'Toole and Raisbeck 1995).

Alterations or cirrhosis of the liver in experimental animals following intermediate or chronic oral exposure to selenium compounds have been reported by Bioulac-Sage et al. (1992), Fitzhugh et al. (1944), Halverson et al. (1970), Harr et al. (1967), Kolodziejczyk et al. (2000), Nelson et al. (1943), and Schroeder and Mitchener (1972). Halverson et al. (1966) reported reduced liver-to-body-weight ratios and increased bilirubin in rats administered 0.44 mg selenium/kg/day for 6 weeks as naturally occurring selenium in wheat. At this level, five of eight rats died. At a dose of 0.84 mg selenium/kg/day administered as sodium selenate in drinking water for 4–6 weeks, rats developed cirrhosis of the liver (Palmer and Olson 1974). At this level, two of six rats died.

Hepatic damage was observed in mice exposed to 0.2 mg selenium/kg/day as sodium selenite in food for 12 weeks (Skowerski et al. 1997a), and ultrastructural examination showed that the cytoplasm of the hepatocytes contained extremely large and irregularly-shaped vacuoles. Wistar rats administered 0.324 mg selenium/kg/day as sodium selenite in food for 12–14 weeks showed degenerative changes to the liver (not fully described in text) (Turan et al. 1999a). Livers of rats fed 0.002 or 0.005 mg selenium/kg/day as sodium selenite for 3 months showed damage that increased with dose (Kolodziejczyk et al. 2000). Rats from the 0.002 mg selenium/kg/day group had a distinct swelling of Küpffer cells in dilated sinusoidal vessels, mainly in the proximity of portal fields, and occasional necrotic areas comprising groups of hepatocytes, while livers from rats receiving 0.005 mg selenium/kg/day showed activation and swelling of the Küpffer cells in widened sinusoidal vessels, relatively abundant infiltrations of mononuclear cells into portal canals, and sporadic areas of necrosis within individual lobules.

Young rats treated with sodium selenite in the feed for 2 months had nodular hyperplasia at a dose of 0.2 mg selenium/kg/day. However, clinical tests of liver function (bilirubin, alanine aminotransferase,

aspartate aminotransferase, alkaline phosphatase, and gamma-glutamyltransferase activities) showed no significant changes (Bioulac-Sage et al. 1992). Diffuse panlobular vacuolar changes were reported in rats fed sodium selenite in the diet for 8 weeks at 0.45 mg/kg/day (Chen et al. 1994).

In a 2-year cancer study, acute toxic hepatitis was common among rats fed sodium selenite or sodium selenate at 0.25 mg selenium/kg/day or higher (Harr et al. 1967; Tinsley et al. 1967). Liver surfaces were mottled, and parenchymatous degeneration was present. Hepatic lesions occurred at a dose as low as 0.10 mg selenium/kg/day. Absolute liver weights decreased with increasing levels of sodium selenate or sodium selenite in the diet. The average liver weight of animals administered selenate (14.5 g) was twice the average liver weight of animals administered selenite (7.2 g); however, the average liver weight of control animals was not reported, and possible dose-related hepatic effects were not discussed by these authors.

Increased serum bile acids, suggesting cholestasis, were observed in rats treated with 1.57 mg selenium/kg/day as sodium selenate in drinking water for 13 weeks, but no effects were noted at 0.92 mg/kg/day (NTP 1994). In a 13-week drinking water study, hepatic effects were not observed in mice treated with sodium selenate at 7.17 mg selenium/kg/day, in mice treated with sodium selenite at doses up to 3.83 mg selenium/kg/day, or in rats treated with sodium selenite at doses up to 1.67 mg selenium/kg/day (NTP 1994). Increased serum aspartate aminotransferase and alanine aminotransferase activities were observed in mice treated by gavage with selenocystine at doses of 9.4 mg selenium/kg/day for 30 days (Sayato et al. 1993) or 4.7 mg selenium/kg/day for 90 days (Hasegawa et al. 1994). No effects on liver enzymes were observed in mice treated with selenocystine at 4.7 mg selenium/kg/day for 30 days (Sayato et al. 1993) or at 2.5 mg selenium/kg/day for 90 days (Hasegawa et al. 1994). Chronic dietary administration of selenium as seleniferous corn or wheat at doses ranging from 0.25 to 0.50 mg/kg/day for 24 months produced cirrhosis of the liver in rats (Nelson et al. 1943).

An increased incidence of amyloidosis of the major organs, including the liver, was observed in mice following lifetime exposure to sodium selenate or sodium selenite in drinking water at a level of 0.57 mg selenium/kg/day (Schroeder and Mitchener 1972). This effect was noted in 30% of control mice and 58% ($p<0.001$) of selenium-treated mice. Data for individual organs were not provided.

Selenium sulfide administered to rats daily by gavage for 13 weeks produced focal coagulation necrosis in the liver with infiltration by inflammatory cells. These changes developed at a dose of 31.6 mg selenium sulfide/kg/day, but not at a dose of 17.8 mg selenium sulfide/kg/day (NTP 1980c). In mice, on

the other hand, oral intubation of selenium sulfide at 464 mg selenium sulfide/kg/day did not produce hepatic effects (NTP 1980c).

Renal Effects. No studies were located regarding renal effects in humans after oral exposure to selenium or selenium compounds.

In domestic and experimental animals, renal effects have been observed following both acute and chronic oral exposures to selenium compounds. Administration of a single oral (subroute not specified) dose of sodium selenite at 5 mg selenium/kg/day produced hydropic degeneration of the kidney in sheep (Smyth et al. 1990). In a study of the toxicity of L-selenomethionine to long-tailed macaques by nasogastric intubation, two animals administered 0.24 mg selenium/kg/day aspirated vomitus secondary to emesis, developed obvious gastritis, and died of anorexia, one after 10 days and the other after 15 days of administration (Cukierski et al. 1989). Histopathologic examination of the kidneys of these animals revealed glomerulonephritis and proximal convoluted tubule nephropathy. The study authors indicated that these changes were consistent with macaque fatal fasting syndrome and may not have resulted from the direct effects of L-selenomethionine. Following long-term ingestion of plants high in selenium, livestock suffering from alkali disease exhibited nephritis (Shamberger 1986). However, treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any changes in kidney weight or histology (O'Toole and Raisbeck 1995).

A dose-related increase in degeneration of the renal papilla (described as mild to minimal) was observed in rats treated with selenate or selenite in the drinking water at about 0.5 mg selenium/kg/day for 13 weeks (NTP 1994). No evidence of renal toxicity was observed in rats given 0.3 mg selenium/kg/day in this study. In contrast to rats, the only kidney effect noted in mice treated with sodium selenate or selenite in the drinking water was increased relative kidney weights (NTP 1994). This effect, which occurred at $1.87 mg selenium/kg/day as selenate and $1.61 mg selenium/kg/day as selenite, was only noted at doses at which drinking water intake was decreased, leading the investigators to suggest the effect may have been a result of dehydration. A similar increase in relative kidney weight associated with decreased water consumption was observed in mice consuming approximately 0.38 mg selenium/kg/day as selenite in drinking water, but no effect on kidney weight or water consumption was observed in mice consuming up to 1.36 mg selenium/kg/day as selenomethionine (Johnson et al. 2000). No renal effects were observed in pigs fed selenium at 1.25 mg selenium/kg as organic selenium found in the plant *A. bisulcatus* for up to 5 days, or D,L-selenomethionine or selenate for up to 6 weeks (Panter et al. 1996). No effects on the kidneys were observed in rats treated with selenite in the diet for 8 weeks at a dose of

0.45 mg selenium/kg/day (Chen et al. 1993). Gavage treatment of mice with selenocystine for 30 days at a dose of 9.4 mg selenium/kg/day had no adverse effect on the kidneys (Sayato et al. 1993).

Rats chronically fed selenite in the diet were reported to exhibit more frequent and more severe nephritis than those given equivalent amounts of selenate (Harr et al. 1967); however, the study authors did not quantify these observations or statistically compare data from the two groups. An increased incidence of amyloidosis of the major organs, including the kidneys, was observed in mice following lifetime exposure to sodium selenate or sodium selenite in drinking water at a level of 0.57 mg selenium/kg/day (Schroeder and Mitchener 1972). This effect was noted in 30% of control mice and 58% ($p<0.001$) of selenium-treated mice. Data for individual organs were not provided.

A mixture of selenium sulfide and selenium disulfide administered to mice daily by gavage for 13 weeks at a dose of 464 mg selenium sulfides/kg/day produced an increase in the incidence and severity of interstitial nephritis compared with the controls, whereas a daily dose of 216 mg selenium sulfides/kg did not elicit renal toxicity (NTP 1980c). In rats, selenium sulfide by oral intubation at 31.6 mg selenium sulfides/kg/day for 13 weeks did not produce renal effects (NTP 1980c).

Endocrine Effects. Thyroid hormone metabolism is the result of a balance in iodine and selenium levels. Selenium is a component of the deiodinase enzymes, including the Type I and Type II iodothyronine 5'-deiodinases, which convert the prohormone thyroxine (T4) to the active form, triiodothyronine (T3) (Köhrle 1994; St Germain and Galton 1997). Iodine deficiency can lead to hypothyroidism; but if iodine deficiency is accompanied by selenium deficiency, thyroid gland destruction may also occur (Contempré et al. 1991a; Hofbauer et al. 1997). Supplementation of individuals deficient in both iodine and selenium with selenium produces a further decrease in thyroid function, but if selenium supplementation is preceded by normalization of iodine levels, normal thyroid function is restored (Contempré et al. 1991a, 1992).

Selenium intake was reported to affect thyroid hormone levels in a dietary study of 11 men in a metabolic unit (Hawkes and Keim 1995). This study was available only as an abstract. The men (20–45 years old) were fed a diet of foods naturally high or low in selenium (double blind) for 120 days. Dietary selenium was 80 µg/day for the first 21 days, then either 13 (n=6) or 356 (n=5) µg/day for 14 weeks. Dietary energy nutrient level was equivalent to 2,800 kcal/day. Serum T3 levels decreased in the high selenium group (-11%), increased in the low selenium group (+7.5%), and were significantly different in the two groups from the third week of treatment onwards. Thyroid stimulating hormone (TSH) concentration

increased in the high selenium group (+37%) and was significantly different from the baseline levels ($p<0.06$). Despite minor adjustments of energy nutrient intake to maintain body weight, by the sixth week the high selenium group started to gain weight relative to the low selenium group. The difference between the two groups became significant after the tenth week. There was a similar increase in lean body mass in both groups. The levels of T3 and TSH remained within the normal human range for the duration of the study.

An examination of thyroid hormone levels in lactating women residing in areas of Venezuela with high levels of selenium in the soil (selenium intake ranged from 250 to 980 μg per day as estimated from selenium content of breast milk) revealed a significant decrease in serum T3 levels, as compared with women having normal selenium intakes (90–350 μg/day), but these hormone levels remained within the normal range (Brätter and Negretti De Brätter 1996). Additionally, a significant inverse correlation for selenium and serum T3 concentration was found using the Spearman Rank test. The study authors noted that the effect of selenium on T3 levels became significant at dietary intake levels of 350–450 μg/day. No significant alterations in serum T4 or TSH levels or correlations with selenium intake were found.

Twenty weeks of selenium supplementation (10, 20, 30, or 40 μg/day) of New Zealanders who normally consume a diet low in selenium (unsupplemented intake of 28–29 μg/day), but show no signs of deficiency, produced a reduction in T4 concentration in all groups (Duffield et al. 1999). However, only the differences between the 10 μg-group and controls and the combined supplemented individuals and controls were significant. T3 and TSH levels were not measured. Thyroglobulin concentration did not change significantly with supplementation.

In a study of 68 male Latvian fish consumers (Hagmar et al. 1998), a significant inverse correlation was found between serum levels of selenium and TSH. No correlation was found between serum selenium concentration and the serum concentrations of T3 or T4. No measurements were made of dietary selenium intake.

Selenium supplementation has been shown to affect type-I-deiodinase activity in male rats (Behne et al. 1992; Eder et al. 1995; Hotz et al. 1997). Exposure to 0.055 or 0.27 mg selenium/kg/day as sodium selenite in food for 40 days produced a significant decrease (approximately 50%) in serum levels of T3 and a nonsignificant reduction in type-I-deiodinase activity compared with rats receiving 0.009 or 0.026 mg selenium/kg/day (Eder et al. 1995). Exposure to 0.27 mg selenium/kg/day did not produce any

other adverse signs, such as weight loss or decreased food consumption, and serum T4 levels were similar in all groups.

Exposure of weanling male Sprague-Dawley rats to 0.09 mg selenium/kg/day as sodium selenate in food for 6 weeks produced a significant (~30%) increase in TSH, compared with controls receiving 0.009 mg selenium/kg/day (Hotz et al. 1997). Serum T3 and T4 levels and thyroid glutathione peroxidase levels were unaffected by dietary selenium. Kidney type-I-deiodinase levels were decreased (~10%) in high selenium animals compared with controls, but the differences were not significant, and liver type-I-deiodinase levels were unaffected by dietary selenium. Iodine-deficient diets produced greater thyroid glutathione peroxidase activity at each dietary level of selenium, and the greatest activity was in rats with high selenium.

No significant changes in thyroid levels of T3 or T4 were found in male Wistar rats fed diets containing high selenium (0.105 mg selenium/kg/day as sodium selenite or 0.118 mg selenium/kg/day as L-selenomethionine) for 3 months, compared with controls receiving adequate selenium (0.0015 mg selenium/kg/day as sodium selenite) (Behne et al. 1992). However, rats eating the high selenium diet showed a significant reduction in hepatic type I deiodinase activity, compared with controls, with a 29% reduction in the production rate of T3 from T4 and a 45% reduction in the production rate of 3,3'-diiodothyronine from T4.

Many studies have documented reduced body weight gain in young animals treated with selenium compounds, and abnormal weight loss in older animals (Grønback et al. 1995; Halverson et al. 1966; Harr et al. 1967; Jacobs and Forst 1981a; Johnson et al. 2000; Nelson et al. 1943; NTP 1994; Palmer and Olson 1974; Panter et al. 1996; Schroeder 1967; Tarantal et al. 1991; Tsunoda et al. 2000). There is evidence to suggest these effects may be due in part to the interactions of selenium or selenium compounds with hormones that regulate normal growth and body weight. In a 14-day study suggesting selenium may inhibit pituitary function, Thorlacius-Ussing (1990) treated nursing rats with sodium selenite in drinking water (0.64 or 0.96 mg/kg/day). The resulting decrease in the body weight gain of the pups observed at both doses may be associated with a reduction in somatomedin C levels (no other hormone levels were tested), and the weight deficiency could be reversed by administration of a growth hormone. Postweanling female Wistar rats treated with sodium selenite (0.64 mg selenium/kg/day) in drinking water for 3 or 6 weeks exhibited decreased weight gain and decreased somatomedin C serum concentrations. When the selenium supplement was removed after 3 weeks, body weight gain returned to normal, but the serum somatomedin C concentrations did not return to control levels. Growth hormone

secretion in response to growth hormone releasing factor was also reduced in the selenium-exposed group (Thorlacius-Ussing et al. 1988). A 10% reduction in body weight and a reduction in tibia lengths, compared to pair-fed controls, were found in rats provided with sodium selenite in the drinking water at 0.46 mg selenium/kg/day for 35 days (Gronbaek et al. 1995). A significant reduction in insulin-like growth factor-binding protein-3 was also noted. The investigators concluded that the reduction in growth caused by excess selenium is not due to reduced caloric intake.

Selenium administered in drinking water for 13 weeks at doses up to 1.57 and 7.17 mg selenium/kg/day as selenate in rats and mice, respectively, and 1.67 and 3.83 mg selenium/kg as selenite in rats and mice, respectively, failed to cause changes in the weights or histology of the thyroid, adrenal glands, parathyroid, or pancreas (NTP 1994).

Lambs given a single oral (subroute not specified) dose of 5 mg selenium/kg as sodium selenite exhibited cytoplasmic flocculation of the pancreas (Smyth et al. 1990). Increased pancreas weights were observed in rats fed organic selenium (seleniferous wheat) at a dose of 0.4 mg selenium/kg/day for 6 weeks (Halverson et al. 1966). Chronic exposure of rats fed sodium selenite or sodium selenate in their diet for a lifetime was associated with pancreatic damage. Although Harr et al. (1967) reported a dose-related increase in the incidence and severity of pancreatic lesions in treated rats, they did not specify the lowest dose at which pancreatic lesions were observed.

Treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any changes in weight or histology of the pancreas, adrenal glands, thyroid, or pituitary gland (O'Toole and Raisbeck 1995).

An increased incidence of amyloidosis of the major organs, including the adrenal gland, was observed in mice following lifetime exposure to sodium selenate or sodium selenite in drinking water at a level of 0.57 mg selenium/kg/day (Schroeder and Mitchener 1972). This effect was noted in 30% of control mice and 58% ($p<0.001$) of selenium-treated mice. Data for individual organs were not provided.

Dermal Effects. Jensen et al. (1984) described both marked alopecia and the deformity and loss of fingernails in a woman who had consumed a selenium supplement containing 31 mg total selenium (in the form of sodium selenite and elemental selenium) per tablet for 77 days. The woman consumed one tablet each day in addition to vitamin supplements (vitamins C, A, D, E, B complex) and a mineral supplement "labeled as containing all 72 trace elements in undefined quantities." In epidemiological studies of

populations chronically exposed to high levels of selenium in food and water, investigators have reported discoloration of skin, pathological deformity and loss of nails, loss of hair, and excessive tooth decay and discoloration (Smith et al. 1936; Yang et al. 1983, 1989a, 1989b). The 1989 studies by Yang et al. follow up their original 1983 study of Chinese populations living in areas classified as having low-, medium-, and high-selenium exposure based on local soils and food supplies. The average and standard error of selenium intakes in the low-, medium-, and high-intake regions were 0.0012±0.00009, 0.0037±0.0004, and 0.025±0.001 mg/kg/day, respectively. The whole blood (average ± standard error) concentrations of selenium in the low-, medium-, and high-intake regions were 0.16±0.00, 0.35±0.02, and 1.51±0.05 mg/L, respectively. The estimated daily dietary selenium intake required to produce these symptoms in an area of China characterized by endemic selenosis was at least 0.016 mg selenium/kg/day (Yang et al. 1989a). This corresponds to a blood concentration of 1.054 mg/L and an estimated daily intake of 0.91 mg/day, assuming a 55-kg Chinese man or woman and using the regression analysis provided by Yang et al. (1989b). The NOAEL from the highest intake population not affected by nail disease is 0.015 mg selenium/kg/day, which corresponds to a blood concentration of 0.97 mg/L. Foods that contributed the greatest levels of selenium were smoked pork, coal-dried corn, chestnuts, pumpkin seeds, dried fruits, and garlic. It has been noted that the selenosis problem in China began when coal with high levels of selenium was burned as the main source of fuel (Whanger 1989). Food was cooked and dried over the open flame, adding selenium to the food. In addition, the people breathed large amounts of smoke, but the contribution of volatilized selenium to the total dose of selenium has not been adequately characterized (Whanger 1989). Coal was also burned on the fields as a fertilizer source. Environmental selenium concentrations in the low-, medium-, and high-intake regions were 0.37–0.48, 0.73–5.66, and 7.06–12.08 mg/kg in soil, and 370, 1,720, and 12,270 mg/L in water, respectively (Yang et al. 1989b).

No evidence of nail disease was observed in a population living on selenium-rich ranches in the western United States (Longnecker et al. 1991). Doses of selenium were calculated to be between 0.001 and 0.01 mg/kg/day, corresponding to a maximum intake of 0.724 mg/day. Whole blood selenium concentrations were 0.18–0.67 mg/kg. Although these values for the United States are consistent with studies of the Chinese population, only one or a few individuals ingested the highest doses.

The highest selenium intake for villagers in a high-selenium area of China in which endemic selenosis did not occur was estimated at 1.51 mg selenium/person/day (0.027 mg selenium/kg/day), with the average dietary selenium intake in this area of selenosis occurrence estimated to be 3.2 mg selenium/person/day (0.058 mg selenium/kg/day) (Yang et al. 1983). The lowest daily dietary selenium intake associated with

dermal effects, 0.91 mg selenium/day, was converted to equivalent daily doses from food (0.016 mg/kg/day) for presentation in Table 3-2.

Five individuals from the high selenium region of China described by Yang et al. (1989a) who had been diagnosed with overt signs of selenosis (hair loss and nail sloughing) in 1986 were reexamined in 1992 (Yang and Zhou 1994). The results of this examination showed that these individuals had recovered from selenosis (overt symptoms of nail sloughing were absent) and that the average selenium concentrations in their blood had fallen from 1,346 to 968 μg/L. The corresponding dietary intakes of selenium were 1,270 and 819 μg/day. This study has been used to establish a LOAEL of 0.023 mg selenium/kg/day and a NOAEL of 0.015 mg selenium/kg/day. Based on the occurrence of these dermal effects, a chronic oral MRL of 0.005 mg selenium/kg/day has been derived from the NOAEL, as described in the footnote in Table 3-2. This MRL is 5 times greater than the recommended dietary allowance (RDA) for selenium of 0.001 mg/kg/day.

In a 30-day study of oral administration of L-selenomethionine to long-tailed macaques, skin lesions appeared on the forearm of one of two macaques given 0.01 mg selenium/kg/day. However, the limited number of animals precludes identifying the dose as a LOAEL for dermal effects (Cukierski et al. 1989). Pigs receiving dietary administration of the same doses of selenium for 35 days exhibited hoof cracking (Mahan and Magee 1991). Symmetrical hair loss, dry scaling skin, and cracked overgrown hooves were observed in one of five pigs and three of five pigs fed sodium selenate or D,L-selenomethionine at a dose of 1.25 mg selenium/kg for up to 6 weeks, respectively (Panter et al. 1996). In an experiment limited to a duration of 5 days because of severe paralysis, similar dermal effects were not observed in pigs fed 1.25 mg selenium/kg/day as selenium contained in the plant *A. bisulcatus*. The form of selenium in *A. bisulcatus* is unknown, although Panter et al. (1996) indicate that it is nonprotein.

Exposure of pigs to feed containing 54 mg/kg selenium for 1–7 days resulted in severe toxicity and death of several animals; however none of the pigs developed coronitis or hoof separation (Penrith and Robinson 1996). Skin from four pigs with alopecia was examined about a month after exposure and was found to have epidermal thickening due to acanthosis and hyperkeratosis, vacuolar degeneration of the basal cells and acanthocytes, necrosis of individual keratinocytes, and serocellular crusts.

In the late 19th and the early 20th century, livestock grazing on plants growing on seleniferous soils in the Great Plains of the United States suffered from alkali disease attributed to the high selenium content of the plants. Alkali disease in horses, cattle, and swine is characterized by alopecia, inflammation at the

coronary band, followed by cracked or malformed hooves and rough hair coat (Draize and Beath 1935). Daily selenium intakes associated with these effects were not quantified. However, treatment of steers with selenomethionine in food at doses of 0.288 mg selenium/kg body weight/day or selenite at doses of 0.808 mg selenium/kg/day for 120 days produced hoof lesions (O'Toole and Raisbeck 1995). In intermediate-duration studies, cracked hoof walls have been observed in pigs fed selenate, selenite, or an unspecified form of selenium at doses of 0.25 mg selenium/kg/day and greater (Baker et al. 1989; Mahan and Magee 1991; Mihailovic et al. 1992; Wahlstrom and Olson 1959b). Poor quality of the hair coat has also been reported in mice administered sodium selenite or selenate in the diet at 0.57 mg selenium/kg/day (Schroeder and Mitchener 1972). Exposure of female BALB/c mice to 0.21 mg selenium/kg/day for 6 months from diets containing selenium as sodium selenite resulted in alopecia around the nose (Boylan et al. 1990).

No studies were located regarding dermal effects in humans or other animals after oral exposure to selenium sulfide or selenium disulfide.

Ocular Effects. A case control study using a hospital discharge register indicated there was no correlation between serum selenium concentrations and cataract occurrence in humans (Knekt et al. 1992). Since this is a case control study, it does not provide information on the potential dietary factors, exposure to specific selenium compounds, or duration of exposure.

Treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days did not produce any changes in the histology of the eyes (O'Toole and Raisbeck 1995). Selenium given to rats and mice in drinking water for 13 weeks at up to 1.6 and 7.2 mg selenium/kg as selenate, respectively, or 1.7 and 3.8 mg selenium/kg as selenite, respectively, did not cause any ocular effects (NTP 1994).

Body Weight Effects. Two studies reported body weight effects in humans after oral exposure to selenium. Selenium intake was found to affect body weight in a dietary study of 11 men in a metabolic unit (Hawkes and Keim 1995). The men (20–45 years old) were fed a diet of foods naturally high or low in selenium (double blind) for 120 days. Dietary selenium was 80 μg/day for the first 21 days, then either 13 (n=6) or 356 (n=5) μg/day for 14 weeks at 2,800 kcal/day. Despite minor adjustments of intake to maintain body weight, by the 6th week the high selenium group started to gain weight relative to the low selenium group, and the difference between the two groups became significant after the 10th week. A similar increase in lean body mass was observed in both groups.

A study that compared children from seleniferous and nonseleniferous areas of Venezuela found slightly reduced height and weight (no statistical analysis was performed) for the children from the seleniferous area (Jaffe et al. 1972). However, the children from the seleniferous zone had a poorer diet, consumed less milk and meat, and had a greater incidence of intestinal parasites, which may account for the differences observed.

In contrast, reduced growth rates of young animals and reduced body weight in older animals are common observations associated with oral administration of excess sodium selenate, sodium selenite, or organic selenium compounds to experimental animals (Boylan et al. 1990; Cukierski et al. 1989; Donaldson and McGowan 1989; Gronbaek et al. 1995; Halverson et al. 1966; Harr et al. 1967; Hasegawa et al. 1994; Johnson et al. 2000; Nelson et al. 1943; NTP 1994, 1996; Palmer and Olson 1974; Panter et al. 1996; Penrith and Robinson 1996; Raisbeck et al. 1996; Sayato et al. 1993; Schroeder 1967; Tarantal et al. 1991; Thorlacius-Ussing 1990, Tsunoda et al. 2000; Turan et al. 1999a). This reduction in growth is often accompanied by reduced food and water consumption, and in dietary or drinking water studies may be an effect of poor palatability of selenium compounds. However, reduced growth is also observed in gavage studies (Cukierski et al. 1989; Hasegawa et al. 1994; Sayato et al. 1993) and, as discussed under endocrine and neurological effects, the growth retardation may have an endocrine or neurotransmitter component. Selenium effects on the levels of thyroid hormones (Behne et al. 1992; Behne and Kyriakopoulos 1993; Eder et al. 1995; Hotz et al. 1997), dopamine metabolites (Tsunoda et al. 2000), somatomedin C (Thorlacius-Ussing 1990) and insulin-like growth factor–binding protein-3 (Gronbaek et al. 1995) have been observed in selenium-treated animals.

Other Systemic Effects. Urinary excretion of selenium was about twice as great in children with a high incidence of dental caries than in children with a low incidence of caries (Hadjimarkos 1969b). Possible confounding factors (e.g., fluoride status and socioeconomic status) were not considered, however. In Yang et al. (1989a), the incidence of mottled teeth in the medium- and high-selenium groups was increased, but the effect was attributed to interactions between selenium and fluoride.

3.2.2.3 Immunological and Lymphoreticular Effects

No studies were located regarding adverse immunologic or lymphoreticular effects in humans after oral exposure to selenium or selenium compounds. One study indicates lymphocyte response was enhanced, as measured by the T-lymphocyte prolific response to pokeweed mitogen in elderly people taking a selenium-enriched yeast supplement (0.0014 mg/kg/day for 6 months). However, the elderly as a group

generally tend to have both lower blood selenium concentrations and lower lymphocyte proliferation than the general population (Peretz et al. 1991).

Studies of mice, rats, and cattle suggest that exposure to high doses of sodium selenite, but not selenomethionine, may reduce immunological responses (Johnson et al. 2000; Koller et al. 1986; Raisbeck et al. 1998; Yaeger et al. 1998). BALB/c mice (five males/group) administered sodium selenite in drinking water at 0, 1, 3, and 9 ppm selenium (0.024, 0.17, 0.38, and 0.82 mg selenium/kg/day) for 14 days showed significant decreases in the relative spleen weight at 9 ppm and the relative thymus weight at 3 and 9 ppm (Johnson et al. 2000). The number of splenocytes in the spleens of the 9 ppm group was reduced by 62%. Single-cell splenocyte cultures were made from the spleens of treated animals and used to determine the effects of selenium treatment on mitogen-induced lymphocyte blastogenesis and cytokine production. Cultured splenic lymphocytes from mice exposed to 9 ppm selenite showed a significant (260%) increase in the basal rate of proliferation and a nonsignificant increase in mitogen-induced proliferation. Exposure to 9 ppm selenite also produced a significant increase in the amount of tumor necrosis factor α (TNFα) and interleukin-1β (IL-1β) produced by lipopolysaccharide (LPS)-stimulated splenic macrophages. However, the results of this experiment must be interpreted with caution as treatment with 9 ppm selenium also produced a large and statistically significant decrease in food (21%) and water (43%) consumption so that some of the effects observed (e.g., changes in organ weights) may reflect effects of dehydration rather than selenium toxicity. In contrast, similar groups of mice treated with up to 9 ppm selenium (1.36 mg selenium/kg/day) as seleno-L-methionine showed no significant changes in body weight gain, organ weights, water consumption, or food consumption compared with controls. There were no changes in the basal or mitogen-stimulated lymphocyte proliferation following treatment with seleno-L-methionine and no alteration in the production of TNFα or IL-1β from splenic macrophages.

Another study in BALB/c mice examined the effects of consumption for 47 days of drinking water containing 7 ppm selenium as selenocystine, selenomethionine, or sodium selenite on immune function (Raisbeck et al. 1998). On the 14th day of the experiment, the mice received a subcutaneous injection of ovalbumin (OVA). Examination of mitogen-stimulated blastogenesis, B-cell function, and IgG concentrations at the end of the experimental period showed a significant decrease in B-cell function for mice treated with the two organic forms of selenium and a significant reduction in the concentration of OVA-specific antibodies for animals treated with any of the three forms of selenium. Total IgG concentration and OVA-stimulated blastogenesis did not vary between groups.

Rats given sodium selenite in drinking water at 0.7 mg selenium/kg/day for 10 weeks exhibited reduced humoral antibody (IgG) production in response to an administered antigen, and reduced prostaglandin synthesis, but there was no effect on natural killer cell (NKC) cytotoxicity (Koller et al. 1986). At lower doses (0.07 or 0.28 mg selenium/kg/day), NKC cytotoxicity was significantly increased, enhancing the immune response to antigenic stimulation, although the delayed-type hypersensitivity (DTH) and prostaglandin E_2 synthesis were significantly reduced. Selenium administration did not affect the ability of resident peritoneal cells to produce interleukin IL-1. Given the enhanced NKC activity at 0.07 and 0.28 mg selenium/kg/day, but not at 0.7 mg selenium/kg/day, and given the reduced antibody and prostaglandin synthesis at 0.7 mg selenium/kg/day, the dose of 0.7 mg selenium/kg/day is identified as the lowest LOAEL. A NOAEL cannot be identified because of the conflict between enhanced NKC activity and reduced DTH and prostaglandin E_2 synthesis occurring at the same dose levels in this study.

Antibody responses to ovalbumin were significantly lower in five male antelope (*Antilocapra americana*) fed a diet containing 15 ppm selenium (a mixture of alfalfa and hay naturally high in selenium) for 164 days than in controls fed a similar diet containing only 0.3 ppm selenium, but there was no difference in total globulin concentration between groups (Raisbeck et al. 1996). No clinical signs of selenosis or treatment-associated lesions were observed in these animals.

Treatment of steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days produced symptoms of selenosis (hoof lesions), but did not produce any changes in the weight or histology of the spleen or thymus or in the histology of the lymph nodes (O'Toole and Raisbeck 1995).

Leukocyte function was significantly reduced in pregnant cows supplemented with 0.135 mg/kg/day selenium for 3 months from diets that contained 0.25 (control), 6, or 12 ppm selenium as sodium selenite. Treated animals showed a significant decrease in forced antibody production and a depression in mitogenic response compared with controls (0.005 mg selenium/kg/day) (Yaeger et al. 1998). No clinical signs of selenium toxicosis were observed in any of the cows during the experiment.

As selenium can enhance some immune system functions, selenium may have a normal physiological function in the immune system. This is supported by an 8-week study in which treatment of mice with selenium as sodium selenite (0.33 mg selenium/kg/day, dietary) resulted in enhanced ability of cytotoxic T-lymphocytes to destroy tumor cells (Kiremidjian-Schumacher et al. 1992).

No studies were located concerning immunological or lymphoreticular effects in humans or experimental animals following oral exposure to selenium sulfide or selenium disulfide.

The highest NOAEL values and all LOAEL values from each reliable study for immunological effects following oral exposure to selenium or selenium compounds for each species and duration category are recorded in Table 3-2 and plotted in Figure 3-2.

3.2.2.4 Neurological Effects

Following acute oral exposure to selenium compounds in humans, aches and pains and irritability (Civil and McDonald 1978), as well as chills and tremors (Sioris et al. 1980) have been reported. The dizziness associated with selenium inhalation exposure has not been documented after selenium ingestion.

In a 1964 study, Rosenfeld and Beath reported listlessness, a general lack of mental alertness, and other symptoms of selenosis in a family exposed for approximately 3 months to well water containing 9 mg selenium/L (0.26 mg selenium/kg/day from drinking water). All of the symptoms resolved after use of the seleniferous water was discontinued. Because Rosenfeld and Beath (1964) did not estimate the family's exposure to dietary selenium, it is not possible to identify the total daily selenium dose associated with the symptoms of selenosis in this family.

In a dietary study of 11 men in a metabolic unit, selenium intake (80 µg/day for the first 21 days, then either 13 (n=6) or 356 (n=5) µg/day for 14 weeks) was reported to have no significant effect on mood, as measured using the Bi-Polar form of the profile of mood states (POMS) (Hawkes and Hornbostel 1996). However, subjects with initially low selenium levels did show significantly greater decreases in mood scores during selenium depletion.

In areas of the People's Republic of China where populations suffer from chronic selenosis, peripheral anesthesia and pain in the limbs were reported (Yang et al. 1983). In extreme cases, exaggerated tendon reflexes, convulsions, and some paralysis and hemiplegia occurred (Yang et al. 1983). These latter cases were associated with an estimated daily dietary intake of selenium of at least 3.22 mg selenium/person/day, averaging 4.99 mg selenium/person/day (Yang et al. 1983). Assuming a weight of 55 kg for Chinese men (Yang 1989b), these dietary levels represent 0.027 mg selenium/kg/day and 0.09 mg selenium/kg/day, respectively. In another high selenium area, no neurological effects were observed in individuals who consumed up to 1.51 mg selenium/day (0.027 mg/kg/day) (Yang et al. 1983). Danish

geriatric patients with a mean age of 75.3 years were given daily either a placebo or an antioxidant cocktail containing 0.004 mg/kg/day of selenium as L-selenomethionine along with zinc, vitamins C, A, B6, and E, and gamma-linolenic acid. After 1 year, whole blood selenium concentrations increased in the treated group, and slight but significant improvements in psychological scores were observed (Clausen et al. 1989). Because a mixture of nutrients was administered, the improvement in the patients cannot be attributed to selenium. People living on ranches with high selenium soils where selenium toxicity in livestock had historically been observed were compared to randomly selected residents in Wyoming and South Dakota. Daily selenium intake was measured by analysis of duplicate food portions. Subjects received a complete physical exam with a symptom questionnaire and laboratory tests. There were no biologically significant changes in clinical signs or blood chemistry. Calculated doses ranged from 0.001 to 0.01 mg selenium/kg/day in the diet (Longnecker et al. 1991).

An increased incidence (4 observed cases, 0.97 expected, standardized incidence ratio = 4.14, 95% confidence interval [CI]=1.13–10.60) of amyotrophic lateral sclerosis, a human motor neuron disease of unknown origin, was reported for a cohort of 5,182 residents of Reggio Emilia, Italy who had been exposed to drinking water with high selenium content (7 μg/L) from 1972 to 1988 compared with the incidence among residents of the surrounding area who had received municipal water containing <1 μg/L selenium (Vinceti et al. 1996). A subcohort of 2,065 if these individuals who had been exposed from 1974 (the earliest date for which a chemical analysis of the municipal tap water was available) was also examined and found to have an increased incidence ratio (4 observed cases, 0.47 expected, standardized incidence ratio = 8.59, 95% CI=2.34–21.98).

In a 30-day study of the administration of L-selenomethionine to long-tailed macaques, severe hypothermia was observed in two of five animals administered 0.12 mg selenium/kg/day, but not in any of the eight animals receiving #0.08 mg selenium/kg/day (Cukierski et al. 1989). However, the increased incidence of hypothermia was not statistically significant. Following 1 week of treatment, all animals administered L-selenomethionine, including the two macaques treated with 0.01 mg selenium/kg/day, exhibited increased drowsiness and lethargy (Cukierski et al. 1989).

Symmetrical focal poliomyelomalacia and other forms of paralysis were seen in swine exposed to 0.58–2.1 mg selenium/kg/day after both acute and intermediate exposures (Baker et al. 1989; Goehring et al. 1984; Harrison et al. 1983; Mihailovic et al. 1992; Panter et al. 1996; Penrith and Robinson 1996; Stowe et al. 1992; Wilson et al. 1983, 1988, 1989). This lesion was noted in animals that showed ataxia, inability to stand, and paralysis of the hind limbs. Additionally, bilateral lesions were noted in the ventral

horns of the cervical and lumbar/sacral intumescences of the spinal cord. Necrosis and cavitation were evident in the larger lesions (Harrison et al. 1983). Bilateral lesions were also observed in several nuclei of the brain stem and in the reticular formation (Wilson et al. 1983). Wilson et al. (1983) reproduced the syndrome in growing pigs by feeding them sodium selenite at 50 mg selenium/kg in the diet for 20–40 days. The study authors did not provide sufficient information to calculate doses on a mg selenium/kg body weight basis, but assuming that young swine consume approximately 4% of their body weight each day, this dose was approximately 2.1 mg/kg/day.

In a study of weaned 5-week-old pigs, a dose of 1.3 mg selenium/kg/day given as sodium selenite in capsules killed all eight pigs within 10 days. Histopathological lesions were found in the brain and spinal cord (Wilson et al. 1989). In a study in which pigs were fed 1.25 mg selenium/kg/day in the form of *A. bisulcatus*, D,L-selenomethionine, or selenate, the selenium in *A. bisulcatus* was the most potent neurotoxin, resulting in complete paralysis in four of five pigs after 5 days of treatment, and in the last pig after 3 weeks of treatment (Panter et al. 1996). In pigs fed selenate, three of five developed complete paralysis, and one pig developed posterior paralysis after 4–21 days of treatment. Although D,L-selenomethionine resulted in the greatest incidence of selenosis, it was the least potent neurotoxicant, resulting in posterior paralysis in two of five pigs after 9 and 24 days of treatment; the pigs that did not develop paralysis were fed D,L-selenomethionine for approximately 31 days. The form of selenium in *A. bisulcatus* is unknown, although Panter et al. (1996) indicate that it is nonprotein.

It has long been believed that blind staggers in livestock results from consumption of plants high in selenium (100–10,000 mg selenium/kg plant) (Rosenfeld and Beath 1964). These plants, which include *A. bisulcatus*, are known as selenium-indicator plants. Blind staggers is characterized by impaired vision, aimless wandering behavior, reduced consumption of food and water, and finally paralysis and death (Rosenfeld and Beath 1964; Shamberger 1986). Trembling of the skeletal muscles was observed in steers fed sodium selenite mixed in the feed at doses between 0.6 and 1.1 mg selenium/kg/day (Maag et al. 1960). At necropsy, two of six steers exhibited neuronal degeneration of the cerebral and cerebellar cortices. However, more recent studies in which cattle were treated with known amounts of selenium have not replicated these effects , and it is likely that blind staggers is not solely the result of selenium toxicity, but may also be attributable to other unidentified causes. For example, treatment of 20 steers with selenomethionine or selenite in food at doses up to 0.808 mg selenium/kg/day for 120 days produced symptoms of selenosis (hoof lesions), but did not produce any neurological signs associated with blind staggers and did not produce any treatment-related changes in the histology of the central nervous system (O'Toole and Raisbeck 1995).

Neurological effects have also been reported for mice after acute or intermediate exposures to selenium. A single oral dose of selenium dioxide dissolved in water given to mice at $1/10^{th}$ the LD_{50} (1.7 mg/kg) caused moderate reductions in alertness, spontaneous activity, touch response, muscle tone, and respiration. Pentobarbital sleeping time was also significantly increased, and there was moderate hypothermia (Singh and Junnarkar 1991). Brain tissue from male BALB/c mice administered sodium selenite or seleno-L-methionine in drinking water at 0, 1, 3, or 9 ppm selenium (sodium selenite : 0.03, 0.24, 0.58, or 1.34 mg selenium/kg/day; seleno-L-methionine: 0.03, 0.26, 0.63, or 1.96 mg selenium/kg/day) for 14 days was examined for changes in the concentrations of norepinephrine (NE), dopamine (DA), dihydroxyphenylacetic acid (DOPAC), homovanillic acid (HVA), serotonin (5-HT), and 5-hydroxyindolacetic acid (5-HIAA) (Tsunoda et al. 2000). Treatment with seleno-L-methionine produced no significant changes in the concentrations of any of the neurotransmitters or their metabolites. DOPAC, DA, and HVA were increased in the striatum of mice receiving 3 or 9 ppm selenium as selenite. The increase was significant at both concentrations for DOPAC and at 3 ppm (but not 9 ppm) for HVA, but was not significant at either concentration for DA. No changes were observed for levels of NE, 5-HT, or 5-HIAA levels in any brain region of mice treated with sodium selenite.

Exposure of female BALB/c mice to 0.21 mg selenium/kg/day for 6 months from diets containing selenium as sodium selenite resulted in significant changes in behavior during open field testing (Boylan et al. 1990). Open field testing measures the arousal level of small rodents and can differentiate between fear-related behavior and general arousal. Mice receiving excessive selenium had reduced sniffing behavior and exhibited greater activity entering more squares, and more interior squares than mice receiving normal selenium diets. These behaviors are indicative of a general state of arousal rather than fear-motivated activity.

No studies were located concerning neurological effects in humans or experimental animals following oral exposure to selenium sulfide or disulfide.

The highest NOAEL values and all LOAEL values from each reliable study for neurological effects following oral exposure to selenium or selenium compounds for each species and duration category are recorded in Table 3-2 and plotted in Figure 3-2.

3.2.2.5 Reproductive Effects

No studies were located regarding adverse effects on human reproduction following oral exposure to elemental selenium or to selenium compounds. No significant correlation was found between the seminal plasma selenium concentration and sperm count or mobility in sperm samples from 211 men (Roy et al. 1990). A nonsignificant increase in spontaneous abortions (relative risk [RR]=1.73; 95% CI=0.62–4.80) was reported among births in the municipality of Reggio Emilia, Italy, where women had been exposed to drinking water containing 7–9 µg selenium (as selenate) between 1972 and 1988 (Vinceti et al. 2000a). A preliminary study completed in China suggests that selenium supplementation (100 µg/day, form not stated) during pregnancy may reduce the incidence of pregnancy-induced hypertension (Li and Shi-mei 1994).

Selenium administered in the diet or in drinking water over short exposure periods (e.g., 1 month) does not appear to affect the fertility of female animals unless the intake is sufficiently high to cause general toxicity (Nobunaga et al. 1979). Despite a small increase in the number of abnormal length estrous cycles, Nobunaga et al. (1979) found no adverse effect on the fertility of female mice from administration of sodium selenite at doses of 0.34 mg selenium/kg/day in drinking water for 30 days before mating and for 18 days during pregnancy. On the other hand, chronic exposure of mice and rats to otherwise nontoxic doses has been shown to reduce fertility and to reduce markedly the viability of the offspring of pairs that are able to conceive (Schroeder and Mitchener 1971b; Wahlstrom and Olson 1959b).

A study of supplementation of female pigs with 0.1 or 0.3 ppm selenium (doses not available) administered as a selenium-enriched yeast or sodium selenite in the diet, from 60 days before breeding until weaning found no adverse effects on reproductive performance (measured by number of offspring) or growth (Mahan and Kim 1996). In another study of the effect of selenium on fertility in pigs, females fed sodium selenite at 0.4 mg selenium/kg/day from 8 weeks of age exhibited reduced rates of conception and also produced offspring with significantly reduced birth weight and weaning weights in the first and second litters (Wahlstrom and Olson 1959b). An altered menstrual cycle was reported in monkeys administered 0.08 mg selenium/kg/day as L-selenomethionine for 30 days (Cukierski et al. 1989).

Vaginal cytology of female rats provided with drinking water containing selenate or selenite indicated that the rats spent more time in diestrus and less time in proestrus and estrus than the controls (NTP 1994). This effect occurred following treatment with 0.31 mg selenium/kg/day as selenate or 0.86 mg selenium/kg/day as selenite. The animals in these studies were not mated, so it is not known if the effects

on the estrous cycle had any effect on fertility. Effects on the estrous cycle were not observed in mice treated with selenate or selenite in the drinking water at doses up to 7.17 mg selenium/kg/day for selenate, or at doses up to 3.83 selenium/kg/day for selenite (NTP 1994).

In a three-generation reproduction study, selenium administered as sodium selenate (0.57 mg selenium/kg/day) in the drinking water of breeding mice produced adverse effects on reproduction (Schroeder and Mitchener 1971b). The most notable observed effects included the failure of about half of the F3 generation pairs to breed successfully. In a two-generation study using rats, selenium administered as potassium selenate had no effect on reproduction at a dose of 0.21 mg selenium/kg/day for 1 year; however, decreased fertility and pup survival were noted at 1.05 mg selenium/kg/day (Rosenfeld and Beath 1954). At 0.35 mg selenium/kg/day for 1 year, the number of young successfully reared by the females was reduced by 50%, and the body weight of the females was approximately 20% less than that of the control females (Rosenfeld and Beath 1954).

A short-term reproductive study of the effects of sodium selenate in drinking water on rats reported some female reproductive toxicity (reduced corpora lutea, reduced implants per litter, shorter estrous cycle), but only at doses (0.418 mg selenium/kg/day) that produced signs of severe maternal toxicity, including a large reduction in water consumption (NTP 1996).

Data from animal studies suggest that exposure to excessive selenium has adverse effects on testosterone levels and sperm production and increases the percentage of abnormal sperm (El-Zarkouny et al. 1999; Kaur and Parshad 1994; NTP 1994). A significant reduction (49%) in serum testosterone levels was reported for New Zealand White rabbits gavaged with 0.3 mg selenium/kg (0.001 mg selenium/kg/day) as sodium selenite once a week for 6 weeks (El-Zarkouny et al. 1999). The percentage of spermatozoa without an acrosome was also increased in treated rabbits compared with controls, but the difference was not significant. Sperm motility, ejaculate volume, sperm concentration, and total sperm output were all reduced by selenium treatment, but statistical analysis of these data was not presented.

Exposure of Wistar rats to 0.234 mg selenium/kg/day as sodium selenite in water produced testicular hypertrophy (Turan et al. 1999a). A dose-related increase in abnormal sperm and a decrease in live sperm were observed in wild-caught rats exposed to selenite in the diet at 0.1 and 0.2 mg selenium/kg/day (Kaur and Parshad 1994). The percentage of abnormal sperm was 3.9% at 0.1 mg/kg/day and 24.6% at 0.2 mg/kg/day. The abnormalities observed were principally in the midpiece region of the sperm, the region that contains a selenoprotein (Sunde 1990). Decreased sperm counts were observed in rats

provided with selenate or selenite in drinking water for 13 weeks at a dose of 0.29 mg selenium/kg/day for selenate and a dose of 0.17 mg selenium/kg/day for selenite (NTP 1994). Effects on sperm were not observed in mice treated with selenate or selenite in the drinking water at doses up to 5.45 mg selenium/kg/day for selenate, and at doses up to 3.31 mg selenium/kg/day for selenite (NTP 1994). The administration of 1.05 mg selenium/kg/day as potassium selenate to rats in drinking water for 1 year did not affect male fertility (Rosenfeld and Beath 1954), and the administration of 0.57 mg selenium/kg/day as sodium selenate for three generations did not reduce male fertility in mice (Schroeder and Mitchener 1971b). A short-term reproductive study of the effects of sodium selenate in drinking water on rats at doses (0.418 mg selenium/kg/day) that produced signs of systemic toxicity did not cause any increase in sperm abnormalities or lesions of the testis or epididymis (NTP 1996).

In a review of selenium poisoning in domestic animals, Harr and Muth (1972) noted a decreased conception rate and an increased fetal resorption rate in cattle, sheep, and horses fed diets naturally containing organic selenium compounds at 25–50 mg selenium/kg diet. Assuming that large animals consume an amount of food equal to about 2–3% of their body weight daily, the doses would have been approximately 0.5–1.5 mg selenium/kg/day. These levels of selenium also produced other signs of toxicity, including hair loss, lameness, and degeneration and fibrosis of the heart, liver, and kidneys. In a case control study of 136 Holstein cows from four herds, an association of cystic ovaries with blood selenium concentrations greater than 108 ng/mL was found (Mohammed et al. 1991). The concentration of progesterone in the milk was significantly higher in the controls than in the cows receiving selenium supplementation, but no information on the selenium dose was presented. No change in estrus cycle length, estrus behavior, progesterone, or estrogen profiles or pregnancy rate was observed in a study of the reproductive response of ewes fed alfalfa pellets containing sodium selenate (24 ppm selenium) or *A. bisculcatus* (29 ppm selenium) as a selenium source for 88 days, from >52 days before pregnancy up to day 28 of gestation (Panter et al. 1995). Doses could not be calculated as food consumption was not listed, and the paper states that the food supply was limited to match that of the group with the lowest intake.

The highest NOAEL value for reproductive effects following intermediate oral exposure to sodium selenite and all reliable LOAEL values for reproductive effects following intermediate or chronic oral exposure to selenium compounds other than selenium sulfide are recorded in Table 3-2 and plotted in Figure 3-2.

3.2.2.6 Developmental Effects

No studies have demonstrated that selenium or its compounds are teratogenic in humans. Robertson (1970) reported on the outcome of pregnancies in a laboratory in which workers handled sodium selenite. Of the five pregnancies, four ended in spontaneous abortion and one resulted in an infant with bilateral clubfoot. The urinary selenium levels in all subjects were similar to those in other individuals living in the same area. The limited number of cases, possible exposure to other toxic agents, and other confounding factors leave the relationship between sodium selenite and developmental effects inconclusive.

No significant increase in spontaneous abortions (RR=1.73; 95% CI=0.62–4.80) was reported among births in the municipality of Reggio Emilia, Italy, where women had been exposed to drinking water containing 7–9 μg selenium (as selenate) between 1972 and 1988 (Vinceti et al. 2000a). Body weight and length at birth were similar in infants of exposed and unexposed women, and no significant increase in the prevalence of congenital abnormalities was found for 353 infants of exposed mothers compared with the 14,481 births among unexposed women.

Zierler et al. (1988) performed a case control study of 270 children born in Massachusetts with severe congenital heart disease and 665 controls randomly selected from birth certificates. The study compared the selenium concentrations in the public drinking water supply used by the mothers close to the time of conception to the selenium concentrations in the water consumed by the controls. The results indicated that selenium exposure via drinking water was associated with a reduction in the risk of congenital heart defects (cono-truncal defects, venticular septal defects, coarctation of the aorta, and patent ductus arteriosis), but many variables are unknown, including other possible confounders (no adjustment for age, parity, tobacco, alcohol, drug use, or socioeconomic status), other sources of selenium in the mothers' diet and environment, the amount of drinking water consumed, and the selenium concentrations in the water during the first trimester.

Excess selenium is a demonstrated teratogen in birds. However, there is no clear evidence linking selenium exposures to teratogenic effects in mammals. Several studies have documented the sensitivity of chick embryos to selenium poisoning. Hatchability of eggs is reduced by dietary levels of organic selenium in grain that are too low to cause toxicity in other farm animals. The eggs are fertile but often produce grossly deformed embryos lacking eyes and beaks and having deformed wings and feet (Franke and Tully 1935; Franke et al. 1936; Gruenwald 1958; Palmer et al. 1973). Deformed embryos have also

been produced by injection of aqueous sodium selenite or sodium selenate into the air cell of the normal, fertile eggs of chickens (Franke et al. 1936; Khan and Gilani 1980). The incidence of malformation among coot, duck, stilt, and grebe embryos from eggs of birds ingesting plant and other food from irrigation drainwater ponds in the San Joaquin Valley of California was much higher than expected (10–42%, depending on the species, versus less than 1% based on data from other areas) (Ohlendorf et al. 1986a, 1988). Selenium concentrations in these ponds were greater than 0.3 mg/L.

The consumption of naturally high seleniferous diets by sheep (Rosenfeld and Beath 1964) and cattle (Dinkel et al. 1963) may interfere with normal fetal development and produce malformations. Malformations were associated with alkali disease and occurred at dietary levels that produced other toxic manifestations, but it is not clear if these reports took into account consumption of other toxic range plants. The specific selenium compound or compounds possibly associated with livestock developmental toxicity have not been identified. No change in the outcome of pregnancy was observed in a study of the reproductive response of ewes fed alfalfa pellets containing sodium selenate (24 ppm selenium) or *Astragalus bisculcatus* (29 ppm selenium) as a selenium source for 88 days, from >52 days before pregnancy up to day 28 of gestation (Panter et al. 1995). All lambs appeared normal; and there was no significant difference in the number or weight of lambs born to treated and control ewes. Doses could not be calculated, as food consumption was not listed, and the paper states that the food supply was limited to match that of the group with the lowest intake

In an intermediate-duration study, an increased number of deaths between birth and weaning, reduced birth weight, and reduced body weight at weaning were observed in offspring of pigs fed selenite at 0.4 mg/kg/day for an unstated period of time (Wahlstrom and Olson 1959b). Treatment of 15 pregnant cows with diets containing 0.25 (control), 6, or 12 ppm selenium (0.005, 0.135, or 0.265 mg Se/kg/day) as sodium selenite beginning at 80 to 110 days gestation and continuing for 3 months resulted in no abnormalities among the offspring apart from one calf in the 12 ppm group which was born weak and subsequently died (Yaeger et al. 1998). This calf had myocardial lesions similar to those described for selenium toxicosis and had markedly elevated hepatic selenium levels, although selenium levels in blood and hair of this calf and its dam were lower than average for the 12 ppm group.

In studies of laboratory mammals, the administration of inorganic selenium compounds at levels that are not maternally toxic has not produced terata (Bergman et al. 1990; Chiachun et al. 1991; Ferm et al. 1990; NTP 1996; Poulsen et al. 1989; Rosenfeld and Beath 1954; Schroeder and Mitchener 1971b; Thorlacius-Ussing 1990). Ferm et al. (1990) administered a single dose of sodium selenate, sodium selenite, or

L-selenomethionine (0, 1.8, 2.2, 2.7, 4.0, 5.0, or 5.9 mg selenium/kg/day) to pregnant Syrian hamsters on gestation day 8. Pathological examination of the fetuses on day 13 showed that the percentage of abnormal litters was significantly increased at a dose of 2.7 mg/kg and greater. Encephalocele was the major malformation noted, and incidences were as follows: 0/71 controls; 4/55 (7.3%) at 1.8 mg/kg; 1/49 (2%) at 2.2 mg/kg; 7/66 (10.6%) at 2.7 mg/kg; 15/70 (21.4%) at 4 mg/kg; 9/38 (23.7%) at 5 mg/kg; and 6/16 (37.5%) at 5.9 mg/kg. Nobunaga et al. (1979) found that administration of sodium selenite in drinking water at 0.34 mg selenium/kg/day for 30 days before mating and for 18 days during pregnancy slightly, but significantly, reduced fetal growth in mice. However, there was no effect on fetal growth in the same study at a dose of 0.17 mg selenium/kg/day. A short-term developmental study (from gestation day 6 until birth) of the effects of sodium selenate in drinking water on rats produced some developmental toxicity (decreased number of live births, reduced pup weight, increased gestation period), but only at doses (0.418 mg selenium/kg/day) that produced signs of severe maternal toxicity including a large reduction in water consumption (NTP 1996). Selenium administered as potassium selenate in drinking water to male and female rats at a dose of 1.05 mg selenium/kg/day for 1–8 months for two successive generations did not cause congenital malformations (Rosenfeld and Beath 1954). Similarly, administration of 0.57 mg selenium/kg/day as sodium selenate in the drinking water of breeding mice for three generations did not have teratogenic effects, although there was an increased incidence in fetal deaths, and a high proportion of the surviving offspring were runts (Schroeder and Mitchener 1971b).

Poulsen et al. (1989) demonstrated that pigs exposed to 42.4 mg/day of selenium as sodium selenite in feed throughout pregnancy produced normal litters, with no adverse effect on piglet survival, litter size, or body weight at birth. Body weights of the pigs during pregnancy were not provided, and therefore mg/kg/day doses could not be calculated. Body weight gains of pigs fed selenium as selenite at a dose of 0.4 mg selenium/kg/day after weaning (duration not specified) were reduced (Wahlstrom and Olson 1959a). The reduction in body weight gain was greater among pigs from dams not fed selenium during gestation and lactation compared to pigs fed selenium (0.4 mg/kg/day) during gestation and lactation. Without providing data, the study authors indicated that there was a greater loss of pigs at birth and during lactation from sows fed selenium, which may have eliminated susceptible pigs.

In a teratology study of long-tailed macaques, no gross abnormalities or growth retardations were observed in fetuses from mothers administered L-selenomethionine at levels of 0.003, 0.025, 0.15, or 0.30 mg selenium/kg/day on gestational days 20–50 (10 animals per group); the mid and high doses were maternally toxic (Tarantal et al. 1991).

The highest NOAEL value and all reliable LOAEL values for developmental effects following intermediate or chronic oral exposure to selenium compounds are recorded in Table 3-2 and plotted in Figure 3-2.

3.2.2.7 Cancer

Early studies reporting that selenium was carcinogenic in mammals after being provided as seleniferous corn or wheat in the diet (Nelson et al. 1943), as sodium selenite or sodium selenate in drinking water (Schroeder and Mitchener 1971a), or as sodium selenate in the diet (Volgarev and Tscherkes 1967) were flawed. The majority of subsequent studies of humans and animals have revealed either no association between selenium intake and the incidence of cancer (Azin et al. 1998; Beems 1986; Coates et al. 1988; Harr et al. 1967; Ma et al. 1995; Menkes et al. 1986; Ratnasinghe et al. 2000; Thompson and Becci 1979; Viacetti et al. 1995; Virtamo et al. 1987) or a chemopreventive association (Birt et al. 1982; Clark et al. 1996a, 1999; Finley et al. 2000; Ip 1981, 1983; Ip and Lisk 1995; Ip et al. 1996, 1997, 1998, 2000a, 2000b; Jiang et al. 1999; Ma et al. 1995; Medina and Shepherd 1981; Overvad et al. 1985; Schrauzer et al. 1976, 1977; Shamberger et al. 1976; Soullier et al. 1981; Thompson and Becci 1980; Woutersen et al. 1999; Yoshizawa et al. 1998).

The only selenium compound that has been shown to be carcinogenic in animals is selenium sulfide (NTP 1980c), although there is also some evidence for carcinogenicity due to ethyl selenac (Innes et al. 1969; NCI 1968). These compounds are very different chemically from the organic and inorganic forms found in foods and the environment. Human dietary studies generally do not identify the selenium form specifically; both organic (from grains and other plant and animal products) and inorganic (from drinking water) forms are ingested. Animal bioassays in which selenium was administered as sodium selenate, sodium selenite, or organic forms of selenium have all shown similar negative results.

Excess incidence of melanoma was reported for a cohort of 2,065 individuals exposed to selenium as selenate that contaminated the municipal water supply in Reggio Emilia, Italy from 1972 until 1988 (Vinceti et al. 1998). Eight individuals among the exposed cohort developed melanoma compared with 128 in the remainder of the municipal population (total number of individuals not given). The SMRs were 5.0 (95% CI=1.6–12.0) for males and 3.2 (95% CI=1.0–7.7) for females. The authors estimate the general dietary intake of selenium in the area to be 45–50 μg/day and the excess selenium supplied in the contaminated tap water to be 10–20 μg/day. However, the study is limited by the fact that no individual

measurements of selenium exposure were made, and individuals were classed as exposed or unexposed depending on their place of residence.

A study of the effects of nutritional supplementation with selenium found a significant reduction in overall cancer mortality and in the incidence of lung (RR=0.54, 95% CI=0.3–0.98, P=0.04), colorectal (RR=0.42, 95% CI=0.18–0.95, P=0.03), and prostate (RR=0.37, 95% CI=0.18–0.71, P=0.003) cancer (Clark et al. 1996a, 1999). The original intent of the study was to assess the effects of selenium supplementation on nonmelanoma skin cancer. Patients with a history of skin carcinoma (1,312 individuals) were randomized into two groups; one group received a selenium supplement of 200 μg/day, and the other received a placebo. Groups were treated for an average of 4.5 years and followed for an average of 6.5 years. Supplementation produced no difference in skin cancer incidence; however, secondary end point analyses of the data found a protective effect for selenium for the cancers described above.

Supporting evidence for an antiprostate cancer effect of selenium was obtained for a nested case-control design within the Health Professional Follow-up study (Yoshizawa et al. 1998), which found that higher prediagnostic selenium levels were associated with reduced prostate cancer incidence. This study included 33,737 male health professionals aged 40–75 years who provided toenail clippings in 1987. The cohort was assessed by questionnaire for incidence of new cases of prostate cancer from 1989 to 1994. Higher levels of selenium in toenail clippings were significantly associated with a reduced risk of prostate cancer. After controlling for factors such as a family history of prostate cancer, body mass index, calcium intake, lycopene intake, saturated fat intake, vasectomy, and geographical region, the odds ratio (OR) was 0.35 (95% CI=0.16–0.78, P for trend=0.03).

Epidemiological studies that focused on the selenium concentration of forage crops as an indicator of available dietary selenium indicated an inverse association between selenium levels and cancer occurrence, with few exceptions. In the United States, male mortality due to cancer of the tongue, esophagus, stomach, intestine, rectum, liver, pancreas, larynx, lungs, kidneys, and bladder was significantly lower in states with high selenium levels in forage crops (concentrations in excess of 0.10 mg selenium/kg) (Shamberger et al. 1976). For females in states with high selenium levels, significantly lower cancer death rates were found for cancer of the esophagus, stomach, intestine, rectum, liver, pancreas, lungs, bladder, thyroid, breast, and uterus (Shamberger et al. 1976). Only male and female mortality due to cancer of the skin and eye, male mortality due to cancer of the lip and aleukemia (absence or deficiency of leukocytes in the blood), and female mortality due to dermal melanoma were

associated with high selenium levels in forage crops. Many of the high selenium areas are in the southwestern portion of the United States, and therefore, exposures to ultraviolet light may have contributed to the skin cancers observed in these areas (Shamberger et al. 1976). In a comparison of selenium intake and cancer mortality rates in different countries, Schrauzer et al. (1977) detected a cancer chemopreventive association between the selenium content of the diet and age-corrected cancer mortality from leukemia and cancers of the intestine, rectum, breast, ovary, prostate, lung, pancreas, skin, and bladder.

In a case control study of lung cancer patients, Menkes et al. (1986) found that the risk of lung cancer was not associated with serum selenium levels (0.113 and 0.110 mg selenium/L in cases and controls, respectively), but significantly increased with decreasing serum levels of vitamins A and E. The study authors suggested that high serum selenium levels were significantly associated with an increased incidence of squamous cell carcinoma as compared to other cellular tumor types, but the statistical analysis used was flawed. In a region of China with high rates of stomach cancer and low intake of several micronutrients (selenium not specifically stated), an intervention trial in 29,584 adults for 5.25 years demonstrated a 21% decrease in stomach cancer in the randomly selected group receiving a nutritional supplement of beta carotene (15 mg/day), vitamin E (30 mg/day), and selenium (50 μg selenium/day as selenium yeast) (Blot et al. 1993). However, because the three nutrients were given in combination, it is not possible to determine what part of this effect (if any) was due to selenium.

Low serum selenium levels have been associated with an increased incidence of cancer in some prospective epidemiological studies (Salonen et al. 1984, 1985; Willet et al. 1983). In the United States, Willet et al. (1983) found that the risk of cancer for subjects in the lowest quintile (fifth) of serum selenium concentrations (<0.115 mg selenium/L) was twice that of subjects in the highest quintile (>0.154 mg selenium/L). In Finland, Salonen et al. (1985) found the risk of fatal cancer for subjects in the lowest tertile (third) of serum selenium concentrations (<0.047 mg selenium/L) was 5.8 times higher than that of the remaining subjects. Mean serum selenium levels in Americans (0.129 mg selenium/L cases; 0.136 mg selenium/L controls) (Willett et al. 1983) are more than twice the mean serum selenium levels in the Finns (0.0505 mg selenium/L for cases; 0.0543 mg selenium/L for controls) (Salonen et al. 1984). Although the age-specific risk of fatal cancers in the two populations cannot be calculated from the data reported, the overall incidence of cancer in the 4-year Finnish study was less than half that in the 5-year U.S. study. In addition, other prospective studies of Americans have found no correlation between fatal cancer and blood selenium concentrations (Coates et al. 1988). Thus, one cannot predict relative cancer risks with serum selenium levels in one population based on data from another population.

A 9-year prospective follow-up study was conducted by Virtamo et al. (1987) on a group of men in Finland. At the beginning of the study, blood samples were taken as part of a study of coronary heart disease and other atherosclerotic diseases. Cancer data were collected from central registries for the years 1976 through 1983. The results indicated no association between low serum selenium levels (<0.045 mg selenium/L) and an increased risk of cancer. Evidence suggests that combined dietary deficiencies of vitamin E and selenium may be associated with increased cancer risk (Salonen et al. 1985).

Epidemiological studies of breast cancer have found inverse correlations, positive correlations, and no correlations between tissue selenium concentrations and cancer incidence (recently reviewed by Garland et al. 1993). In a case control study of plasma selenium and breast cancer in which the controls had benign breast disease, a preventive effect of selenium was found only among individuals who had higher plasma selenium and were not taking selenium supplementation (Hardell et al. 1993). This effect was significant (odds ratio 0.38) at a serum selenium concentration range of 0.08–0.09 mg/L in women 50 years old or more. GSH-Px activity in erythrocytes was not found to be a marker for the risk of breast cancer. A case control study of 162 cases of breast cancer in Dutch women did not find a significant difference in dietary, plasma, erythrocyte, or toenail selenium between cases and 529 controls when multivariate-adjusted odds ratios were calculated. Dutch women have lower selenium intake than women in the United States and one of the highest incidences of breast cancer in Western Europe. The authors of this study surmised that other studies reporting an inverse relationship between selenium levels and breast cancer may be seeing an effect of the cancer (e.g., decreased uptake of selenium or anorexia), rather than lower selenium level contributing to the development of cancer (van't Veer et al. 1990). Similarly, a large prospective study of 434 cases in the United States found no correlation between selenium content in nails, established as a long-term marker of selenium (Hunter et al. 1990a), and breast cancer (Hunter et al. 1990b). It is interesting to note that a more recent investigation of the same cancer cases found an inverse correlation between vitamin A (retinoids) in the diet and breast cancer (Hunter et al. 1993). Retinoids are believed to have chemopreventive activity (Clausen et al. 1989; Hunter at al. 1993). Although the data as a whole for breast cancer and tissue selenium concentrations do not support a clear chemopreventive effect for selenium, it is possible that very high selenium concentrations or very low selenium concentrations outside the ranges observed in these studies could play a role in human cancer risk (Garland et al. 1993).

There were several inadequacies in the early studies that reported carcinogenic effects in animals following oral administration of selenium-containing compounds. Nelson et al. (1943) (also reported as Fitzhugh et al. 1944) administered naturally seleniferous corn or wheat diets containing 5, 7, or 10 mg

selenium/kg diet (0.25, 0.35, or 0.50 mg selenium/kg/day) to female rats for 2 years. Selenium administration produced high mortality (69%) in all treatment groups by the end of the first 12 months, and the first tumors appeared after 18 months of treatment. Tumors developed only in animals with cirrhotic livers, and the tumors were reported to be nonmalignant. The possible contribution of overt hepatotoxicity to the development of liver tumors is not known. The incidences of tumors in the surviving animals in the three dose groups were 6/25, 3/21, and 2/7, respectively. The investigators had difficulty discerning malignant from nonmalignant tumors, and most animals had died of cirrhosis of the liver before the appearance of liver tumors. These difficulties cast doubt on the conclusion of the investigators that selenium induced tumor formation in these rats.

A statistically significant increase was reported in the incidence of all tumors and malignant tumors in rats administered 0.28–0.42 mg selenium/kg/day as sodium selenite or sodium selenate in drinking water for a lifetime (Schroeder and Mitchener 1971a). Not all autopsied animals were examined histologically, however, and high mortality in all groups occurred as a result of a virulent pneumonia epidemic that occurred during the study. In addition, the statistical analysis failed to account for the fact that the selenium-treated rats lived longer than did the control rats. Analysis of the incidence of tumors among animals with equal longevities indicates that the incidence of tumors in the selenate-treated rats was not significantly different from that in the controls.

A series of dietary studies assessed the effects of various dietary supplements on selenium tumor induction in male rats (Volgarev and Tscherkes 1967), but the conclusions that can be drawn from these experiments are limited since they did not include controls. Tumors (primarily liver) were found in 10/23 male rats administered sodium selenate in the diet at a dose of 0.34 mg selenium/kg/day for more than 18 months (Volgarev and Tscherkes 1967). The first tumors appeared after 18 months of selenium administration, by which time, 43% of the animals had already died (group started with 40 animals). Tumors were also found in 3/16 male rats administered sodium selenate in the diet at an initial dose of 0.34 mg selenium/kg/day for 6 months, followed by 0.68 mg/kg/day until the animals death (Volgarev and Tscherkes 1967). In a third group of experiments, no tumors were found in 200 male rats administered sodium selenate in the diet (0.34 mg selenium/kg/day) for 26 months. However, there was very high mortality among these rats, and survival time was 10 months shorter than among the similarly fed animals in the first experiment. The authors noted that an additional 200 male rats were maintained in their laboratory during these experiments and fed stock rations. The life spans of these animals exceeded those used in the experiments and no tumors were found at autopsy.

More recent animal bioassays have failed to demonstrate any association between excessive selenium exposure and carcinogenesis. Chen et al. (2000) reported a significant increase in rat esophageal adenocarcinogenesis in response to supplementation with 0.06 mg selenium/kg/day as sodium selenite for 40 weeks. However, selenium supplementation has generally been shown to significantly inhibit tumors induced by chemicals, viruses, or ultraviolet light (Birt et al. 1982; Finley et al. 2000; Ip 1981, 1983; Ip and Lisk 1995; Ip et al. 1996, 1997,1998, 2000a, 2000b; Jabobs 1983; Jacobs et al. 1977a, 1977b, 1979, 1981; Jiang et al. 1999; Medina and Shepherd 1981; Overvad et al. 1985; Schrauzer et al. 1976; Soullier et al. 1981; Thompson and Becci 1980; Woutersen et al. 1999). Results following administration of selenium as sodium selenate, sodium selenite, and organic forms of selenium are similar. Additional research reviewed in El-Bayoumy (1991, 1995, 1997) indicates that synthetic organoselenium compounds may be more potent cancer preventive agents than selenate, selenite, or the selenoamino acids.

Two sources reported the results of a study of rats administered sodium selenate or sodium selenite in the diet for a lifetime (Harr et al. 1967; Tinsley et al. 1967). A vehicle control and two positive control groups (administered a known hepatocarcinogen, *N*-2-fluorenyl-acetamide [FAA]) were included. Mortality was high in the highest dose group (0.8 mg selenium/kg/day), and therefore, selenium administration was discontinued. Longevity was reduced in animals fed 0.4 mg selenium/kg/day, but not in the animals administered lower doses. Of the original 1,437 experimental animals, 1,126 were necropsied. Half of the 88 FAA-fed rats developed neoplasms, half of which were hepatic carcinomas, indicating that the strain of rat and dietary conditions were compatible with the development of hepatic carcinogenesis. The incidence of cancer of all types in the necropsied control rats (11 out of 482, or 2.3%) was somewhat higher than the incidence of cancer in the selenium-treated animals that were necropsied (9 out of 553, or 1.6%). A statistical analysis of the data from this study was not reported. Although the reduced longevity of animals administered 0.4 mg selenium/kg/day might have prevented the observation of some late-developing cancers, the large number of rats necropsied, the end points examined, and the doses administered provide credible evidence of the lack of carcinogenic potential of sodium selenate or selenite.

Mice were fed tortula yeast diets containing up to 1.0 mg selenium/kg diet (equivalent to 0.13 mg selenium/kg body weight/day) as sodium selenite for 2 weeks prior to a single application of 0.125 mg 7,12-dimethylbenz[a]anthracene (DMBA) to the skin or repeated daily applications of 0.25 mL of a 0.03% solution of benzo[a]pyrene in acetone for 27 weeks (Shamberger 1970). The highest dose of selenite used, 0.13 mg selenium/kg/day, significantly decreased the number of tumors induced by both aromatic compounds. No significant increase was found in the incidence of spontaneous tumors in mice

following administration of 3 mg selenium/L in drinking water as either sodium selenite or sodium selenate for a lifetime (Schroeder and Mitchener 1972). This level corresponds to doses of 0.31–0.34 mg selenium/kg/day for the males and 0.42 mg selenium/kg/day for the females. The single dose administered, however, might not have been the maximal dose that could be tolerated. There were 7% more malignant tumors in the selenium-treated animals (13 out of 88 sectioned, or 15%) than in the controls (10 out of 119 sectioned, or 8%), but the difference was not statistically significant. The forms of selenium administered did not influence the incidence of tumors. In this study, only 88 out of 211 selenium-treated animals and 109 out of 209 control animals were examined histologically.

The only selenium compound that has been shown to be carcinogenic in animals is selenium sulfide (NTP 1980c), although there is some inconclusive evidence that ethyl selenac may also be carcinogenic (Innes et al. 1969; NCI 1968). A statistically significant increase in hepatomas (0/16 controls; 12/16 treated) was observed in male mice of one strain (C57BL/6 x C3H/Anf)F_1) receiving 2 mg selenium/kg as ethyl selenac, but not in male or female mice of another strain (C57BL/6 x AKR)F_1) receiving the same dose (Innes et al. 1969; NCI 1968).

Statistically significant increases in hepatocellular carcinomas and adenomas in rats and hepatic carcinomas and adenomas, as well as alveolar/bronchiolar carcinomas and adenomas, in female mice have been observed following chronic oral exposure to selenium sulfide (NTP 1980c). The incidence of hepatocellular carcinomas in rats was 1/50, 0/50, and 15/49 in males and 0/50, 0/50, and 21/50 in females at 0, 3, and 15 mg selenium sulfide/kg/day, respectively. In mice, the incidences of hepatocellular carcinomas and adenomas were 15/50, 14/50, and 23/50 in males, and 0/49, 2/50, and 25/49 in females at 0, 20, and 100 mg selenium sulfide/kg/day, respectively. Selenium sulfide is a pharmaceutical compound used in some antidandruff shampoos and is not administered orally. Because selenium sulfide is not absorbed through the skin, use of shampoos containing this compound should be safe, unless one intentionally consumes the product or has open cuts or sores on the scalp or hands. Chemically, selenium sulfide and ethyl selenac are very different from the organic and inorganic selenium compounds found in foods and in the environment.

In 1975, the International Agency for Research on Cancer (IARC) evaluated the literature relating selenium to carcinogenesis in both humans and animals. The Agency stated that the available data provided no suggestion that selenium is carcinogenic in humans (IARC 1975a), and IARC subsequently assigned selenium to Group 3: not classifiable as to its carcinogenicity to humans (IARC 1987). The forms of selenium considered included sodium selenate, sodium selenite, and the organic forms of

selenium contained in plant materials. Separate evaluations of ethyl selenac and methyl selenac assigned them to group 3, also (IARC 1975a, 1987). According to EPA, selenium is not classifiable as to its carcinogenicity in humans and is rated as Group D (IRIS 2001). The evidence for selenium sulfide, however, is sufficient to classify it as group B2 (probable human carcinogen)(IRIS 2001).

3.2.3 Dermal Exposure

3.2.3.1 Death

No studies were located concerning death in humans after dermal exposure to selenium or selenium compounds. In a range-finding study using mice dermally exposed to selenium sulfide for a maximum of 17 applications, 8 out of 20 animals died at 714 mg selenium sulfide/kg (NTP 1980a). However, the effects noted in this study were equivocal since there was no indication that the application sites were covered to prevent ingestion. Further, severe skin damage developed, and this may have led to direct systemic absorption of the compound.

3.2.3.2 Systemic Effects

No studies were located concerning respiratory, cardiovascular, gastrointestinal, hematological, musculoskeletal, hepatic, renal, endocrine, or body weight effects in humans or other animals following dermal exposure to selenium or selenium compounds.

Dermal Effects. Skin toxicity in humans, notably skin rashes, burns, and contact dermatitis, has been reported for both acute and chronic exposure to selenium fumes and acute exposures to selenium dioxide (Middleton 1947). No effects were detected in a study of eight women exposed daily for 2 weeks to an experimental sunscreen lotion containing up to 0.003 mg/kg/day selenium as L-selenomethionine (Burke et al. 1992a). A case report of a severe allergic skin response following intermediate exposure to sodium selenite (Senff et al. 1988) is discussed under immunological effects. Single topical exposures to selenious acid resulted in purpura, inflammation around hair follicles, and a pustular rash with some ulceration in exposed workers (Pringle 1942). However, these effects may have been due to the caustic effects of the acid. A single case report of hyperpigmentation and hair loss after use of a shampoo containing 1% selenium sulfide was located (Gillum 1996), but a study of the efficacy of an antidandruff shampoo containing 1% selenium sulfide found no adverse effects after 6 weeks of use by 150 individuals (Neumann et al. 1996).

Application of 100 μL of a lotion (oil-in-water emulsion) containing 0.02% selenium as selenomethionine three times a week to the shaved backs of mice for 39 weeks did not result in significant dermal effects (Burk et al. 1992b). Dermal effects were also not observed in hairless mice treated in the same manner for 49 weeks.

In mice, topical application of selenium sulfide resulted in erythema and skin irritation at 29 mg/kg, acanthosis at 143 mg/kg, and severe skin damage at 714 mg/kg (NTP 1980a).

Ocular Effects. No studies were located regarding ocular effects in humans after dermal exposure to selenium or selenium compounds. However, older reports on eye contact with selenium or selenium compounds indicate that acute exposure to selenium dioxide caused ocular and conjunctival irritation, and caused severe pain, lacrimation, blurred vision, and dulled corneas upon contact (Middleton 1947). Brief exposure to clouds of selenium fumes resulted in lacrimation, irritation, and redness of the eyes (Clinton 1947).

No studies were located regarding ocular effects in laboratory animals after dermal exposure to selenium or selenium compounds.

3.2.3.3 Immunological and Lymphoreticular Effects

A 1988 case report describes a female laboratory technician who developed severely pruritic vesicles between the fingers after 6 months of exposure to a medium containing selenium. After 2 years, the severity of the symptoms increased to include eczema on the face and neck, watering eyes, and two asthma attacks within a 2-month period. Sodium selenite or the medium containing selenium were the only positive patch tests (Senff et al. 1988).

No studies were located regarding immunological and lymphoreticular effects in laboratory animals after dermal exposure to selenium or selenium compounds.

No studies were located regarding the following health effects in humans or laboratory animals after dermal exposure to selenium or to selenium compounds:

3.2.3.4 Neurological Effects

3.2.3.5 Reproductive Effects

3.2.3.6 Developmental Effects

3.2.3.7 Cancer

No studies were located regarding carcinogenic effects in humans after dermal exposure to selenium or selenium compounds.

The results of most animal studies have not indicated that elemental selenium or selenium compounds are carcinogenic when topically applied to the skin of experimental animals (NTP 1980a, 1980b; Shamberger 1970). Several studies indicate that selenium compounds may protect against effects of known dermal carcinogens (polynuclear aromatic hydrocarbons [PAHs] and ultraviolet light). Shamberger (1970) reported that topical application of a solution containing 0.0005% sodium selenide significantly reduced the number of dermal papillomas induced by painting DMBA on the shaved backs of mice. More recently, Burke et al. (1992b) reported orally and topically administered L-selenomethionine decreased ultraviolet burns and skin cancer in albino (BALB:c) and hairless pigmented (Skh:2) female mice.

Only one study was found in which tumor development was reported after topical administration of selenium ointment (Tsuzuki et al. 1960). An unspecified number of mice were exposed to an unspecified amount of ointment containing 2.5, 5.0, 7.5, or 10% elemental selenium 6 days a week for an unspecified period of time by topical administration to hip skin. Tumors developed on the base of the necks of two female mice. Ingestion of the compound was possible because the animals may have licked the ointment. No other details were reported. The study is inconclusive because of the lack of controls, short duration, and inadequate description of the study protocol and results.

The National Toxicology Program (NTP 1980a) conducted a dermal application study of selenium monosulfide. The compound was applied to the skin of groups of 50 male and 50 female Swiss mice at 0, 0.5, or 1.0 mg selenium sulfide/mouse 3 days/week for 86 weeks. The application sites were not covered; therefore, ingestion of the test compounds was possible. The incidence of tumors in the treated groups did not differ significantly from that in the control group.

The NTP (1980b) also tested Selsun, a prescription dandruff shampoo containing 2.5% selenium sulfide (also a mixture of the mono- and disulfides), for carcinogenic properties. Groups of 50 male and

50 female Swiss mice were dermally exposed to a 0, 25, or 50% solution of Selsun in distilled water 3 days/week for 86 weeks. These doses were equivalent to 0, 0.31, or 0.625 mg selenium sulfide/mouse/day. The incidences of alveolar or bronchiolar adenomas or carcinomas in male mice were significantly increased over vehicle control values, but not over untreated control values. There was no significant effect in female mice. Some ingestion of the compound was possible, because the application sites were not covered. Also, the male mice may have been susceptible to another ingredient in the shampoo (the chemical composition of the shampoo was not reported), or the bioassay may have been too short due to decreased survival to detect a carcinogenic effect in females. Male mice that received dermal application of slightly larger doses of selenium sulfide (NTP 1980a) did not develop significantly more cancers than the controls.

3.2.4 Other Routes of Exposure

Endocrine Effects. Intraperitoneal injection of diabetic rats with sodium selenate has been reported to have insulin-like effects, producing a decrease in plasma glucose concentrations (McNeill et al. 1991). However, it is not clear that this is due to an effect of selenium on insulin metabolism, since food and water consumption were also decreased, and this is likely to have produced the decreased glucose levels.

Neurological Effects. Intraperitoneal injection of selenium (3.0 mg Se/kg as sodium selenite) into male Sprague-Dawley rats produced a significant increase (70%) in dopamine overflow from the striatum (as measured by an implanted dialysis probe) with a concomitant significant reduction in HVA levels (Rasekh et al. 1997). DOPAC levels were not changed. Direct infusion of 10 mM selenium into the striatum also produced a significant increase in dopamine overflow accompanied by slight, but significant decrease in HVA and DOPAC. Direct infusion of 10 mM selenium into the nucleus accumbens also produced a rapid and significant increase in dopamine overflow, but with no changes in DOPAC or HVA concentrations. The selenium induced changes in dopamine overflow were suppressed by the dopamine receptor agonist quinpirole.

3.3 GENOTOXICITY

Inorganic selenium compounds have been observed to have both genotoxic and antigenotoxic effects. The antigenotoxic effects generally occur at lower selenium exposure levels than the frank genotoxicity. This discussion will focus on genotoxic effects only. *In vitro* studies of the genotoxicity of selenium

compounds are summarized in Table 3-4, and *in vivo* studies of the genotoxicity of selenium compounds are summarized in Table 3-5.

Selenium dioxide was found to be mutagenic in both the Ames and the VITO-TOX *Salmonella typhimurium* tests of genotoxicity (van der Lelie et al. 1997).

In general, sodium selenite and sodium selenate have produced mixed results in bacterial mutagenicity test systems (Table 3-4). Sodium selenite induced base-pair substitution mutations using *S. typhimurium* and was also positive in the transformation assay using *Bacillus subtilis* (Kramer and Ames 1988; Nakamuro et al. 1976; Noda et al. 1979). However, negative results have also been reported for sodium selenite both in *S. typhimurium* and the rec assay using *B. subtilis* (Lofroth and Ames 1978; Noda et al. 1979). Sodium selenate, on the other hand, has tested positive in *S. typhimurium* (base-pair substitution) and in the rec assay using *B. subtilis* (Lofroth and Ames 1978; Noda et al. 1979), but has tested negative using the transformation assay in *B. subtilis* (Nakamuro et al. 1976).

Results with mammalian cell systems are also mixed, although sodium selenite is more consistently genotoxic in these systems. Sodium selenite has been observed to induce unscheduled deoxyribonucleic acid (DNA) synthesis (UDS), chromosomal aberrations, and sister chromatid exchange in cultured human fibroblasts (Lo et al. 1978; Ray et al. 1978; Whiting et al. 1980); UDS in Chinese hamster V79 cells (Sirianni and Huang 1983); and chromosomal aberrations in cultured Chinese hamster ovary cells (Whiting et al. 1980). However, sodium selenate induced chromosomal aberrations in Chinese hamster ovary cells (Whiting et al. 1980) and UDS in Chinese hamster V79 cells (Sirianni and Huang 1983), but did not induce chromosomal aberrations in human leukocytes or cultured human fibroblasts (Lo et al. 1978; Nakamuro et al. 1976). A comparison of cytotoxicity and induction of tetraploidy in Chinese hamster V79 cells induced by sodium selenite or its major excretory product trimethylselenonium found that sodium selenite was about 1,000 times more cytotoxic than trimethylselenonium, but that neither compound produced a significant change in mitotic index (Ueda et al. 1997).

The addition of glutathione to test mixtures enhances the genotoxicity of sodium selenite, sodium selenate, and sodium selenide in bacterial test systems, indicating that production of a reactive species mutagenic for bacteria occurs via a reductive mechanism following concomitant exposure to these compounds (Whiting et al. 1980). This finding is supported by results in mammalian test systems. For example, in cultured human leukocytes, sodium selenite induces chromosome aberrations and sister chromatid exchanges (Nakamuro et al. 1976; Ray and Altenburg 1978; Ray et al. 1978). Sister chromatid

Table 3-4. Genotoxicity of Selenium *In Vitro*

Species (test system)	End point	Result: With activation	Result: Without activation	Reference
Prokaryotic organisms:	Mutation			
Salmonella typhimurim	(Na_2SeO_3)		–	Lofrothand Ames 1978
	(Na_2SeO_4)	NT	+	
S. typhimurim	(SeO_2)	NT	+	van der Lelie et al. 1997
S. typhimurim TA100	(Na_2SeO_3)	NT	+	Noda et al. 1979
S. typhimurium TA98, TA1537		NT	–	
S. typhimurium TA100	(Na_2SeO_4)	NT	+	
S. typhimurium TA98, TA1537		NT	–	
Bacillus subtilis rec assay	(Na_2SeO_3)	NT	–	Noda et al. 1979
B. subtilis rec assay	(Na_2SeO_4)	NT	+	Kanematsu et al. 1980
	(SeO_2)	NT	+	
B. subtilis transformation	(SeO_2)	NT	+	Nakamuro et al. 1976
	(Na_2SeO_3)	NT	+	
	(Na_2SeO_4)	NT	–	
Eukaryotic organisms:				
Mammalian cells				
	Chromosomal aberrations			
Chinese hamster ovary	(Na_2SeO_3)	NT	+	Whiting et al. 1980
	(Na_2SeO_4)	NT	+	
Human leukocytes	(SeO_2)	NT	+	Nakamuro et al. 1976
	(Na_2SeO_3)	NT	+	
	(Na_2SeO_4)	NT	–	
Human lymphocytes	(Na_2SeO_4)	NT	+	Biswas 1997
Human lymphocytes	(Na_2SeO_3)	NT	+	Biswas et al. 2000
	(Na_2SeO_4)	NT	+	

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Table 3-4. Genotoxicity of Selenium *In Vitro (continued)*

Species (test system)	End point	Result: With activation	Result: Without activation	Reference
Human lymphocytes	(Na_2SeO_3)	NT	+	Khalil 1989
	(Selenomethionine)	NT	+	
Cultured human fibroblasts	(Na_2SeO_3)	+	+	Lo et al. 1978
	(Na_2SeO_4)	–	–	
	Tetraploidy			
Chinese hamster V79 cells	(Na_2SeO_3)	NT	+	Ueda et al. 1997
	(Trimethylselenonium)	NT	+	
	DNA strand breaks			
Mouse mammary carcinoma cells	Na_2SeO_3)	NT	+	Lu et al. 1995b
	(Na_2SeO_4)	NT	+	
	(Methylselenocyanate)	NT	–	
	(Se-methylselenocysteine)	NT	–	
	Unscheduled DNA synthesis			
Cultured human fibroblasts	(Na_2Se)	NT	+	Whiting et al. 1980
	(Na_2SeO_3)	NT	+	
	(Na_2SeO_4)	NT	+	
	Sister chromatid exchange			
Cultured human fibroblasts	(Na_2SeO_3)	NT	+	Ray et al. 1978
Human lymphocytes	(Na_2SeO_3)	NT	+	Khalil 1989
	(Selenomethionine)	NT	+	
	(Selenocystine)	NT	+	Khalil 1994

– = negative result; + = positive result; DNA = deoxyribonucleic acid; NT = not tested; (Na_2Se) = sodium selenide; (Na_2SeO_3) = sodium selenite; (Na_2SeO_4) = sodium selenate; (SeO_2) = selenium dioxide

Table 3-5. Genotoxicity of Selenium *In Vivo*

Species (test system)	End point	Results	Reference
Human lymphocytes (Na_2SeO_3)	Chromosomal aberrations, sister chromatic exchanges	–	Norppa et al. 1980a
Monkey (*Macaca fisciculari*) bone marrow (L-selenomethionine)	Micronuclei	+ (adult toxic dose) – (fetal at maternally toxic doses)	Choy et al. 1989 Choy et al. 1993
Mouse bone marrow (Na_2SeO_3) (Na_2SeO_4)	Chromosome breaks and spindle disturbances	 + +	Biswas et al. 1997
Mouse bone marrow (Na_2SeO_3) (Na_2SeO_4)	Chromosome breaks and spindle disturbances	 + +	Biswas et al. 1999a
Mouse bone marrow (H_2SeO_3) (Na_2SeO_4)	Micronucleus induction	 + +	Itoh and Shimada 1996
Mouse bone marrow (H_2SeO_3)	Micronucleus induction	+	Rusov et al. 1996
Rat bone marrow (Na_2SeO_3)	Chromosomal aberrations	+	Newton and Lilly 1986
Rat bone marrow (SeS)	Chromosomal aberrations	–	Moore et al. 1996b
Rat bone marrow (SeS)	Micronucleus induction	+	Moore et al. 1996b
Rat spleen (SeS)	Chromosomal aberrations	–	Moore et al. 1996b

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Table 3-5. Genotoxicity of Selenium *In Vivo (continued)*

Species (test system)	End point	Results	Reference
Rat spleen (SeS)	Micronucleus induction	–	Moore et al. 1996b
Rat lymphocytes (Na_2SeO_3)	Chromosomal aberrations	–	Newton and Lilly 1986

+ = positive result; – = negative result

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exchange was not observed at similar sodium selenite concentrations in a human lymphoblastoid cell line; however, exchanges were observed when these same cells were incubated with sodium selenite and red blood cell lysate (Ray and Altenburg 1978). The observation that internal constituents of red blood cells may contribute to the genotoxicity of sodium selenite supports the suggestion that metabolism is involved in the production of an active species following exposure to sodium selenite in these test systems. The active species responsible for the genotoxic effects is not known.

At high concentrations, sodium selenite induces unscheduled DNA synthesis and chromosome aberrations in cultured human fibroblasts (Lo et al. 1978). The addition of a metabolic activator (S9 fraction) or glutathione increased both the number of aberrations and the toxicity of sodium selenite (Whiting et al. 1980) and sodium selenate (Lo et al. 1978; Whiting et al. 1980).

Sodium selenite, sodium selenide, methylselenocyanate, and Se-methylselenocysteine were all found to be cytotoxic to cells of a mouse mammary carcinoma line; however, only sodium selenite and sodium selenide induced DNA strand breaks (Lu et al. 1995b).

Selenomethionine (Khalil 1989) and selenocystine (Khalil 1994) have tested positive for sister chromatid exchanges in cultured human lymphocytes. Selenomethionine, sodium selenite, and sodium selenate tested positive for chromosomal aberrations in cultured human lymphocytes (Biswas 1997; Biswas et al. 2000; Khalil 1989). Sodium selenite was considerably more clastogenic than sodium selenate (Biswas et al. 2000).

The genotoxicity of selenium monosulfide was assessed in an *in vivo*/*in vitro* micronucleus and chromosome aberration assay in rats (Moore et al. 1996b). Male Wistar rats (4/dose) were administered 25, 50, or 100 mg/kg selenium monosulfide in corn oil. Negative control rats received corn oil by gavage and positive controls were injected intraperitoneally with 20 mg/kg cyclophosphamide. Animals were sacrificed 24 hours after treatment and the femur marrow and spleen were removed and cultured. Spleen and marrow cultures were examined 24 or 48 hours after establishment, respectively. No increase in chromosome aberrations or micronucleus formation in cells from treated rats was observed.

Results of *in vivo* genotoxicity tests have been both negative and positive (Table 3-5). Chromosomal aberrations and sister chromatid exchanges in lymphocytes were not increased in nine neuronal ceroid lipofuscinosis patients treated with intramuscular sodium selenite injections or tablets (0.005–0.05 mg selenium/kg/day) for 1–13.5 months, or in five healthy persons given selenite (0.025 mg/kg/day) for

2 weeks (Norppa et al. 1980a). Among the treated patients, there was no distinction between route of exposure.

Compared to untreated controls, a significant increase in the number of micronuclei was observed in bone marrow cells of macaques treated by nasogastric intubation with L-selenomethionine at a dose of 0.24 mg selenium/kg/day for 15 days (Choy et al. 1989). No effect on the number of micronuclei was observed in macaques treated with L-selenomethionine at a dose of 0.12 mg selenium/kg/day for 19 days. A significant increase in the number of micronuclei in bone marrow cells was not observed in the offspring of macaques treated by nasogastric intubation with L-selenomethionine at a dose of 0.12 mg selenium/kg/day on gestation days 20–50 (Choy et al. 1993). The doses of L-selenomethionine used in these studies produced obvious signs of toxicity (loss of body weight, poor appetite, constipation, depression, weakness) in the macaques.

Chromosomal aberrations were not increased in the lymphocytes of rats given two intravenous doses of sodium selenite at 2.3–2.7 mg selenium/kg (Newton and Lilly 1986). Chromosomal aberrations in bone marrow cells were significantly increased in these rats, but the total dose of selenium was near the intravenous LD_{50} for selenite which has been reported as 5.7 mg selenium/kg in rats (Olson 1986).

Bone marrow cells of male mice gavaged with sodium selenate or sodium selenite showed a significant increase in chromosome breaks and spindle disturbances compared with untreated controls (Biswas et al. 1997, 1999a). The number of chromosomal aberrations increased with dose and was slightly greater with sodium selenite than with sodium selenate. A significant increase in micronucleus formation was observed in bone marrow cells of male mice intraperitoneally injected with selenous acid and in female mice intramuscularly injected with sodium selenite, but not in male mice intraperitoneally injected with sodium selenate (Itoh and Shimada 1996; Rusov et al. 1996).

Selenium appears to affect the ability of liver enzymes to activate some chemical mutagens. Studies in animals exposed orally to sodium selenite in the diet at doses between 0.05 and 0.125 mg selenium/kg/day indicate that selenium may inhibit the mutagenic effect of other chemical agents (Gairola and Chow 1982; Schillaci et al. 1982). In these studies, *S. typhimurium* was used to assess the mutagenicity of DMBA, benzo[a]pyrene (BAP), and 2-aminoanthracene (2AA) in the presence of liver microsomal enzymes from rats fed either a basal diet (0.02–0.15 mg selenium/kg diet or 0.001–0.0075 mg selenium/kg/day) or a sodium selenate-supplemented diet (basal diet plus 1–5 mg selenium/kg/diet or 0.05–0.25 mg selenium/kg/day) for 3–20 weeks. DMBA and 2AA were found to be less mutagenic in the

presence of liver microsomal enzymes taken from rats fed the selenium-supplemented diets than in the presence of microsomal enzymes taken from rats fed the basal diet; BAP mutagenicity was not changed.

The genotoxicity of selenium monosulfide was assessed in an *in vivo* micronucleus and chromosome aberration assays in rats (Moore et al. 1996b). Male Wistar rats (5/dose/timepoint) were administered 12.5, 25, or 50 mg/kg selenium monosulfide in corn oil. Negative control rats received corn oil by gavage and positive controls were injected intraperitoneally with 20 mg/kg cyclophosphamide. Animals were sacrificed 24, 36, or 48 hours after treatment and the femur marrow and spleen cells were examined. A small, but significant increase in micronucleated bone marrow cells was observed 24 hours after treatment with 50 mg/kg selenium monosulfide and 48 hours after treatment with 12.5 mg/kg selenium monosulfide. Selenium monosulfide was cytotoxic at the 50 mg/kg dose after 24 hours. No increase in micronucleus formation was observed in the spleen. No increase in chromosome aberrations was observed in the bone marrow or spleen.

3.4 TOXICOKINETICS

Occupational studies indicate that humans absorb elemental selenium dusts and other selenium compounds, but quantitative inhalation toxicokinetic studies in humans have not been done. Studies in dogs and rats indicate that following inhalation exposure, the rate and extent of absorption vary with the chemical form of selenium. Studies in humans and experimental animals indicate that, when ingested, several selenium compounds including selenite, selenate, and selenomethionine are readily absorbed, often to greater than 80% of the administered dose. Although a study of humans did not detect evidence of dermal absorption of selenomethionine, one study of mice indicates selenomethionine can be absorbed dermally. There is little or no information available on the absorption of selenium sulfides, but selenium disulfides are not believed to be absorbed through intact skin.

Selenium accumulates in many organ systems in the body; in general, the highest concentrations are found in the liver and kidney (Table 3-6). Selenium concentrations in tissues do not seem to be correlated with effects. Tissue concentrations were highest in pigs fed D,L-selenomethionine, while a similar dose of selenium (form not stated) given as *A. bisulcatus* was a more potent neurotoxin. Blood, hair, and nails also contain selenium, and selenium has been found in human milk (Table 3-7). In addition, selenium is subject to placental transfer.

Table 3-6. Selenium Concentrations in Human Tissues[a,b]

Selenium concentration				
Mean	SD	Range	Country	Reference
Fetal tissues				
Liver (μg selenium/g)				
2.8	0.2		United States	Robkin et al. 1973[c]
Blood (mg selenium/L)				
0.12	0.008		United States	Hadjimarkos et al. 1959
1.04	0.28		United States	Baglan et al. 1974[c]
0.070	0.017		New Zealand	Thompson and Robinson 1980
0.061	0.014		Scandanavia	Korpela et al. 1984
Erythrocytes (mg selenium/L)				
0.39	0.08		United States	Rudolph and Wong 1978
0.149			Scandanavia	Haga and Lunde 1978
0.104	0.026		New Zealand	Thompson and Robinson 1980
Plasma (mg selenium/L)				
0.13	0.03		United States	Rudolph and Wong 1978
0.033	0.008		New Zealand	Thompson and Robinson 1980
Serum (mg selenium/L)				
0.052			Scandanavia	Haga and Lunde 1978
Adult/Infant tissues				
Adrenal gland (μg selenium/g)				
0.46	0.03		United States	Blotcky et al. 1979
0.21 (infant)			Canada	Dickson and Tomlinson 1967
0.36 (adult)				
Brain (μg selenium/g)				
		0.114–0.171	Denmark	Larsen et al. 1979
0.110	0.021		Germany	Oster et al. 1988c
0.16 (infant)			Canada	Dickson and Tomlinson 1967
0.27 (adult)				
		0.115–0.222	Japan	Ejima et al. 1996
Fat (μg selenium/g)				
0.09 (infant)			Canada	Dickson and Tomlinson 1967
0.12 (adult)				

Table 3-6. Selenium Concentrations in Human Tissues[a,b] (*continued)*

Selenium concentration				
Mean	SD	Range	Country	Reference
Adult/Infant tissues (*cont.*)				
Gonad (μg selenium/g)				
0.46 (infant)			Canada	Dickson and Tomlinson 1967
0.47 (adult)				
Heart (μg selenium/g)				
0.33	0.13		United States	Blotcky et al. 1979
0.170	0.032		Germany	Oster et al. 1988c
0.155	0.030 (LV)			
0.55 (infant)			Canada	Dickson and Tomlinson 1967
0.22 (adult)				
Intestine (μg selenium/g)				
0.31 (infant)			Canada	Dickson and Tomlinson 1967
0.22 (adult)				
Kidney (μg selenium/g)				
0.89	0.11		United State	Blotcky et al. 1979
0.771	0.169		Germany	Oster et al. 1988c
0..92 (infant)			Canada	Dickson and Tomlinson 1967
0.63 (adult)				
0.78	0.19	0.36–1.29	Sweden	Muramatsu and Parr 1988
Liver (μg selenium/g)				
0.62	0.04		United States	Blotcky et al. 1979
0.50	0.08	0.35–0.65	United States	Zeisler et al. 1984
1.73	0.24		United States	McConnell et al. 1975[c]
		0.27–0.51	Denmark	Larsen et al. 1979
0.291	0.078		Germany	Oster et al. 1988c
0.995	0.308		Finland	Alfthan et al. 1991[c]
0.45	0.11		Finland	Aaseth et al. 1990
0.06			Bulgaria	Damyanova 1983
0.33	0.12	0.082–0.64	Sweden	Muramatsu and Parr 1988
0.19	0.05	0.10–0.27	New Zealand	Casey et al. 1983
0.34 (infant)			Canada	Dickson and Tomlinson 1967
0.39 (adult)				

Table 3-6. Selenium Concentrations in Human Tissues[a,b] (*continued)*

Selenium concentration				
Mean	SD	Range	Country	Reference
Adult/Infant tissues (*cont.*)				
Lung (μg selenium/g)				
0.30	0.02		United States	Blotcky et al. 1979
0.132	0.033		Germany	Oster et al. 1988c
0.17 (infant)			Canada	Dickson and Tomlinson 1967
0.21 (adult)				
Pancreas (μg selenium/g)				
0.55	0.13		United States	Blotcky et al. 1979
0.63	0.07		United States	McConnell et al. 1975[c]
0.05 (infant)			Canada	Dickson and Tomlinson 1967
0.13 (adult)				
Prostate (μg selenium/g)				
0.26	0.02		United States	Blotcky et al. 1979
0.150	0.035		Germany	Oster et al. 1988c
Skeletal muscle (μg selenium/g)				
0.40	0.20		United States	Blotcky et al. 1979
0.111	0.017		Germany	Oster et al. 1988c
0.31 (infant)			Canada	Dickson and Tomlinson 1967
0.40 (adult)				
		0.13–0.21	Denmark	Larsen et al. 1979
Skin (μg selenium/g)				
0.24	0.02		United States	Blotcky et al. 1979
Spleen (μg selenium/g)				
0.37	0.03		United States	Blotcky et al. 1979
0.226	0.044		Germany	Oster et al. 1988c
0.37 (infant)			Canada	Dickson and Tomlinson 1967
0.27 (adult)				
Stomach (μg selenium/g)				
0.19 (infant)			Canada	Dickson and Tomlinson 1967
0.17 (adult)				

Table 3-6. Selenium Concentrations in Human Tissues[a,b] (*continued)*

Selenium concentration				
Mean	SD	Range	Country	Reference
Adult/Infant tissues (*cont.*)				
Testis (μg selenium/g)				
0.28	0.03		United States	Blotcky et al. 1979
0.274	0.048		Germany	Oster et al. 1988c
Thyroid (μg selenium/g)				
1.02	0.20		United States	Blotcky et al. 1979
0.72	0.44	0.15–1.90	Finland	Aaseth et al. 1990
0.64 (infant) 1.24 (adult)			Canada	Dickson and Tomlinson 1967

[a]General population measures unless otherwise noted
[b]Selenium concentrations in adult blood and blood components, urine, hair, nails, milk, placenta, and semen are found in Table 3-7
[c]Dry weight

LV = left ventricle; RV = right ventricle; SD = standard deviation

Table 3-7. Biomarkers: Selenium Concentrations in Human Tissues and Fluids[a]

Selenium concentration				
Mean	SD	Range	Country	Reference
Whole blood (mg selenium/L)				
0.132	0.029	0.08–0.13	United States	Corden et al. 1989
0.206		0.10–0.30	United States	Allaway et al. 1968
0.109	0.015		United States	Dworkin et al. 1986
0.157		0.103–0.191	United States	Shamberger 1983
0.182	0.037		Canada	Dickson and Tomlinson 1967
0.095	0.009		China	Zhu 1981
0.095	0.091		China	Yang et al. 1983
0.164	0.032		Greece	Bratakos et al. 1990
0.108	0.006	0.076–0.140	Italy	Minoia et al. 1990
		0.079–0.103	Finland	Jaakkola et al. 1983
		0.080–0.089	Finland	Kumpusalo et al. 1990[b]
		0.077–0.089	Finland	Kumpusalo et al. 1990[c]
0.069	0.018		New Zealand	Thomson and Robinson 1980[d]
0.059	0.012		New Zealand	Rea et al. 1979
0.092	0.001	0.06–0.013	Germany	Oster et al. 1988b
Erythrocyte (mg selenium/L)				
0.174		0.11–0.28	United States	Meyer and Verreault 1987
0.13	0.02		United States	Dworkin et al. 1986
0.52	0.05		United States	Rudolph and Wong 1978[d]
0.131	0.002	0.060–0.210	Germany	Oster et al. 1988b
		0.057–0.087	New Zealand	Watkinson 1981
0.074	0.016		New Zealand	Rea et al. 1979
0.103	0.030		New Zealand	Thomson and Robinson 1980
Plasma (mg selenium/L)				
0.155		0.081–0.225	United States	Clark et al. 1984
0.095	0.016		United States	Dworkin et al. 1986
0.21	0.03		United States	Rudolph and Wong 1978[d]
0.148			United States	Coates et al. 1988
0.081	0.016		Canada	Dickson and Tomlinson 1967
0.153	0.021		Japan	Hojo 1987
0.089	0.014		Netherlands	van't Veer et al. 1990
0.081	0.001	0.056–0.105	Italy	Minoia et al. 1990
0.118	0.027	0.064–0.173	Italy	Sesana et al. 1992
0.048	0.010		New Zealand	Rea et al. 1979
0.041	0.011		New Zealand	Thomson and Robinson 1980[d]

Table 3-7. Biomarkers: Selenium Concentrations in Human Tissues and Fluids (*continued*)

Selenium concentration				
Mean	SD	Range	Country	Reference
Serum (mg selenium/L)				
0.136	0.002		United States	Willett et al. 1983
0.110	0.016		United States	Menkes et al. 1986
0.162	1.48		United States	Coates et al. 1988
0.07			United States	McConnell et al. 1975[e]
0.198	0.055	0.123–0.363	United States	Longnecker et al. 1991
0.143	0.016		Canada	Lalonde et al. 1982[f]
0.081	0.001	0.033–0.121	Italy	Minoia et al. 1990
		0.087–0.093	Italy	Morisi et al. 1989
0.118	0.016	0.087–0.308	Finland	Luoma et al. 1992
0.055	0.001		Finland	Virtamo et al. 1987
0.073	0.015		South Africa	Heese et al. 1988[d]
0.207		0.07–0.81	Venezuela	Brätter et al. 1991a
		0.229–0.621	Venezuela	Brätter and Negretti De Brätter 1996
Urine (mg selenium/L)				
0.034	0.024		England	Glover 1970
0.058	0.026	0.020–0.113	Japan	Hojo 1981a
0.026	0.012		China	Yang et al. 1983
0.024	0.002		Greece	Bratakos et al. 1990
0.022	0.002	0.002–0.031	Italy	Minoia et al. 1990
Hair (µg selenium/g)				
0.64	0.02		United States	Thimaya and Ganapathy 1982
0.359	0.004		China	Zhu 1981
0.36	0.17		China	Yang et al. 1983
0.42	0.88		Greece	Bratakos et al. 1990
0.42	0.10	0.21–0.63	Sweden	Muramatsu and Parr 1988[e]
3.40	2.0	0.95–9.6 (female)	Japan	Imahori et al. 1979[e]
3.70	2.3	0.06–14.2 (male)		
1.02	1.04	maternal	England	Razagui and Haswell 1997
0.63	0.52	neonatal		
0.54	0.34	maternal	Spain	Bermejo Barrera et al. 2000
0.77	0.24	child		

Table 3-7. Biomarkers: Selenium Concentrations in Human Tissues and Fluids (*continued*)

Selenium concentration				
Mean	SD	Range	Country	Reference
Nails (μg selenium/g)				
1.56	0.58	0.083–3.82	United States	Longnecker et al. 1991
0.82	0.174		United States	Hunter et al. 1990a
0.63	0.12		Netherlands	van't Veer et al. 1990
0.54	0.91		Greece	Bratakos et al. 1990
0.78		0.085–2.75	Netherlands	Van Noord et al. 1992
Milk (μg selenium/mL)				
		0.208–0.256	Africa	Funk et al. 1990[b]
0.018	0.004	0.007–0.033	United States	Shrearer and Hadjimarkos 1975
0.021		0.013–0.053	United States	Hadjimarkos 1963
0.016	0.005			Smith et al. 1990
0.026			United States	Ellis et al. 1990
0.062	0.055	0.015–0.214	Chile	Cortez 1984
0.010	0.002	0.006–0.013	Finland	Kumpulainen 1983
0.011			Austria	Li et al. 1999
0.012			Germany	Michalke and Schramel 1998
		0.025–0.250	Venezuela	Brätter et al. 1991a
		0.043–0.112	Venezuela	Brätter and Negretti De Brätter 1996
Placenta (mg selenium/L)				
1.70	0.61		United States	Baglan et al. 1974[e]
0.193	0.016		United States	Korpela et al. 1984
0.18	0.007		United States	Hadjimarkos et al. 1959
Semen (μg selenium/g)				
0.063	0.020	0.016–0.131	Singapore	Roy et al. 1990
1.80	0.11		Finland	Suistomaa et al. 1987[e]

[a]General population measures unless otherwise noted
[b]Range of mean concentration
[c]Range of mean concentrations for multivitamin users
[d]Only women were sampled
[e]Dry weight
[f]Only men were sampled

SD = standard deviation

As a component of glutathione peroxidase and the iodothyronine 5'-deiodinases, selenium is an essential micronutrient for humans. Its role in the deiodinase enzymes may be one reason that children require more selenium for growth than adults. Selenium is also a component of the enzyme thioredoxin reductase, which catalyses the NADPH-dependent reduction of the redox protein thioredoxin. Other selenium-containing proteins of unknown functions, including selenoprotein P found in the plasma, have also been identified. Excess selenium administered as selenite and selenate has been shown to be metabolized to methylated compounds and excreted.

Selenium is primarily eliminated in the urine and feces in both humans and laboratory animals. The distribution of selenium between the two routes seems to vary with the level of exposure and time after exposure. The form of selenium excreted is dependent on the form of selenium that was ingested. In cases of acute exposure to toxic concentrations of selenium or selenium compounds, significant amounts of selenium can be eliminated in the breath, causing the characteristic "garlic breath."

3.4.1 Absorption

3.4.1.1 Inhalation Exposure

Studies regarding the absorption of selenium in humans following inhalation exposure are limited to occupational studies. Glover (1970) examined urinary selenium levels of workers employed in a selenium rectifier plant. Workers exposed to higher levels of unspecified inorganic selenium compounds in the air excreted higher levels of selenium in their urine than workers in other areas of the plant with lower concentrations of selenium in the air. Although the study indicates that selenium was absorbed from the lungs of the workers, the nonspecific exposure levels and lack of compound identification precluded an estimate of the extent and rate of absorption from the lungs. Significantly increased serum selenium levels were reported for workers at a rubber tire repair shop in Toluca City, Mexico compared with a group of unexposed individuals from the same city (Sánchez-Ocampo et al. 1996). The workers in this study were exposed to selenium (no levels reported) from vulcanized rubber, both as dust in the air and from handling the tires; thus, it is not possible to attribute absorption to a single route.

Studies using dogs and rats indicate that absorption of selenium following inhalation exposure is extensive, although the rate of absorption depends on the chemical form of selenium. In rats (Medinsky et al. 1981a) and dogs (Weissman et al. 1983), the absorption of selenium following inhalation exposure to selenious acid aerosol is approximately twice as rapid as the absorption of selenium following

inhalation exposure to elemental selenium aerosol. However, Medinsky et al. (1981a) found that with either form after 4 days most of the selenium was absorbed following inhalation exposure and that the distribution of selenium in the body tissues was identical, suggesting that selenium entered the same body pool following pulmonary uptake (Medinsky et al. 1981a).

3.4.1.2 Oral Exposure

Selenium compounds are readily absorbed from the human gastrointestinal tract. The bioavailability of ingested selenium can be affected by the physical state of the compound (e.g., solid or solution), the chemical form of selenium (e.g., organic, inorganic), and the dosing regimen. However, in general, it appears that the degree of selenium absorption (i.e., percent of administered dose absorbed) in humans is independent of the exposure level, but that in some cases, absorption is greater when selenium deficiency exists.

In humans, absorption of sodium selenite or selenomethionine can exceed 80% for both small and relatively large doses (Griffiths et al. 1976; Thomson 1974; Thomson and Stewart 1974; Thomson et al. 1977). A total of 90–95% of a small amount of sodium selenite (0.010 mg selenium/person) administered in aqueous solution was absorbed (Thomson 1974). Absorption of a large dose (1.0 mg/person) of either sodium selenite or selenomethionine was 90–95 and 97% of the administered dose, respectively (Thomson et al. 1977). These data indicate a lack of homeostatic control over the dose range tested. However, Martin et al. (1989a) reported increased absorption of selenium from sodium selenite in an aqueous solution by healthy male volunteers kept on a selenium-deficient diet, indicating that a lower-bound homeostatic control on absorption exists. Griffiths et al. (1976) reported 96–97% absorption of a single dose of 0.002 mg selenium administered as selenomethionine in solution. Similarly, Thomson et al. (1977) reported 97% absorption of a single large dose of 1.0 mg selenium administered as selenomethionine in solution to one subject. The subjects in these studies were New Zealand women.

Other studies have indicated that humans might absorb selenomethionine more efficiently than sodium selenite (Moser-Veillon et al. 1992; Swanson et al. 1991). Young et al. (1982) studied human absorption of dietary selenium in young men in the United States. The men ate either ^{75}Se-labeled chicken alone (0.013 mg selenium/person) or the chicken plus supplemental labeled sodium selenite (0.071 mg selenium/person in a solution mixed with the meal). Eighty percent of the selenium in the chicken meat was absorbed, but less than 30% of the selenium administered as sodium selenite was absorbed. Similarly, Robinson et al. (1978) found that 75% of selenomethionine, but only 46% of selenite, was

absorbed during a 10–11-week administration of solutions providing 0.0013–0.0023 mg selenium/kg/day to New Zealand women. It is not clear why the estimated absorption of sodium selenite varied between 46 and 30% in these trials.

Experimental animals also efficiently absorb selenium compounds from the gut independent of the level of selenium exposure. Several studies have reported absorption of 80–100% in rats given dietary selenium administered as sodium selenite, sodium selenate, selenomethionine, or selenocystine (Furchner et al. 1975; Thomson and Stewart 1973). Other animal species also readily absorb orally administered selenium compounds. Furchner et al. (1975) estimated that over 90% of an oral dose of selenious acid was absorbed in mice and dogs, although monkeys absorbed less of the administered dose (amount unspecified). Using an *in vivo* perfusion method in which selenite was added directly to the duodenal end of the small intestine, the absorption of selenite was found to be linear (slope=0.0386) over the concentration range of 1–200 μM (Chen et al. 1993).

In one study of rats, absorption of selenite or selenomethionine into the blood stream following oral exposure occurred primarily in the duodenum and, to a lesser extent, in the jejunum and the ileum (Whanger et al. 1976). Compared to the small intestine, little selenium was absorbed from the stomach (Whanger et al. 1976), and it was not determined whether absorption occurred in the large intestine. In an *in vitro* study using everted intestinal sacs from hamsters, Spencer and Blau (1962) found that selenomethionine was transported against a concentration gradient with the same characteristics as methionine. Selenomethionine was not found to be degraded during transport. This study suggests that in the intestines, methionine and selenomethionine share the same transport mechanism.

A comparison of absorption of selenium by selenium-depleted rats after oral administration of sodium selenate, selenomethionine, or methyl selenocysteine (from high-selenium broccoli) found that gross absorption of selenium from methyl selenocysteine was significantly lower (85%) than from sodium selenate or selenomethionine (91%), further, true selenium absorption adjusted for urinary excretion was significantly different for methyl selenocysteine, sodium selenate, and selenomethionine, with the lowest absorption for methyl selenocysteine and the highest for selenomethionine (Finley 1998). Absorption of selenium from selenomethionine was not significantly lower than from sodium selenate.

In vivo experiments with ligated rat intestines have shown that there is significantly higher absorption and transfer to the body of selenium as selenocystine or selenodiglutathione than selenium as selenite from ligated loops of ileum, but that absorption of the three forms of selenium in the jejunum was

approximately similar (Vendeland et al. 1992). *In vitro* experiments with brush border membrane vesicles derived from rat intestines have shown dramatic differences in the uptake and binding of selenium depending on the form in which it is presented, with absorption of organic forms being much more efficient than absorption from selenite or selenate (Vendeland et al. 1992, 1994). Selenium from selenocystine or selenodiglutathione was absorbed 10 times more quickly than selenium from sodium selenite (Vendeland et al. 1992). Similarly, selenium was much more efficiently absorbed from selenomethionine than from selenite or selenate (Vendeland et al. 1994). Binding also varied between selenomethionine, selenite, and selenate, with selenite binding exceeding that of selenate by 37-fold and selenomethionine exceeding selenite by 14-fold (Vendeland et al. 1994).

3.4.1.3 Dermal Exposure

Dermal absorption was tested in eight women at a maximum dose of 0.0029 mg selenium/kg as selenomethionine (0.05% L-selenomethionine in a lotion). No detectable increase in serum selenium concentrations was observed; but because the concentrations tested were so low, absorption cannot be ruled out (Burke et al. 1992a). Absorption of selenium disulfide was examined using a monthly 24-hour urine specimen in 16 persons who washed their hair weekly with a 1% selenium disulfide shampoo. No differences were found from control urinary selenium levels over the 1-year exposure period (Cummins and Kimura 1971). No absorption of selenium from selenium sulfide was seen in 15 persons who applied a 2.5% selenium sulfide suspension to their torsos and allowed it to remain on the body overnight (Kalivas 1993).

Mice were treated with a maximum of 0.02% selenium as selenomethionine by topical application of a lotion three times per week for 39 weeks to the shaved back and ears (size of area not specified). The applied dose was 0.29 mg/kg/day. Controls received the lotion without selenium. Dermal effects were not observed in the selenomethionine-treated mice. However, treated animals had significantly higher concentrations of selenium than the controls in the liver and ventral skin away from the application site (Burke et al. 1992b). These data suggest that mice can absorb topically applied selenomethionine, but since the areas were not occluded, some oral absorption during grooming is also possible.

3.4.1.4 Other Routes of Exposure

3.4.2 Distribution

Most studies report similar distribution patterns for both organic and inorganic selenium compounds tested. Normal levels of selenium found in various human tissues are shown in Table 3-6. Selenium concentrations in human fluids and tissues that are easily collected (e.g., placenta) are provided in Section 3.8.2, Biomarkers Used to Identify or Quantify Exposure to Selenium. Selenium from sodium selenite and sodium selenate is found at the highest concentrations in the liver and kidney of humans and other animals following oral administration or intravenous or subcutaneous injection (Cavalieri et al. 1966; Heinrich and Kelsey 1955; Jereb et al. 1975; Thomson and Stewart 1973). Similarly, monkeys receiving high doses of L-selenomethionine orally for up to 30 days accumulated the highest concentrations of selenium in the liver and kidneys (Willhite et al. 1992). Selenium from selenomethionine tends to be retained in tissues at higher concentrations (3–10-fold greater) and for longer periods of time than inorganic selenium compounds. The increased selenium tissue concentrations are not due to the slightly greater absorbance of selenomethionine (Butler et al. 1990; Gronbaek and Thorlacius-Ussing 1992; Ip and Hayes 1989; Salbe and Levander 1990b), but rather to the slower elimination as a consequence of its incorporation into body proteins (Stadtman 1983, 1987, 1990).

3.4.2.1 Inhalation Exposure

No studies were located regarding the distribution of selenium in humans after inhalation of elemental selenium or selenium compounds.

Weissman et al. (1983) reported that selenium concentrated in the liver, kidney, spleen, and lungs of dogs following inhalation exposure to selenious acid or elemental selenium aerosols.

3.4.2.2 Oral Exposure

A study of 100 paired samples of maternal and neonate hair found that the concentration in neonatal hair (0.63±0.52 μg/g) was lower than in maternal hair (1.02±1.04 μg/g), but the results were not analyzed statistically (Razagui and Haswell 1997). Levels of selenium in 30 paired samples of the hair of a mother and her child found no correlation between the selenium concentration of the hair of the mother and her child (Bermejo Barrera et al. 2000). The average level of selenium in the children's hair (0.77±0.24 μg/g)

was higher than that of their mothers (0.54±0.34 μg/g). The higher concentration of selenium in the children's hair could represent increased absorption or retention, but no information was provided in the study as to the age of the children or to possible differences in dietary intake of selenium between mother and child.

In rats and dogs, the selenium arising from sodium selenite administered in drinking water or in the diet is widely distributed in the body, although concentrated primarily in the liver and kidney (Furchner et al. 1975; Sohn et al. 1991; Thomson and Stewart 1973).

In most studies, selenium from selenomethionine accumulates in tissues to a greater extent than equal administered doses of selenium from selenite or selenate. Behne et al. (1991) reported higher liver and muscle selenium concentrations in rats receiving selenium orally as selenomethionine for 3 or 6 weeks than as selenite for the same length of time. Ip and Hayes (1989) reported similar results for blood, liver, kidney, and skeletal muscle. Salbe and Levander (1990b) compared distribution of dietary selenomethionine and selenate in rats and found higher selenium concentrations in plasma, erythrocytes, liver, muscle, hair, and nails in animals receiving selenomethionine. (Hair and nails have been used to gauge long-term human selenium exposure and were, therefore, included in this study.) Monkeys receiving selenomethionine in drinking water for 11 months had selenium concentrations in plasma, erythrocytes, liver, muscle, and hair that were 3–10-fold greater than monkeys receiving selenite (Butler et al. 1990). The higher levels of selenium found after selenomethionine compared to selenite treatment are likely a result of a greater retention of selenium from selenomethionine, rather than a difference in absorption. Butler et al. (1990) indicate that dietary ascorbic acid can reduce selenite absorption, but not selenomethionine absorption. Therefore, the differential effect of ascorbic acid on selenium absorption may have contributed to the difference in selenium content of tissues observed in monkeys treated with selenite, compared to monkeys treated with selenomethionine. Two studies of rats indicate that the central nervous system also concentrates more selenium when administered as selenomethionine than when administered as inorganic selenium compounds (Gronbaek and Thorlacius-Ussing 1992; Zi-Jian Jie 1992).

A comparison of distribution of selenium in selenium-depleted rats after oral administration of sodium selenite, sodium selenate, selenomethionine, or methyl selenocysteine (from high-selenium broccoli) revealed that the rate of restoration of selenium in the liver and muscle was significantly slower for methyl selenocysteine than other forms of selenium (Finley 1998). The rate of repletion in muscle was significantly faster for selenomethionine than other groups, but kidney and plasma showed no significant

difference in the rate of repletion for any form of selenium. The rate of repletion of glutathione peroxidase activity in the tissues was similar to the rate of repletion of the tissue itself and was slowest when methyl selenium was the administered form.

Another study of distribution of selenium in selenium-deficient rats fed either sodium selenite or selenomethionine found that concentration of selenium in blood and hair increased with administered dose, but was higher for selenium administered as selenomethionine (Shiobara et al. 1998).

A study of dietary supplementation of female pigs with 0.1 or 0.3 ppm selenium from a selenium-enriched yeast or from sodium selenite (doses not given) from 60 days before breeding until weaning found that the concentration of selenium in milk, dam, and offspring tissues increased with the dose of selenium administered and was higher when the source of selenium was the selenium-enriched yeast (Mahan and Kim 1996).

A study using pigs indicates that tissue levels of selenium do not correlate with effects. Tissue concentrations of selenium were higher in pigs fed 1.25 mg selenium/kg/day as D,L-selenomethionine than in pigs fed the same dose of selenium as *A. bisulcatus* or selenate, although neurological effects were more severe and occurred after fewer days of treatment with *A. bisulcatus* (Panter et al. 1996). The form of selenium in *A. bisulcatus* is unknown, although Panter et al. (1996) indicate that it is nonprotein.

In poultry, selenium is concentrated in the pancreas to a greater extent following oral administration of selenomethionine than following oral administration of sodium selenite (Cantor et al. 1975). The differential ability of the two compounds to concentrate in the pancreas of birds may explain why selenium administered as selenomethionine is more effective than the same dose of selenium administered as sodium selenite in preventing pancreatic fibrosis in chicks, a condition indicative of selenium deficiency (Cantor et al. 1975).

The distribution profiles of single oral or intravenous doses of selenium (2 mg selenium/kg as sodium selenite) administered to Wistar rats were dependent on the route of administration (Kaneko et al. 1999). Selenium concentration was highest in the kidney or liver, followed by the heart, lung, or spleen; then plasma and the brain. Oral administration produced lower doses of selenium than injection in all organs except the kidney where levels produced by the two routes were comparable (this may reflect the importance of urine as a route of excretion).

Following oral exposure, selenium is found in human milk (Brätter and Negretti De Brätter 1996; Brätter et al. 1991b; Li et al. 1999; Michalke and Schramel 1998; Moser-Veillon et al. 1992; Rodríguez Rodríguez et al. 1999; Viitak et al. 1995; Yang 1989b). Selenium is also found in the milk of mice, rats, dogs, pigs, cows, and monkeys (Abdelrahman and Kincaid 1995; Archimbaud et al. 1992; Banuelos and Mayland 2000; Chhabra and Rao 1994; Hawkes et al. 1994; Mahan and Kim 1996; Parizek et al. 1971a). This supplies offspring with selenium during the time period in which they are fed exclusively on milk (about 6 months for humans). Transplacental transfer of selenium has been demonstrated in humans, rats, hamsters, dogs, pigs, and monkeys (Archimbaud et al. 1992; Choy et al. 1993; Hawkes et al. 1994; Jandial et al. 1976; Mahan and Kim 1996; Parizek et al. 1971a; Willhite et al. 1990).

3.4.2.3 Dermal Exposure

Although unable to detect increased selenium in human females exposed to selenomethionine dermally, Burke and coworkers found elevated liver and skin selenium concentrations in mice treated with a topical lotion containing selenomethionine applied to the shaved back and ears (size of area not specified), although since the areas were not occluded, some oral absorption during grooming is also possible (Burke et al. 1992a, 1992b). In rats, between 9 and 27% of dermally applied selenious acid was absorbed, as measured in ^{75}Se radioisotope studies (Medinsky et al. 1981b).

3.4.2.4 Other Routes of Exposure

In humans, selenium has been found to be widely distributed to organs and tissues following injection of sodium selenite, sodium selenate, and selenomethionine, with the highest concentrations generally found in the liver and kidneys (Ben-Porath and Kaplan 1969; Cavalieri et al. 1966; Jereb et al. 1975; Lathrop et al. 1972). In studies involving injection of radiolabelled selenium, the pancreas accumulated high concentrations of radiolabelled selenium immediately following injection; but within hours the selenium rapidly disappeared from this organ (Lathrop et al. 1972). Using an *in vitro*, dually perfused, human term placenta, selenite has also been shown to cross the human placenta (Eisenmann and Miller 1994). Further, following intravenous injection, ^{75}Se from selenomethionine was found to cross the near-term human placenta (Jandial et al. 1976).

There is a rapid decline in serum selenium levels 1 hour after intravenous administration of sodium selenite or sodium selenate to humans (Burk 1974; Nelp and Blumberg 1965). Burk (1974) found that 50% of the plasma selenium was protein-bound within the first 2 hours after administration; 85% was

bound within 4–6 hours after administration; and 95% was bound after 24 hours. Circulating alpha-2 globulins have been reported to have the greatest affinity for selenium (Hirooka and Galambos 1966a). Burk (1974) found that lipoproteins, primarily the very low density lipoprotein (VLDL) and the low-density lipoprotein (LDL) fractions, were also involved in selenium binding.

In vitro studies of human plasma and whole blood incubated with sodium selenite have indicated that selenite is accumulated in erythrocytes by an active transport mechanism (Lee et al. 1969). Several studies indicate that the selenite is chemically altered in the erythrocyte and then transported back into the plasma, where the selenium metabolite binds to plasma proteins (Burk 1974; Hirooka and Galambos 1966a; Lee et al. 1969).

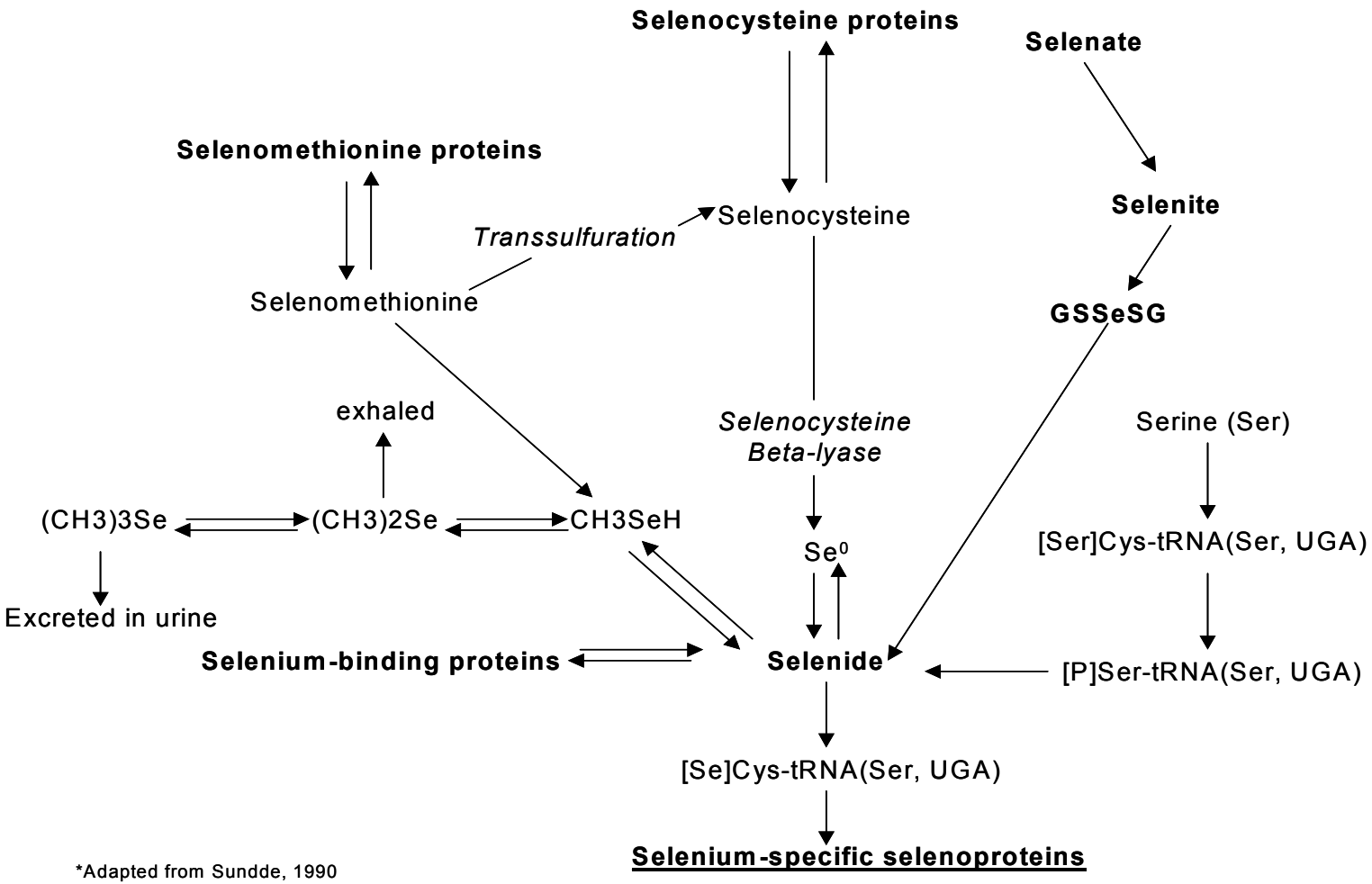
A high degree of protein binding of selenium in the plasma has also been demonstrated in experimental animals. Sandholm (1973) found that selenite administered intravenously to mice can be metabolically altered by erythrocytes to a form that binds to plasma proteins. In mice, rats, and dogs, selenite initially binds to albumin. Later, selenite can be found bound to alpha and gamma globulins in rats and to alpha-2 and beta-1 globulins in dogs (Imbach and Sternberg 1967; Sternberg and Imbach 1967).

3.4.3 Metabolism

Selenium is an essential element, and the metabolism of selenium resulting in incorporation into selenoproteins is outlined in Figure 3-4. Sunde (1990) has defined four classes of selenoproteins: selenium-specific proteins, proteins incorporating selenocystiene at cysteine codons, proteins incorporating selenomethionine at methionine position in those proteins, and proteins that bind selenide nonspecifically. The selenium-specific proteins, which include the enzymes glutathione peroxidase, thyoxine reductase, and iodothyronine 5'-deiodinase, constitute the most physiologically important class of selenoproteins. These proteins contain selenocysteine that is incorporated cotranslationally using selenide and serine as the precursors. This process is specified by a uracil-guanine-adenine (UGA) codon, which usually functions as a stop codon. A stem-loop structure in the 3' untranslated region is required for UGA to specify selenocysteine (Berry et al. 1991). This cotranslational process is the only known pathway for the production of selenocysteine in humans. In contrast to selenocysteine, selenomethionine cannot be biosynthesized by human tissues (Levander 1986).

The second and third classes of selenoproteins form in a similar manner: selenomethionine bound to the transfer ribonucleic acid (tRNA) for methionine competes with methionine bound to the tRNA for

Figure 3-4. Metabolic Pathways for Selenium*



*Adapted from Sundde, 1990

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methionine at methionine codons, and selenocysteine bound to the tRNA for cysteine competes with cysteine bound to the tRNA for cysteine at cysteine codons (Sunde 1990). The amount of selenoamino acids incorporated into protein is dependent on the ratio of the selenoamino acid and the amino acid bound to the amino acid tRNA.

The last class of selenoproteins contains the selenium binding proteins. This is an operational class defined by Sunde (1990) as "selenoproteins with selenium bound tightly enough so that the selenium remains attached during standard protein purification procedures that produce discrete selenium labeled species." This class contains selenoproteins that have not been fully characterized.

As indicated in Figure 3-4, selenide, which can nonspecifically bind to proteins, is a central selenium species in the pathways leading to the formation and degradation of selenium proteins. Selenide is also formed from selenite by reduction via glutathione. This reaction occurs in rat (Gasiewicz and Smith 1978) and human (Lee et al. 1969) red blood cells, as well as in human plasma containing added glutathione (Mas and Sarker 1989).

Selenocysteine can also be metabolized to selenide. This reaction requires a specific enzyme, selenocysteine β-lyase, which catalyzes the decomposition of selenocysteine to alanine and hydrogen selenide. The enzyme requires pyridoxal 5-phosphate as a cofactor. In humans, the highest levels of selenocysteine β-lyase activity are found in the liver, followed by the kidney, heart, adrenal gland, and muscle (Daher and Van Lente 1992). In mice orally exposed to selenocysteine, an intermediate metabolite selenocysteine-glutathione selenyl sulfide is formed in the small intestine and transported to the liver via the blood plasma (Hasegawa et al. 1995, 1996a). This compound can be nonenzymatically reduced by excess glutathion or enzymatically reduced by glutathione reductase in liver cytosol extracts to reform selenocysteine, which can be further metabolized.

Selenomethionine metabolism to selenide and the incorporation into selenium-specific proteins may occur by two pathways: metabolism to methane selenol and selenide or via selenocysteine. Evidence that the incorporation of selenium from selenomethionine into protein is by the transsulfuration pathway (methionine to cysteine) comes from studies of selenomethionine metabolism in lymphoblast cell lines deficient in cystathionine lyase and cystathionine synthetase, enzymes of the transsulfuration pathway (Beilstein and Whanger 1992). Deficiency in these enzymes greatly reduces the incorporation of selenomethionine into glutathione peroxidase.

Similar to other metals, selenium can be methylated, and the extent of methylation is dose-dependent. Dimethyl selenide is exhaled, and the trimethylselenonium ion is a major urinary metabolite of selenium. Methylation of selenium is a detoxification pathway that is especially important at high selenium doses.

Humans accidentally exposed to high levels of selenium have been reported to have a noticeable garlic odor of the breath, probably as a result of excretion of dimethyl selenide in expired air (Bopp et al. 1982; Wilbur 1980). Garlic odor of the breath has been noted in humans following ingestion of toxic levels of sodium selenate (Civil and McDonald 1978) and following inhalation of elemental selenium dust or selenium dioxide (Glover 1970).

In human populations with sufficient levels of selenium, dietary selenium is apparently partitioned into a selenite-exchangeable storage pool and a selenite-nonexchangeable storage pool. The selenite-exchangeable pool shows saturation kinetics. After this pool is filled, dietary selenium as seleno-methionine may be the primary determinant of selenium bioavailability and serum selenium concentrations (Meltzer et al. 1990, 1992). Data from both humans and Rhesus monkeys indicate that the selenium concentration in glutathione peroxidase is independent of the form of selenium administered and suggest a metabolic saturation at average intake rates (Butler et al. 1990; Meltzer et al. 1990).

In macaques that were orally administered doses of 0.025–0.3 mg selenium/kg as L-selenomethionine for up to 30 days, both erythrocyte selenium and glutathione peroxidase–specific activity showed a delay before increasing in a dose-related manner (Hawkes et al. 1992). At 0.15 and 0.3 mg selenium/kg, glutathione peroxidase–specific activity in erythrocytes continued to increase for 15 days after cessation of treatment and remained elevated through the end of the study (40 days after the end of treatment). The investigators attributed this effect to an initial deposition of selenium into a nonspecific pool (such as substitution for methionine in serum proteins), followed by slow release into the erythrocyte. Wistar rats also show incorporation of selenomethionine into proteins (Behne et al. 1991).

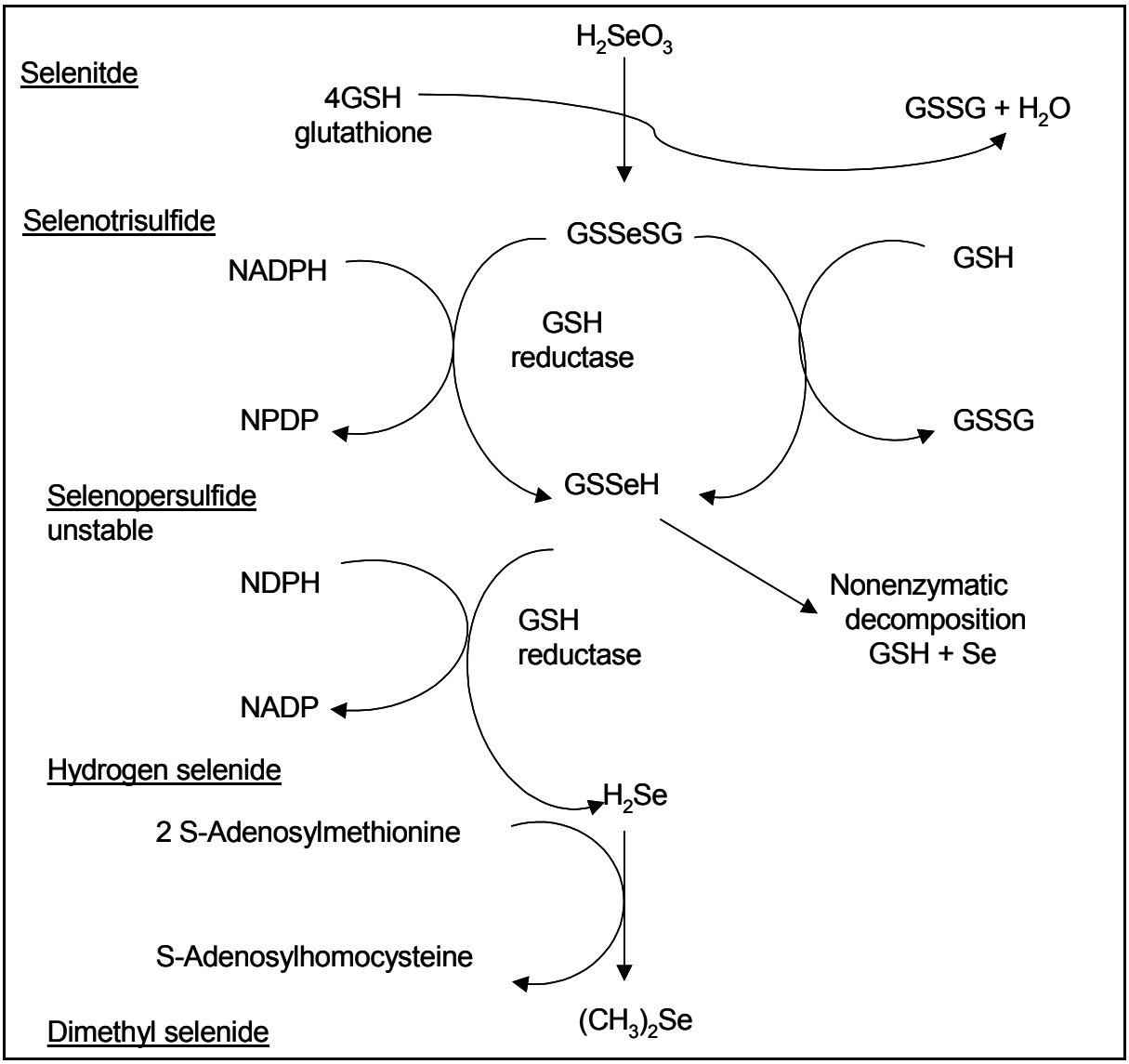
In rats, dimethyl selenide has been identified as the primary respiratory metabolite following injection of sodium selenite or sodium selenate (Hirooka and Galambos 1966b) and appears to be produced in the liver (Nakamuro et al. 1977). In mice, dimethyl selenide and dimethyldiselenide have been detected in expired air following the addition of unspecified amounts of sodium selenite, D,L-selenomethionine, or D,L-selenocystine to their drinking water (Jiang et al. 1983). A third unidentified volatile selenium compound was detected in expired air of the mice following D,L-selenomethionine injection (Jiang et al. 1983).

In rats, the trimethylselenonium ion has been identified as the predominant urinary metabolite following intraperitoneal administration of sodium selenite (Byard and Baumann 1967), sodium selenate, seleno-methionine, selenocystine, or methylselenocysteine, or following ingestion of seleniferous wheat (Palmer et al. 1970). A total of 30.8% of the urinary selenium was in the form of trimethylselenonium after administration of 15 ppm selenium in wheat. Another major selenium metabolite that appeared in the urine more slowly than the trimethylselenonium ion was identified chromatographically, but the chemical structure of that metabolite was not defined (Palmer et al. 1970).

Similarly, the trimethylselenonium ion was the major urinary metabolite of selenium excreted by rats after intraperitoneal injection of either methylselenocysteine (4 mg/kg) or selenocysteine (3 mg/kg) (Palmer et al. 1970). The amounts of trimethylselenonium ion excreted were 50.6 and 49.7% of the total urinary metabolites after methylselenocysteine and selenocysteine administration, respectively. In both cases, urinary metabolism accounted for only 10–15% of the administered dose. As selenium was not measured in feces or expired air, recovery of the dose was incomplete. In a review of the metabolic pathways resulting in the production of dimethyl selenide from selenite in rodents, Ganther (1979) indicated that reduction of selenite or selenate to dimethyl selenide requires glutathione and the methylating agent *S*-adenosylmethionine. NADPH, coenzyme A, ATP, and magnesium (II) salts are also required to provide optimal conditions for this reaction (Ganther 1979). Ganther (1971) and Hsieh and Ganther (1975) found that selenite initially reacts nonenzymatically with glutathione to form a seleno-trisulfide derivative. The selenotrisulfide is then reduced nonenzymatically in the presence of glutathione or enzymatically by glutathione reductase in the presence of NADPH to a selenopersulfide (GSSeH). The selenopersulfide is unstable and decomposes to glutathione and selenium or is enzymatically reduced by glutathione reductase in the presence of NADPH to hydrogen selenide (Ganther 1971; Hsieh and Ganther 1975). Hydrogen selenide can be methylated by *S*-adenosylmethionine in the presence of selenium methyltransferase to form dimethyl selenide (Figure 3-5).

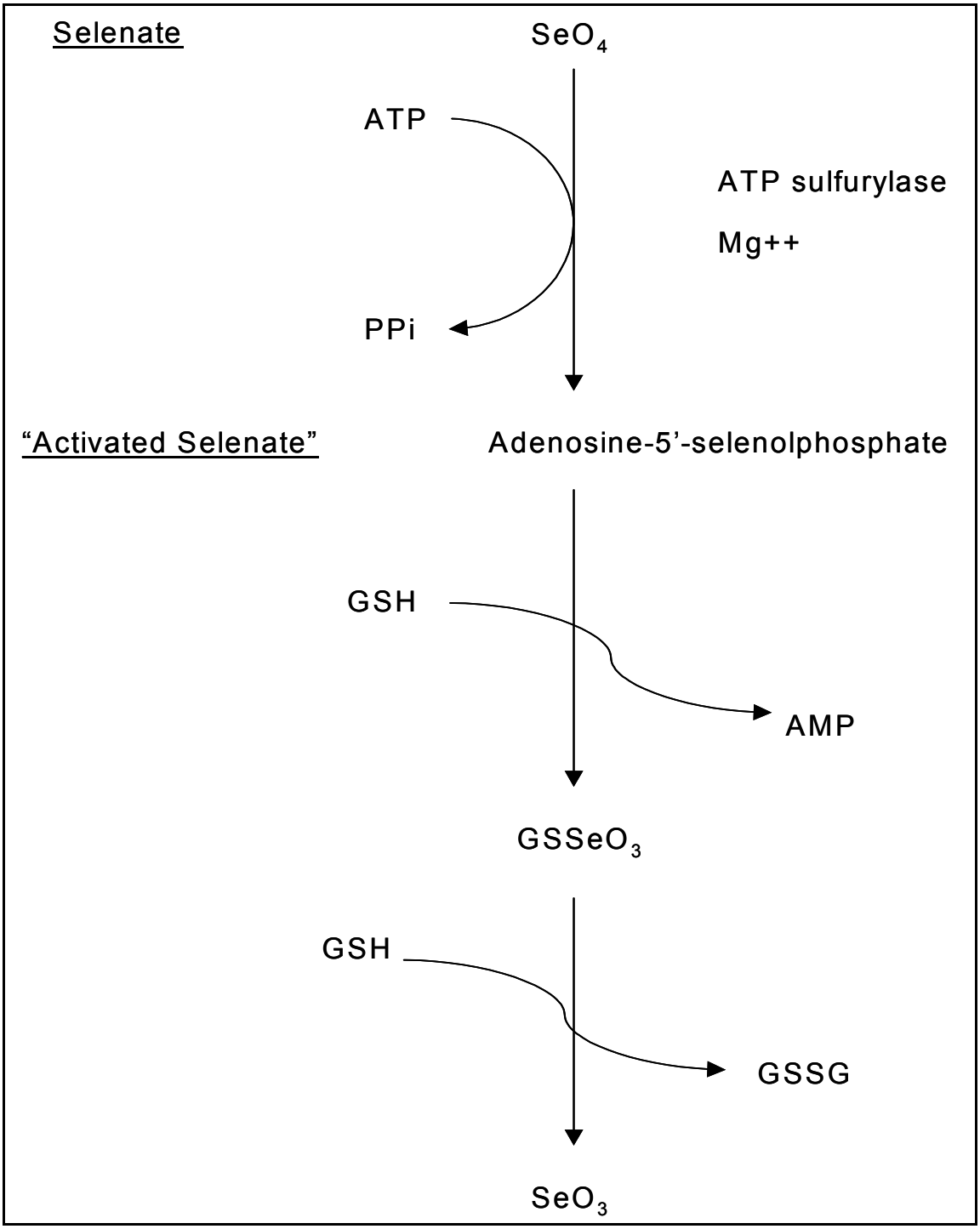
Selenate apparently is not converted to dimethyl selenide as readily as is selenite. Studies of selenate metabolism are limited in mammals, but studies using bacteria indicate that selenate must be activated prior to conversion to selenite (Bopp et al. 1982). Dilworth and Bandurski (1977) demonstrated that in the presence of ATP, magnesium (II) salts, and ATP-sulfurylase, yeast could convert selenate to eventually yield selenite (Figure 3-6). Data regarding the metabolism of selenium sulfide after administration to humans or other animals were not located in the literature.

Figure 3-5. Proposed Pathway for Formation of Dimethyl Selenide from Selenite in Animals*



*Adapted from Hsieh and Ganther 1975 and Ganther 1971

Figure 3-6. Activation and Reduction of Selenate to Selenite in Yeast *Saccharomyces cerevisiae**



*Adapted from Dilworth and Bandurski 1977

3.4.4 Elimination and Excretion

Excretion of selenium can occur in the urine, feces, and expired air (Griffiths et al. 1976; Hawkes et al. 1992, 1994; Lathrop et al. 1972; McConnell and Roth 1966; Thomson and Stewart 1974). Sweat is a minor pathway of selenium excretion in humans (Levander et al. 1987). Moreover, the initial rate of excretion appears to be dose dependent (Lathrop et al. 1972; McConnell and Roth 1966; Thomson and Stewart 1974). Some researchers have found that urinary excretion and fecal excretion of selenium are similar, with each route contributing approximately 50% of the total output (Stewart et al. 1978). However, the proportion excreted via each route seems dependent on several factors, including the level of exposure, the time since exposure, and the level of exercise. Lactating women and subjects depleted of selenium and have decreased excretion of selenium in the urine and feces (Martin et al. 1989a, 1989b; Moser-Veillon et al. 1992). At high selenium exposure levels, excretion of selenium in expired air becomes more significant (McConnell and Roth 1966; Olson et al. 1963).

3.4.4.1 Inhalation Exposure

Following acute inhalation exposures to selenium compounds, humans excrete some of the absorbed dose in the expired air (Glover 1970), but no studies were located that actually quantified the rate of excretion or identified the selenium compounds in the expired air of humans.

3.4.4.2 Oral Exposure

Several human studies have indicated that the rate of urinary excretion is most rapid in the first 24 hours following oral administration or intravenous injection of sodium selenite (Kuikka and Nordman 1978; Thomson and Stewart 1974). Thomson and Stewart (1974) found <6% of a trace dose (0.01 mg selenium) of orally administered sodium selenite was excreted in the urine within 24 hours of administration, whereas 64–73% of a 1-mg dose of selenium was excreted in the first 24 hours (Thomson 1974). Thomson et al. (1977) also found that a lower proportion of the selenium from an oral dose of 0.1 mg selenium administered as selenomethionine was excreted in the 24-hour urine than from a larger dose (1.0 mg selenium). Thus, when higher amounts of selenium are administered, a higher proportion of the selenium is excreted in the urine during the first 24 hours following exposure.

Decreasing urinary or fecal excretion appears to be the homeostatic mechanism by which the body retains greater amounts of selenium. Martin et al. (1989a) observed greater retention of selenium by individuals

maintained on a selenium-deficient diet. This increase in retention was correlated with a decrease in fecal elimination. Similarly, the increased retention of selenium from selenomethionine compared to selenite was correlated with decreased elimination (Swanson et al. 1991). Lactating women have a greater retention of selenium from selenomethionine compared to selenite and a decreased urinary elimination (Moser-Veillon et al. 1992). Muscle activity seems to influence urinary excretion of selenium as demonstrated by the doubling of selenium concentration in the urine of women following vigorous exercise (Oster and Prellwitz 1990).

Less information is available regarding the elimination of selenium in the feces of humans than in the urine of humans. However, levels of fecal excretion of selenium have been reported to be similar to levels of urinary selenium excretion when dietary levels of selenium are not excessive (Patterson et al. 1989). Over a 14-day period, Stewart et al. (1978) found urinary elimination of selenium to average 0.013 mg selenium/day and fecal elimination of selenium to average 0.011 mg selenium/day in four New Zealand women exposed to 0.024 mg selenium/day in their normal diets. Levander and Baumann (1966a, 1966b) have suggested that some of the selenium in the feces can be attributed to biliary excretion. Higher dietary levels of selenium appear to enhance fecal elimination (Gortner and Lewis 1939).

In humans, whole body retention studies following oral administration of sodium selenite have indicated that selenium elimination is triphasic (Thomson and Stewart 1974). During the initial phase, which lasted about 1 week, elimination of selenium was rapid, with a half-life of approximately 1 day (Thomson and Stewart 1974). In the second phase, which also lasted approximately 1 week, selenium elimination was slower, with a half-life of 8–9 days. In the third phase, selenium elimination was much slower, with a half-life estimated to be 115–116 days. The first two elimination phases correspond to the fecal elimination of nonabsorbed selenium and the urinary excretion of absorbed but unutilized selenium (Thomson and Stewart 1974). Selenomethionine elimination is also triphasic; however, its terminal half-life is longer than that of sodium selenite. The average half-lives of selenomethionine for the three phases were measured to be approximately 0.4–2, 5–19, and 207–290 days, respectively (Griffiths et al. 1976).

An examination of elimination data from 44 pigs exposed to excess selenium as sodium selenite in feed was found to fit a one-compartment model of selenium elimination (Davidson-York et al. 1999). Serum selenium levels were monitored over a period of 46 days beginning 1–14 days after termination of exposure to the feed containing excess selenium. Data were not adequate to depict the initial distribution phase, but a geometric mean elimination half-life of 12 days was calculated. It is likely that the period of

elimination included in this study corresponds to the second phase described by Thomson and Stewart (1974).

The chemical form of selenium may play a role in determining how rapidly selenium is excreted in the urine. In rats, the rate of urinary excretion of selenium has been found to be greater following oral administration of sodium selenite than of selenomethionine (Thomson and Stewart 1973). A comparison of excretion of selenium by selenium-depleted rats after oral administration of sodium selenate, selenomethionine, or methyl selenocysteine (from high-selenium broccoli) found that excretion of selenium from methyl selenocysteine or selenomethionine was significantly lower than from sodium selenate; further, that there was no significant difference between secretion of selenium from methyl selenocysteine and selenomethionine (Finley 1998). This may contribute to the greater retention of selenium from selenomethionine, than from inorganic selenium (Martin et al. 1989a). However, another study of excretion of selenium from rats fed selenium as either sodium selenite or selenomethionine found that excretion of selenium increased with administered dose, but was similar for both forms of selenium (Shiobara et al. 1998).

As exposure to oral L-selenomethionine increased in macaques, the amount of selenium eliminated in the urine/day increased, as did the maximum rate of urinary excretion. However, the percentage of administered dose appearing in the urine decreased with an increase in dose (Hawkes et al. 1994).

3.4.4.3 Dermal Exposure

No studies were located regarding the excretion of selenium by humans or other animals after dermal exposure to elemental selenium or selenium compounds.

3.4.4.4 Other Routes of Exposure

Whole body retention studies in sheep following injection of selenium have indicated that selenium excretion in animals follows a triexponential profile (Blodgett and Bevill 1987b; Ewan et al. 1967). In a 2-week study, Blincoe (1960) estimated the half-life for ^{75}Se in rats following intraperitoneal injection of ^{75}Se-labeled sodium selenite (0.93 mg selenium/kg). Initially, the excretion of selenium was rapid, with a half-life of approximately 0.8 day; the second phase of excretion was slower, with a half-life of 13 days. These results parallel the initial phases of selenium excretion seen in humans. The abbreviated duration of the Blincoe (1960) study did not permit the determination of a terminal elimination phase half-life. In

rats, Ewan et al. (1967) found the final phase of elimination of selenium following a single subcutaneous injection of sodium selenite to be dose independent (from 0.008 mg selenium/kg to 2 mg selenium/kg), with a half-life of 65–78 days. Blodgett and Bevill (1987b) found the elimination rate of selenium in sheep during the second phase following a single intramuscular injection of sodium selenite to be dose dependent, with larger doses resulting in longer half-lives (i.e., doses of 0.4, 0.6, 0.7, or 0.8 mg selenium/kg resulting in half-lives for selenium elimination of 6.3, 8.8, 15.1, and 20.4 hours, respectively). The reasons for the decreasing elimination rate with increasing dose during the second phase is not clear.

Dietary levels of selenium and the individual's selenium nutritional status are the most important factors that influence the route and rate of selenium excretion. Selenium excretion in expired air is only significant when exposures to selenium are high. Rats injected subcutaneously with sodium selenite at doses of 2.2–5.4 mg selenium/kg excreted 41–62% of the administered selenium in exhaled air, whereas rats injected with sodium selenite at doses of 0.005–0.9 mg selenium/kg excreted only 0.2–11% of the administered selenium in expired air (McConnell and Roth 1966; Olson et al. 1963). As the amount of administered sodium selenite increased, the percent of the administered selenium excreted in the urine decreased (from approximately 22–33% of the administered selenium at doses of 0.005–0.9 mg selenium/kg to 3–14% of the administered selenium at doses of 3.1–5.4 mg selenium/kg) (McConnel and Roth 1966). Selenium in the feces was not measured in this study. Burk et al. (1972) found that as the dietary level of sodium selenite was increased, a larger proportion of an injected tracer dose of selenium (as sodium selenite) was excreted. At a dietary level of 0.005 mg selenium/kg, approximately 60% of the injected selenium had been excreted in the first 35 days following administration. At a dietary level of 0.05 mg selenium/kg, over 94% of the injected selenium had been excreted over the same period of time.

In experimental animals, other factors that can cause an increase in selenium levels in expired air are higher dietary levels of selenium, protein, or methionine (Ganther et al. 1966). Phenobarbital induction of microsomal enzymes has also led to increased exhalation of selenium following intravenous administration of sodium selenite (Sternberg et al. 1968).

3.4.5 Physiologically Based Pharmacokinetic (PBPK)/Pharmacodynamic (PD) Models

Physiologically based pharmacokinetic (PBPK) models use mathematical descriptions of the uptake and disposition of chemical substances to quantitatively describe the relationships among critical biological processes (Krishnan et al. 1994). PBPK models are also called biologically based tissue dosimetry models. PBPK models are increasingly used in risk assessments, primarily to predict the concentration of potentially toxic moieties of a chemical that will be delivered to any given target tissue following various combinations of route, dose level, and test species (Clewell and Andersen 1985). Physiologically based pharmacodynamic (PBPD) models use mathematical descriptions of the dose-response function to quantitatively describe the relationship between target tissue dose and toxic end points.

PBPK/PD models refine our understanding of complex quantitative dose behaviors by helping to delineate and characterize the relationships between: (1) the external/exposure concentration and target tissue dose of the toxic moiety, and (2) the target tissue dose and observed responses (Andersen et al. 1987; Andersen and Krishnan 1994). These models are biologically and mechanistically based and can be used to extrapolate the pharmacokinetic behavior of chemical substances from high to low dose, from route to route, between species, and between subpopulations within a species. The biological basis of PBPK models results in more meaningful extrapolations than those generated with the more conventional use of uncertainty factors.

The PBPK model for a chemical substance is developed in four interconnected steps: (1) model representation, (2) model parametrization, (3) model simulation, and (4) model validation (Krishnan and Andersen 1994). In the early 1990s, validated PBPK models were developed for a number of toxicologically important chemical substances, both volatile and nonvolatile (Krishnan and Andersen 1994; Leung 1993). PBPK models for a particular substance require estimates of the chemical substance-specific physicochemical parameters, and species-specific physiological and biological parameters. The numerical estimates of these model parameters are incorporated within a set of differential and algebraic equations that describe the pharmacokinetic processes. Solving these differential and algebraic equations provides the predictions of tissue dose. Computers then provide process simulations based on these solutions.

The structure and mathematical expressions used in PBPK models significantly simplify the true complexities of biological systems. If the uptake and disposition of the chemical substance(s) is adequately described, however, this simplification is desirable because data are often unavailable for

many biological processes. A simplified scheme reduces the magnitude of cumulative uncertainty. The adequacy of the model is, therefore, of great importance, and model validation is essential to the use of PBPK models in risk assessment.

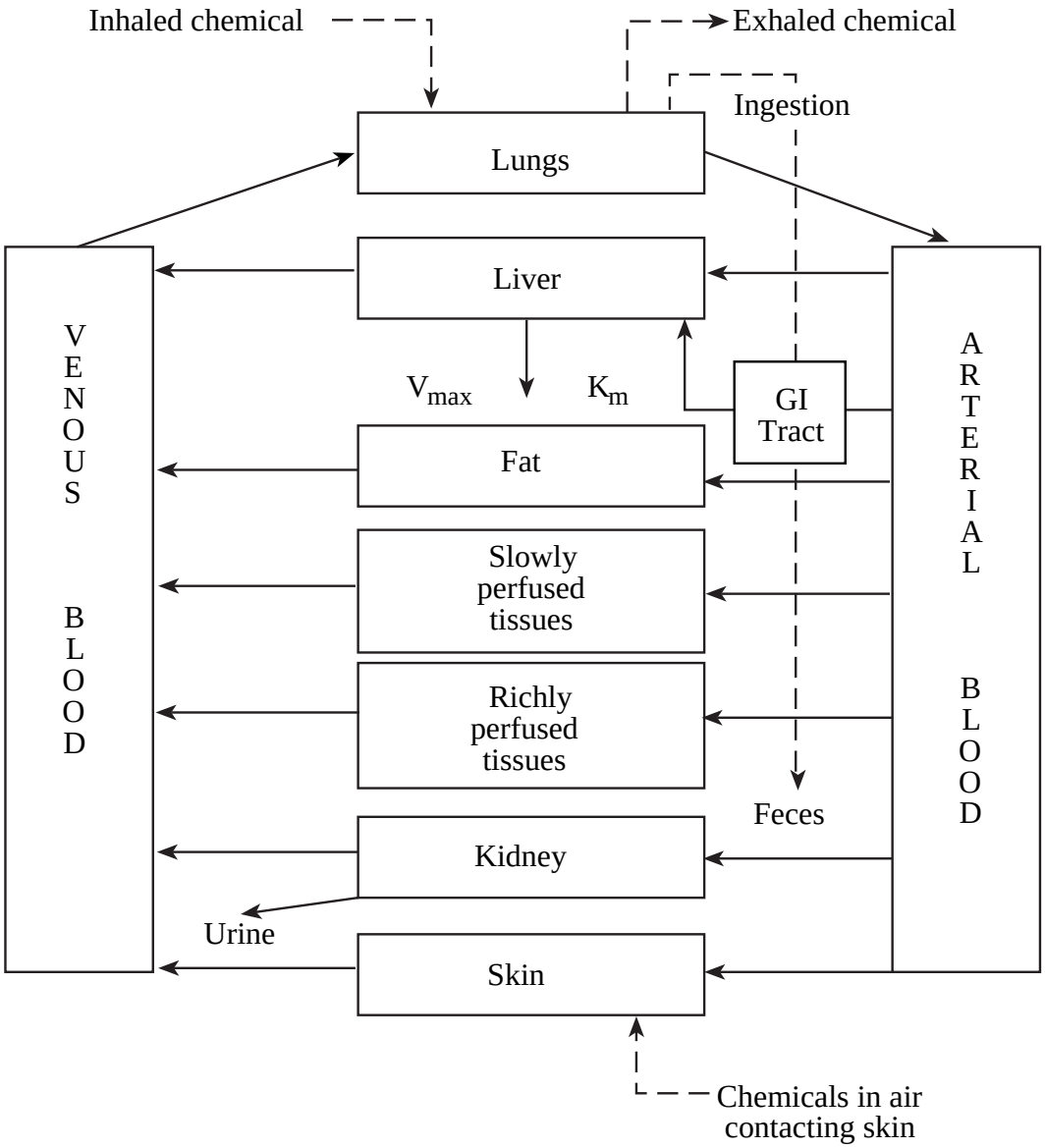
PBPK models improve the pharmacokinetic extrapolations used in risk assessments that identify the maximal (i.e., the safe) levels for human exposure to chemical substances (Andersen and Krishnan 1994). PBPK models provide a scientifically sound means to predict the target tissue dose of chemicals in humans who are exposed to environmental levels (for example, levels that might occur at hazardous waste sites) based on the results of studies where doses were higher or were administered in different species. Figure 3-7 shows a conceptualized representation of a PBPK model.

Two models for selenium were located in the literature. Patterson and coworkers (Patterson and Zech 1992; Patterson et al. 1989, 1993) have developed compartmental models of the kinetics of selenium orally administered as selenite or selenomethionine in adult humans.

Patterson et al. (1989) Selenite Model

Description of the model. Patterson and coworkers (Patterson and Zech 1992; Patterson et al. 1989, 1993) developed a compartmental model of the kinetics of ingested selenite in adult humans based on data from human subjects who consumed a single oral dose of 200 μg ^{74}Se as selenite. The model assumes that 84% of the administered selenium is absorbed and that absorption is rapid. Absorbed selenite is assumed to distribute to six compartments: gastrointestinal tract, plasma, hepatopancreatic/ lymphatic system, liver/pancreas, bile, and tissues (Figure 3-8). Unabsorbed selenium is excreted in the feces. Absorption occurs from the gastrointestinal compartment (probably the small intestine, but also possibly the stomach) into a rapidly turning-over pool (the intestinal cells or enterocytes) from which it leaves by two pathways. The central compartment is represented as four kinetically distinct plasma pools, P1 (the portal circulation), P2 (before passage through the liver), P3 (after passage through the liver), and P4 (after passage through the tissues). In the first pathway, selenium enters P1. The second pathway is to a liver/pancreatic compartment. Transport into and out of P1 is very rapid ($T_{1/2}$ approximately 0.36 hours) and this may represent selenium in the portal circulation passing through the liver before appearing in P3, but not removed in the first pass. The second pathway is via the hepatopancreatic/lymphatic system compartment to a second plasma pool (P2). Appearance of selenium in P2 is delayed ($T_{1/2}$ approximately 0.55 hours), representing the time needed to move through

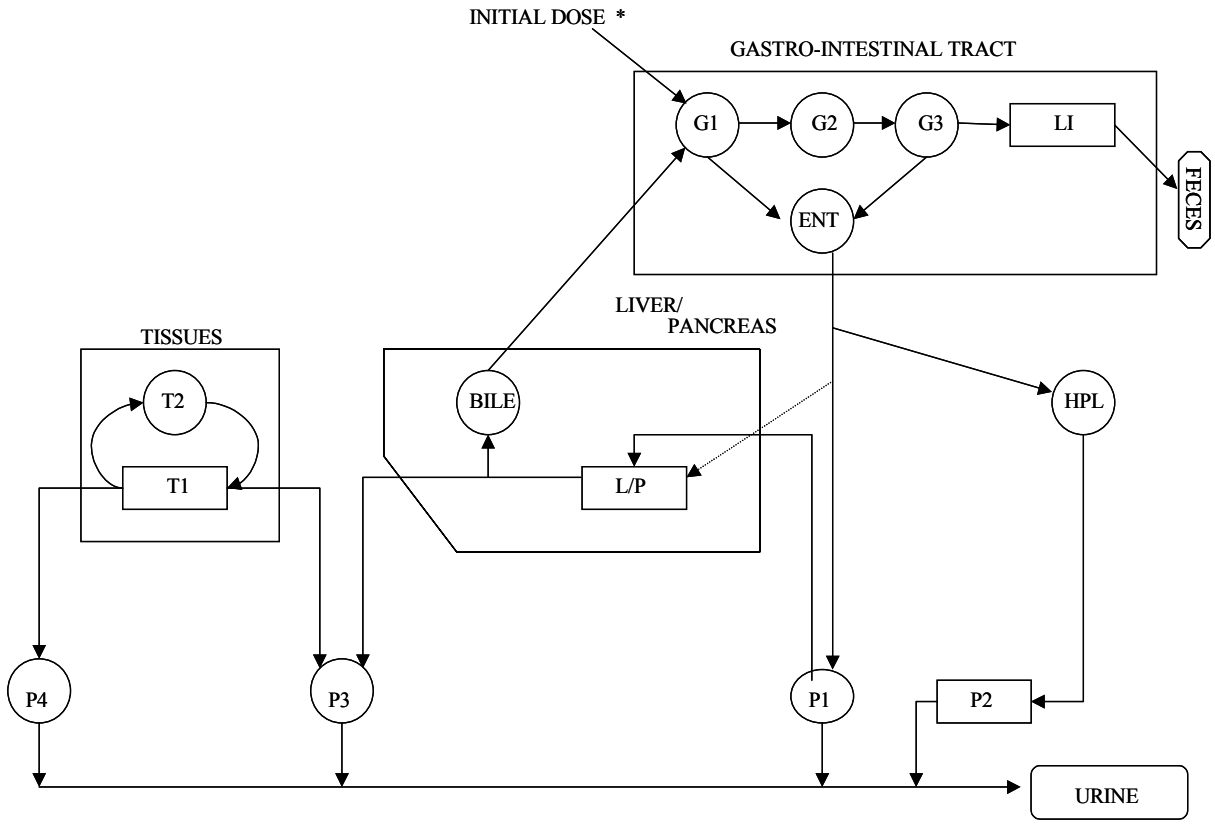
Figure 3-7. Conceptual Representation of a Physiologically Based Pharmacokinetic (PBPK) Model for a Hypothetical Chemical Substance



Source: adapted from Krishnan et al. 1994

Note: This is a conceptual representation of a physiologically based pharmacokinetic (PBPK) model for a hypothetical chemical substance. The chemical substance is shown to be absorbed via the skin, by inhalation, or by ingestion, metabolized in the liver, and excreted in the urine or by exhalation.

Figure 3-8. Selenite Model, a Kinetic Model for Selenite Metabolism



The arrow with an asterisk indicates the site of entry of the oral Se tracer. Arrows between compartments represent pathways of fractional transport. Compartments depicted as rectangles represent delays. Compartments G1, G2, G3, 3 gut compartments, probably the small intestine; ENT, enterocytes (intestinal cells); HPL, compartment in hepato-pancreatic subsystem or lymphatic system; L/P, liver and pancreas; LI, large intestine; T1, T2, peripheral tissues, e.g., skeletal muscle, bone, kidney. Feces and urine compartments are drawn in the shape of test tubes to represent fractional (single) collections. The model includes absorption distributed along the gastrointestinal tract, enterohepatic recirculation, four kinetically distinct plasma pools, P1–P4, a subsystem consisting of liver and pancreas, and a slowly turning-over tissue pool.

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the hepatopancreatic/lymphatic system compartment. From the two plasma pools (P1 and P2), selenium can be excreted in the urine ($T_{1/2}$ approximately 3.94 and 1.96 hours, respectively) or it can move into the liver/pancreas compartment. After a delay of 4–6 hours, the selenium leaves the liver/pancreas either to a bile compartment ($T_{1/2}$ approximately 0.13 hours) and thence to the gut (G1) for excretion in feces or to a third plasma pool (P3) ($T_{1/2}$ approximately 0.19 hours). From P3, selenium can be excreted in the urine ($T_{1/2}$ approximately 4.15 hours) or can move into a large, slowly turning-over tissue compartment. Finally, selenium is transferred very slowly ($T_{1/2}$ approximately 1.27 hours) from the tissues (probably final metabolic products) to a fourth plasma pool (P4) and hence to the urine ($T_{1/2}$ approximately 6.54 hours).

Validation of the model. The extent to which this model has been validated is not described in Patterson and coworkers (Patterson and Zech 1992; Patterson et al. 1991, 1993).

Risk assessment. The model was designed to simulate the pharmacokinetics of selenium orally administered as selenite to humans as a preparation for a larger anticancer supplementation study jointly undertaken by the National Cancer Institute (NCI) and the U.S. Department of Agriculture (USDA) (Patterson and Zech 1992; Patterson et al. 1991, 1993).

Target tissues. The model is designed to simultaneously account for the appearance and disappearance of selenium in plasma, urine, and feces after administration of a single oral dose of ^{74}Se as selenite (Patterson and Zech 1992; Patterson et al. 1991, 1993).

Species extrapolation. The model is designed for applications to human dosimetry and cannot be applied to other species without modification.

Interroute extrapolation. The model is designed to simulate oral exposures to selenite and cannot be applied to other routes of exposure without modification.

Extrapolation to other forms of selenium. The model is designed to simulate oral exposures to selenite and cannot be applied to other forms of selenium without modification.

Swanson et al. (1991) Selenomethionine Model

Description of the model. Swanson and coworkers (Patterson et al. 1993; Swanson et al. 1991) produced a model for ingested selenomethionine in adult humans based on data from human subjects who consumed a single oral dose of 200 μg ^{74}Se as selenomethionine and the model of the kinetics of ingested selenite described above. Four major changes (indicated by bold lines in Figure 3-9) were made to the selenite model to achieve an adequate fit to the selenomethionine data: (1) the amount of label absorbed into the enterocyte was increased (the absorption of ^{74}Se was 98% for selenomethionine compared with 84% for selenite), (2) the amount of label removed from the plasma in the first pass through the liver was increased, (3) a pathway from P4 back to the liver was added, providing for conservation and reutilization of amino acids (estimated 95% of material from P4 is recycled), and (4) a second tissue subgroup was added to the model and rate constants were adjusted so that the subgroups had different turnover times.

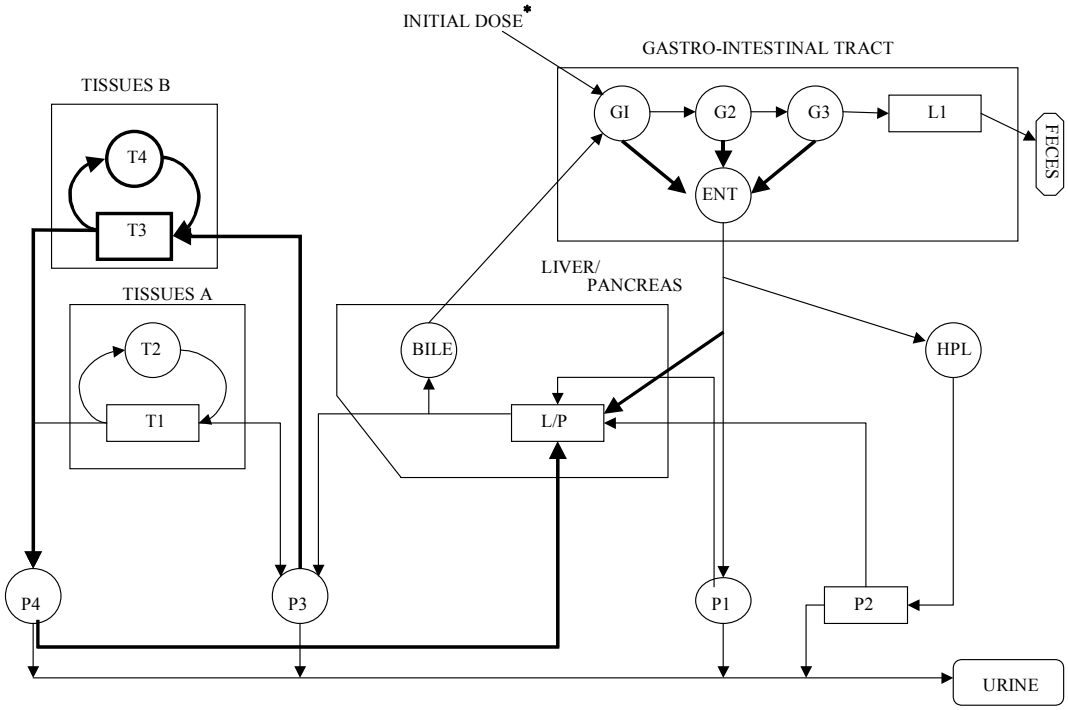
The most important differences between the selenite and selenomethionine models lie in the turnover times. The estimated turnover times in the plasma, liver/pancreas, and tissues are shorter for selenomethionine than for selenite, but the estimated turnover time for the whole body is more than twice as long for selenomethionine as for selenite. This is probably because selenite is not recirculated, whereas selenomethionine is extensively recycled, passing through the individual organs and tissues many times before being excreted.

Validation of the model. The extent to which this model has been validated is not described by the authors (Patterson et al. 1993; Swanson et al. 1991).

Risk assessment. The model was designed to simulate the pharmacokinetics of selenium orally administered as selenomethionine to humans as a preparation for a larger anti-cancer supplementation study jointly undertaken by the NCI and the USDA (Patterson et al. 1993; Swanson et al. 1991).

Target tissues. The model is designed to simultaneously account for the appearance and disappearance of selenium in plasma, urine, and feces after administration of a single oral dose of ^{74}Se as selenomethionine (Patterson et al. 1993; Swanson et al. 1991).

Figure 3-9. Selenomethionine Model, a Kinetic Model for Selenomethionine Metabolism



The arrow with an asterisk indicates the site of the oral Se tracer. Arrows between compartments represent pathways of fractional transport. Compartments depicted as rectangles represent delays. G1, G2, G3, three gut compartments, probably small intestine; ENT, enterocytes (intestinal cells); HPL, compartment in hepatopancreatic subsystems or lymphatic system; L/P, liver and pancreas; LI, large intestine; T1, T2, T3, T4, peripheral tissues, e.g., skeletal muscle, bone, kidney. Feces and urine along the gastrointestinal tract, enterohepatic recirculation, four kinetically distinct plasma pools, P1–P4, a subsystem consisting of the liver and pancreas, two tissue subsystems that are slowly turning-over, and a pathway for reutilization of selenium metabolites from peripheral tissues. The bold lines indicate the major modifications to the Selenite Model (Figure 3-8).

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Species extrapolation. The model is designed for applications to human dosimetry and cannot be applied to other species without modification.

Interroute extrapolation. The model is designed to simulate oral exposures to selenomethionine and cannot be applied to other routes of exposure without modification.

Extrapolation to other forms of selenium. The model is designed to simulate oral exposures to selenomethionine and cannot be applied to other forms of selenium without modification.

3.5 MECHANISMS OF ACTION

3.5.1 Pharmacokinetic Mechanisms

As discussed in Section 3.4.1, selenium is readily absorbed by inhalation or ingestion when present in any of several compounds. Inhalation and oral absorption are extensive, although the rate of absorption varies depending on the form of selenium (Medinsky et al. 1981a; Moser-Veillon et al. 1992; Swanson et al. 1991; Weissman et al. 1983; Young et al. 1982). Oral bioavailability is generally independent of the exposure level, but may be increased in some selenium-deficient individuals (Griffiths et al. 1976; Martin et al. 1989a; Thomson 1974; Thomson et al. 1977). Selenate and selenomethionine appear to be absorbed by the intestine largely unchanged, while selenite and selenocysteine are metabolized during absorption (Hasegawa et al. 1995, 1996a; Spencer and Blau 1962; Whanger et al. 1976, 1996). No evidence of significant dermal absorption of selenium by humans was located, although mice can absorb topically-applied selenomethionine (Burke et al. 1992b).

Absorbed selenium is carried throughout the body in the blood, eventually being distributed to all tissues. Injection studies in humans have shown that after selenium enters the blood, it rapidly becomes protein-bound (Burk 1974; Hirooka and Galambos 1966a), while *in vitro* studies have shown that selenite is accumulated in erythrocytes via an active transport mechanism (Lee et al. 1969). Selenium is an essential element and is incorporated into selenoproteins (e.g., glutathione peroxidase, iodothyronine deiodinases) as selenocysteine. Most studies report similar distribution patterns for selenium, regardless of the form in which it was administered; however, the concentration reached is generally higher for doses delivered as an organic forms of selenium than for the same dose delivered as an inorganic form (Behne et al. 1991; Butler et al. 1990; Gronbaek and Thorlacius-Ussing 1992; Ip and Hayes 1989; Salbe and Levander 1990b; Shiobara et al. 1998; Zi-Jian Jie 1992). In humans, the highest levels of selenium are found in the

liver and kidney (see Table 3-6 for normal levels of selenium in human tissues). Selenomethionine is not synthesized by humans, but can be incorporated into proteins in the place of methionine; because of this, selenomethionine is retained for a longer time within the body than inorganic forms, and it may therefore represent a storage form of the element.

Excretion of selenium by humans occurs in the urine, feces, expired air, and sweat, but urine and feces are the major routes of elimination. Some of the selenium in feces may be due to bilary excretion (Levander and Baumann 1966a, 1966b). Elimination is reduced in selenium-deficient individuals and may represent a mechanism by which selenium levels are regulated (Martin et al. 1989a; Swanson et al. 1991). Methylation is an important mechanism of detoxification for selenium; dimethyl selenide is exhaled, and the trimethylselenonium ion is the major urinary metabolite of selenium. Experiments in mice suggest that the hepatic toxicity of selenium may be at least partly due to depression of selenium methylation in the liver, resulting in the accumulation of excess selenides (Nakamuro et al. 2000).

3.5.2 Mechanisms of Toxicity

Little is known about the specific biochemical mechanism(s) by which selenium and selenium compounds exert their acute toxic effects. Long-term effects on the hair, skin, nails, liver, and nervous system are also well documented, and a general theory has been developed to explain the toxicity of exposure to excess selenium, as discussed below. Generally, water-soluble forms are more easily absorbed and are generally of greater acute toxicity. Mechanisms of absorption and distribution for dermal and pulmonary uptake are unknown and subject to speculation, but an active transport mechanism for selenomethionine absorption in the intestine has been described (Spencer and Blau 1962). The mechanisms by which selenium exerts positive effects as a component of glutathione peroxidase, thioredoxin reductase, and the iodothyronine 5'-deiodinases are better understood, but the roles of other selenium-containing proteins in mammalian metabolism have not been clarified.

One theory of the mechanism of acute selenium toxicity concerns inactivation of the sulfhydryl enzymes necessary for the oxidative reactions in cellular respiration (Lombeck et al. 1987; Mack 1990; Shamberger 1981). Acute toxic effects, such as pulmonary edema, can result in respiratory failure and death. The lung, however, does not appear to be a target organ at levels of exposure less than the occupational permissible exposure limits (PELs) or threshold limit values (TLV).

Selenium can replace sulfur in biomolecules, especially when the concentration of selenium is high and the concentration of sulfur is low in the organism. This substitution may be a mechanism of selenium toxicity (Stadtman 1983). Selenocysteine is specifically found in some proteins (e.g., glutathione peroxidase); selenomethionine appears to randomly substitute for methionine in protein synthesis. This appears to be an additional mechanism for intermediate or chronic duration toxicity (Stadtman 1983; Tarantal et al. 1991). Skin, hair, and nail damage are significant indicators of chronic selenium overexposure. Although the mechanism causing these integumentary effects is not known, it is probably related to the high selenium concentrations in these tissues as a consequence of the substitution of selenium for sulfur in certain amino acids. Researchers have concluded that the nails and hair are routes to excrete excess selenium (Yang et al. 1989b). Other studies suggest that selenium may also influence thyroid hormone function via the deiodinase enzymes (Brätter and Negretti De Brätter 1996; Hawkes and Keim 1995).

3.5.3 Animal-to-Human Extrapolations

No studies were located that specifically examined species-related differences in selenium pharmacokinetics. Similar patterns of absorption, distribution, and elimination have been reported for human and animal systems and the dermal, endocrine, and neurological effects of chronic exposure in humans are similar to those reported for animals exposed to very high doses of selenium. However, species-specific differences in toxicity are present (e.g., the main effect of selenium toxicity in rodents is damage to the liver, which is not observed in humans) and this may represent evidence of underlying differences in how selenium is metabolized.

3.6 TOXICITIES MEDIATED THROUGH THE NEUROENDOCRINE AXIS

Recently, attention has focused on the potential hazardous effects of certain chemicals on the endocrine system because of the ability of these chemicals to mimic or block endogenous hormones. Chemicals with this type of activity are most commonly referred to as *endocrine disruptors*. However, appropriate terminology to describe such effects remains controversial. The terminology *endocrine disruptors*, initially used by Colborn and Thomas (1992) and again by Colborn (1993), was also used in 1996 when Congress mandated the Environmental Protection Agency (EPA) to develop a screening program for "...certain substances [which] may have an effect produced by a naturally occurring estrogen, or other such endocrine effect[s]...". To meet this mandate, EPA convened a panel called the Endocrine Disruptors Screening and Testing Advisory Committee (EDSTAC), which in 1998 completed its

deliberations and made recommendations to EPA concerning *endocrine disruptors*. In 1999, the National Academy of Sciences released a report that referred to these same types of chemicals as *hormonally active agents*. The terminology *endocrine modulators* has also been used to convey the fact that effects caused by such chemicals may not necessarily be adverse. Many scientists agree that chemicals with the ability to disrupt or modulate the endocrine system are a potential threat to the health of humans, aquatic animals, and wildlife. However, others think that endocrine-active chemicals do not pose a significant health risk, particularly in view of the fact that hormone mimics exist in the natural environment. Examples of natural hormone mimics are the isoflavinoid phytoestrogens (Adlercreutz 1995; Mayr et al. 1992; Livingston 1978). These chemicals are derived from plants and are similar in structure and action to endogenous estrogen. Although the public health significance and descriptive terminology of substances capable of affecting the endocrine system remains controversial, scientists agree that these chemicals may affect the synthesis, secretion, transport, binding, action, or elimination of natural hormones in the body responsible for maintaining homeostasis, reproduction, development, and/or behavior (EPA 1997c). Stated differently, such compounds may cause toxicities that are mediated through the neuroendocrine axis. As a result, these chemicals may play a role in altering, for example, metabolic, sexual, immune, and neurobehavioral function. Such chemicals are also thought to be involved in inducing breast, testicular, and prostate cancers, as well as endometriosis (Hoel et al. 1992; Giwercman et al. 1993; Berger 1994).

Selenium is a component of all three members of the deiodinase enzyme family, the enzymes responsible for deiodination of the thyroid hormones (Kohrle 1994; St. Germain and Galton 1997). The deiodinases contain a selenocysteine at the active site, which is required for catalytic activity. There are three types of deiodinases and they differ in terms of tissue distribution, reaction kinetics, efficiency of substrate utilization, and sensitivity to inhibitors. The first to be recognized as a selenoprotein was type I iodothyronine 5'-deiodinase which converts the prohormone thyroxine (T4) to the active form, triiodothyronine (T3) and to date, studies of the effects of excess selenium have focused on this protein. Under normal circumstances the human thyroid produces only 20–30% of its hormone as T3; the remainder is T4 (a minute amount of reverse T3 (rT3) is also produced), which is largely converted to active T3 by type I deiodinase located within the liver, euthyroid pituitary, kidney, thyroid, and brain. Type I deiodinase is a membrane bound protein and, thus, its activity has not been directly measured in studies of humans supplemented with selenium. Human studies have instead measured serum levels of T3, rT3, T4, and TSH.

Two human studies have demonstrated a decrease in T3 levels in response to increased dietary selenium although the hormone levels remained within the normal human range (Brätter and Negretti De Brätter 1996; Hawkes and Keim 1995). The effect of increased dietary selenium on other thyroid hormones is unclear. No significant correlation between selenium intake and serum T4 or TSH levels was found in the study of Brätter and Negretti De Brätter (1996). While in the study of Hawkes and Keim (1995), TSH concentration increased in the high selenium group (+37%) and was significantly different relative to baseline levels ($p<0.06$). In a third study of the effects of selenium supplementation, New Zealanders with normally low selenium intake (unsupplemented intake of 28–29 μg/day), showed a reduction in T4 concentration in all groups after 20 weeks (Duffield et al. 1999). A significant inverse correlation was found between serum levels of selenium and TSH among fish consumers; however, it is not known if this population had a high selenium intake (Hagmar et al. 1998).

Male rats receiving diets supplying $0.05mg selenium/kg/day for 6–12 weeks has been shown to have reductions in type-I-deiodinase activity (Behne et al. 1992; Eder et al. 1995; Hotz et al. 1997). However, the levels of thyroid hormones in these animals have not shown a consistent pattern. Exposure to 0.055 mg selenium/kg/day as sodium selenite for 40 days produced a significant decrease in serum levels of T3 (Eder et al. 1995). While in another study, a dose of 0.09 mg selenium/kg/day as sodium selenate in food for 6 weeks produced a significant (~30%) increase in TSH (Hotz et al. 1997), and no significant changes in thyroid levels of T3 or T4 were found in rats receiving 0.105 mg selenium/kg/day as sodium selenite or 0.118 mg selenium/kg/day as L-selenomethionine for 3 months (Behne et al. 1992).

Many studies have documented reduced body weight gain in young animals treated with selenium compounds and abnormal weight loss in older animals (Grønback et al. 1995; Halverson et al. 1966; Harr et al. 1967; Jacobs and Forst 1981a; Nelson et al. 1943; NTP 1994; Johnson et al. 2000; Palmer and Olson 1974; Panter et al. 1996; Schroeder 1967; Tsunoda et al. 2000). There is evidence to suggest that these effects may be due in part to the interactions of selenium or selenium compounds with hormones that regulate normal growth and body weight. Reduced somatomedin C levels, reduced insulin-like growth factor–binding protein-3, and a reduction in growth hormone secretion in response to growth hormone releasing factor have been reported for rats exposed to sodium selenite in drinking water (Gronbaek et al. 1995; Thorlacius-Ussing et al. 1988).

No studies were located regarding adverse effects on human reproduction following oral exposure to elemental selenium or to selenium compounds. However, data from animal studies suggest that oral exposure to selenium may be associated with male infertility. Adverse effects associated with selenium

exposure include decreased sperm counts in rats and rabbits (El-Zarkouny et al. 1999; Kaur and Parshad 1994; NTP 1994), sperm abnormalities in rats and rabbits (El-Zarkouny et al. 1999; Kaur and Parshad 1994), testicular hypertrophy in rats (Turan et al. 1999a), and a significant reduction in serum testosterone in rabbits (El-Zarkouny et al. 1999). However, it is not clear what effect, if any, this had on the ability of the animals to reproduce, as chronic administration of selenate did not affect male fertility in rats or mice (Rosenfeld and Beath 1954; Schroeder and Mitchener 1971b).

Chronic exposure of mice and rats to otherwise nontoxic doses has been shown to reduce fertility and to markedly reduce the viability of the offspring of pairs that are able to conceive (Schroeder and Mitchener 1971b; Wahlstrom and Olson 1959b). Selenium exposure has been shown to alter the length of the estrous cycle in female mice (Nobunaga et al. 1979) and to alter the menstrual cycle in monkeys (Cukierski et al. 1989). Vaginal cytology of female rats provided with drinking water containing selenate or selenite indicated that the rats spent more time in diestrus and less time in proestrus and estrus than the controls (NTP 1994). However, it is not clear what effect, if any, this had on the ability of the animals to reproduce.

Fertility studies in mice, rats, and pigs have demonstrated reduced rates of conception after oral treatment with selenium as selenate or selenite (Rosenfeld and Beath 1954; Schroeder and Mitchener 1971b; Wahlstrom and Olson 1959b). Decreased conception rates and increased resorption rates have been reported for cattle, sheep, and horses fed diets naturally containing organic selenium compounds and exhibiting symptoms of selenosis (Harr and Muth 1972). An increased concentration of progesterone in the milk and an association of cystic ovaries with elevated blood selenium concentrations was observed in cows receiving selenium supplementation (Mohammed et al. 1991).

Other possible examples of endocrine disruption due to selenium exposure include pancreatic damage in sheep and rats fed selenium as sodium selenite, sodium selenate, or seliniferous wheat (Halverson et al. 1966; Harr et al. 1967; Smyth et al. 1990) and decreased plasma glucose (an insulin-like effects) in rats injected with sodium selenate. However, these are isolated reports and it is not clear what relevance they have for selenium toxicity in humans.

3.7 CHILDREN'S SUSCEPTIBILITY

This section discusses potential health effects from exposures during the period from conception to maturity at 18 years of age in humans, when all biological systems will have fully developed. Potential effects on offspring resulting from exposures of parental germ cells are considered, as well as any indirect effects on the fetus and neonate resulting from maternal exposure during gestation and lactation. Relevant animal and *in vitro* models are also discussed.

Children are not small adults. They differ from adults in their exposures and may differ in their susceptibility to hazardous chemicals. Children's unique physiology and behavior can influence the extent of their exposure. Exposures of children are discussed in Section 6.6 Exposures of Children.

Children sometimes differ from adults in their susceptibility to hazardous chemicals, but whether there is a difference depends on the chemical (Guzelian et al. 1992; NRC 1993). Children may be more or less susceptible than adults to health effects, and the relationship may change with developmental age (Guzelian et al. 1992; NRC 1993). Vulnerability often depends on developmental stage. There are critical periods of structural and functional development during both prenatal and postnatal life and a particular structure or function will be most sensitive to disruption during its critical period(s). Damage may not be evident until a later stage of development. There are often differences in pharmacokinetics and metabolism between children and adults. For example, absorption may be different in neonates because of the immaturity of their gastrointestinal tract and their larger skin surface area in proportion to body weight (Morselli et al. 1980; NRC 1993); the gastrointestinal absorption of lead is greatest in infants and young children (Ziegler et al. 1978). Distribution of xenobiotics may be different; for example, infants have a larger proportion of their bodies as extracellular water and their brains and livers are proportionately larger (Altman and Dittmer 1974; Fomon 1966; Fomon et al. 1982; Owen and Brozek 1966; Widdowson and Dickerson 1964). The infant also has an immature blood-brain barrier (Adinolfi 1985; Johanson 1980) and probably an immature blood-testis barrier (Setchell and Waites 1975). Many xenobiotic metabolizing enzymes have distinctive developmental patterns. At various stages of growth and development, levels of particular enzymes may be higher or lower than those of adults, and sometimes unique enzymes may exist at particular developmental stages (Komori et al. 1990; Leeder and Kearns 1997; NRC 1993; Vieira et al. 1996). Whether differences in xenobiotic metabolism make the child more or less susceptible also depends on whether the relevant enzymes are involved in activation of the parent compound to its toxic form or in detoxification. There may also be differences in excretion, particularly in newborns who all have a low glomerular filtration rate and have not developed efficient

tubular secretion and resorption capacities (Altman and Dittmer 1974; NRC 1993; West et al. 1948). Children and adults may differ in their capacity to repair damage from chemical insults. Children also have a longer remaining lifetime in which to express damage from chemicals; this potential is particularly relevant to cancer.

Certain characteristics of the developing human may increase exposure or susceptibility, whereas others may decrease susceptibility to the same chemical. For example, although infants breathe more air per kilogram of body weight than adults breathe, this difference might be somewhat counterbalanced by their alveoli being less developed, which results in a disproportionately smaller surface area for alveolar absorption (NRC 1993).

Selenium is known to be an essential micronutrient for humans and animals; therefore, inadequate as well as excessive selenium intake can cause adverse health effects. The Food and Nutrition Board of the National Research Council has established adequate intakes (AI) of 15–20 μg/day for infants based on the selenium content of milk of well nourished, but unsupplemented, mothers (NAS 2000). No data were available on which to base RDAs for children or adolescents and so the RDAs for children and adolescents are extrapolated from adult values. Studies of selenium deficient populations suggest that children are more susceptible to the effects of selenium deficiency and have the highest need for selenium of any individuals in the population (Chen et al. 1980; Yang et al. 1988).

Limited information is available on the toxicity of selenium in children, but the available information suggests that children may be less susceptible to toxic effects of selenium than adults. Most data come from children living in areas of chronic high dietary selenium intake (Yang et al. 1989a, 1989b). Children (aged 3–12 years) in a seleniferous area of China were found to have a significantly higher intake of selenium than the adults in their community, but a corresponding increase in blood levels of selenium appeared only in the children aged 7–12. When the incidence of selenosis in different age groups was examined, it was found that 97% of cases were older than 18 years, and no cases were observed in children below 12 years of age, even though selenium intakes per kg body weight and blood selenium levels in these age groups were found to be either higher than or equal to those of affected adults. One study of children living in a seleniferous area of Venezuela found a significant increase in the percentage of children showing lower than normal height compared with controls from a nonseleniferous area (Brätter et al. 1991a). However, these children also had very low intakes of zinc compared with controls (10–25% of controls) and it is likely that their reduced growth rate is due to inadequate intake of zinc. Another study that compared children from seleniferous and non-seleniferous areas of Venezuela

found slightly reduced height, weight, hemoglobin levels, and hematocrit values for the children from the seleniferous area (no statistical analysis was performed), although no clinical signs of selenosis were observed (Jaffe et al. 1972). However, the children from the seleniferous zone had a poorer diet, consumed less milk and meat, and had a greater incidence of intestinal parasites, which may account for the differences observed.

No adverse developmental effects of excess selenium have been reported for humans. Excess selenium is a demonstrated teratogen in birds (Franke and Tully 1935; Franke et al. 1936; Gruenwald 1958; Kahn and Gilani 1980; Palmer et al. 1973), but there is no clear evidence linking selenium exposures to developmental effects in mammals. Malformations have been reported for livestock that consumed naturally high seleniferous diets (Dinkel et al. 1963; Rosenfeld and Beath 1964), but it is not clear that these reports took into account consumption of other toxic range plants. Other studies of developmental effects in livestock receiving controlled diets with known amounts of selenium have generally not observed abnormalities, reduced birth weights, or increased mortality (Panter et al. 1995; Yaeger et al. 1998). Likewise, studies of laboratory animals have not observed developmental effects, except at levels of selenium administration that produce maternal toxicity (Bergman et al. 1990; Chiachun et al. 1991; Ferm et al. 1990; NTP 1996; Poulsen et al. 1989; Rosenfeld and Beath 1954; Schroeder and Mitchener 1971b; Thorlacius-Ussing 1990). In a teratology study of long-tailed macaques, no gross abnormalities or growth retardations were observed in fetuses from mothers administered doses that produced maternal toxicity.

No studies were located that compared pharmacokinetic properties of selenium in humans or animals of different ages. Selenium is transferred to fetuses via the placenta (Archimbaud et al. 1992; Choy et al. 1993; Hawkes et al. 1994; Jandial et al. 1976; Mahan and Kim 1996) and to infants via breast milk (Brätter and Negretti De Brätter 1996; Brätter et al. 1991b; Li et al. 1999; Michalke and Schramel 1998; Moser-Veillon et al. 1992; Rodríguez Rodríguez et al. 1999; Viitak et al. 1995; Yang 1989b). Studies of lactating women have shown a clear relationship between levels of selenium in the mother's diet and the concentration of selenium in her breast milk (Brätter et al. 1991b). However, there were no reports located of adverse effects in infants breast-fed by mothers in regions with high selenium diets.

3.8 BIOMARKERS OF EXPOSURE AND EFFECT

Biomarkers are broadly defined as indicators signaling events in biologic systems or samples. They have been classified as markers of exposure, markers of effect, and markers of susceptibility (NAS/NRC 1989).

Due to a nascent understanding of the use and interpretation of biomarkers, implementation of biomarkers as tools of exposure in the general population is very limited. A biomarker of exposure is a xenobiotic substance or its metabolite(s) or the product of an interaction between a xenobiotic agent and some target molecule(s) or cell(s) that is measured within a compartment of an organism (NAS/NRC 1989). The preferred biomarkers of exposure are generally the substance itself or substance-specific metabolites in readily obtainable body fluid(s), or excreta. However, several factors can confound the use and interpretation of biomarkers of exposure. The body burden of a substance may be the result of exposures from more than one source. The substance being measured may be a metabolite of another xenobiotic substance (e.g., high urinary levels of phenol can result from exposure to several different aromatic compounds). Depending on the properties of the substance (e.g., biologic half-life) and environmental conditions (e.g., duration and route of exposure), the substance and all of its metabolites may have left the body by the time samples can be taken. It may be difficult to identify individuals exposed to hazardous substances that are commonly found in body tissues and fluids (e.g., essential mineral nutrients such as copper, zinc, and selenium). Biomarkers of exposure to selenium are discussed in Section 3.8.1.

Biomarkers of effect are defined as any measurable biochemical, physiologic, or other alteration within an organism that, depending on magnitude, can be recognized as an established or potential health impairment or disease (NAS/NRC 1989). This definition encompasses biochemical or cellular signals of tissue dysfunction (e.g., increased liver enzyme activity or pathologic changes in female genital epithelial cells), as well as physiologic signs of dysfunction such as increased blood pressure or decreased lung capacity. Note that these markers are not often substance specific. They also may not be directly adverse, but can indicate potential health impairment (e.g., DNA adducts). Biomarkers of effects caused by selenium are discussed in Section 3.8.2.

A biomarker of susceptibility is an indicator of an inherent or acquired limitation of an organism's ability to respond to the challenge of exposure to a specific xenobiotic substance. It can be an intrinsic genetic or other characteristic or a preexisting disease that results in an increase in absorbed dose, a decrease in

the biologically effective dose, or a target tissue response. If biomarkers of susceptibility exist, they are discussed in Section 3.10 “Populations That Are Unusually Susceptible”.

3.8.1 Biomarkers Used to Identify or Quantify Exposure to Selenium

Selenium can be detected in the blood, feces, urine, hair, and nails of exposed individuals. Both selenium deficiency and excessive levels of selenium are associated with several disorders. For purposes of comparison, reported mean selenium concentrations in whole blood, blood constituents, urine, hair, nails, and the placenta for healthy individuals living in the United States and several other countries are listed in Table 3-7. The analytical methods used to measure selenium (described in Chapter 7) have improved, and the more recent studies may be more reliable. The values for the Chinese populations studied by Yang et al. (1983, 1989b) were those reported for individuals living in the selenium "adequate" regions included in the study. "Normal" selenium concentrations in blood constituents and other tissues in people from some countries (e.g., Finland) are generally lower than those in people living in the United States. In general, urinary excretion rates of 20–200 μg selenium/day are not associated with either selenium deficiency or toxicity (Sanz Alaejos and Diaz Romero 1993).

In the United States and other developed countries, hair selenium concentrations are not necessarily indicative of dietary exposure to environmental selenium. Users of therapeutic dandruff shampoos containing selenium sulfide may have high levels of selenium in their hair because the externally deposited selenium adsorbs to hair (Alfthan 1985). However, due to minimal levels of dermal absorption of selenium from shampoo, blood and urine levels are not significantly affected by selenium-containing shampoos (Howe 1979). Toenail samples have also been used as biomarkers of selenium exposure (Hunter et al. 1990a). Selenium levels in toenails were measured in volunteers who ate bread containing selenium for 1 year (Longnecker et al. 1993). During this time period, selenium in the large toenail did not reach a steady state, while a steady state was reached in the other toenails. After conclusion of the 1-year exposure, levels of selenium continued to decline until they reached baseline levels in 2 years.

At plasma and whole blood selenium concentrations of #0.10 mg selenium/L, a positive correlation has been reported between blood selenium levels and both erythrocyte and whole blood GSH-Px activity (Perona et al. 1977; Thomson 1977; Valentine et al. 1988). GSH-Px is an enzyme that acts as a scavenger of peroxides and protects cells from oxidative damage. However, whole blood selenium levels less than or equal to 0.10 mg selenium/L represent the lower end of the range of whole blood selenium concentrations reported by Allaway et al. (1968) for American males.

A correlation between blood selenium levels and GSH-Px activity was not observed when plasma and whole blood selenium levels were above 0.10 mg selenium/L. Therefore, GSH-Px activity is likely to be a biomarker for selenium deficiency but not for overexposure. Neve et al. (1988), on the other hand, found no relationship between erythrocyte or plasma GSH-Px activity levels and plasma selenium levels in a group of Belgian subjects with plasma selenium levels between 0.087 and 0.13 mg selenium/L. However, platelet GSH-Px activity levels did correlate with plasma selenium levels within this range (Neve et al. 1988). Valentine et al. (1980) measured the level of selenium in whole blood, urine, and hair of 33 residents from a Mexican village that consumed drinking water contaminated with selenium (0.026–1.8 mg selenium/L) from a uranium mill tailing pond. Blood levels ranging from 0.133 to 0.248 mg selenium/L, urine excretion rates ranging from 14.4 to 337.5 μg selenium/day, and hair selenium levels ranging from 0.02 to 1.98 μg selenium/g were not correlated with GSH-Px activity. In examining the relationship between selenium and GSH-Px activity, selenium-dependent GSH-Px activity must be distinguished from nonselenium-dependent GSH-Px activity (Edwards and Blackburn 1986).

Selenoprotein P which contains 10 selenocysteines, is the principal selenoprotein found in plasma (Sunde 1990). Selenoprotein P in plasma also does not continue to increase with increasing selenium and has been suggested as an alternative to GSH-Px as a biomarker for selenium status (Duffield et al. 1999; Huang et al. 1995). The function of selenoprotein P has still not been determined.

Field studies have used primarily blood or urine levels to indicate the degree of selenium exposure. Valentine et al. (1978) found a significant correlation between selenium levels in well water used for drinking and urine selenium excretion measured for 35 residents in a New Mexico community. However, no correlation was found between selenium levels in well water and the blood selenium levels of the 35 residents (Valentine et al. 1978). The correlation coefficients between the log of urine-selenium excretion (μg selenium/day) and the log of blood-selenium (mg selenium/L) with the log of the well water selenium concentration (mg selenium/L) were 0.57 ($p<0.01$) and 0.14 ($p>0.05$), respectively. The correlation coefficient between the log of hair selenium concentration (μg selenium/g) and the log of the well water selenium levels (mg selenium/L) was 0.45 ($p<0.01$).

Clinical symptoms have been associated with excessive blood, urine, and hair levels of selenium in exposed patients. Glover (1967) examined workers in a selenium rectifier factory and found that selenium levels in urine from workers exposed to selenium (annual averages from 1954 and 1958 range from 0.076 to 0.109 mg selenium/L urine) were higher than the average urine selenium levels of preemployment applicants (average, 0.034 mg selenium/L urine; range, 0–0.15 mg selenium/L). Garlic

breath, skin rashes, indigestion, lassitude, and irritability were noted, but no increase in mortality among exposed workers was detected. Smith and Westfall (1937) examined urine selenium levels in rural populations in Wyoming, South Dakota, and Nebraska and reported evidence of skin discoloration and lesions, tooth decay, diseased nails, gastrointestinal disturbances, and arthritis in individuals with urine selenium levels of 0.2–1.98 mg selenium/L; however, the authors did not show a significant correlation between clinical signs and the level of selenium in the urine. Longnecker et al. (1991) examined ranchers in the same area of the United States where selenosis of livestock had been observed. No clinical effects were observed with concentrations up to 2.2 mg/L in urine. Yang et al. (1983, 1989a, 1989b) measured mean blood, urine, and hair selenium levels of 3.2 mg selenium/L, 2.68 mg selenium/L, and 32.2 μg selenium/g, respectively, in a high selenium area where chronic selenosis was common in China. The clinical signs of selenium intoxication included loss of hair and nails, skin lesions, tooth decay, and nervous system disorders. In another area of China with high environmental levels of selenium but no signs of chronic selenosis in the population, blood selenium levels averaged 0.44 mg selenium/L (with a range from 0.35 to 0.58 mg selenium/L).

At blood levels of from 0.06 to 0.20 mg selenium/L, Deguchi (1985) found selenium to be positively correlated with grasping power and blood pressure in normal men and women and with hematocrit and hemoglobin concentrations in normal women. Similar correlations were not found in subjects with proteinuria or hypertension. In addition, Gebre-Medhin et al. (1988) found that in healthy children, serum selenium levels of 0.055–0.082 mg selenium/L were positively correlated with serum cholesterol, serum triglycerides, low and very low density lipoproteins, and apolipoproteins. Similar correlations were not found in diabetic children, who have slightly elevated serum selenium levels.

Biomarkers of Deficiency. Two endemic diseases, Keshan disease and Kashin-Beck disease, have been reported in selenium-deficient populations in China in which mean hair, blood, and urine selenium levels are low (Yang et al. 1988). Keshan disease, manifested as nausea, vomiting of yellowish fluid, and necrosis of the myocardium, has been found in a population with an average whole blood selenium concentration of 0.018 mg selenium/L, an average urinary concentration of 0.007 mg selenium/L, and an average hair selenium concentration of 0.123 μg/g (Yang et al. 1988). Kashin-Beck disease, which causes atrophy, degeneration, and necrosis of cartilage tissue, was observed in selenium-deficient areas in China in which the average selenium concentration in hair ranged from 0.077 to 0.165 μg selenium/g and blood selenium concentrations averaged approximately 0.02 mg selenium/L. In nonaffected areas in China, the selenium content in hair is greater than 0.2 μg selenium/g and in blood is greater than 0.06 mg selenium/L (Yang et al. 1988). These selenium-deficiency diseases are unlikely to occur in persons in the

United States. If selenium is not added to parenteral nutrition solutions, persons on long-term total parenteral nutrition are at risk for developing selenium deficiency symptoms which include cardiomyopathies, muscle pain, and weakness (Thomson 1991).

There is also some evidence that low serum selenium levels are associated with increased cancer risk, but this is not conclusive (Hojo 1981a; Willett et al. 1983). Salonen et al. (1984) concluded that an increased risk of cancer (a combination of gastrointestinal, respiratory, urogenital, hematologic, dermal, and skeletal cancers) in humans in Finland is associated with serum selenium levels of 0.045 mg selenium/L and below. Virtamo et al. (1987) found that cancer patients in Finland, including individuals with gastrointestinal, respiratory, skin, skeletal, urogenital, and hematological cancers, had slightly but not significantly lower serum selenium levels (mean and standard error of 0.0539±0.0015 mg selenium/L) compared with noncancer patients (0.0553±0.0005 mg selenium/L). However, serum selenium is generally an indicator only of very recent selenium status. As such, serum selenium may indicate an effect of cancer (malabsorption or anorexia) rather than a cause (Lockitch 1989; van't Veer et al. 1990).

A deficiency of selenium is also associated with cardiomyopathy (Johnson et al. 1981; Oster et al. 1983). Salonen et al. (1982) noted a statistically significant association between serum selenium concentrations of less than 0.045 mg selenium/L and the adjusted relative risk of coronary death, cardiovascular death, and myocardial infarction. Hojo (1981a) noted that patients with epilepsy had significantly lower urinary selenium levels than controls.

3.8.2 Biomarkers Used to Characterize Effects Caused by Selenium

Specific biomarkers of selenium effects were not found. Garlic breath is a marker of over-exposure to selenium compounds. However, as other metals that are methylated (e.g., arsenic) also result in garlic odor of the breath, this effect is not a unique marker of selenium over-exposure. Hair and nail effects may be the most frequent effects of overexposure to selenium. Hair becomes dry and brittle and breaks off at the scalp. Nails are also brittle and have white spots and longitudinal streaks, and break off easily (Lockitch 1989). Although these effects may not be specific to selenium, if they are observed, a determination of selenium status may be useful.

Yang et al. (1989b) have used increased prothrombin time (increased clotting time), a measure of hepatic damage, as a biomarker for selenium but their interpretation of their observations may be unwarranted. The difference they saw in affected humans was very small (1 second); prothrombin time has not been

previously demonstrated to correlate with symptoms of selenosis nor used to detect selenosis; finally, since the test has not been widely used, the results reported for the small number of affected individuals may be within the range of normal values for the general population or a subpopulation (IRIS 1996).

In humans and in animal studies, high concentrations of selenium have been demonstrated to cause neurological effects. Biomarkers of effect for the neurological system have been reviewed by ATSDR (OTA 1990).

3.9 INTERACTIONS WITH OTHER CHEMICALS

A wide variety of interactions of selenium with essential and nonessential elements, vitamins, xenobiotics, and sulfur-containing amino acids have been demonstrated in numerous studies. Selenium has been reported to reduce the toxicity of many metals including mercury, cadmium, lead, silver, and to some extent, copper (Frost 1972). Most forms of selenium and arsenic interact to reduce the toxicity of both elements (Levander 1977). Because of selenium's role in the antioxidant glutathione peroxidase enzymes, selenium also reduces the toxicity of metals in vitamin E–deficient animals (Diplock et al. 1967).

The interactions of selenium with other elements and compounds are complex and not well understood (Naganuma et al. 1983; NAS 1976a). The degree to which selenium is toxic, is taken up by tissues, or is excreted can be influenced by these interactions. Some of the major interactions of selenium compounds with other elements and compounds are described below.

Arsenic. In general, arsenic antagonizes selenium toxicity (Levander 1977). This effect extends to selenium in sodium selenite and selenate, seleniferous wheat, selenocystine, and selenomethionine (Levander 1977). However, a very pronounced synergistic toxicity exists between arsenic and two methylated selenium metabolites, trimethylselenonium ion and dimethyl selenide (Obermeyer et al. 1971). One of the more striking demonstrations is the antagonism of arsenic-induced terata in rodents by concomitant selenium exposure (Holmberg and Ferm 1969), and pretreatment of mice with sodium selenite reduced the clastogenic effects of a subsequent dose of sodium arsenite (Biswas et al. 1999b). Moxon et al. (1945) found that arsenic could reduce selenium toxicity when compounds of both elements were injected subcutaneously, thereby indicating that arsenic did more than interfere with the gastrointestinal absorption of selenium. Kamstra and Bonhorst (1953) found that arsenic reduced the excretion of volatile selenium compounds in expired air following the injection of compounds of both

elements into rats at acutely toxic levels. Levander and Baumann (1966a) found that the amount of selenium retained in the liver decreased and the amount of selenium appearing in the gastrointestinal tract increased as the dose of administered arsenic was increased. Experiments with rats and guinea pigs with cannulated bile ducts confirmed that arsenic increased the biliary excretion of selenium and that selenium increased the biliary excretion of arsenic (Levander and Baumann 1966b). It has recently been suggested that the mutual reduction in toxicity of arsenic and selenium administered together is due to the formation of an arsenic-selenium compound, seleno-bis(S-glutathionyl)arsinium (Gailer et al. 2000b). This compound was isolated from the bile of rabbits injected with selenium and arsenic and identified by X-ray spectroscopy.

Cadmium. Selenium can antagonize the nephrotoxic and hepatotoxic effects of cadmium in rats (Flora et al. 1982; Lindh et al. 1996; Nehru and Bansal 1996; Stajn et al. 1997), the inflammation, atrophy, and necrosis induced by cadmium in testes of rats (Jones et al. 1997; Mason and Young 1967; Ohta and Imamiya 1986; Wlodarczyk et al. 1995; Yiin et al. 1999), and the cardiotoxicity of cadmium in rats (Jamall et al. 1989). The protective effects are thought to occur as a result of the formation of a selenium-cadmium complex of high molecular weight (Chen et al. 1975; Jamall et al. 1989; Jamba et al. 1997; Ohta and Imamiya 1986).

Fluoride. Fluoride ion may interact with selenium; however, the degree and types of interaction depend upon the chemical form of selenium (i.e., organic or inorganic) and the dose. Moxon and DuBois (1939) reported that fluoride increased the toxicity of selenium in rats at 5 mg fluoride/L in the drinking water of young rats fed a diet containing 11 ppm selenium (0.55 mg selenium/kg/day) as seleniferous wheat. Selenium decreased growth and increased mortality in rats drinking fluoridated water compared to rats drinking deionized water. These results were disputed by Hadjimarkos (1969a) who administered 3 mg selenium/L as sodium selenite (0.15 mg selenium/kg/day) either with or without 50 mg fluoride/L as sodium fluoride in the drinking water of rats. The growth and mortality data indicated that the combined administration of selenium and fluoride under the conditions used did not increase selenium toxicity. However, the amount of administered fluoride was significantly higher and the amount of administered selenium significantly lower in the Hadjimarkos (1969a) study than the amounts administered by Moxon and DuBois (1939). No additional studies were located that reexamined the possible interaction between fluoride and selenium.

Iodine. Iodine and selenium interact to affect thyroid function. There are at least two aspects to this interaction. First, selenium is an important component of the iodothyronine 5'-deiodinases which

functions in the control of circulating triiodothyronine. Second, during the production of the thyroid hormone, hydrogen peroxide is produced, which is detoxified by GSH-Px. One possible consequence of the interaction between iodine and selenium has been noted in human populations deficient in iodine. In some iodine-deficient regions, a reversible hypothyroidism with goiter formation is observed. In other iodine-deficient areas, hypothyroidism is accompanied by thyroid gland destruction. The thyroid gland destruction is thought to result in populations that are deficient in both iodine and selenium (Contempré et al. 1991a, 1992, 1995). This hypothesis has been supported by a study in rats which found that iodine and selenium deficiency increased necrosis and induced fibrosis in the thyroid gland. Supplementation with iodine in selenium-deficient rats resulted in greater toxicity, possibly as a result of hydrogen peroxide which could not be detoxified because of the lack of GSH-Px (Contempré et al. 1995). This study suggests that iodine supplements may produce adverse effects in selenium-deficient individuals.

Mercury. Simultaneous administration of mercury and selenium in equimolar doses to animals resulted in decreased toxicity of both elements in acute and chronic studies with inorganic and organic mercury and with either inorganic or organic selenium compounds, although inorganic forms of selenium appear to be more effective than organic forms (Chang 1983; Rao et al. 1998; Skerfving 1978). Selenium protects against the acute nephrotoxicity of the mercuric ion and methylmercuric ion in rats (Ganther et al. 1972; Hansen 1988; Magos et al. 1987; Parizek and Ostadalova 1967) and possibly against acute neurotoxicity of the methylmercuric ion in rats (Ohi et al. 1980). The protective effect of selenium has been associated with a higher whole body retention of mercury rather than with increased mercury excretion (Hansen 1988; Magos et al. 1987). Selenium has been shown to inhibit biliary excretion of methyl mercury in rats (Urano et al. 1997), while mercury exposure reduces urinary selenium excretion in humans (Ellingsen et al. 1995). Although the mechanism of the interaction has not yet been elucidated, selenium and mercury appear to form a metabolically inert compound by reaction with GSH (Gailer et al. 2000b). Further support for the role of this compound comes from the observation that selenium-treated animals can remain unaffected despite an accumulation of mercury in tissues to levels that are otherwise associated with toxicity (Skerfving 1978). Additional support comes from the 1:1 ratio of selenium and mercury found in the livers of marine mammals and in the bodies of experimental animals injected with mercury and selenium, regardless of the ratio of the administered doses (Hansen 1988).

Although the fetotoxicity of methylmercuric chloride has been enhanced in selenium-deficient mice (Nishikido et al. 1987), additional selenium administration does not appear to protect against teratogenic effects (i.e., cleft palate) of methylmercuric chloride in mice (Lee et al. 1979). High doses of selenium administered as selenite for 30 days prior to gestation and through gestation day 18 to mice fed a diet

containing high doses of methylmercuric chloride increased the incidence of cleft palate (Nobunaga et al. 1979). Concurrent treatment of pregnant or lactating mice receiving nontoxic doses of methyl mercury in drinking water with selenomethionine increased the deposition of mercury in the offspring (Nielsen and Andersen 1995).

Methionine and Vitamin E. Combinations of methionine and vitamin E have been found to be antagonistic to selenium toxicity. In one study, selenium concentrations in the liver and kidneys of rats fed selenium-containing diets with methionine and vitamin E were less than the concentrations found in the livers and kidneys of rats fed selenium with either methionine or vitamin E alone (Levander and Morris 1970). The results are compatible with the hypothesis that methionine detoxifies selenium by forming methylated derivatives of selenium that are eliminated in the urine and in expired air (see Section 3.4.4) or that methionine and selenomethionine are in the same pool of amino acids and that by increasing the amount of methionine relative to selenomethionine, the probability of selenomethionine being randomly inserted into proteins during synthesis decreases (Stadtman 1977, 1980, 1983, 1987, 1990). As discussed in Section 3.11, methionine administered as an antidote for acute selenium toxicity in rats was ineffective (Lombeck et al. 1987). This result supports a mechanism of action involving protein synthesis rather than a methylation mechanism hypothesis.

Silver. Selenium has been shown to be protective against the hepatotoxic effects of silver in vitamin E–deficient rats. A 0.15% solution of silver acetate in the drinking water of rats produced necrotic degeneration of the liver and high mortality. Dietary selenium supplementation at 1 mg selenium/kg food resulted in a significant reduction in the toxic effects of silver (Diplock et al. 1967). One report indicates a nontoxic dose of silver acetate in rats minimizes effects of acute selenium toxicity. However, the body burden of selenium in several organs increased with treatment with silver acetate. It is postulated that this antagonistic effect may be due to the formation and disposition of silver selenides which are relatively insoluble and nontoxic (Eybl et al. 1992).

Sulfate. Sulfate appears to reduce the growth inhibition that results from dietary exposure of rats to high levels of selenite or selenate (Halverson and Monty 1960). Sulfate does not appear to be protective against selenium-induced liver damage (Halverson and Monty 1960).

Antagonistic interactions with several additional metals including antimony, germanium, and bismuth have been reported (Paul et al. 1989). Complex interactions of selenium with other metals, vitamins, and nutrients usually lead to a reduced toxicity of selenium and/or a reduced toxicity of the interacting

substance. However, vitamin C (ascorbic acid) may increase the absorption and toxic effects of selenium in humans (HSDB 2001; Lombeck et al. 1987; Mack 1990; Martin et al. 1989a, 1989b). The relevance of these interactions to selenium exposure of the general public is unknown. Many review articles are available concerning the interactions of selenium and other chemicals, including those by Combs, Jr., and Combs (1987), Hansen (1988), Levander and Morris (1970), Magos and Webb (1980), Naganuma et al. (1983), and Whanger (1981).

3.10 POPULATIONS THAT ARE UNUSUALLY SUSCEPTIBLE

A susceptible population will exhibit a different or enhanced response to selenium than will most persons exposed to the same level of selenium in the environment. Reasons may include genetic makeup, age, health and nutritional status, and exposure to other toxic substances (e.g., cigarette smoke). These parameters result in reduced detoxification or excretion of selenium, or compromised function of organs affected by selenium. Populations who are at greater risk due to their unusually high exposure to selenium are discussed in Section 6.7, Populations With Potentially High Exposures.

Data concerning human subpopulations with unusual susceptibility to the toxic effects of selenium were not located. Epidemiologic studies have identified populations with very low or very high nutritional status, and these groups are expected to have very different responses to selenium exposures. Pregnant and nursing women are believed to require more selenium than the general public (NRC 1989).

It is possible that persons exposed to high fluoride levels in drinking water might be at greater risk of adverse health effects from exposure to excessive levels of selenium (Moxon and DuBois 1939; Yang et al. 1989a), but evidence on this point is equivocal (Hadjimarkos 1969a) and requires further study. Individuals with vitamin E–deficient diets might also be at greater risk of liver damage from exposure to excess selenium (Levander and Morris 1970). Based on studies of chemically induced diabetes in rats, selenium may change insulin needs (McNeil et al. 1991). Therefore, insulin-dependent diabetics may be more sensitive to adverse health effects due to selenium exposure than the general population.

Cretins or other individuals with iodine or thyroid deficiencies may be more sensitive to adverse health effects from selenium exposure (Contempré et al. 1991b, 1992). Iodine supplementation of these individuals without selenium supplementation may further exacerbate the effects. The elderly may be less susceptible to the negative effects of selenium and more prone to selenium deficiencies. A number of

researchers have reported lower absorption of selenium and lower selenium tissue concentrations in the elderly compared to younger adults (Martin et al. 1991; Morisi et al. 1989).

Populations living in the western United States in areas eating produce grown in highly seleniferous soils could be at greater risk of adverse health effects from additional environmental exposure to selenium if their selenium nutritional status is already high (see Section 6.6).

3.11 METHODS FOR REDUCING TOXIC EFFECTS

This section will describe clinical practice and research concerning methods for reducing toxic effects of exposure to selenium. However, because some of the treatments discussed may be experimental and unproven, this section should not be used as a guide for treatment of exposures to selenium. When specific exposures have occurred, poison control centers and medical toxicologists should be consulted for medical advice. The following texts provide specific information about treatment following exposures to selenium:

Nadig RJ. 1994. Cadmium and other metals and metalloids. In: Goldfrank LR, Weisman RS, Flomenbaum N, et al. eds. Goldfrank's toxicological emergencies. 6th ed. Norwalk, CT: Appleton and Lange, 1342-1343.

Mofenson HC, and Caraccio TR. 1998. Toxicity of household products. In: Viccellio P, ed. Emergency toxicology. 2nd ed. Philadelphia, PA: Lippincott-Raven, 519.

3.11.1 Reducing Peak Absorption Following Exposure

No specific recommendations have been reported for reducing absorption following acute high-dose exposure to selenium or selenium compounds via inhalation or dermal exposure (Gosselin et al. 1984; HSDB 2001). There have been very few reported cases of overexposure via inhalation in industrial settings but some have resulted in toxic effects (Lockitch 1989). General procedures suggested for reducing absorption following accidental industrial exposure include moving the exposed person into fresh air, removing contaminated clothing and shoes, and flushing exposed skin or eyes with running water (HSDB 2001).

Oral exposures to toxic quantities of selenious acid, sodium selenate, and selenium dioxide have been reported (Lockitch 1989). In general, only supportive treatment has been recommended (HSDB 2001; Mack 1990). In some cases, gastric lavage and induction of vomiting by use of emetics have been

reported to be useful in reducing absorption, but because selenious acid (in gun blueing, pH 1) is caustic, both procedures could result in additional damage by this compound (Lombeck et al. 1987; Mack 1990). The possibility of a sudden onset of shock, seizures, severe hypotension, and cardiorespiratory arrest has been used to argue against emesis (Mack 1990). It has also been suggested that oils and alcohol are to be avoided in treatment of ingested selenium sulfide because these agents may increase absorption (Gosselin et al. 1984).

3.11.2 Reducing Body Burden

In acute exposure situations, selenium compounds are rapidly absorbed and widely distributed throughout many organ systems following inhalation or ingestion (see Section 3.4.2). Extensive parenteral fluid administration has been used to force the urinary excretion of selenium (Lombeck et al. 1987). Chelating agents have not been effective in experiments, and both calcium disodium ethylene diamine tetraacetate (EDTA) and dimercaprol (British Anti-Lewisite, BAL) may increase the toxic effects of selenium (Lombeck et al. 1987; Mack 1990; Paul et al. 1989). Although vitamin C (ascorbic acid) is used to reduce the body burdens of other metals, it may also increase the absorption and toxic effects of selenium in humans (HSDB 2001; Lombeck et al. 1987; Mack 1990; Martin et al. 1989a, 1989b). Bromobenzene has been reported to increase the urinary excretion of selenium, but because bromobenzene is also a hepatic toxin, its use is dangerous (Gosselin et al. 1984; HSDB 2001).

3.11.3 Interfering with the Mechanism of Action for Toxic Effects

The exact molecular mechanism of toxic action by selenium and selenium compounds is not known. One theory is that at a biochemical level, selenium inactivates sulfhydryl enzymes leading to depression of cellular oxidative processes (Lombeck et al. 1987; Mack 1990; Shamberger 1981). No information was located on established therapies designed to interfere with this possible mechanism of action of selenium. Because selenomethionine is known to randomly insert into proteins, rats were treated with methionine after acute selenosis had developed, but no effect was observed (Lombeck et al. 1987). However, pretreating rats with dietary methionine and vitamin E reduced the toxicity of dietary selenium as measured by decreased liver damage, reduced body weight gain, and decreased liver and kidney concentrations of selenium compared to those in rats that had not received supplements (Levander and Morris 1970). Inorganic sulfate fed simultaneously with selenite or selenate in the diet protected rats from the toxicity of selenium as measured by body weight gain; however, sulfate did not protect against

liver necrosis caused by selenium (Halverson et al. 1962). It would, therefore, seem plausible that another nontoxic sulfur-containing chemical could be found to be effective against acute selenium toxicity.

The search for an agent that both reduces the acute toxicity of selenium and increases the excretion of the selenium compound formed has proved difficult (Paul et al. 1989). In some experimental cases, other metals have been shown to mitigate the toxicity of selenium, possibly by forming metal selenides with low solubility and toxicity (see Section 3.9). Several metal-containing compounds were tested for efficacy in reducing toxic effects and increasing elimination of selenium from sodium selenate injected into rats. Germanium citrate is nontoxic and was found to be effective both at reducing toxic effects and increasing the rate of selenium elimination. However, the germanium compound, bis-carboxyethyl germanium sesquioxide, had no positive effect on toxicity or distribution to organs but did increase the amount of selenium excreted in the urine (Paul et al. 1989). In mice, pretreatment with a nontoxic dose of silver acetate was shown to reduce the toxic effects of sodium selenite. However, this treatment increased the whole body burden of selenium, and the concentrations in several organs were raised compared to those in the controls injected with sodium selenite only (Eybl et al. 1992).

3.12 ADEQUACY OF THE DATABASE

Section 104(i)(5) of CERCLA, as amended, directs the Administrator of ATSDR (in consultation with the Administrator of EPA and agencies and programs of the Public Health Service) to assess whether adequate information on the health effects of selenium is available. Where adequate information is not available, ATSDR, in conjunction with the National Toxicology Program (NTP), is required to assure the initiation of a program of research designed to determine the health effects (and techniques for developing methods to determine such health effects) of selenium.

The following categories of possible data needs have been identified by a joint team of scientists from ATSDR, NTP, and EPA. They are defined as substance-specific informational needs that if met would reduce the uncertainties of human health assessment. This definition should not be interpreted to mean that all data needs discussed in this section must be filled. In the future, the identified data needs will be evaluated and prioritized, and a substance-specific research agenda will be proposed.

3.12.1 Existing Information on Health Effects of Selenium

The existing data on health effects of inhalation, oral, and dermal exposure of humans and animals to selenium are summarized in Figure 3-10. The purpose of this figure is to illustrate the existing information concerning the health effects of selenium. Each dot in the figure indicates that one or more studies provide information associated with that particular effect. The dot does not necessarily imply anything about the quality of the study or studies, nor should missing information in this figure be interpreted as a "data need". A data need, as defined in ATSDR's *Decision Guide for Identifying Substance-Specific Data Needs Related to Toxicological Profiles* (ATSDR 1989), is substance-specific information necessary to conduct comprehensive public health assessments. Generally, ATSDR defines a data gap more broadly as any substance-specific information missing from the scientific literature.

As seen in Figure 3-10, very little quantitative information is available regarding the health effects in humans exposed to selenium compounds via inhalation. The only quantitative inhalation studies in humans that relate selenium exposure levels or selenium body levels to health effects following inhalation exposure are epidemiological cancer studies. Fatalities following inhalation exposure to selenium compounds have not been reported. Despite the large number of cases of reported inhalation exposures in occupational settings, characterization of exposure concentrations and the selenium compounds present in the air are generally lacking. It is therefore not possible to link the degree and types of symptoms reported in workers to selenium exposure levels. There have been no reports of immunological, developmental, reproductive, or genetic effects in humans resulting from inhalation exposure to selenium compounds. Complaints of dizziness and fatigue have accompanied occupational inhalation exposures, but characterization of the exposure levels required to produce neurological symptoms is lacking.

Most of the information concerning the health effects in humans following exposure to selenium and selenium compounds is for the oral exposure route. However, exposure levels associated with the few documented fatalities resulting from accidental or suicidal poisoning with selenium compounds are lacking, as are exposure levels for other nonfatal poisonings by ingestion. A series of epidemiological studies in China have provided the only data about chronic exposure levels to excess dietary selenium that resulted in adverse effects on skin, nails, and hair and in possible neurological effects.

Older reports from the western United States described similar symptomology in the 1930s but did not characterize daily selenium intake. More recent reports show no clinical symptoms in the same area. The

Figure 3-10. Existing Information on Health Effects of Selenium

	Death	Acute	Intermediate	Chronic	Immunologic/Lymphoretic	Neurologic	Reproductive	Developmental	Genotoxic	Cancer
		Systemic								
Inhalation		●				●				●
Oral	●	●		●	●	●	●			●
Dermal		●								

Human

	Death	Acute	Intermediate	Chronic	Immunologic/Lymphoretic	Neurologic	Reproductive	Developmental	Genotoxic	Cancer
		Systemic								
Inhalation	●	●								
Oral	●	●	●	●	●	●	●	●	●	●
Dermal					●					●

Animal

● Existing Studies

possible inverse relationship between dietary selenium intake and the risk of various types of cancer has been examined in numerous epidemiological studies in the United States and in other countries.

Concern for the dermal route of exposure to selenium compounds as a cause of adverse health effects in humans is extremely low except for the acid forms, which owe their dermal effects to their acidity more than to their selenium content. Selenium sulfide, an ingredient in some antidandruff shampoos, does not appear to be absorbed through the skin. Ingestion of large amounts of the compound, however, would be of concern because selenium sulfide has been shown to be carcinogenic in rats and mice following oral exposure.

Data are available for acute inhalation exposures for a few of the volatile selenium compounds that have resulted in the death of animals. These exposures also produced signs of central nervous system toxicity, lung injury, and possible damage to heart and liver. No studies were located concerning health effects in animals following intermediate or chronic inhalation exposures to volatile selenium compounds or selenium dust.

In animals, the focus on the oral toxicity of selenium has taken two routes, one in laboratory animals and the other in studies of selenium toxicity to livestock. In laboratory animals, attention has been directed toward the hepatotoxic properties of selenites, selenates, and selenium contained in grains following early reports that selenium produced hepatic carcinomas in rats. An intermediate-duration study has also shown that selenate and selenite can cause kidney effects in rats while mice are less sensitive to this effect of selenium compounds. In recent years, much of the research in laboratory animals using the oral route of administration of selenium compounds has been directed toward the anticarcinogenic properties of selenium compounds.

In livestock, concern for selenium toxicity and deficiency is high. In areas of the country with selenium-poor soils, dietary selenium supplementation for livestock has been necessary to prevent chronic selenium deficiency diseases. Dietary supplementation programs have resulted in cases of accidental poisonings from misuse of the selenium supplements (Hopper et al. 1985).

3.12.2 Identification of Data Needs

Acute-Duration Exposure. The primary target organ in humans following acute exposure to high concentrations of selenium by inhalation or oral routes is the lung, with cardiovascular, hepatic, and renal systems all affected (lesser systemic effects were observed in all other organ systems except the musculoskeletal system) (Carter 1966; Civil and McDonald 1978; Clinton 1947; Koppel et al. 1986; Wilson 1962). Two case reports of acute dermal exposure were also located; the results revealed effects on the skin and eyes (Middleton 1947; Pringle 1942). Additional epidemiological or occupational studies would be useful to further characterize the effects of acute exposure via all routes and to confirm the target organ data.

Studies regarding single inhalation or oral exposures of rats, guinea pigs, rabbits, and mice have provided information on lethal levels of exposure to selenium compounds (Cummins and Kimura 1971; Dudley and Miller 1941; Hall et al. 1951; Miller and Williams 1940; Olson 1986; Smyth et al. 1990). However, few levels at which sublethal effects first appear have been identified. Clinical observations and gross necropsies have been performed, but no single-dose exposure study has included internal examination of the animals to identify dose-response data for sublethal systemic toxic effects. Such studies might provide information on the thresholds for systemic toxicity following single-dose exposure. Repeated inhalation exposure studies in animals are limited to a few days of exposure (Hall et al. 1951). Although the studies have demonstrated cumulative toxicity following repeated inhalation exposure to inorganic selenium compounds, effects other than lethality have been poorly characterized. Single-dose exposure studies have been conducted with selenium monosulfide in mice (target systems: respiratory and neurological) and selenium disulfide in rats (target organ not specified); however, the results have varied and there is uncertainty about which or how much of each of the compounds was administered. There were no effects in mice following acute dermal exposure. Additional dermal exposure studies in animals would be useful to confirm the effects found in humans. The data were insufficient for the derivation of acute oral and inhalation MRLs.

Intermediate-Duration Exposure. No human studies of intermediate inhalation exposure to selenium were located. Following oral exposure, one study in humans revealed endocrine effects in iodine-deficient individuals (Contempré et al. 1991a, 1992) and others revealed endocrine effects in individuals receiving sufficient levels of iodine (Duffield et al. 1999; Hawkes and Keim 1995). Results from one study in humans revealed dermal effects following intermediate dermal exposure (Pringle

1942). There were insufficient data to derive intermediate MRLs. Additional epidemiological or occupational studies would be useful in elucidating the potential target organs and effect levels.

No intermediate inhalation studies were located in animals. Intermediate-duration inhalation studies, in which selenium is administered as selenium dioxide, hydrogen selenide, or selenium dust, might help to identify air concentrations of these substances that produce sublethal effects not only on the respiratory system, but also on the hepatic, renal, hematological, and cardiovascular systems. As exposure to the selenoamino acids is via ingestion, inhalation studies of these compounds would not be necessary.

Intermediate-duration oral exposure studies have been performed with rats, pigs, mice, and monkeys at several dose levels using several selenium compounds (Baker et al. 1989; Behne et al. 1992; Bioulac-Sage et al. 1992; Chen et al. 1993; Cukierski et al. 1989; Das et al. 1989b; Eder et al. 1995; Halverson et al. 1966; Hasgawa et al. 1994; Hotz et al. 1997; Mahan and Magee 1991; Mihailovic et al. 1992; NTP 1980c, 1994; Palmer and Olson 1974; Panter et al. 1996). The major effects were hepatic, dermal, endocrine and neurological. Additional studies are needed to confirm these data. No intermediate-duration dermal administration studies have been conducted with the environmental forms of inorganic selenium likely to be of concern (e.g., sodium selenate and sodium selenite), although it is unlikely that these forms would be dermally absorbed to a significant degree. Dermal application of selenomethionine to the skin of mice did not result in any direct effects on the skin, or other signs of toxicity, although it was absorbed (Burke et al. 1992b). The organic compounds of selenium are usually not free in the environment but, rather, contained in plant and animal material. Therefore, no further dermal studies would be useful.

Chronic-Duration Exposure and Cancer. Several occupational studies of chronic inhalation exposure to inorganic selenium compounds were located (Glover 1967; Holness et al. 1989; Kinnigkeit 1962). Effects reported in these studies were primarily respiratory, although cardiovascular, gastrointestinal, hematological, musculoskeletal, dermal, ocular, and neurological effects were also noted. Animal data are not available for inhalation exposures of chronic duration. Data in this area would be helpful to establish an animal model for respiratory effects of inorganic selenium compounds, since most human exposure has been occupational and to a variety of compounds. Neurological effects have been documented in animals after chronic oral exposure, but further study of neurological effects in animals after inhalation exposure is needed to provide a model for the effects observed after occupational exposure in humans. Following chronic oral exposure, the primary effects in humans were dermal, neurological and endocrine (Brätter and Negretti De Brätter 1996; Clausen et al. 1989; Longnecker et al.

1991; Yang et al. 1983, 1989a, 1989b; Yang and Zhou 1994). An MRL has been derived for chronic oral exposure to selenium based on a NOAEL for dermal effects. One case report of chronic dermal exposure revealed dermal effects (Senff et al. 1988). Additional epidemiological or retrospective studies of chronic exposure would be helpful for confirming the existing data. Studies examining the role of nutrition in selenium toxicity would be especially useful.

Although the lung does not appear to be a target organ in animals after chronic oral exposure to selenium compounds, data have not been adequately reported (Harr et al. 1967; Henschler and Kerschner 1969; Schroeder and Mitchener 1972), and further studies might be useful to fully rule out these effects. Studies examining possible gastrointestinal and musculoskeletal effects in animals after chronic exposure to selenium or selenium compounds or to seleniferous grains might be helpful in determining the mechanisms of alkali disease whose symptoms have been observed in grazing livestock (Harr et al. 1967; Shamberger 1986). Hepatic and renal lesions following chronic selenium exposure have been adequately characterized. Investigations of systemic effects associated with chronic oral administration of selenium compounds, however, have been limited.

No studies were located regarding carcinogenic effects in animals after chronic inhalation exposure to selenium or selenium compounds. No further investigation is needed since humans have not been shown to have an increased risk of malignancy from selenium exposure. The majority of oral studies have provided information on the absence of carcinogenic effects in humans and animals (Beems 1986; Clark et al. 1996a, 1999; Coates et al. 1988; Harr et al. 1967; Menkes et al. 1986; Thompson and Becci 1979; Virtamo et al. 1987). However, earlier and less complete studies had suggested that selenium was carcinogenic following oral exposure of animals (Nelson et al. 1943; Schroeder and Mitchener 1971a; Volgarev and Tscherkes 1967). Chronic oral exposure studies conducted in mice and rats by gavage administration of a mixture of selenium monosulfide and selenium disulfide produced liver tumors in rats and lung tumors in female mice (NTP 1980c). The relative proportion of the two compounds was not clear, although physical evidence suggested that the dose solution was primarily selenium monosulfide. Further studies utilizing selenium sulfides might be useful in determining possible effects in humans.

Genotoxicity. Chromosomal aberrations and sister chromatid exchanges in lymphocytes were not increased in humans treated (oral or intramuscular injection) with sodium selenite (Norppa et al. 1980a). Compared to untreated controls, a significant increase in the number of micronuclei was observed in bone marrow cells of mice treated orally with selenite or selenate, and macaques treated orally with L-selenomethionine (Biswas et al. 1997, 1999a; Choy et al. 1989; Itoh and Shimada 1996; Rusov et al.

1996). A significant increase in the number of micronuclei in bone marrow cells was not observed in the offspring of macaques treated with L-selenomethionine on gestation days 20–50 (Choy et al. 1993).

Genotoxicity studies (*Salmonella*/microsome assays, sister chromatid exchange, and tests of unscheduled DNA synthesis and of chromosome aberrations in cultured mammalian cells) indicate that selenite, selenate, and selenide have both genotoxic and antigenotoxic effects (Biswas et al. 1997, 2000; Gairola and Chow 1982; Khalil 1994; Lu et al. 1995b; Schillaci et al. 1982; Ueda et al. 1997; van der Lelie et al. 1997). The underlying mechanisms responsible for the varying genotoxicity results remain to be elucidated.

Reproductive Toxicity. One study that measured the concentration of selenium in sperm samples indicated no correlation between selenium concentrations and sperm count or motility (Roy et al. 1990). No significant increase in spontaneous abortions was reported among women chronically exposed to drinking water containing 7–9 μg selenium (Vinceti et al. 2000a). No other human studies were located. A few reproductive toxicity studies in animals (Chowdhury and Venkatakrishna-Bhatt 1983; Harr and Muth 1972; NTP 1996; Schroeder and Mitchener 1971b; Wahlstrom and Olson 1959b) indicate that oral exposure to excess sodium selenite can reduce female fertility, although male fertility appears not to be affected. Oral treatment of rats with sodium selenate or selenite has been shown to increase the number of abnormal sperm in males (El-Zarkouny et al. 1999; Kaur and Parshad 1994; NTP 1994), produce testicular hypertrophy (Turan et al. 1999a), and affect the estrous cycle (NTP 1994, 1996). Fertility was not examined in these studies. Selenium dioxide produced testicular degeneration following intraperitoneal administration to rats (Chowdhury and Venkatakrishna-Bhatt 1983). Disturbances in the menstrual cycle (anovulation, short luteal and follicular phases) were observed in monkeys treated orally with L-selenomethionine (Cukierski et al. 1989) and mice treated orally with sodium selenite (Nobunaga et al. 1979). Studies of both male and female reproductive toxicity of selenium following oral and inhalation exposure in rats and other mammals to selenium dioxide and other forms of selenium, both organic and inorganic, would be useful. Such studies could provide information regarding the reproductive effects of the various forms of selenium that might be encountered in occupational settings, at waste sites, and in the drinking water and food from highly seleniferous areas of the United States.

Developmental Toxicity. No developmental studies were found regarding inhalation or dermal exposure in humans or animals. Developmental studies using the oral route of administration indicate that excessive sodium selenate or sodium selenite intake can result in fetal toxicity and reduced growth in experimental mammals (Dinkel et al. 1963; Ferm et al. 1990; NTP 1996; Rosenfeld and Beath 1964;

Wahlstrom and Olson 1959a), but generally only at doses that produce maternal toxicity. Developmental effects were not observed in macaque fetuses from mothers given toxic oral doses of L-selenomethionine during gestation (Tarantal et al. 1991). Intravenous injection of sodium selenite in mice did not indicate that the compound is teratogenic in rodents (Yonemoto et al. 1984). Intravenous injections of sodium selenate, D,L-selenomethionine, and D,L-selenocystine into neonatal rats indicated that some selenium compounds can contribute to the formation of one type of cataracts (Ostadalova and Babicky 1980). Cataracts were not observed in the offspring of macaques treated orally with L-selenomethionine during gestation (Tarantal et al. 1991). Additional developmental toxicity studies of selenium compounds in mammals do not seem to be necessary at this time.

Immunotoxicity. No studies were located regarding adverse immunological effects in humans following inhalation or oral exposure. One case report describes immunological effects following dermal exposure (Senff et al. 1988). Animal studies of possible adverse immunological effects from excessive exposure to selenium compounds are limited (Dudley and Miller 1941; Glenn et al. 1964a; Hall et al. 1951; Smyth et al. 1990). One study (Koller et al. 1986) included a battery of immunological tests, some of which indicated beneficial effects of sodium selenite administration and others that indicated adverse effects. Additional immunotoxicity tests, including challenges of the immune system, might characterize the significance of the different immunological effects that have been observed following selenium administration.

Other than selenium sulfide, an ingredient in some antidandruff shampoos, selenium compounds have not been tested for sensitization. The potential for dermal contact by humans does exist, however, in occupational settings and to a lesser extent in soil at waste sites.

Neurotoxicity. Data from an epidemiological study of humans and from studies in livestock indicate that the central nervous system is an end point of concern following oral exposure to selenium compounds (Baker et al. 1989; Boylan et al. 1990; Cukierski et al. 1989; Harrison et al. 1983; Panter et al. 1996; Rosenfeld and Beath 1964; Stowe et al. 1992; Tsunoda et al. 2000; Yang et al. 1983). Chronic oral exposure studies of laboratory animals that focus on behavioral effects and histopathological changes in the central nervous system might provide useful dose-response information on central nervous system effects.

Epidemiological and Human Dosimetry Studies. A few human epidemiological studies have identified blood selenium levels indicative of adequate selenium status and indicative of selenium toxicity. However, there are large differences in selenium blood levels in populations from different parts of the world (e.g., China, New Zealand, and Finland) (Salonen et al. 1985; Yang et al. 1989a). For example, blood selenium levels in healthy Finnish populations average 0.055 mg selenium/L (Virtamo et al. 1987), whereas blood selenium levels in healthy U.S. populations are much higher, averaging 0.206 mg selenium/L (Allaway et al. 1968). Extrapolation from the relationship between blood selenium levels and selenium toxicity in populations from one region of the world to populations in another region may not be appropriate. Studies examining the particular forms of selenium and the contribution of diet in determining individual and population selenium status would be useful. The selenium status of an individual will determine the magnitude of additional selenium intake that can be tolerated without resulting in adverse effects. Evidence for adverse effects on the endocrine system has also been found following intermediate and chronic oral exposure to elevated levels of dietary selenium in humans and animals (Brätter and Negretti De Brätter 1996; Behne et al. 1992; Eder et al. 1995; Hawkes and Keim 1995; Hotz et al. 1997). Studies of humans with high dietary intakes of selenium that monitored thyroid hormone levels and iodine intake would be useful. Studies of humans taking selenium supplements would also help further identify the long-term effects of selenium status on human health.

Biomarkers of Exposure and Effect.

Exposure. Selenium exposure can be correlated with concentrations detected in human blood, blood components, urine, hair, and nails. Selenium concentrations found in these biomarkers in the general population can be found in Table 3-7. However, these markers vary greatly among different populations (Longnecker et al. 1991). Levels of plasma, erythrocyte and platelet GSH-Px activity, as well as selenoprotein P may serve as better markers of selenium deficiency than selenium concentrations. Additional research into markers of selenium status in populations and how they may be used to estimate an additional selenium exposure that would be safe would be helpful.

Effect. Perhaps the earliest and most frequent symptoms of selenosis in humans are dry and brittle hair that breaks off, and brittle nails with white spots or streaks. Although these effects may not be specific to selenium, determination of selenium status may be useful if they are observed in a subject. Additional, biomarkers of negative effects that could be detected before clinical signs of selenium toxicity could be helpful in preventing selenium poisoning.

Absorption, Distribution, Metabolism, and Excretion. The absorption of selenium has been investigated in humans following oral exposure and in animals following oral and inhalation exposures (Finley 1998; Glover 1970; Griffiths et al. 1976; Martin et al. 1989a; Medinsky et al. 1981a; Sánchez-Ocampo et al. 1996; Thomson et al. 1977). In humans, no quantitative data exist on either the extent or rate of absorption of selenium from the lung or the skin. Information that selenium is absorbed following inhalation is limited to occupational case studies in which larger quantities of selenium have been measured in the urine of workers occupationally exposed to selenium. In order to understand all possible routes for human overexposure to selenium, information concerning the dermal and inhalation absorption of selenium and its compounds in humans would be useful, even though potential exposures to selenium might be more likely to occur by the oral route for the general public.

The oral absorption of different physical and chemical forms of selenium (e.g., selenite, selenate, and selenomethionine as solids or in aqueous solution) has been investigated in humans (Griffiths et al. 1976; Martin et al. 1989a; Moser-Veillon et al. 1992; Robinson et al. 1978; Swanson et al. 1991; Thomson 1974; Thomson and Stewart 1974; Thomson et al. 1977) and in animals (Finley 1998; Furchner et al. 1975; Thomson and Stewart 1973; Vendeland et al. 1992; Whanger et al. 1976). Oral absorption of naturally occurring selenium and the effects of dietary levels on the absorption of exogenous selenium have also been investigated (Young et al. 1982). These studies have revealed that several selenium compounds appear to be readily absorbed from the gastrointestinal tract of humans and animals. It also appears that the degree of absorption in humans is independent of the exposure level, but that in some cases, absorption is greater when a selenium deficiency exists.

Distribution studies in humans and animals indicate that selenium is widely distributed in the body and is concentrated in the liver and kidney following oral, intravenous, or subcutaneous exposures (Cavalieri et al. 1966; Finley 1998; Heinrich and Kelsey 1955; Jereb et al. 1975; Kaneko et al. 1999; Mahan and Kim 1996; Razagui and Haswell 1997; Shiobara et al. 1998; Thomson and Stewart 1973). Studies of intravenous administration of selenomethionine have indicated that animals and humans concentrate this compound in the pancreas, but it is unlikely that this selenium compound will be encountered in large quantities in the environment except in animals and plants along with other organic selenium compounds. It would be useful to know if selenomethionine concentrates in the pancreas of humans following oral intake. Following oral exposure, the distribution of selenium across the placenta into the fetuses of rats, hamsters, dogs, and monkeys (Archimbaud et al. 1992; Choy et al. 1993; Hawkes et al. 1994; Mahan and Kim 1996; Parizek et al. 1971a; Willhite et al. 1990) and the transfer of selenium from milk to suckling offspring of rats, dogs, and monkeys (Archimbaud et al. 1992; Choy et al. 1993; Hawkes et al. 1944;

Parizek et al. 1971a) have also been investigated. Selenium levels have been measured in human milk (Brätter and Negretti De Brätter 1996; Brätter et al. 1991b; Li et al. 1999; Michalke and Schramel 1998; Moser-Veillon et al. 1992; Rodríguez Rodríguez et al. 1999; Viitak et al. 1995; Yang 1989b) and the concentration of selenium in human milk has been shown to correlate with dietary intake (Brätter et al. 1991b). The uptake of selenium by erythrocytes and its subsequent metabolic alteration and ultimate binding to plasma proteins have been investigated (Sandholm 1973).

The metabolism of selenium is now fairly well understood. To become incorporated into selenium-specific proteins (e.g., glutathione peroxidase, thioredoxin reductase, iodothyronine 5'-deiodinase) through a cotranslational mechanism requires that selenium be in the form of selenide (Sunde 1990). All forms of selenium can be transformed to selenide, although the rates of transformation vary. For example, selenate is not converted to selenide as readily as selenite. The formation of selenide from selenocysteine requires a specific enzyme, selenocysteine β-lyase, which catalyzes the decomposition of selenocysteine to alanine and hydrogen selenide. Excess selenium is methylated and exhaled or excreted in the urine in both humans and animals. Further research is required to determine which selenium metabolites or intermediates lead to toxicity.

In humans and animals, intravenous and oral administration data indicate that the major route of selenium excretion is in the urine (Byard and Baumann 1967; Davidson-York et al. 1999; Finley 1998; Griffiths et al. 1976; Palmer et al. 1970; Patterson et al. 1989; Shiobara et al. 1998; Swanson et al. 1991). Excretion of selenium in feces constitutes a minor pathway immediately following exposure but the amount excreted can be equal to that excreted in urine depending on the chemical form of selenium administered, the size of the dose, and the length of time since dosing. Both human and animal studies indicate that the extent of excretion by any one route is related to the administered dose and the frequency of administration (Finley 1998; Lathrop et al. 1972; McConnell and Roth 1966; Shiobara et al. 1998; Thomson and Stewart 1974). The extent of excretion of selenium compounds in the expired air has been investigated in animals, but no quantitative studies in humans for this route exist; however, it is believed to be a minor pathway especially at lower doses (McConnell and Roth 1966; Olson et al. 1963).

Comparative Toxicokinetics. The target organs and adverse health effects are generally similar across species. However, the liver appears to be the primary target organ for the oral toxicity of selenium in animals following intermediate and chronic exposure (Baker et al. 1989; Biolac-Sage et al. 1992; Fitzhugh et al. 1944; Halverson et al. 1970; Harr et al. 1967; Hasegawa et al. 1994; Kolodziejczyk et al. 2000; Nelson et al. 1943; Palmer and Olson 1974; Sayato et al. 1993; Schroeder and Mitchener 1972;

Skowerski et al. 1997a; Turan et al. 1999a), whereas liver cirrhosis or dysfunction have not been found in reports of chronic selenosis in humans (Longnecker et al. 1991; Yang et al. 1989a). Different metabolites may help explain the cataract formation observed in neonatal rats and the teratogenic activity of selenium seen in birds but not in humans or other mammals (Tarantal et al. 1991). Toxicokinetic studies with some design similarities have been performed in humans and several animal species (Behne et al. 1991; Bopp et al. 1982; Cantor et al. 1975; Ganther 1979; Hawkes et al. 1992; Obemeyer et al. 1971; Palmer et al. 1970; Willhite et al. 1990, 1992). Comparative toxicokinetic studies, per se, have not been performed. PBPK models for selenium administered orally as selenite or selenomethionine have been developed for humans, but no animal models were located. Animal models for the oral route would be useful in assessing toxicokinetic similarities and differences between species.

Methods for Reducing Toxic Effects. Current methods for reducing toxic effects of selenium and selenium compounds after acute exposures are general supportive treatment methods based on those used for other toxic metals (HSDB 2001; Mack 1990). Although currently a rare form of poisoning, the use of selenious acid in gun blueing could make acute selenium exposure more common in households in the future. Additional research aimed at decreasing absorption, speeding excretion, and reducing the body burden of selenium would be valuable.

Children's Susceptibility. Limited information is available on the toxicity of selenium in children, but the available information suggests that children may be less susceptible to toxic effects of selenium than adults and more susceptible to deficiency. Most data comes from children living in areas of chronic high dietary selenium intake (Yang et al. 1989a, 1989b). Additional research on age specific effects of selenium toxicity does not appear necessary at present.

Child health data needs relating to exposure are discussed in 6.8.1 Identification of Data Needs: Exposures of Children.

3.12.3 Ongoing Studies

The American Health Foundation is involved in on-going research to develop new organoselenium chemopreventive agents for cancer having an increased therapeutic ratio compared with some of the historical selenium compounds, such as selenite. Additional federally sponsored research that was reported in the CRIS/USDA (2001) and CRISP (2001) databases is shown in Table 3-8.

Table 3-8. On-going Studies on Selenium Health Effects

Investigator	Institute	Research Area	Reference
Alberts, DS	University of Arizona	Phase III trials of chemopreventive agents on colon carcinogenesis	CRISP 2001
Cohen, HJ	Stanford University	Relationship of the synthesis and secretion of an extracellular selenium dependent glutathione peroxidase to changes in renal function	CRISP 2001
Combs, GF	Cornell University	Metabolic events at extremes of selenium intake; characterization of antioxidant status of a large cohort of free-living Americans	CRIS/UDSA 2001
Beran, M	Vyzkumny Ustav Potravinarsky	Evaluation of combined supplementation with selenium and iodine on levels of selenium-dependent enzymes, thyroidal hormones and other biochemical parameters.	CRIS/UDSA 2001
Berry, MJ	Brigham and Women's Hospital	Mechanism of selenoprotein synthesis in eukaryotes	CRISP 2001
Berry, MJ	Brigham and Women's Hospital	Selenoprotein P function and regulation of expression	CRISP 2001
Block, E	Roswell Park Memorial Institute	Identify selenium compounds from high-selenium garlic	CRISP 2001
Bosland, MC	New York University School of Medicine	Preclinical prostate cancer chemoprevention studies	CRISP 2001
Burk, RF	Vanderbilt University	Selenium supplementation of patients with cirrhosis	CRISP 2001
Burk, RF	Vanderbilt University	Selenoprotein-P structure, function and activity	CRISP 2001
Cassano, PA	Cornell University	Nutritional influences on lung disease	CRIS/UDSA 2001
Coltman, CA	CTRC Research Foundation	Chemoprevention of prostate cancer	CRISP 2001
Combs, GF	Cornell University	Kinetics of organic and inorganic selenium during dietary supplementation	CRIS/UDSA 2001
Davis, CD	Agricultural Research Service	Role of selenium in cancer susceptibility	CRIS/UDSA 2001

Table 3-8. On-going Studies on Selenium Health Effects (*continued*)

Investigator	Institute	Research Area	Reference
Diamond, AM	University of Illinois	Mechanism by which selenium protects against mutagenesis	CRISP 2001
Driscoll, DM	Cleveland Clinic Foundation	Mechanism of selenoperoxidase biosynthesis	CRISP 2001
El-Bayoumy, KE	American Health Foundation	Chemoprevention of oral cancer: model studies	CRISP 2001
El-Bayoumy, KE	American Health Foundation	Chemoprevention of lung cancer: model studies	CRISP 2001
El-Bayoumy, KE	American Health Foundation	Chemoprevention of mammary cancer by organoselenium	CRISP 2001
Fiala, E	American Health Foundation	Organoselenium compounds as modifiers of carcinogenesis	CRISP 2001
Ganther, H	Roswell Park Memorial Institute	Selenium metabolism and anti-carcinogenic action	CRISP 2001
Ganther, H	University of Wisconsin	Organoselenium compounds biosynthesis and function	CRIS/UDSA 2001
Gesteland, RF	University of Utah	Genetic analysis of synthesis of selenium containing proteins	CRISP 2001
Gladyshev, VN	University of Nebraska	Identity of terminator and selenocysteine UGA codons	CRISP 2001
Gladyshev, VN	University of Nebraska	Biochemistry and molecular biology of selenium containing enzymes	CRIS/UDSA 2001
Glauert, HP	University of Kentucky	Effect of dietary antioxidants on hepatic NF-KB activation	CRIS/UDSA 2001
Gottschall, EB	National Jewish Medical and Research Center	Randomized, placebo-controlled, double blind trial of asbestos-exposed workers using high selenium yeast supplementation	CRISP 2001
Guttenplan, JB	New York University	Antimutagenesis by lycopene and selenium in rodents	CRISP 2001
Ip, C	Roswell Park Memorial Institute	Mammary cancer prevention by novel selenium compounds; selenium metabolism and cancer prevention	CRISP 2001
James, LF	Agricultural Research Service	Livestock poisoning from *Astragalus* and *Oxytropis* species	CRIS/UDSA 2001

Table 3-8. On-going Studies on Selenium Health Effects (*continued*)

Investigator	Institute	Research Area	Reference
Kim, J	University of Texas MD Anderson Cancer Center	Feasibility study of L-selenomethionine in prevention of prostate cancer	CRISP 2001
Kolonel, LN	University of Hawaii at Manoa	Epidemiologic studies of diet and cancer in Hawaii	CRISP 2001
Kolonel, LN	University of Hawaii at Manoa	Biomarkers of prostate cancer risk in a multi ethnic cohort	CRISP 2001
Koutnik, V	University of Brno	Selenium in food chains and its impact on human health	CRIS/UDSA 2001
Lei, X	Cornell University	Antioxidative role of glutathione peroxidase in transgenic mice	CRISP 2001
Levander, OA	Agricultural Research Service	Role of vitamin E and selenium in human health promotion	CRIS/UDSA 2001
Levander, OA	University of Maryland	Kinetics of organic and inorganic selenium during dietary supplementation	CRIS/UDSA 2001
Marshall, JR	University of Arizona	Phase III trial of selenium for prostate cancer prevention	CRISP 2001
May, JM	Vanderbilt University	Antioxidant interactions of selenium and vitamins C and E	CRISP 2001
Medina, D	Roswell Park Memorial Institute	Selenoproteins in rat mammary tumorigenesis	CRISP 2001
Nomura, AM	Kuakini Medical Center	Cancer epidemiology of migrant Japanese in Hawaii	CRISP 2001
Pence, BC	Texas Technical University Health Sciences Center	Induction by selenium of the antioxidant and the prooxidant, apoptotic pathways in cultured cells	CRISP 2001
Powis, G	University of Arizona	Thioredoxin reductases and cancer	CRISP 2001
Prolla, TA	University of Wisconsin	Role of dietary selenium in intestinal tumorigenesis	CRISP 2001
Rao, L	University of Wisconsin	Genetic characterization of the selenoenzyme phospholipid -hydroperoxide glutathione peroxidase	CRISP 2001
Reddy, BS	American Health Foundation	Chemoprevention of cancer by organoselenium compounds	CRISP 2001

Table 3-8. On-going Studies on Selenium Health Effects (*continued*)

Investigator	Institute	Research Area	Reference
Reddy, CC	Pennsylvania State University	Antioxidant effects on prostaglandin metabolism, lipid peroxidation and Immunologic defense	CRIS/UDSA 2001
Roberts, JC	University of Utah	Preparation and preliminary studies of selenazoidine carboxylic acids as novel selenium delivery agents	CRISP 2001
Roy, M	New York University	Selenium supplementation and immunocompetence in the elderly	CRIS/UDSA 2001
Smith, AM	Ohio State University	Influence of gender and life cycle on selenium requirements and metabolism	CRIS/UDSA 2001
Sunde, RA	University of Missouri	New essential roles for selenium; regulatory elements of selenium-dependent peroxidases; regulatory elements of the rat glutathione peroxidase gene	CRIS/UDSA 2001
Taylor, EW	University of Georgia	Selenoproteins, NF-KB, and HIV disease in drug users	CRISP 2001
Thompson, I	University of Texas Health Science Center San Antonio	Biomarkers of risk for prostate cancer	CRISP 2001
Thompson, H	Roswell Park Memorial Institute	Mechanisms of selenium anticancer and toxic activities	CRISP 2001
Thompson, H	Roswell Park Memorial Institute	Six month trial of selenium as a chemopreventive agent for lung cancer.	CRISP 2001
Veillon, C	Agricultural Research Service	Metabolism, function, and interactions of selenium using stable isotopes	CRIS/UDSA 2001
Weiss, SL	University of Missouri	Molecular basis for selenium regulation of glutathione peroxidase mRNA	CRIS/UDSA 2001
Whanger, PD	Oregon State University	The metabolic function of selenoprotein W	CRISP 2001

Table 3-8. On-going Studies on Selenium Health Effects (*continued*)

Investigator	Institute	Research Area	Reference
Whanger, P	Oregon State University	Role of selenium and vitamin E in scour and immunity of newborn calves; influence of pregnancy on selenium metabolism in women of low selenium status; metabolic relationship between selenium and myopathy	CRIS/UDSA 2001
Yu, MC	University of Southern California	Singapore cohort study of diet and cancer	CRISP 2001

NCI = National Cancer Institute; NIH = National Institutes of Health

4. CHEMICAL AND PHYSICAL INFORMATION

4.1 CHEMICAL IDENTITY

Information regarding the chemical identity of selenium and some selenium compounds is located in Table 4-1.

4.2 PHYSICAL AND CHEMICAL PROPERTIES

Selenium is a metalloid element with atomic number 34 and an atomic mass of 78.96 (Lide 2000). Selenium belongs to Group 16 (Group VIA) of the periodic table, located between sulfur and tellurium, and resembles sulfur both in its various forms and in its compounds. The six stable isotopes of selenium are ^{74}Se, ^{76}Se, ^{77}Se, ^{78}Se, ^{80}Se, and ^{82}Se. These isotopes occur naturally with approximate abundances of 0.87, 9.02, 7.58, 23.52, 49.82, and 9.19 %, respectively (Hoffmann and King 1997). Artificial radioactive isotopes of selenium have also been created by neutron activation. The gamma emitting isotope ^{75}Se has been used in diagnostic applications of medicine (Hoffmann and King 1997). Selenium exists in several allotropic forms. Three are generally recognized, but as many as six have been claimed (Lide 2000). The stable form at ordinary room temperatures is the grey or hexagonal form with a melting point of 220.5 EC (Lide 2000). The other two important forms are red (monoclinic) with a melting point of 221 EC and amorphous selenium, which exists in black and red forms. Black amorphous selenium is vitreous and is formed by the rapid cooling of liquid selenium. Red amorphous selenium is colloidal and is formed in reduction reactions (Hoffmann and King 1997). Important selenium oxidation states are -2, 0, +4, and +6.

The chemical properties of selenium are similar to sulfur. Selenium combines with metals and many nonmetals directly or in aqueous solution. The selenides resemble sulfides in appearance, composition, and properties (Hoffmann and King 1997). Selenium may form halides by reacting vigorously with fluorine and chlorine, but the reactions with bromine and iodide are not as rapid. Selenium does not react directly with hydrogen fluoride or hydrogen chloride, but decomposes hydrogen iodide to liberate iodine and yield hydrogen selenide (Hoffmann and King 1997). Selenium reacts with oxygen to form a number of oxides, the most stable of which is selenium dioxide.

Information regarding the physical and chemical properties of selenium and selenium compounds is located in Table 4-2.

Table 4-1. Chemical Identity of Selenium and Selected Compounds[a]

Characteristic	Selenium	Hydrogen selenide	Selenic acid	Selenious acid
Synonyms	Elemental selenium; selenium base; selenium dust; colloidal selenium; selenium homopolymer[b]; selenium alloy	Dihydrogen selenide; hydrogen selenide [H_2Se]; selenium anhydride; selenium dihydride; selenium hydride; selane	Selenic acid, liquid[DOT,b]	Monohydrated selenium dioxide; selenous
Registered trade name(s)	C.I. 77805; VANDEX[b]			
Chemical formula	Se	H_2Se	H_2SeO_4	H_2SeO_3
Wisewesser line notation	SE	H2 SE	H2.SE-04	H2SE-03
Identification numbers:				
CAS	7782-49-2	7783-07-5	7783-08-6	7783-00-8
NIOSH RTECS	VS7700000	MX1050000	VS6575000	VS7175000
EPA hazardous waste	No data	No data	No data	U204[c]
OHM/TADS	7216880	No data	No data	No data
DOT/UN/NA/IMCO shipping	UN 2658[b]	UN 2202; hydrogen selenide, ananhydrous	UN 1905; IMCO 8.0	No data
HSDB	4493		675	6065
NCI	No data		No data	No data

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Table 4-1. Chemical Identity of Selenium and Selected Compounds[a] (*continued)*

Characteristic	Sodium selenate	Potassium selenate	Sodium selenide	Sodium selenite
Synonyms	Disodium selenate	Selenic acid, dipotassium salt[b]	Disodium monoselenide[b]	Disodium selenite; disodium selenium trioxide; selenious acid disodium salt; sodium selenium oxide
Registered trade name(s)	P-40[b]; Sel-Tox SS02 and SS-20[c]	No data	No data	
Chemical formula	Na_2SeO_4[d]	K_2SeO_4	Na_2Se[b]	Na_2SeO_3[d]
Wisewesser line notation	NA2 SE-04[b]	KA2 SE-04	NA2 SE	NAS SE-03
Identification numbers:				
CAS	13410-01-0	7790-59-2	1313-85-5[b]	10102-18-8
NIOSH RTECS	No data	VS6600000	WE0350000[b]	VS7350000
EPA hazardous waste	No data	No data	No data	No data
OHM/TADS	No data	No data	No data	7217299
DOT/UN/NA/IMCO shipping	No data	No data	No data	UN 2630[b]
HSDB	No data	No data	No data	768
NCI	No data	No data	No data	No data

Table 4-1. Chemical Identity of Selenium and Selected Compounds[a] (*continued)*

Characteristic	Selenium dioxide	Selenium trioxide	Selenocystine	Selenomethionine
Synonyms	Selenious anhydride; selenium oxide; selenium oxide [SeO_2]; selenous acid anhydride	No data	Selenium cystine[b]; 3,3-diselenodi-DL-alanine [b]; seleno-DL-cystine[b]; DL-selenocystine[b]	Methionine, seleno[b]; 2-amino-4-(methylselenyl) butyric acid; 2-amino-4-(methylseleno)
Registered trade name(s)	No data	No data		
Chemical formula	SeO_2	SeO_3[e]	$C_2H_4NO_2(CH_2)Se_2(CH_2)C_2H_4NO_2$	$(CH_3)Se(CH_2)2C_2H_4NO_2$
Identification numbers:				
CAS	7446-08-04	13768-86-0[f]	1464-43-3[b]	1464-42-2
NIOSH RTECS	VS8575000	No data	AY6030000[b]	ES100000
EPA hazardous waste	V204[c]	No data	No data	No data
OHM/TADS	7800105	No data	No data	No data
DOT/UN/NA/IMCO shipping	No data	No data	No data	No data
HSDB	677	No data	No data	No data
NCI	No data	No data	No data	No data

Table 4-1. Chemical Identity of Selenium and Selected Compounds[a] (*continued)*

Characteristic	Selenium sulfide	Selenium disulfide
Synonyms	Selenium monosulfide; selenium sulfide [SeS]; selensulfid (German); sulfur selenide (Sse)	Selenium disulphide; selenium sulfide[b]; sulfur selenide
Registered trade name(s)	No data	Exsel; Selsun Blue; Selsum[b]; Seleen
Chemical formula	SeS	SeS_2[b]
Wisewesser line notation	SE S	SE S2[b]
Identification numbers:		
CAS	744-34-6	7488-56-4
NIOSH RTECS	VTO525OOO	VS8925000
EPA hazard waste	V205[b]	V205
OHM/TADS	8400272	8400272
DOT/UN/NA/IMCO shipping	No data	UN 2657
HSDB	679	No data
NCI	NCI-C50033	No data

[a]All information obtained from HSDB 2001 except where noted
[b]RTECS 2001
[c]EPA 1980a, 1980b (40 CFR 261.33)
[d]Budavari et al. 1996
[e]Lide 2000
[f]Grayson 1982

CAS = Chemical Abstracts Service; DOT/UN/NA/IMCO = Department of Transportation/United Nations/North America/International Maritime Dangerous Goods Code; EPA = Environmental Protection Agency; HSDB = Hazardous Substances Data Bank; NCI = National Cancer Institute; NIOSH = National Institute for Occupational Safety and Health; OHM/TADS = Oil and Hazardous Materials/Technical Assistance Data System; RTECS = Registry of Toxic Effects of Chemical Substances

Table 4-2. Physical and Chemical Properties of Selenium and Selenium Compounds[a]

Property	Selenium	Hydrogen selenide	Selenic acid	Selenious acid
Molecular weight	78.96	80.98	144.97	128.97
Color/form	Red, grey, or black	Colorless[b]	White hexagonal prisms; hygroscopic[b]	White hygroscopic prisms[b]
Physical state	Solid	Gas	Solid	Solid
Melting point	221 EC(red); 220.5 EC (grey); 180 EC (black)[b]	-65.73 EC	58 EC	70 EC (decomposes)[b]
Boiling point	685 EC	-41.3 EC	260 EC	None, loses water upon heating
Density (g/cm^3)	4.39 (red); 4.81 (grey); 4.28 (black)[b]	2.12 (-42 EC)	2.9508 (15 EC)	3.004 (15 EC)
Odor	Unknown; upon combustion, smells like rotten horseradish	Disagreeable odor	No data	No data
Odor threshold:				
Water (mg/m^3)	No data	No data	No data	No data
Air	No data	No data	No data	No data
Solubility:				
Water (g/100 mL)	Insoluble	377 mL/100 mL at 4 EC; 270 mL/100 mL at 22.5 EC; 0.73 mL/100 mL at 20 EC[c]	Very soluble in hot water	90 parts dissolve in 100 parts of water at 0 EC; 400 parts in 100 parts at 90 EC
Organic solvents	Insoluble in alcohol, slightly soluble in carbon disulfide (2 mg/100 mL, room temperature), soluble in ether	Soluble in carbon disulfide, carbonyl chloride	Soluble in sulfuric acid Decomposes in alcohol[b]	Very soluble in alcohol
Partion coefficients:				
Log K_{ow}	No data	No data	No data	No data
Log K_{oc}	No data	No data	No data	No data

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Vapor pressure (mmHg)	1 mmHg at 356 EC (grey)	1,330 mmHg at -30 EC; 3,420 mmHg at 0.2 EC; 9,120 mmHg at 30.8 EC	No data	2 mmHg at 15 EC; 4.5 mmHg at 33 EC; 7mmHg at 40.3 EC

Table 4-2. Physical and Chemical Properties of Selenium and Selenium Compounds[a] *(continued)*

Property	Selenium	Hydrogen selenide	Selenic acid	Selenious acid
Henry's law constant	Not applicable	No data	Not applicable	Not applicable
Autoignition temperature	No data	No data	No data	No data
Flashpoint	No data	Not applicable	No data	No data
Flammability limits	No data	No data	No data	No data
Conversion factors	No data	ppm selenium to mg Selenium/m^3 in air (20 EC): ppm selenium x 3.23=mg selenium/m^3 to ppm selenium in air (20 EC): mg selenium/m^3 x 0.31=ppm selenium (v/v)	No data	No data
Explosive limits	Unknown[d]	No data	No data	No data

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Table 4-2. Physical and Chemical Properties of Selenium and Selenium Compounds[a] (*continued*)

Property	Sodium selenate	Potassium selenate	Sodium selenide	Sodium selenite
Molecular weight	188.94	221.15	124.94	172.94
Color/form	White crystals	Colorless crystals or white powder	Crystalline; turns red on exposure to air and deliquesces	White tetragonal crystals[b]
Physical state	Solid	Solid	Solid	Solid
Melting point	No data	No data	>875 EC	No data
Boiling point	No data	No data	No data	No data
Density (g/cm^3)	1.61[b]	3.07	2.625 (10 EC)	No data
Odor	No data	No data	No data	No data
Odor threshold:				
Water (mg/m^3)	No data	No data	No data	No data
Air	No data	No data	No data	No data
Solubility:				
Water (g/100 mL)	Very soluble in water	Soluble in about 1 part of water	Decomposes in water	Freely soluble in water
Organic solvent(s)				
Partition coefficients:				
Log K_{ow}	No data	No data	No data	No data
Log K_{oc}	No data	No data	No data	No data
Vapor pressure (mmHg)	No data	No data	No data	No data
Henry's law constant	No data	No data	No data	No data
Autoignition temperature	No data	No data	Not flammable[e]	Not flammable[e]
Flashpoint	No data	No data	Not flammable[e]	Not flammable[e]
Flammability limits	No data	No data	Not flammable[e]	Not flammable[e]
Conversion factors	No data	No data	No data	No data
Explosive limits	No data	No data	No data	No data

Table 4-2. Physical and Chemical Properties of Selenium and Selenium Compounds[a] (*continued*)

Property	Selenium dioxide	Selenium trioxide	Selenocystine	Selenomethionine
Molecular weight	110.96	126.96[b]	334.12[c]	196.11
Color/form	Lustrous, tetragonal needles; yellowish-green vapor	White crystals[b]	No data	Transparent, hexagonal sheets or plates; metallic luster or crystals
Physical state	Solid	Solid	No data	Solid
Melting point	340 EC; sublimes at 315 EC[b]	118 EC[b]	No data	DL form: 265 EC (decomposes); L form: 266–268 EC
Boiling point	None[b]	sublimes[b]	No data	Not applicable
Density (g/cm^3)	3.954 (15 EC)	3.44 [b]	No data	No data
Odor	Pungent sour smell	No data	No data	No data
Odor threshold:				
Water (mg/m^3)	0.0002[e]	No data	No data	No data
Air	No data	No data	No data	No data
Solubility:				
Water (g/100 mL)	38.4 at 14 EC;	Soluble in water	No data	No data
Organic solvent(s)	In methanol: 10.16 at 11.8 EC; in 93% ethanol: 6.67 at 14 EC; in acetone: 4.35 at 15.3 EC; in acetic acid: 1.11 at 13.9 EC; soluble in benzene		No data	No data
Partition coefficients:				
Log K_{ow}	No data	No data	No data	No data
Log K_{oc}	No data	No data	No data	No data
Vapor pressure (mmHg)	12.5 mm Hg at 70 EC; 20.2 mm Hg at 94 EC; 39.0 mm Hg at 181 EC; 760 mm Hg at 315 EC; 848 mm Hg at 320 EC	No data	No data No data	No data No data
Henry's law constant	Not applicable	Not applicable	No data	No data
Autoignition temperature	Not flammable[e]	Not flammable[e]	No data	No data

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Table 4-2. Physical and Chemical Properties of Selenium and Selenium Compounds[a] (*continued*)

Property	Selenium dioxide	Selenium trioxide	Selenocystine	Selenomethionine
Flashpoint	Not flammable[e]	Not flammable[e]	No data	No data
Flammability limits	Not flammable[e]	Not flammable[e]	No data	No data
Conversion factors	ppm (v/v) to mg/m^3 in air (20 EC): ppm (v/v) x 4.53 = mg/m^3; mg/m^3 to ppm (v/v) in air (20 EC): mg/m^3 x 0.22 = ppm (v/v)	No data	No data	No data
Explosive limits	No data	No data	No data	No data

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Table 4-2. Physical and Chemical Properties of Selenium and Selenium Compounds[a] (*continued*)

Characteristic	Selenium sulfide	Selenium disulfide
Molecular weight	111.02[f]	143.08[g]
Color/form	Orange-yellow tablets or powder[f]	Bright red-yellow powder[g]
Physical state	Solid[f]	Solid[g]
Melting point	118–119 EC (decomposes)[f]	<100 EC [g]
Boiling point	No data	Decomposes [g]
Density (g/cm^3)	3.056 (0 EC)[f]	No data
Odor	No data	No data
Odor threshold:		
Water (mg/m^3)	No data	No data
Air	No data	No data
Solubility:		
Water (g/100 mL)	Insoluble	Insoluble
Organic solvent(s)	Insoluble in ether; decomposes in alcohol[f]	No data
Partition coefficients:		
Log K_{ow}	Not applicable	No data
Log K_{oc}	Not applicable	No data
Vapor pressure (mmHg)	Not applicable	Not applicable
Henry's law constant	Not applicable	Not applicable
Autoignition temperature	No data	No data
Flashpoint	No data	No data
Flammability limits	No data	No data
Conversion factors	No data	No data
Explosive limits	No data	No data

Note: The gray metallic form is the most stable form of selenium (Budavari et al. 1996).

[b]Lide 2000
[c]RTECS 2001
[d]NIOSH/OSHA 1981
[e]Weiss 1986
[f] Lide 1993
[g] Windholz et al. 1983

[a]All information obtained from Budavari et al. 1996, unless otherwise noted.

5. PRODUCTION, IMPORT/EXPORT, USE, AND DISPOSAL

5.1 PRODUCTION

Selenium is distributed widely in nature and is found in most rocks and soils at concentrations between 0.1 and 2.0 ppm (Fishbein 1983). However, elemental selenium is seldom found naturally, but it is obtained primarily as a byproduct of copper refining (Fishbein 1983). Selenium is contained in the constituents of the copper anode that are not solubilized during the copper refining process and ultimately accumulate on the bottom of the electrorefining tank. These constituents, usually referred to as slimes, contain roughly 5–25% selenium and 2–10% tellurium. Selenium is commercially produced by either soda ash roasting or sulfuric acid roasting of the copper slimes.

Soda Ash Roasting. A soda ash binder is mixed with the slimes and water to form a stiff paste. The paste is extruded or pelletized and allowed to dry and then roasted at 530–650 EC. The roasted product is then ground and leached into water. The resultant hexavalent selenium dissolves as sodium selenate, Na_2SeO_4. The sodium selenate may be reduced by controlled heating to sodium selenide, which is leached with water to form a liver-red solution of sodium selenide that is readily oxidized to the elemental form by blowing air through the solution (Hoffmann and King 1997). A second process for the reduction of hexavalent selenium involves the use of concentrated hydrochloric acid or ferrous iron salts catalyzed by chloride ions as the reductant (Hoffmann and King 1997).

Sulfuric Acid Roasting. In this method, the copper slimes are mixed with sulfuric acid and roasted at 500–600 EC to produce selenium dioxide, which volatilizes readily at the roasting temperature. The selenium dioxide is reduced to elemental selenium during the scrubbing process with sulfur dioxide and water. The resultant commercial grade selenium can be purified to 99.5–99.7% (Hoffmann and King 1997).

The U.S. production of selenium was 373 and 379 metric tons in 1995 and 1996, respectively (USGS 2001). No production data were reported for the years 1997–1999. All of the primary selenium producers in the United States are electrolytic copper refiners. Asarco Incorporated and Kennecott Utah Copper Corporation produce refined selenium in the United States (Hoffmann and King 1997; SRI 2000). Two other copper refiners, Phelps Dodge Corporation and Magma Copper Company, send selenium or selenium bearing copper slimes outside of the United States for final processing (Hoffmann and King 1997).

Tables 5-1 and 5-2 list facilities in each state that produce, process, or import selenium and selenium compounds, respectively, for commercial use. The data do not include facilities such as electric power generating plants that release selenium unintentionally as a byproduct. The intended use and the range of maximum amounts of these substances that are stored on site are also included. The data listed in these tables are derived from the Toxics Release Inventory (TRI99 2001). Only certain types of facilities were required to report. Therefore, this is not an exhaustive list.

5.2 IMPORT/EXPORT

The import volume of selenium was 324, 428, 346, 339, and 320 metric tons for 1995, 1996, 1997, 1998, and 1999, respectively (USGS 2001). U.S. exports of selenium were 270, 322, 127, 151, and 240 metric tons for 1995, 1996, 1997, 1998, and 1999, respectively (USGS 2001).

5.3 USE

In electronics, selenium's semiconductor and photoelectric properties make it useful in "electric eyes," photographic exposure meters, and rectifiers for home entertainment equipment. In addition, a large proportion of the available selenium is used to coat the metal cylinders from which a photographic image is transferred in xerography (Fishbein 1983). Selenium is widely used in the glass industry to counter coloration that results from iron impurities. It is also used in the production of both red and black glasses (Fishbein 1983). Selenium contained in pigments are used in plastics, paints, enamels, inks, and rubber (Fishbein 1983). Selenium is used as a catalyst in the preparation of pharmaceuticals including niacin and cortisone, as an ingredient in antidandruff shampoos (selenium sulfide), and as a constituent of fungicides (selenium sulfide) (IARC 1975a). Radioactive selenium is used in diagnostic medicine and aids in the visualization of difficult-to-study malignant tumors (Fishbein 1983; Jereb et al. 1975). Selenium is contained in some dietary supplements at concentrations in the range of 10–25 μg/tablet (Goodman et al. 1990). Selenium is also used as a nutritional feed additive for poultry and livestock, in pesticide formulations, and as an accelerator and vulcanizing agent in rubber production (Fishbein 1983; NAS 1976a). Table 5-3 lists some specific uses of selected selenium compounds.

The 1995 consumption patterns for selenium by industry were as follows: glass, 35%; electrical, 30%; pigments, 10%; metallurgy, 10%; agricultural/biological, 5%; and miscellaneous, 10% (Hoffmann and King 1997).

Table 5-1. Facilities that Produce, Process, or Use Selenium

State[a]	Number of facilities	Range of maximum amounts on site in pounds[b]	Activities and uses[c]
CA	2	100–99,999	13
IL	2	1,000–99,999	1, 5, 8, 9
IN	1	100,000–999,999	9
MI	1	0–99	13
NY	1	0–99	8
OK	1	100–999	1, 5
OR	1	100,000–999,999	13
PA	4	0–99,999	1, 2, 3, 5, 7, 8, 9
SC	1	100,000–999,999	13
WY	1	0–99	1, 6

Source: TRI99 2001

[a]Post office state abbreviations used
[b]Range represents maximum amounts on site reported by facilities in each state
[c]Activities/Uses:

1. Produce
2. Import
3. Onsite use/processing
4. Sale/Distribution
5. Byproduct
6. Impurity
7. Reactant
8. Formulation Component
9. Article Component
10. Repackaging
11. Chemical Processing Aid
12. Manufacturing Aid
13. Ancillary/Other Uses

Table 5-2. Facilities that Produce, Process, or Use Selenium Compounds

State[a]	Number of facilities	Range of maximum amounts on site in pounds[b]	Activities and uses[c]
AL	3	0–99,999	1, 5, 9, 13
AR	1	10,000–99,999	13
AZ	3	10,000–9,999,999	1, 2, 3, 4, 5, 6, 8, 10
FL	1	1,000–9,999	1, 5, 6, 10, 13
GA	5	0–99,999	1, 2, 3, 4, 5, 6, 7, 9, 10
IA	2	10,000–99,999	3, 4, 8, 9
ID	1	1,000,000–9,999,999	1, 5
IL	5	1,000–999,999	1, 5, 6, 8, 10, 13
IN	3	1,000–99,999	1, 5, 6, 10, 13
KS	1	100–999	13
KY	4	100–99,999	1, 3, 4, 5, 6, 10, 13
LA	2	100–99,999	1, 3, 4, 5, 6, 7, 9, 13
MA	1	10,000–99,999	1, 5
MD	2	1,000–99,999	1, 3, 4, 5, 6, 7,
MI	5	0–999,999	1, 2, 3, 4, 5, 6, 7, 9, 10
MN	1	1,000–9,999	1, 6, 7, 10
MT	1	10,000–99,999	1, 3, 5, 6, 13
NC	3	1,000–99,999	1, 3, 4, 5, 6, 8, 10
NM	4	10,000–49,999,999	1, 3, 4, 5, 6, 10, 13
NV	5	10,000–9,999,999	1, 5, 6, 7, 11
NY	1	10,000–99,999	9
OH	5	1,000–99,999	1, 3, 4, 5, 6, 8, 10, 13
OK	2	1,000–9,999	1, 5, 9, 13
PA	7	100–999,999	1, 4, 5, 6, 7, 10, 13

Table 5-2. Facilities that Produce, Process, or Use Selenium Compounds (*continued*)

State[a]	Number of facilities	Range of maximum amounts on site in pounds[b]	Activities and uses[c]
SC	1	10,000–99,999	1, 3, 4, 5, 6, 10, 13
TN	1	10,000–99,999	1, 5
TX	6	1,000–999,999	1, 2, 3, 4, 5, 6, 7, 9, 10
UT	4	1,000–9,999,999	1, 3, 4, 5, 6, 7, 8, 10
WV	4	100–99,999	1, 3, 4, 5, 6, 10, 13
WY	2	0–99,999	1, 3, 4, 5, 6, 10, 13

Source: TRI99 2001

[a]Post office state abbreviations used
[b]Range represents maximum amounts on site reported by facilities in each state
[c]Activities/Uses:

1. Produce
2. Import
3. Onsite use/processing
4. Sale/Distribution
5. Byproduct
6. Impurity
7. Reactant
8. Formulation Component
9. Article Component
10. Repackaging
11. Chemical Processing Aid
12. Manufacturing Aid
13. Ancillary/Other Uses

Table 5-3. Some Selenium Compounds and Their Uses[a]

Compound	Use
Elemental selenium	In rectifiers, photoelectric cells, blasting caps, xerography, stainless steel; as dehydrogenation-catalyst
Sodium selenate (Na_2SeO_4)	As insecticide; in glass manufacture; in medicinals to control animal diseases
Sodium selenite (Na_2SeO_3)	In glass manufacture; as soil additive for selenium-deficient areas
Selenium diethyldithio-carbamate	Fungicide; vulcanizing agent
Selenium disulfide (SeS_2)	In veterinary medicine
Selenium sulfide (SeS)	In anti-dandruff shampoos and in veterinary medicine
Selenium dioxide (SeO_2)	Catalyst for oxidation, hydrogenation, or dehydrogenation of organic compounds
Selenium hexafluoride (SeF_6)	As gaseous electric insulator
Selenium oxychloride ($SeOCl_2$)	Solvent for sulfur, selenium, tellurium, rubber, bakelite, gums, resins, glue, asphalt, and other materials
Aluminum selenide (Al_2Se_3)	Preparation of hydrogen selenide for semi-conductors
Ammonium selenite [$(NH_4)_2SeO_3$]	Manufacture of red glass
Cadmium selenide	Photoconductors, photoelectric cells, rectifiers
Cupric selenate ($CuSeO_4$)	In coloring copper and copper alloys
Tungsten diselenide (WSe_2)	In lubricants

[a]Adapted from Fishbein 1983

5.4 DISPOSAL

Selenium was listed by EPA in 1973 as a nonradioactive hazardous element and, as such, is subject to many regulations (Dawson and Mercer 1986). Selenium compounds should be stored in a dry area to avoid contamination of water with selenium and to decrease the hazards that may result from human exposure to selenium-contaminated water (ITII 1976).

Disposal and waste treatment consist of treating an acidified solution of selenium with sodium sulfite to form the reducing agent sulfur dioxide. The selenium solution is then heated to produce elemental selenium, which is less mobile in the environment and less bioavailable, and the solution is filtered and washed (ITII 1976).

According to the TRI, in 1999, an estimated 13,255 pounds of selenium metal and 694,250 pounds of selenium compounds were transferred off-site, presumably for disposal (TRI99 2001).

6. POTENTIAL FOR HUMAN EXPOSURE

6.1 OVERVIEW

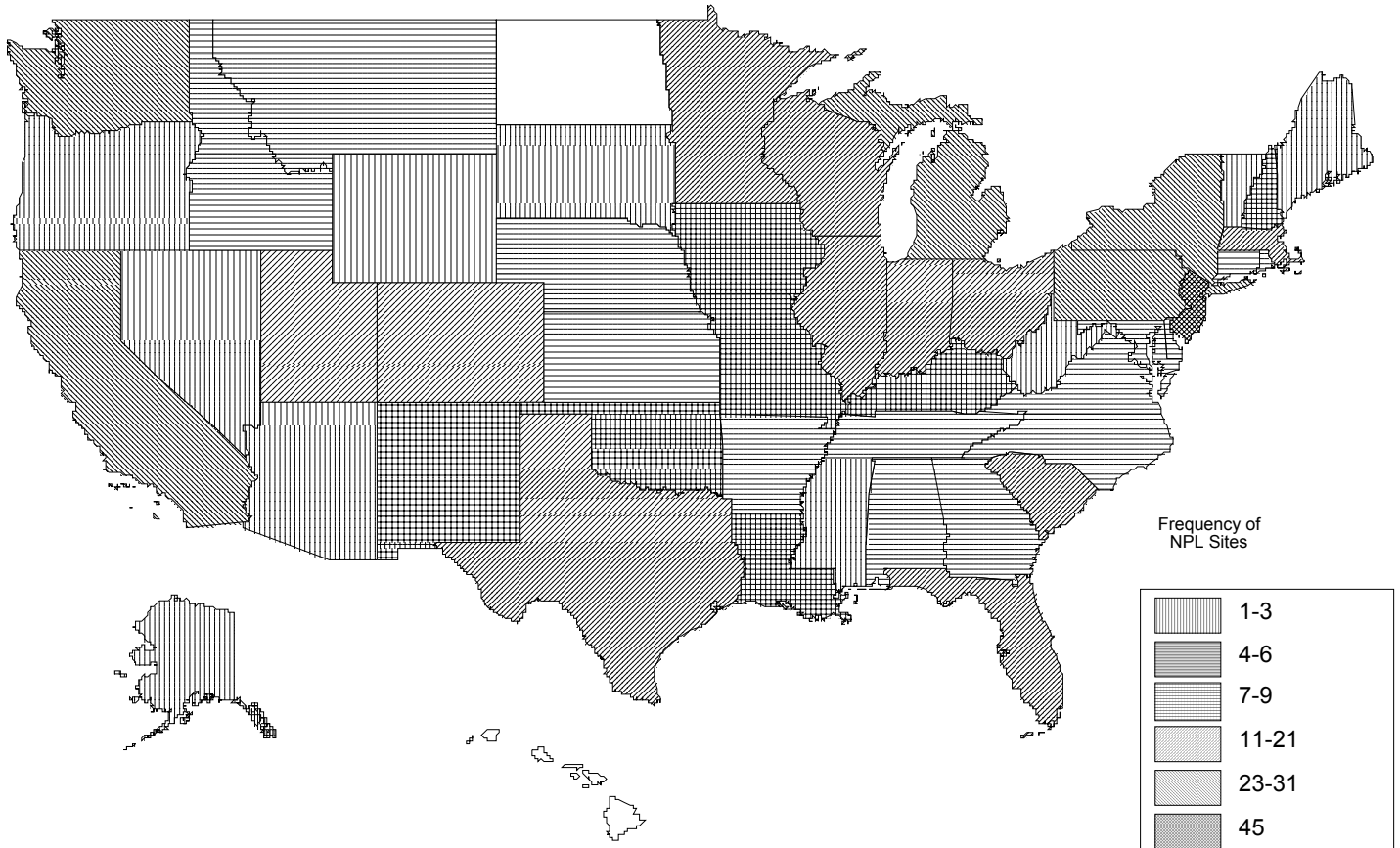
Selenium has been identified in at least 494 of the 1,585 hazardous waste sites that have been proposed for inclusion on the EPA National Priorities List (NPL) (HazDat 2001). However, the number of sites evaluated for selenium is not known. The frequency of these sites can be seen in Figure 6-1. Of these sites, 490 are located within the United States and 4 are located in the Commonwealth of Puerto Rico (not shown).

Selenium is ubiquitous in the environment, being released from both natural and anthropogenic sources. The principal releases of selenium into the environment as a consequence of human activities result from the combustion of coal. Workers in the metals industry and health services, mechanics, and painters may be exposed to higher levels of selenium than the general population or persons employed in other trades. For the general population, the primary exposure pathways, in order of decreasing relative proportions, are food, water, and air. The relative proportions of these exposure pathways at hazardous waste sites is not known. Although selenium has been reported at hazardous waste sites, analysis on specific forms has not been performed. In air, selenium dioxide, methyl selenide, and dimethyl selenide are the most prevalent forms found in the atmosphere. Selenates and selenites are water soluble and, thus, can be found in water sources. Salts of selenic and selenious acids are most likely to be found in surface water and water contained in soil. Selenium sulfides would not be expected to be found at most hazardous waste sites, since they are usually manufactured for use in shampoos. Natural sources of selenium include the weathering of selenium-containing rocks to soils and volcanic eruptions.

The primary factor determining the fate of selenium in the environment is its oxidation state. Selenium is stable in four valence states (-2, 0, +4, and +6) and forms chemical compounds similar to those of sulfur. The heavy metal selenides (-2) are insoluble in water, as is elemental selenium. The inorganic alkali selenites (+4) and selenates (+6) are soluble in water (Weast 1988) and are therefore more bioavailable.

Conditions such as pH (negative log hydrogen ion concentration), Eh (oxidation-reduction potential), and the presence of metal oxides affect the partitioning of the various compounds of selenium in the environment. In general, elemental selenium is stable in soils and is found at low levels in water because of its ability to coprecipitate with sediments. The soluble selenates are readily taken up by plants and converted to organic compounds such as selenomethionine, selenocysteine, dimethyl selenide, and

Figure 6-1. Frequency of NPL Sites with Selenium Contamination



Derived from HazDat 2001

dimethyl diselenide. Selenium is bioaccumulated by aquatic organisms and may also biomagnify in aquatic organisms.

6.2 RELEASES TO THE ENVIRONMENT

The greatest proportion of selenium released to the environment as a consequence of human activities is in coal fly ash, resulting from coal combustion. Manmade emission sources of atmospheric selenium include coal and oil combustion facilities, selenium refining factories, base metal smelting and refining factories, mining and milling operations, and end-product manufacturers (e.g., some semiconductor manufacturers). Natural atmospheric releases of selenium result from volatilization of selenium by plants and bacteria and from volcanic activity. Some selenium is released to water via sewage effluent, agricultural runoff, and industrial waste water. Selenium is released to soil primarily by leaching and weathering of the parent bedrock material, although dry and wet deposition also contribute to soil selenium levels.

According to the Superfund Amendments and Reauthorization Act (SARA), Section 313, Toxic Release Inventory (TRI99 2001), an estimated total of 338,004 pounds of selenium metal was released to air, water, land, or injected underground from manufacturing and processing facilities in the United States in 1999 (see Table 6-1). In addition, 6,622,466 pounds of selenium compounds were released to air, water, land, or injected underground in 1999 (see Table 6-2). These data include all facilities that manufacture, import, and process selenium and selenium compounds as well as facilities (electric generating facilities, petroleum facilities, etc.) with unintentional releases to the environment. The TRI data should be used with caution because only certain types of facilities are required to report. This is not an exhaustive list.

6.2.1 Air

Combustion of coal and other fossil fuels is the primary source of airborne selenium compounds. In air, elemental selenium burns to form selenium dioxide; however, during the combustion of fossil fuels, essentially all of the selenium dioxide produced should be reduced to elemental selenium by the sulfur dioxide that results from the combustion of these materials (NAS 1976a). Estimates of the quantity of selenium released to the air from fossil fuel combustion vary. Estimated annual selenium air emissions from stationary sources in the United States for 1969–1971, 1978, and 1983 were 900, 1,240, and 1,560 tons selenium/year, respectively (EPA 1974; Lee and Duffield 1979). Dulka and Risby (1976) estimated yearly releases of selenium to the air from fossil fuel combustion to be 1,000 tons. Harr (1978)

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Table 6-1. Releases to the Environment from Facilities that Produce, Process, or Use Selenium

		Reported amounts released in pounds per year[a]						
State[b]	Number of facilities	Air[c]	Water	Underground injection	Land	Total on-site release[d]	Total off-site release[e]	Total on and off-site release
CA	2	5	0	No data	29,636	29,641	5	29,646
IL	3	27	No data	No data	No data	27	5,730	5,757
IN	3	0	No data	No data	2,655	2,655	2,445	5,100
LA	1	No data	No data	No data	No data	No data	No data	No data
MI	1	5	No data	No data	No data	5	10	15
NC	1	No data	No data	No data	No data	No data	No data	No data
NY	1	No data	No data	No data	No data	No data	47	47
OH	3	No data	No data	No data	No data	No data	No data	No data
OK	1	No data	285	No data	5	290	No data	290
OR	2	0	No data	No data	118,668	118,668	No data	118,668
PA	4	69,730	No data	No data	453	70,183	5,018	75,201
SC	1	35	No data	No data	116,500	116,535	No data	116,535
UT	1	No data	No data	No data	No data	No data	No data	No data
WY	1	No data	No data	No data	No data	No data	0	0
Total	26	69,802	285	0	267,917	338,004	13,255	351,259

Source: TRI99 2001

[a]Data in TRI are maximum amounts released by each facility.
[b]Post office state abbreviations are used.
[c]The sum of fugitive and stack releases are included in releases to air by a given facility.
[d]The sum of all releases of the chemical to air, land, water, and underground injection wells.
[e]Total amount of chemical transferred off-site, including to publicly owned treatment works (POTW).

Table 6-2. Releases to the Environment from Facilities that Produce, Process, or Use Selenium Compounds

		Reported amounts released in pounds per year[a]						
State[b]	Number of facilities	Air[c]	Water	Underground injection	Land	Total on-site release[d]	Total off-site release[e]	Total on and off-site release
AL	3	4,405	3,800	No data	43,600	51,805	356	52,161
AR	1	474	No data	No data	No data	474	13,946	14,420
AZ	3	1,284	0	No data	640,900	642,184	6,795	648,979
CO	1	No data	No data	No data	No data	No data	No data	No data
FL	2	7,005	No data	No data	350	7,355	15	7,370
GA	6	45,929	461	No data	27,639	74,029	7,470	81,499
IA	5	No data	No data	No data	No data	No data	90	90
ID	1	1,830	No data	No data	97,940	99,770	10	99,780
IL	9	6,333	4,200	No data	23,553	34,086	79,458	113,544
IN	4	4,632	2,185	No data	7,596	14,413	6,012	20,425
KS	1	5	No data	No data	250	255	5	260
KY	4	16,307	5,800	No data	36,896	59,003	No data	59,003
LA	2	193	25	No data	No data	218	27,262	27,480
MA	1	245	0	No data	20	265	5,600	5,865
MD	2	16,101	1,001	No data	730	17,832	4,916	22,748
MI	5	18,306	2,733	No data	799,173	820,212	422,068	1,242,280
MN	1	50	2,300	No data	No data	2,350	90	2,440
MN	1	750	No data	No data	28,000	28,750	1,250	30,000
MO	1	No data	No data	No data	No data	No data	No data	No data
NC	5	65,019	915	No data	32,710	98,644	10	98,654

Table 6-2. Releases to the Environment from Facilities that Produce, Process, or Use Selenium Compounds (*continued*)

		Reported amounts released in pounds per year[a]						
State[b]	Number of facilities	Air[c]	Water	Underground injection	Land	Total on-site release[d]	Total off-site release[e]	Total on and off-site release
NE	4	135	51	0	1,601,025	1,601,211	0	1,601,211
NM	5	1,187	0	No data	157,003	158,190	18,005	176,195
NY	1	0	No data	No data	No data	0	15,151	15,151
OH	7	60,889	9,817	No data	65,491	136,197	35,130	171,327
OK	2	11,595	No data	No data	26,267	37,862	147	38,009
PA	9	52,484	1,596	No data	33,100	87,180	33,616	120,796
SC	1	3,393	No data	No data	8,869	12,262	No data	12,262
TN	2	13,510	5,050	No data	28,305	46,865	500	47,365
TX	13	160,337	177	33,509	287,207	481,230	3,554	484,784
UT	5	9,937	1,255	No data	1,869,154	1,880,346	3,529	1,883,875
WA	1	No data	No data	No data	No data	No data	No data	No data
WI	1	No data	No data	No data	No data	No data	No data	No data
WV	9	77,743	3,700	No data	101,983	183,426	9,265	192,691
WY	2	10,544	No data	No data	35,508	46,052	No data	46,052
Total	121	590,622	45,066	33,509	5,953,269	6,622,466	694,250	7,316,716

Source: TRI99 2001

[a]Data in TRI are maximum amounts released by each facility.
[b]Post office state abbreviations are used.
[c]The sum of fugitive and stack releases are included in releases to air by a given facility.
[d]The sum of all releases of the chemical to air, land, water, and underground injection wells.
[e]Total amount of chemical transferred off-site, including to publicly owned treatment works (POTW).

estimated that 1,500 tons were released annually, with additional air releases from industrial and municipal wastes totaling 2,700 tons and 360 tons, respectively. Selenium releases to the air are likely to increase as more coal is burned in the future. The estimated selenium emissions from Canadian non-ferrous smelters (stack plus fugitive) were 3.02 tons in 1993 (Skeaff and Dubreuil 1997).

Incineration of rubber tires, paper, and municipal waste is an additional source of atmospheric selenium. Hashimoto et al. (1970) reported selenium concentrations in rubber tires to be 1.3 mg/kg. Seventy different kinds of paper have been found to contain selenium (West 1967). Combustion of municipal solid waste results in stack emissions ranging from 0.00098 to 0.00216 pounds (0.44–0.98 g) of selenium per ton of refuse (Johnson 1970).

The amount of selenium contributed to the air by other sources is not known. Microbial action within the soil may also contribute selenium to the air (Fishbein 1983). Selenium biomethylation volatilizes about 3,000 tons of selenium per year into the atmosphere, which eventually returns to earth in rainfall (NAS 1976a). Volcanic gas is suspected to be the major natural source of atmospheric selenium. Certain plants metabolize inorganic selenium compounds to volatile selenium in the forms of dimethyl selenide (Lewis et al. 1971) and dimethyl diselenide (Evans et al. 1968). Animals are also capable of volatilizing selenium and releasing dimethyl selenide in expired air (Schultz and Lewis 1940).

Fly ash settling ponds (which contain high concentrations of selenium) and hazardous waste sites where selenium compounds were disposed of in the past are potential sources of atmospheric selenium through fugitive dust emissions. Selenium emissions from these potential sources have not been quantified.

According to TRI, an estimated total of at least 69,802 pounds of selenium metal and 590,622 pounds of selenium compounds were discharged to the air from manufacturing and processing facilities in the United States in 1999 (TRI99 2001) (see Tables 6-1 and 6-2). The data listed in the TRI tables should be used with caution since only certain types of facilities are required to report. This is not an exhaustive list.

Selenium has been identified in air at 13 sites collected from the 494 NPL hazardous waste sites where it was detected in some environmental media (HazDat 2001).

6.2.2 Water

Surface waters can receive selenium from the atmosphere by dry and wet deposition, from adjoining waters that may contain selenium, from surface runoff, and from subsurface drainage. Sewage treatment plants are another source of selenium releases to water. Effluents from sewage treatment plants and oil refineries appear to be the major sources of selenium in the San Francisco estuarine system (Cutter 1989). In a study of direct discharges from oil refineries in San Francisco Bay, the average selenium concentration in the effluent was 0.067 mg/L with a range of 0.0066–0.156 mg/L (Barceloux 1999; Cutter 1989). Approximately 50–76% of the total selenium in the effluents was selenite. This proportion of selenite is higher than that found in natural estuary sources in the San Francisco Bay (Cutter 1989). About 150,000–460,000 tons of selenium per year are deposited in coal fly ash (Andren and Klein 1975; Doran 1982). Selenium from fly ash settling ponds and hazardous waste sites could reach surface water via runoff or could reach groundwater via leaching. Concentrations of 0.10–0.25 mg/L in a settling basin effluent from coal fly ash in North Carolina were reported by Lemly (1985). Overflow from the ash basin of a coal fired electric generating facility to Belews Lake resulted in surface water selenium concentrations of 0.005–0.020 mg/L in the lake basin. These levels have been reduced considerably since 1986 when the discharge of selenium laden waste water to the lake was discontinued. The peak selenium concentration in 1996 was <0.001 mg/L (Lemly 1997). Selenium concentrations as high as 0.28 mg/L have been reported for raw sewage, 0.045 mg/L for primary effluent, and 0.050 mg/L for secondary effluent (Baird et al. 1972). Irrigation drainage from seleniferous soils can increase selenium concentrations in surface water and has resulted in levels that are toxic to wildfowl at Kesterson National Wildlife Refuge in California (Ohlendorf et al. 1986a, 1988).

Selenium was found to be released during coal mining because of the oxidation of selenium-bearing pyrite (Dreher and Finkelman 1992).

According to the TRI, an estimated total of 285 pounds of selenium metal and 45,066 pounds of selenium compounds were discharged to water from manufacturing and processing facilities in the United States in 1999 (TRI99 2001) (see Tables 6-1 and 6-2). The data listed in the TRI tables should be used with caution since only certain types of facilities are required to report. This is not an exhaustive list.

Selenium has been identified in groundwater at 266 sites and surface water at 102 sites collected from the 494 NPL hazardous waste sites where it was detected in some environmental media (HazDat 2001).

6.2.3 Soil

The primary factor that controls selenium concentrations in soil is the selenium content of the parent bedrock materials that release selenium via weathering processes and leaching (NAS 1976a). Natural weathering processes are thought to release about 100,000–200,000 metric tons of selenium per year (Andren and Klein 1975). Atmospheric deposition of selenium also contributes to selenium in the soil. In the past, selenium was used in pesticide products, but because of its stability in soils and subsequent contamination of food crops, its use in pesticide products is now restricted. The release of selenium to soil from fly ash settling ponds and hazardous waste sites has not been quantified.

According to the TRI, an estimated total of 267,917 pounds of selenium metal and 5,953,269 pounds of selenium compounds were discharged to land from manufacturing and processing facilities in the United States in 1999 (TRI99 2001). In addition, 33,509 pounds of selenium compounds were injected underground (see Table 6-2). The data listed in the TRI tables should be used with caution since only certain types of facilities are required to report. This is not an exhaustive list.

Selenium has been identified in soil at 182 sites and sediment at 108 sites collected from the 494 NPL hazardous waste sites where it was detected in some environmental media (HazDat 2001).

6.3 ENVIRONMENTAL FATE

The behavior of selenium in the environment is influenced to a large degree by its oxidation state and the consequent differences in the behavior of its different chemical compounds (Callahan et al. 1979c; NAS 1976a). The oxidation state of selenium in the environment is dependent on ambient conditions, particularly on pH, pE, and biological activity (Maier et al. 1988).

6.3.1 Transport and Partitioning

The volatile selenium compounds that partition into the atmosphere include the inorganic compounds selenium dioxide and hydrogen selenide and the organic compounds dimethyl selenide and dimethyl diselenide. Hydrogen selenide is highly reactive in air and is rapidly oxidized to elemental selenium and water (NAS 1976a), but the other compounds can persist in air.

Selenium compounds released to the atmosphere can be removed from it by dry or wet deposition to soils or to surface water. The annual wet deposition rate of selenium at two rural/agricultural sites in Queenstown, Maryland and St. Mary's, Maryland were 287 and 140 μg/m^2-year, respectively (Scudlark et al. 1994). Selenium concentrations ranging from 0.04 to 1.4 μg/L have been detected in rain and snow (Hashimoto and Winchester 1967). Kubota and coworkers (1975) reported selenium concentrations of 0.02–0.37 μg/L in rainwater at several locations in the United States and Denmark. Selenium was detected at average concentrations of 5.60–7.86 μg/L during four rainfall events in Riyadh, Saudi Arabia (Alabdula'aly and Khan 2000).

The forms of selenium expected to be found in surface water and the water contained in soils are the salts of selenic and selenious acids. Selenic acid (H_2SeO_4) is a strong acid. The soluble selenate salts of this acid are expected to occur in alkaline waters. Sodium selenate is one of the most mobile selenium compounds in the environment because of its high solubility and inability to adsorb onto soil particles (NAS 1976a). Selenious acid (H_2SeO_3) is a weak acid, and the diselenite ion predominates in waters between pH 3.5 and 9. Most selenites are less soluble in water than the corresponding selenates (NAS 1980b).

Selenium in an aquatic environment is bioaccumulated by aquatic organisms (Chau and Riley 1965; Ohlendorf et al. 1986a; Rudd and Turner 1983a; Saiki and Lowe 1987). Lemly (1985) has reported bioconcentration factors (BCFs) of 150–1,850 and bioaccumulation factors (BAFs) of 1,746–3,975 for selenium in freshwater. In the Kesterson National Wildlife Refuge in the San Joaquin Valley of California, elevated levels of selenium have been measured (dry weight) in algae (average 35 mg/kg), midge larvae (139 mg/kg), dragonfly and damselfly nymphs (average 122 and 175 mg/kg, respectively), and mosquito fish (170 mg/kg) (Ohlendorf et al. 1986b). For comparison, the mean concentrations of selenium found in fish throughout the United States in the 1976–1977, 1978–1979, and 1980–1981 National Pesticide Monitoring Program were 0.56, 0.46, and 0.47 mg/kg wet weight, respectively (Lowe et al. 1985; May and McKinney 1981; Ohlendorf et al. 1986b). Similarly, Lemly (1985) found elevated selenium concentrations in aquatic organisms living in a power plant cooling reservoir in North Carolina. The degree of bioaccumulation of selenium exhibited a stable pattern over several years, with selenium concentrations (wet weight) as follows: fish (6–35 mg/kg) > benthic insects (12–15 mg/kg) > annelids (10–12 mg/kg) > molluscs and crustaceans (5–9 mg/kg) > periphyton (4–6 mg/kg) (Ohlendorf et al. 1986a). In fish, selenium was concentrated in visceral tissue (25–35 mg/kg wet weight) more than in skeletal muscle (6–11 mg/kg wet weight). Adams (1976) reported BCFs of 62.1, 14.3, 6.3, 3.2, and 10.5 for selenium in the viscera, gill, head and tail, muscle, and whole trout, respectively. The BCFs and

BAFs for selenium in visceral tissue (i.e., heart, hepatopancreas, spleen, and gonads) of fish have been estimated to range from 35 to 1,850 and from 1,058 to 3,980, respectively (Lemly 1982, 1985). Lemly (1985) also estimated BAFs for selenium in skeletal muscle of fish to range from 485 to 1,746, depending on the species. Maier et al. (1988) estimated selenium BAFs for algae to range from 100 to 2,600, and Besser et al. (1993) estimated BCFs of 16,000 for algae, 200,000 for daphnids, and 5,000 for bluegills from exposures to 1 μg/L selenomethionine. Selenite was more concentrated than selenate for algae and daphnids, whereas bluegills concentrated both inorganic species about equally (Besser et al. 1993). Selenium accumulation from selenomethionine occurred more readily than from selenite or selenate (Besser et al. 1989).

Some evidence indicates that under natural conditions selenium might biomagnify in aquatic organisms (Lemly 1985; Maier et al. 1988; NCDNR 1986; Sandholm et al. 1973). Biomagnification is evidenced by progressively higher concentrations of an element or substance in organisms at successively higher trophic levels. More than 50% of the selenium contained in sediments in the ponds and the reservoir in the Kesterson National Wildlife Refuge in California occur in organic forms (Maier et al. 1988), resulting from the synthesis and bioaccumulation of organic selenium before the plants die and decay on the bottom.

In soils, pH and Eh are determining factors in the transport and partitioning of selenium. Elemental selenium is essentially insoluble and may represent a major inert "sink" for selenium introduced into the environment under anaerobic conditions (NAS 1976b). Heavy metal selenides and selenium sulfides, which are also insoluble, predominate in acidic (low pH) soils and soils with high amounts of organic matter. Selenium in this form is immobile and will remain in the soil. The selenides of other metals such as copper and cadmium are of low solubility (NAS 1976b). Sodium and potassium selenites dominate in neutral, well-drained mineral soils, where some soluble metal selenites may be found as well. In alkaline (pH>7.5), well-oxidized soil environments, selenates are the major selenium species. Because of their high solubility and low tendency to adsorb onto soil particles, the selenates are very mobile (Kabatas-Pendias and Pendias 1984) and are readily taken up by biological systems (Klaassen et al. 1986) or leached through the soil. Gerritse et al. (1982) found selenium to be very mobile in sewage sludge leachate. They reported a K_ds (distribution coefficient = [concentration of selenium sorbed on soil or rock]/[concentration of selenium in solution]) of 14.9 mL/g for sandy loam and 5.91 mL/g for sludge-treated sandy soils. Selenite forms stable ferric oxide-selenite adsorption complexes in acid or neutral soils (Geering et al. 1968).

When environments favor the soluble forms of selenium (alkaline and oxidizing conditions), these forms can be accumulated by plants. In addition, although both selenite (Se^{4+}) and selenate (Se^{6+}) are soluble forms of selenium, selenate was found to be the preferred form of selenium taken up by plants (Banuelos and Meek 1990). Preferential uptake of selenate may be caused by its tendency to be less strongly adsorbed to soil particles and organic matter than selenite (Banuelos and Meek 1990). Selenium uptake by plants is influenced by many factors including soil type, pH, colloidal content, concentration of organic material, oxidation-reduction potentials in the root-soil environment, and total level of selenium in the soil (Fishbein 1983; Robberecht et al. 1982). In acidic soils (pH 4.5–6.5) and under high moisture conditions, selenium is in the form of selenite and is bound to colloids as iron hydroxide selenium complexes. These complexes are insoluble and generally not bioavailable to plants (Galgan and Frank 1995). In basic soils (pH 7.5–8.5), selenium is present as soluble selenate. Soluble selenates (principally sodium selenate) appear to be responsible for most of the naturally occurring accumulation of high levels of selenium by plants, although much of the total selenium in soil may be present in other forms (NAS 1976a). The use of lime and plant ash as fertilizers, which would raise the pH of the soil and favor the formation of selenate, has been implicated as a contributing factor in the accumulation of selenium in crops grown in high selenium soil found in certain regions of China (Yang et al. 1988).

6.3.2 Transformation and Degradation

6.3.2.1 Air

Selenium dioxide released to the air from the combustion of fossil fuels should be largely reduced to elemental selenium by sulfur dioxide formed during the combustion (NAS 1976b). During a 1991 study, Oehm et al. found that selenium dioxide reacting with atmospheric moisture generates selenious acid aerosols. Hydrogen selenide is unstable in air and is oxidized to elemental selenium and water (NAS 1976a). Hazards from hydrogen selenide are expected, therefore, to be confined to occupational settings where the confined gas might build up to hazardous levels despite oxidative losses (NAS 1976a). Dimethyl selenide and methyl selenide are volatile organic compounds that can partition into and persist in the atmosphere. Other selenium compounds released to the atmosphere as dust can be removed by wet deposition (in rain or snow) or by dry deposition.

6.3.2.2 Water

In general, the more soluble and mobile forms of selenium (e.g., selenite and selenate) dominate under aerobic (high oxygen concentrations) and alkaline (high pH) conditions (NAS 1976a; Shamberger 1981). Selenates have been predicted thermodynamically to predominate under aerobic conditions, but a review of the literature indicates that both selenites and selenates are equally common in surface waters (Robberecht and Van Grieken 1982). For selenites in solution, equilibria will be set up between H_2SeO_3, $HSeO_3^-$, SeO_3^-, $HSe_2O_5^-$, and $Se_2O_5^{2-}$. The relative concentrations of these species will be determined by the pH of the solution and the total concentration of the electrolytes. Between pH 3.5 and 9, dissolved selenite would be expected to be present predominantly as the diselenite ion, whereas dissolved selenate would occur predominantly as SeO_4^{2-}. Sodium predominates as the counter ion of selenate and selenite in most surface waters.

A study completed by Bender et al. (1991) using a simulated laboratory pond found that bacteria and cyanobacteria have two possible mechanisms for the uptake and transformation of selenate. The uptake mechanism involves the reduction of selenate to elemental selenium that will be physically held within the biological mat. The microorganisms were also found to cause the transformation of soluble selenium into volatile alkyl selenium compounds (Bender et al. 1991).

In some deep aquifers, selenium transport in groundwater was found to be strongly retarded (White et al. 1991). This phenomena is thought to be caused by chemical reduction and precipitation mediated by microbial activity.

Under acidic conditions, selenite can be rapidly reduced to elemental selenium by mild reducing agents such as ascorbic acid or sulfur dioxide (NAS 1980b). Selenate can be converted to selenite or elemental selenium in aquatic systems, but this reaction is slow relative to other transformations. Once formed, elemental selenium is stable over a wide range of pH values and a range of mildly oxidizing to reducing conditions. The formation of various metal selenides is favored by acidic and reducing conditions (NAS 1976b), as found in organic-rich sediments.

Aquatic organisms can convert selenium to both inert and soluble forms. Duckweed, phytoplankton, bacteria, and fungi have been demonstrated to synthesize selenoamino acids from absorbed inorganic selenium compounds (Maier et al. 1988). These selenoamino acids are not likely to be found at significant dissolved concentrations in water, however, because amino acids are rapidly catabolized by

bacteria. Benthic bacteria and fungi are capable of methylating elemental and inorganic selenium salts (Chau et al. 1976). Hydrogen selenide can be formed in a reducing environment (Cutter 1982; NAS 1976a). Both hydrogen selenide and the methylated forms of selenium are unstable in water and would be expected to rapidly volatilize to the atmosphere (Fishbein 1983).

6.3.2.3 Sediment and Soil

In soils, elemental selenium and inorganic selenium compounds such as sodium selenite can be methylated by microorganisms and subsequently volatilized to the atmosphere (Doran 1982; Fishbein 1983; Shamberger 1981). Microorganisms such as *Aeromonas*, *Flavobacterium*, and *Pseudomonas* are suspected of methylating inorganic and organic selenium compounds to dimethyl selenide and dimethyl diselenide (Doran and Alexander 1976; Fishbein 1983; Reamer and Zoller 1980). Microbes cultured from rhizosphere of bulrush (*Scirpus robustus*) plants were shown to biomethylate soluble selenate and selenite and substantially volatilize these compounds over a 15-day incubation period (Azaizeh et al. 1997). Temperature plays a significant role in the microorganism-mediated volatilization of selenium compounds; temperature reduction from 20 to 10E C resulted in a 25% reduction, and from 20 to 4 EC resulted in a 90% reduction in the dimethyl selenide produced (Chau et al. 1976). Reamer and Zoller (1980) examined microbial transformation of selenium in aerobic sewage contaminated with elemental selenium and selenite. They found dimethyl selenide to be the principal microbial product at low selenite concentrations (1–10 mg/kg), whereas dimethyl diselenide and dimethyl selenone were the principal products at higher selenite concentrations (100–1,000 mg/kg). Dimethyl selenide was the only product recovered from sludge contaminated with elemental selenium (Reamer and Zoller 1980). In general, microorganisms appear to methylate organic selenium compounds more readily than either selenite or selenate (Maier et al. 1988). Elemental selenium is converted to methylated selenium compounds the least rapidly (Maier et al. 1988). Selenium methylation and subsequent return from the atmosphere as selenite in rainwater is likely to be the major natural process by which selenium cycling occurs in the environment (Doran 1982).

Demethylation of the trimethylselenonium ion can also occur in soil. Microorganisms are evidently required for this reaction since it did not occur in autoclaved soil (Yamada et al. 1994). Selenium added to the soil as trimethylselenonium was not recovered in the soil, suggesting that trimethylselenonium was demethylated to gaseous selenium compounds, for example, dimethylselenide.

Terrestrial plants take up soluble selenate and selenite and biosynthesize organic selenium compounds, predominantly selenomethionine and, to a lesser extent, selenocysteine. Selenates tend to be taken up by plants from soils more readily than selenites, in part because selenites tend to adsorb more strongly to soils (Dimes et al. 1988; Zhang et al. 1988). These compounds can be released to the soils once the plants die and decay. Water-soluble organic selenium compounds are also probably readily taken up by plants (Shamberger 1981; Shrift 1964).

6.4 LEVELS MONITORED OR ESTIMATED IN THE ENVIRONMENT

Selenium can be detected in most biological and nonbiological materials in the environment. Selenium occurs in aquatic and terrestrial organisms as well as in water, air, and soil. Among foods consumed by humans, meat products generally contain the highest concentration of selenium while vegetables and fruits contain the lowest. Brazil nuts contain extremely high levels of selenium since they grow in the foothills of the Andes Mountains, where the soils are high in selenium (Secor and Lisk 1989). Cereals contain intermediate levels of selenium.

6.4.1 Air

Background ambient air concentrations of selenium are generally in the nanogram per cubic meter (ng/m^3) range (Harrison et al. 1971; John et al. 1973; Peirson et al. 1973). Dams et al. (1970) found concentrations of selenium in suspended air particulate matter of 2.5 ng/m^3 in Niles, Michigan, and 3.8 ng/m^3 in East Chicago, Indiana. During 1968–1969, 18 air samples collected around Buffalo, New York, showed a range of 3.7–9.7 ng/m^3 (Pillay et al. 1971). Based on these results, the National Academy of Sciences (NAS 1976a) has estimated that the average selenium concentration in the air is well below 10 ng/m^3. A monitoring study to determine the seasonal variation of pollutants in the air of Alaska was conducted from 1984 to 1987 (Sturges and Shaw 1994). The average concentration of selenium in Poker Flats, Alaska was 0.035 ng/m^3 (June 1 through January 31, 1984–1987) and 0.067 ng/m^3 (February 1 through May 31, 1984–1987). The nearly 2-fold increase in concentration during the spring months were attributed to local marine biogenic volatilization of selenium, and not a coal burning origin (Sturges and Shaw 1994). Selenium was detected in the ambient atmosphere at seven sites in the United Kingdom at concentrations ranging from 0.1 to 42.3 ng/m^3 (Lee et al. 1994). The lowest levels were observed in the rural areas of Chilton and Windermere with mean concentrations of 1.3 and 0.9 ng/m^3, respectively. The highest mean concentration of 16.7 ng/m^3 was observed in the industrial area of Walsall.

6.4.2 Water

Selenium has been detected in surface waters and groundwaters in the United States at generally low concentrations. Selenium has been detected in oceans at an average value of $9x10^{-5}$ mg/L (0.09 µg selenium/L) (Schutz and Turekiam 1965). In a study of selenium concentrations in major watersheds of the United States, selenium was detected in only 2 of 535 samples (<0.5%) at a concentration greater than the lowest detection limit of 0.010 mg/L (Lakin and Davidson 1967). Examination of the EPA STORET database for the state of North Carolina revealed that only 3.3% of 657 samples of surface water contained more than 0.001 mg/L, and the highest value was 0.012 mg/L (NCDNR 1986). Watersheds that receive selenium contaminated waste water have high levels of selenium in surface water samples. The selenium concentration in Lake Belews, North Carolina has dropped from a maximum value of about 0.020 mg/L (pre-1986) to <0.001 mg/L in 1996, due to the discontinued release of selenium laden waste water from a local coal fired power plant (Lemly 1997). The selenium concentration in portions of Pigeon River and Pigeon Lake, Michigan that receive waste water input from a coal fly ash disposal facility were <0.001–0.0075 mg/L (Besser et al. 1996).

High selenium levels are more likely to be found in irrigation return waters, seeps, springs, and shallow wells where seleniferous soils may contribute to the selenium content of the water. Glover et al. (1979) found that under unusual geological conditions, selenium concentrations in groundwater may reach 0.60 mg/L. In another study conducted in a seleniferous area of South Dakota, 34 of 44 wells did not show any measurable selenium; however, the remaining 10 had concentrations ranging from 0.050 to 0.339 mg/L (Smith and Westfall 1937). Selenium concentrations determined in 107 irrigation and 44 livestock well waters in the San Joaquin area of California exceeded 0.010 mg/L in 26 wells, but exceeded 0.020 mg/L in only 11 wells (Oster et al. 1988a). The maximum concentration was 0.272 mg/L (Oster et al. 1988a).

Selenium accumulation in agricultural drainage waters and basins has been documented in the western United States, particularly in California. The problem was first discovered in the Kesterson Wildlife Refuge in the San Joaquin Valley of California. In 1975, the U.S. Bureau of Reclamation finished construction of an 85-mile subsurface agricultural water drain that terminated in a series of evaporation ponds called Kesterson Reservoir. By 1983, however, it was confirmed that the drain waters contained high concentrations of selenium (>1.35 mg/L in some areas) leached from the soil by application of irrigation water (Maier et al. 1988). Because the high selenium levels produced death and deformities in fish and waterfowl, delivery of subsurface water to Kesterson was terminated in 1986 (Lewis 1988).

Measurements of trace elements in the 27 other evaporation basins in the San Joaquin Valley have revealed only 3 basins with total selenium exceeding 0.10 mg/L and only 50 acres of evaporation basin cells with selenium concentrations in excess of 1.0 mg/L (CRWQCB 1988).

6.4.3 Sediment and Soil

Selenium is estimated to be the 69th most abundant element in the earth's crust, with an average concentration of 0.05–0.09 mg/kg (Glover et al. 1979). Chemically, selenium closely resembles sulfur. Consequently, sulfides of bismuth, iron, mercury, silver, copper, lead, and zinc have been found to contain selenium (Shamberger 1981). Selenium is concentrated in the sulfide minerals galena, chalcopyrite, arsenopyrite, sphalerite, pyrite, marcasite, and pyrrhotite (Coleman and Delevaux 1957). Jarosite and barite have also been found to contain selenium at low levels. The sulfides containing the highest selenium concentrations are those associated with uranium ores in sandstone-type deposits in the western United States. In the immediate vicinity of sandstone-type uranium deposits, selenium concentrations as high as 1,000 mg/kg have been found (Shamberger 1981). Hydrothermal ore is also known to contain high concentrations of selenium. The best known are epithermal gold, silver, antimony, and mercury deposits (Shamberger 1981). Selenium has been found in volcanic rocks in the western United States at concentrations as high as 120 mg/kg (Glover et al. 1979).

Various studies estimated selenium concentration of most soils to be between 0.01 and 0.2 mg/kg (Sindeeva 1964). One study analyzed several thousand soil samples in the United States and found that most seleniferous soils contained <2 mg/kg, with a maximum concentration of <100 mg/kg (Rosenfeld and Beath 1964). The highest U.S. soil levels of selenium are found in areas of the West and Midwest.

Atmospheric deposition of selenium from mining and smelting activities also appears to be a source to soils and plants (Glooschenko and Arafa 1988). In this study, an indirect relationship between distance from smelters and selenium concentration was shown using *Sphagnum fuscum* as an indicator. Washout of atmospheric selenium by precipitation appeared to be the primary mechanism for accumulating selenium in soils and plants in the vicinity of smelters (Glooschenko and Arafa 1988).

Sandstone has been found to contain selenium in varying concentrations, but most probably contains <1 mg/kg (Rosenfeld and Beath 1964). However, sandstone in Wyoming has been found to contain >100 mg/kg (Beath et al. 1946). Generally, the selenium concentration of limestone is low; however,

shales of the Niobrara formation in South Dakota have been found to contain over 40 mg/kg. The range of selenium concentrations in phosphate rocks is <1–300 mg/kg (NAS 1976a).

The shales appear to contain consistently higher concentrations of selenium than limestone or sandstone. Despite the fact that the shales vary so widely in their selenium concentration, they are fairly reliable indicators of soils high in selenium (NAS 1976a).

The disposal of selenium contaminated waste water has resulted in elevated selenium levels in sediments of Lake Belews, North Carolina. The concentration of selenium in sediments ranged from 4 to 12 μg/g (pre-1986), but has dropped to 1–4 μg/g (1996) due to the discontinued release of selenium laden waste water from a local coal fired power plant (Lemly 1997). Selenium was measured in 445 surface soil samples from Florida with a concentration range of 0.01–4.62 μg/g and an arithmetic mean of 0.25 μg/g (Chen et al. 1999). Selenium was detected in soils and bed sediment from the South Platter River Basin at concentrations of 0.30–3.80 μg/g (Heiny and Tate 1997). The highest levels were observed in areas consisting of a high degree of Precambrian rock formation.

6.4.4 Other Environmental Media

Coal and Oil. Petroleum has been found to contain 500–950 mg/kg crude petroleum and 500–1,650 mg/kg heavy petroleum (Hashimoto et al. 1970). An average of 2.8 mg/kg coal has been reported for 138 samples from U.S. deposits (Pillay et al. 1969).

Plants. Several species of grasses and herbaceous plants accumulate selenium, and some of these are endemic to the western United States. Primary accumulators are *Astragalus*, *Oonopsis*, *Stanelya*, *Xylorhiza*, and *Machaeranthera*. Secondary accumulators are *Astor*, *Gatierreaia*, *Atriplex*, *Grindelia*, *Castillaja*, and *Comandra*. Primary accumulators can contain 100–100,000 mg/kg of plant tissue, whereas secondary accumulators contain 25–100 mg/kg of plant tissue (dry weight). Nonaccumulator plants generally contain less than 25 mg of selenium/kg of plant tissue (dry weight) (Rosenfeld and Beath 1964). In some plants, including the leaves of beets and cabbage, and in garlic, as much as 40–50% of the selenium may be in the form of selenate (Cappon 1981).

A study by Arthur et al. (1992) showed an increased uptake of selenium by terrestrial plants growing on soil-capped fly ash landfill sites. The selenium concentrations rarely exceeded 5 mg/kg, and there were no signs of selenium toxicity to plants. A similar study by Shane et al. (1988) on greenhouse vegetables

established that the uptake of selenium by these vegetables is proportional to the percentage of selenium in the growth medium. Another greenhouse study showed that four floating aquatic plants, *Azolla caroliniana*, *Eichjornia crassipes*, *Salvinia rotundifolia*, and *Lemna minor*, absorbed selenium quickly upon exposure (Horne 1991).

Animals. Aquatic animals accumulate selenium from lakes and rivers high in selenium content. Fish in the Kesterson National Wildlife Refuge in California had selenium concentrations up to 96 mg/kg, and aquatic birds had levels up to 130 mg/kg (Barceloux 1999). Selenium was detected in fish from three sites of the Pigeon River and Pigeon Lake in Michigan (Besser et al. 1996). It was determined that selenium concentrations in fish at sites receiving seepage and effluents from a coal fly ash disposal facility were considerably higher than for fish upstream from the facility. Mean concentrations of selenium in white sucker and northern pike ranged from 0.46 to 0.88 μg/g in a noncontaminated portion of the river, while concentrations in a contaminated portion of the river and lake were 1.1–2.4 μg/g (Besser et al. 1996). The mean concentrations of selenium in the feathers of five species of birds at Clear Lake, California were 3.20 μg/g (osprey), 1.38 μg/g (western grebe), 2.51 μg/g (great blue heron), 0.94 μg/g (turkey vulture), and 1.05 μg/g (mallard) (Cahill et al. 1998). The ospreys, which consume large mature fish had the highest selenium levels, while turkey vultures which rarely interact with the contaminated aquatic system and mallards which are semi-domesticated had the lowest levels. Selenium was detected in 24 of 24 black-crowned night herons from the Delaware Bay at concentrations of 2.84–5.95 μg/g (Rattner et al. 2000). The greatest levels were observed in the herons from Pea Patch Island, an island adjacent to a shipping channel for the petrochemical industry. Selenium was detected in the liver of 70 out of 70 redheads (*Athya americana*) in Louisiana and Texas at concentrations of 1.56–5.86 μg/g (Michot et al. 1994). The selenium concentration in moose liver from 12 areas of Sweden ranged from 0.0027 to 3.054 μg/g (Galgan and Frank 1995). The highest levels were observed in areas with a high degree of selenium deposition from industrial sources.

Food. In a review of the foods that contribute the highest proportion of the daily selenium intake of human populations in the United States, Schubert et al. (1987) estimated selenium concentrations in over 100 food items on the basis of 65 articles published after 1960. Table 6-3 presents the selenium concentrations for some of the food items analyzed. The quality of the data was evaluated on the basis of sample size, analytic method, sample handling, sampling plan, and analytic quality control. Schubert et al. (1987) chose not to present standard deviations or standard errors of the samples because of the different sampling biases present in the studies.

Table 6-3. Selenium Concentrations in Foods in the United States[a] (mg selenium/kg, wet weight)

Food item	Average	Minimum	Maximum	Number of acceptable samples
Fruits and vegetables				
Apples, raw	0.004	0.003	0.006	5
Carrots, raw	0.017	0.006	0.029	5
Oranges	0.015	0.013	0.018	3
Potatoes	0.013	0.004	0.023	7
Grains, nuts, and cereals				
Bread, white	0.32	0.23	0.54	6
Bread, whole wheat	0.44	0.28	0.67	3
Corn flakes	0.063	0.026	0.12	4
Special K	0.063	0.35	0.94	4
Egg noodles, dry	0.66	0.43	1.35	7
Egg noodles, cooked	0.19	0.14	0.42	2
Nuts, Brazil[b]	14.7	0.20	253	72
Dairy products				
Whole mile	0.016	0.011	0.025	4
Swiss cheese	0.083	0.062	0.10	2
Cottage cheese	0.060	0.052	0.068	2
Meat				
Chicken, cooked	0.21	0.17	0.26	2
Beef, cooked	0.26	0.15	0.52	3
Pork/ham, fresh/cured	0.33	0.19	0.51	6
Salami	0.20	0.13	0.33	2
Seafood[c]				
Salmon, canned	0.75	0.31	1.49	3
Shrimp, canned/cooked	0.64	0.21	1.61	4
Swordfish	2.84	2.54	3.44	4
Organ meats				
Beef liver, cooked	0.56	0.43	0.71	2
Beef kidney, raw	1.70	1.45	2.32	4

[a]Food is the normal source of selenium which is essential for human health. Concentrations from Schubert et al. (1987) except where noted.
[b]Secor and Lisk (1989)
[c]Bioavailability of selenium from some fish may be lower than from other foods.

In general, fruits and vegetables were found to contain less than 0.01 mg/kg, whereas root vegetables contained higher concentrations of selenium (Table 6-3). Beale et al. (1990) found milk and meat to have the same range of selenium concentrations as Schubert et al. (1987). In another study, no apparent correlation existed between the selenium concentration of canned versus fresh fruits and vegetables (Morris and Levander 1970).

Grain products varied greatly in their selenium concentration. Wheat bread and flour were high in selenium, whereas white bread and white flour contained considerably less selenium. Very low levels of selenium were found in certain processed cereals, such as corn flakes, but not in others, such as oat cereal (Morris and Levander 1970; Schubert et al. 1987).

Dairy products contained variable concentrations of selenium as well but in general contained lower levels than did meat products. Organ meats (e.g., liver and kidney) and seafoods contained higher levels of selenium than poultry or beef (Morris and Levander 1970; Schubert et al. 1987). The U.S. Fish and Wildlife Service collected 315 whole fish samples from 109 stations nationwide and analyzed them for selenium. The resulting selenium concentrations were as follows (wet weight): geometric mean of 0.42 μg/g, maximum 2.3 μg/g, and 85th percentile concentration of 0.73 μg/g (Schmitt and Brumbagh 1990). Consumption of the foods with higher selenium levels contributes to the daily intake of adequate amounts of selenium.

Analysis of commercial baby foods indicated that processing may reduce selenium levels of the food (Morris and Levander 1970).

A recent survey conducted by the U.S. Food and Drug Administration (FDA) which analyzed foods consumed in the United States during the period of 1991–1999, detected selenium in 3,654 out of 6,671 food samples analyzed (FDA 2000). The results of this survey are summarized in Table 6-4.

6.5 GENERAL POPULATION AND OCCUPATIONAL EXPOSURE

Because selenium is ubiquitous in the environment and has been detected in so many media, exposure of the general population to selenium can occur in a variety of ways, including occupational exposure, inhalation, and ingestion of selenium via drinking water, foods, and selenium supplements. For exposure via the food pathway, Schubert et al. (1987) estimated that beef, white bread, pork or ham, chicken, and eggs provide over 50% of the daily selenium intake in the U.S. population. The FDA (1982a) estimated

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
	Overall:	6,671	3,025	1,206	0.07	0.12	0	1.8	0.017
Whole milk, fluid	1	25	5	18	0.019	0.012	0	0.044	0.02
Lowfat (2% fat) milk, fluid	2	25	6	14	0.022	0.015	0	0.056	0.025
Chocolate milk, fluid	3	25	5	15	0.021	0.014	0	0.054	0.023
Skim milk, fluid	4	25	5	14	0.024	0.016	0	0.058	0.025
Plain yogurt, lowfat	6	25	5	9	0.031	0.019	0	0.068	0.033
Chocolate milk shake, fast-food	7	25	5	17	0.023	0.014	0	0.051	0.026
Evaporated milk, canned	8	25	4	4	0.043	0.024	0	0.102	0.047
American, processed cheese	10	25	0	3	0.183	0.025	0.097	0.231	0.178
Cottage cheese, 4% milkfat	11	25	2	4	0.083	0.039	0	0.178	0.08
Cheddar cheese	12	25	0	4	0.198	0.045	0.1	0.318	0.194
Ground beef, pan-cooked	13	25	0	1	0.197	0.052	0.127	0.333	0.187
Beef chuck roast, baked	14	25	0	0	0.251	0.058	0.15	0.379	0.24
Beef steak, loin, pan-cooked	16	25	0	1	0.256	0.063	0.13	0.439	0.24
Ham, baked	17	25	0	1	0.29	0.077	0.12	0.42	0.278
Pork chop, pan-cooked	18	25	0	0	0.46	0.16	0.245	0.808	0.448
Pork sausage, pan-cooked	19	25	0	4	0.215	0.094	0.066	0.556	0.207
Pork bacon, pan-cooked	20	25	0	0	0.38	0.15	0.186	0.836	0.323
Pork roast, baked	21	25	0	1	0.34	0.11	0.13	0.692	0.333
Lamb chop, pan-cooked	22	25	0	2	0.25	0.13	0.095	0.74	0.22
Chicken, fried (breast, leg, and thigh) homemade	24	25	0	2	0.25	0.1	0.067	0.465	0.243
Turkey breast, roasted	26	25	0	0	0.34	0.14	0.095	0.583	0.329
Liver, beef, fried	27	25	0	0	0.65	0.25	0.089	1.22	0.67
Frankfurters, beef, boiled	28	25	2	3	0.098	0.037	0	0.155	0.102
Bologna, sliced	29	25	0	5	0.134	0.037	0.07	0.239	0.13

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Salami, sliced	30	25	0	3	0.202	0.046	0.079	0.313	0.197
Tuna, canned in oil	32	26	0	0	0.69	0.13	0.498	1.013	0.655
Fish sticks, frozen, heated	34	26	0	1	0.168	0.035	0.076	0.257	0.171
Eggs, scrambled	35	26	0	1	0.217	0.073	0.076	0.405	0.206
Eggs, fried	36	25	0	0	0.278	0.084	0.149	0.454	0.259
Eggs, boiled	37	25	0	2	0.27	0.1	0.023	0.477	0.274
Pinto beans, dry boiled	38	25	2	6	0.076	0.043	0	0.13	0.064
Pork and beans, canned	39	25	5	10	0.034	0.023	0	0.076	0.038
Lima beans, immature, frozen, boiled	42	25	17	8	0.005	0.009	0	0.036	0
Green peas, fresh/frozen, boiled	46	25	18	5	0.007	0.013	0	0.044	0
Peanut butter, smooth	47	25	2	8	0.086	0.068	0	0.271	0.073
Peanuts, dry roasted	48	25	5	5	0.075	0.068	0	0.272	0.063
White rice, cooked	50	25	3	4	0.057	0.035	0	0.17	0.055
Oatmeal, quick (1–3 minutes), cooked	51	25	2	4	0.058	0.034	0	0.18	0.052
Wheat cereal, farina, quick (1–3 minutes), cooked	52	25	3	3	0.076	0.047	0	0.205	0.069
Corngrits, regular, cooked	53	25	6	13	0.025	0.025	0	0.095	0.019
Corn, fresh/frozen, boiled	54	25	17	6	0.007	0.012	0	0.034	0
Cream style corn, canned	56	25	18	7	0.005	0.008	0	0.029	0
Popcorn, popped in oil	57	26	5	7	0.083	0.071	0	0.267	0.073
White bread	58	25	0	3	0.211	0.075	0.05	0.363	0.197
White roll	59	25	0	0	0.265	0.076	0.144	0.41	0.266
Cornbread, homemade	60	25	1	3	0.124	0.04	0	0.194	0.123

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Biscuit, from refrigerated dough, baked	61	24	0	3	0.127	0.038	0.073	0.22	0.119
Whole wheat bread	62	25	0	0	0.32	0.079	0.198	0.48	0.32
Tortilla, flour	63	25	0	1	0.227	0.099	0.032	0.469	0.229
Rye bread	64	25	0	0	0.26	0.061	0.155	0.4	0.246
Blueberry muffin, commercial	65	25	0	3	0.113	0.04	0.065	0.246	0.108
Saltine crackers	66	26	1	5	0.098	0.036	0	0.197	0.1
Corn chips	67	25	5	8	0.04	0.032	0	0.099	0.034
Pancake from mix	68	25	0	5	0.136	0.074	0.05	0.39	0.129
Egg noodles, boiled	69	25	0	1	0.218	0.082	0.052	0.373	0.232
Macaroni, boiled	70	26	0	1	0.242	0.087	0.034	0.43	0.245
Corn flakes	71	26	5	4	0.057	0.048	0	0.195	0.05
Fruit-flavored, sweetened cereal	72	25	0	5	0.075	0.026	0.031	0.14	0.079
Shredded wheat cereal	73	26	7	5	0.046	0.04	0	0.13	0.044
Raisin bran cereal	74	26	5	9	0.049	0.059	0	0.297	0.035
Crisped rice cereal	75	25	1	10	0.085	0.071	0	0.216	0.044
Granola cereal	76	26	0	2	0.144	0.053	0.066	0.244	0.14
Oat ring cereal	77	26	1	0	0.23	0.078	0	0.335	0.235
Apple, red, raw	78	26	25	1	0	0.002	0	0.011	0
Orange, raw	79	26	24	2	0.001	0.003	0	0.012	0
Banana, raw	80	26	16	8	0.009	0.014	0	0.054	0
Watermelon, raw	81	26	25	1	0	0.002	0	0.012	0
Peach, raw	83	26	25	1	0	0.002	0	0.012	0
Applesauce, bottled	84	26	26	0	0	0	0	0	0
Pear, raw	85	26	26	0	0	0	0	0	0
Strawberries, raw	86	25	23	2	0.001	0.003	0	0.012	0

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Fruit cocktail, canned in heavy syrup	87	26	26	0	0	0	0	0	0
Grapes, red/green, seedless, raw	88	26	26	0	0	0	0	0	0
Cantaloupe, raw	89	26	16	10	0.007	0.009	0	0.025	0
Plums, raw	91	26	25	1	0	0.002	0	0.012	0
Grapefruit, raw	92	26	24	2	0.001	0.003	0	0.011	0
Pineapple, canned in juice	93	26	25	1	0.001	0.003	0	0.017	0
Sweet cherries, raw	94	20	20	0	0	0	0	0	0
Raisins, dried	95	25	24	1	0.001	0.003	0	0.014	0
Prunes, dried	96	25	25	0	0	0	0	0	0
Avocado, raw	97	25	24	1	0.001	0.006	0	0.028	0
Orange juice, from frozen concentrate	98	25	24	2	0.001	0.003	0	0.015	0
Apple juice, bottled	99	25	24	1	0.002	0.008	0	0.04	0
Grapefruit juice, from frozen concentrate	100	26	26	0	0	0	0	0	0
Prune juice, bottled	103	26	25	1	0	0.002	0	0.011	0
Lemonade, from frozen concentrate	105	26	25	0	0.002	0.009	0	0.047	0
Spinach, fresh/frozen, boiled	107	25	18	7	0.003	0.005	0	0.015	0
Collards, fresh/frozen	108	25	17	7	0.005	0.009	0	0.041	0
Iceberg lettuce, raw	109	26	24	2	0.001	0.004	0	0.014	0
Cabbage, fresh, boiled	110	26	21	4	0.003	0.007	0	0.03	0
Coleslaw with dressing, homemade	111	26	17	8	0.011	0.016	0	0.047	0
Sauerkraut, canned	112	26	14	11	0.009	0.015	0	0.071	0
Broccoli, fresh/frozen, boiled	113	26	16	8	0.011	0.027	0	0.134	0

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Celery, raw	114	26	24	2	0.001	0.003	0	0.012	0
Asparagus, fresh/frozen, boiled	115	26	5	11	0.042	0.045	0	0.217	0.034
Cauliflower, fresh/frozen, boiled	116	26	17	7	0.009	0.022	0	0.103	0
Tomato, red, raw	117	25	22	3	0.002	0.005	0	0.019	0
Tomato sauce, plain, bottled	119	26	23	3	0.003	0.008	0	0.037	0
Green beans, fresh/frozen, boiled	121	26	23	3	0.001	0.004	0	0.013	0
Cucumber, raw	123	26	25	1	0	0.002	0	0.011	0
Summer squash, fresh/frozen, boiled	124	26	22	4	0.002	0.005	0	0.019	0
Green pepper, raw	125	26	26	0	0	0	0	0	0
Winter squash, fresh/frozen, baked, mashed	126	26	24	2	0.001	0.003	0	0.012	0
Onion, mature, raw	128	26	18	8	0.006	0.01	0	0.039	0
Radish, raw	132	26	25	1	0	0.002	0	0.011	0
French fries, frozen, heated	134	26	25	1	0.001	0.003	0	0.016	0
Mashed potatoes, from flakes	135	26	21	5	0.004	0.009	0	0.035	0
White potato, boiled without skin	136	26	25	1	0.001	0.005	0	0.028	0
White potato, baked with skin	137	26	20	6	0.004	0.007	0	0.02	0
Potato chips	138	26	13	8	0.026	0.046	0	0.217	0.006
Scalloped potatoes, homemade	139	26	14	10	0.012	0.015	0	0.048	0
Sweet potato, fresh, baked	140	26	22	3	0.004	0.009	0	0.033	0
Spaghetti with tomato sauce and meatballs, homemade	142	26	0	4	0.123	0.035	0.048	0.2	0.116

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Beef stew with potatoes, carrots, and onion, homemade	143	26	1	5	0.07	0.026	0	0.12	0.071
Macaroni and cheese, from box mix	146	26	0	2	0.195	0.055	0.076	0.339	0.189
Quarter-pound hamburger on bun, fast-food	147	26	0	2	0.177	0.046	0.091	0.3	0.173
Meatloaf, homemade	148	26	0	2	0.191	0.048	0.074	0.3	0.195
Spaghetti with tomato sauce, canned	149	26	0	6	0.106	0.028	0.06	0.187	0.1
Lasagna with meat, homemade	151	26	0	4	0.147	0.032	0.093	0.213	0.147
Chicken potpie, frozen, heated	152	26	3	3	0.071	0.032	0	0.127	0.076
Chicken noodle soup, canned, condensed, prepared with water	155	26	5	9	0.028	0.018	0	0.06	0.03
Tomato soup, canned, condensed, prepared with water	156	26	22	4	0.002	0.005	0	0.017	0
Vegetable beef soup, canned, condensed, prepared with water	157	26	11	15	0.01	0.01	0	0.026	0.013
White sauce, homemade	160	26	6	6	0.032	0.022	0	0.076	0.035
Dill cucumber pickles	161	26	24	2	0.001	0.004	0	0.017	0
Margarine, stick, regular (salted)	162	25	24	1	0	0.002	0	0.012	0
Butter, regular (salted)	164	26	21	5	0.003	0.007	0	0.021	0
Mayonnaise, regular, bottled	166	26	11	12	0.021	0.021	0	0.078	0.024
Half & half cream	167	26	6	18	0.019	0.013	0	0.042	0.021
Cream substitute, frozen	168	26	26	0	0	0	0	0	0

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
White sugar, granulated	169	26	26	0	0	0	0	0	0
Pancake syrup	170	26	26	0	0	0	0	0	0
Honey	172	26	26	0	0	0	0	0	0
Tomato catsup	173	26	22	4	0.002	0.005	0	0.016	0
Chocolate pudding, from instant mix	175	26	6	15	0.027	0.024	0	0.087	0.025
Vanilla flavored light ice cream	177	24	6	10	0.026	0.016	0	0.046	0.03
Chocolate cake with chocolate icing, commercial	178	26	6	11	0.035	0.022	0	0.077	0.041
Yellow cake with white icing, prepared from cake and icing mixes	179	26	7	5	0.035	0.024	0	0.075	0.042
Sweet roll/Danish, commercial	182	26	0	5	0.128	0.04	0.043	0.22	0.123
Chocolate chip cookies, commercial	183	26	6	5	0.043	0.032	0	0.123	0.045
Sandwich cookies with creme filling, commercial	184	26	5	15	0.032	0.022	0	0.081	0.029
Apple pie, fresh/frozen, commercial	185	26	17	9	0.007	0.011	0	0.033	0
Pumpkin pie, fresh/frozen, commercial	186	26	6	11	0.033	0.021	0	0.076	0.037
Milk chocolate candy bar, plain	187	26	4	4	0.046	0.025	0	0.11	0.047
Caramel candy	188	26	10	15	0.017	0.015	0	0.05	0.022
Gelatin dessert, any flavor	190	26	25	1	0.001	0.003	0	0.017	0
Cola carbonated beverage	191	26	25	1	0.001	0.003	0	0.014	0
Fruit drink, from powder	193	26	25	0	0.001	0.006	0	0.032	0

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Low-calorie cola carbonated beverage	194	26	26	0	0	0	0	0	0
Coffee, decaffeinated, from instant	196	26	24	2	0.001	0.006	0	0.032	0
Tea, from tea bag	197	26	25	1	0.001	0.006	0	0.032	0
Beer	198	26	21	5	0.002	0.005	0	0.015	0
Dry table wine	199	26	24	2	0.002	0.008	0	0.04	0
Whiskey	200	26	25	1	0	0.001	0	0.007	0
Tap water	201	26	25	1	0	0	0	0.002	0
Milk-based infant formula, high iron, ready-to-feed	202	25	6	18	0.017	0.011	0	0.03	0.021
Milk-based infant formula, low iron, ready-to- feed	203	25	6	18	0.018	0.011	0	0.037	0.021
Beef, strained/junior	205	26	6	12	0.028	0.02	0	0.075	0.026
Chicken, strained/junior, with/without broth or gravy	207	25	0	1	0.129	0.024	0.063	0.181	0.134
Chicken/turkey with vegetables, high/lean meat, strained/junior	208	2	0	2	0.064	0.004	0.061	0.066	0.064
Beef with vegetables, high/lean meat, strained/junior	209	2	2	0	0	0	0	0	0
Ham with vegetables, high/lean meat, strained/junior	210	2	0	2	0.102	0.033	0.079	0.125	0.102
Vegetables and beef, strained/junior	211	25	14	11	0.007	0.009	0	0.033	0
Vegetables and chicken, strained/junior	212	26	7	19	0.015	0.015	0	0.073	0.012
Vegetables and ham, strained/junior	213	26	7	18	0.016	0.012	0	0.041	0.018

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Chicken noodle dinner, strained/junior	214	26	6	10	0.029	0.018	0	0.064	0.032
Macaroni, tomatoes, and beef, strained/junior	215	26	5	10	0.028	0.017	0	0.06	0.032
Turkey and rice, strained/junior	216	26	7	14	0.025	0.022	0	0.095	0.025
Carrots, strained/junior	218	26	25	1	0.001	0.005	0	0.026	0
Green beans, strained/junior	219	26	24	2	0.001	0.005	0	0.019	0
Mixed vegetables, strained/junior	220	26	17	7	0.009	0.018	0	0.081	0
Sweet potatoes, strained/junior	221	26	26	0	0	0	0	0	0
Creamed corn, strained/junior	222	26	11	9	0.017	0.021	0	0.074	0.012
Peas, strained/junior	223	26	23	3	0.001	0.004	0	0.016	0
Creamed spinach, strained/junior	224	25	6	13	0.022	0.017	0	0.068	0.026
Applesauce, strained/junior	225	26	24	2	0.001	0.003	0	0.012	0
Peaches, strained/junior	226	26	26	0	0	0	0	0	0
Pears, strained/junior	227	25	24	1	0	0.002	0	0.012	0
Apple juice, strained	230	25	25	0	0	0	0	0	0
Orange juice, strained	231	26	26	0	0	0	0	0	0
Custard pudding, strained/junior	232	26	5	10	0.032	0.019	0	0.071	0.035
Fruit dessert/pudding, strained/junior	233	26	25	1	0	0.002	0	0.011	0
Fruit-flavored yogurt, lowfat (fruit mixed in)	235	26	7	14	0.022	0.015	0	0.047	0.024
Swiss cheese	236	26	0	5	0.18	0.054	0.109	0.368	0.174

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Cream cheese	237	26	4	8	0.053	0.031	0	0.099	0.054
Veal cutlet, pan-cooked	238	26	0	1	0.165	0.045	0.098	0.354	0.162
Ham luncheon meat, sliced	239	26	0	0	0.237	0.078	0.096	0.374	0.22
Chicken breast, roasted	240	25	0	1	0.27	0.12	0.09	0.623	0.228
Chicken nuggets, fast-food	241	25	0	1	0.2	0.1	0.052	0.595	0.177
Chicken, fried (breast, leg, and thigh), fast-food	242	25	0	1	0.218	0.065	0.131	0.353	0.21
Haddock, pan-cooked	243	19	0	0	0.397	0.076	0.256	0.503	0.4
Shrimp, boiled	244	25	0	0	0.38	0.1	0.2	0.574	0.369
Kidney beans, dry, boiled	245	26	8	12	0.02	0.017	0	0.051	0.019
Peas, mature, dry, boiled	246	26	10	4	0.05	0.09	0	0.457	0.034
Mixed nuts, no peanuts, dry roasted	247	25	1	0	0.53	0.39	0	1.8	0.44
Cracked wheat bread	248	26	0	0	0.285	0.065	0.209	0.448	0.269
Bagel, plain	249	26	0	0	0.311	0.085	0.165	0.518	0.299
English muffin, plain, toasted	250	26	0	0	0.263	0.068	0.144	0.402	0.25
Graham crackers	251	26	4	3	0.055	0.03	0	0.1	0.057
Butter-type crackers	252	26	4	2	0.061	0.031	0	0.102	0.069
Apricot, raw	253	21	19	2	0.001	0.004	0	0.015	0
Peach, canned in light/medium syrup	254	26	26	0	0	0	0	0	0
pear, canned in light syrup	255	26	26	0	0	0	0	0	0
Pineapple juice, from frozen concentrate	256	26	26	0	0	0	0	0	0
Grape juice, from frozen concentrate	257	26	26	0	0	0	0	0	0
French fries, fast-food	258	26	22	4	0.003	0.007	0	0.023	0
Carrot, fresh, boiled	259	26	22	4	0.002	0.006	0	0.027	0

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Tomato, stewed, canned	260	26	24	2	0.001	0.003	0	0.014	0
Tomato juice, bottled	261	26	20	6	0.004	0.007	0	0.023	0
Beets, fresh/frozen, boiled	262	25	22	3	0.002	0.006	0	0.023	0
Brussels sprouts, fresh/frozen, boiled	263	26	16	8	0.009	0.013	0	0.044	0
Mushrooms, raw	264	26	2	3	0.108	0.054	0	0.227	0.095
Eggplant, fresh, boiled	265	26	26	0	0	0	0	0	0
Turnip, fresh/frozen, boiled	266	26	23	3	0.002	0.006	0	0.025	0
Okra, fresh/frozen, boiled	267	26	22	3	0.003	0.008	0	0.03	0
Mixed vegetables, frozen, boiled	268	26	20	6	0.004	0.008	0	0.032	0
Beef stroganoff, homemade	269	26	0	0	0.191	0.043	0.121	0.311	0.183
Green peppers stuffed with beef and rice, homemade	270	26	3	4	0.065	0.029	0	0.113	0.066
Chili con carne with beans, homemade	271	26	3	6	0.052	0.025	0	0.09	0.057
Tuna noodle casserole, homemade	272	26	0	1	0.173	0.042	0.107	0.281	0.166
Salisbury steak with gravy, potatoes, and vegetable, frozen meal, heated	273	26	6	4	0.034	0.022	0	0.062	0.041
Turkey with gravy, dressing, potatoes, and vegetable, frozen meal, heated	274	26	0	5	0.093	0.025	0.051	0.17	0.091
Quarter-pound cheeseburger on bun, fast-food	275	26	0	0	0.18	0.041	0.108	0.331	0.18
Fish sandwich on bun, fast-food	276	26	0	0	0.184	0.04	0.109	0.281	0.189

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Frankfurter on bun, fast-food	277	26	0	3	0.199	0.048	0.096	0.315	0.197
Egg, cheese, and ham on English muffin, fast-food	278	26	0	0	0.263	0.079	0.095	0.451	0.256
Taco/tostada, from Mexican carry-out	279	26	2	3	0.103	0.039	0	0.161	0.104
Cheese pizza, regular crust, from pizza carry-out	280	26	0	0	0.239	0.053	0.138	0.332	0.235
Cheese and pepperoni pizza, regular crust, from pizza carry-out	281	26	0	0	0.229	0.067	0.068	0.381	0.225
beef chow mein, from Chinese carry-out	282	26	3	5	0.068	0.043	0	0.192	0.071
Bean with bacon/pork soup, canned, condensed, prepared with water	283	26	7	19	0.015	0.013	0	0.052	0.014
Mushroom soup, canned, condensed, prepared with whole milk	284	26	5	18	0.021	0.017	0	0.061	0.019
Clam chowder, New England, canned, condensed, prepared with whole milk	285	26	4	12	0.032	0.018	0	0.06	0.036
Vanilla ice cream	286	26	6	19	0.019	0.012	0	0.043	0.021
Fruit flavor sherbet	287	26	21	4	0.005	0.013	0	0.059	0
Popsicle, any flavor	288	26	25	1	0.001	0.006	0	0.03	0
Chocolate snack cake with chocolate icing	289	26	9	15	0.02	0.017	0	0.056	0.025
Cake doughnuts with icing, any flavor, from doughnut store	290	26	0	7	0.097	0.036	0.032	0.164	0.097

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
Brownies, commercial	291	26	4	7	0.045	0.026	0	0.096	0.049
Sugar cookies, commercial	292	26	4	13	0.039	0.025	0	0.091	0.035
Suckers, any flavor	293	26	24	1	0.003	0.014	0	0.07	0
Pretzels, hard, salted, any shape	294	26	5	7	0.04	0.025	0	0.094	0.043
Chocolate syrup dessert topping	295	26	20	5	0.006	0.013	0	0.054	0
Jelly, any flavor	296	26	25	1	0.001	0.005	0	0.025	0
Sweet cucumber pickles	297	26	22	4	0.002	0.005	0	0.017	0
Yellow mustard	298	26	0	0	0.33	0.13	0.103	0.724	0.308
Black olives	299	26	25	0	0.001	0.007	0	0.038	0
Sour cream	300	26	4	20	0.027	0.022	0	0.1	0.022
Brown gravy, homemade	301	26	7	9	0.031	0.025	0	0.094	0.032
French salad dressing, regular	302	26	14	10	0.017	0.036	0	0.184	0
Italian salad dressing, low-calorie	303	26	24	2	0.002	0.006	0	0.023	0
Olive/safflower oil	304	26	25	1	0.001	0.003	0	0.014	0
Coffee, from ground	305	26	26	0	0	0	0	0	0
Fruit-flavored carbonated beverage	306	26	25	1	0.001	0.004	0	0.022	0
Fruit drink, canned	307	26	24	2	0.001	0.005	0	0.022	0
Martini	308	26	26	0	0	0	0	0	0
Soy-based infant formula, ready-to-feed	309	26	8	17	0.013	0.009	0	0.023	0.016
Egg yolk, strained/junior	310	12	0	0	0.293	0.026	0.253	0.33	0.292
Rice infant cereal, instant, prepared with whole milk	311	26	5	3	0.051	0.03	0	0.093	0.06

Table 6-4. U.S. Food and Drug Administration—Total Diet Study (TDS)—Market Baskets 91-3 through 99-1 (*continued)*

TDS food description	TDS food number	Selenium—summary of results: Number of results	Number of not detected	Number of traces	Mean (mg/kg)	Standard deviation (mg/kg)	Minimum (mg/kg)	Maximum (mg/kg)	Median (mg/kg)
rice cereal, strained/junior									
Bananas with tapioca, strained/junior	312	20	9	11	0.011	0.011	0	0.032	0.016
Beets, strained/junior	313	26	25	1	0	0.002	0	0.01	0
Split peas with vegetables and ham/bacon, strained/junior	314	15	15	0	0	0	0	0	0
Teething biscuits	316	26	8	18	0.015	0.011	0	0.038	0.017
Salmon, steaks or filets, fresh or frozen, baked									
Rice cereal with apple, strained/junior	317	26	0	0	0.192	0.052	0.109	0.356	0.188
Squash, strained/junior	318	6	0	0	0.285	0.052	0.205	0.341	0.29

that the greatest portion of daily selenium intake occurs from the ingestion of grains and cereals (51.8%) (FDA 1982a). Meat, fish, and poultry were estimated to contribute 36.4% and dairy products to contribute 9.7% (FDA 1982a).

Various estimates of the selenium intake for Americans have ranged from 0.071 to 0.152 mg/day (FDA 1982a; Levander 1987; Pennington et al. 1989; Schrauzer and White 1978; Schubert et al. 1987; Welsh et al. 1981). Schubert et al. (1987) estimated the intake of selenium for the U.S. population to be 0.071 mg/day. They based their estimate on their review of selenium concentrations in different types of foods and the amount of each type of food eaten. The amount of each food type eaten daily was estimated from the U.S. Department of Agriculture's 1977–1978 Nationwide Food Consumption Survey (NFCS). Welsh et al. (1981) estimated the mean daily selenium intake of a group of 22 Maryland residents to be 0.081 mg/day (the median was 0.074 mg/day). In California, the mean daily selenium intake of eight individuals was estimated to be 0.127 mg/day (Schrauzer and White 1978). The FDA (1982a) estimated the average daily selenium intake of the U.S. population to be 0.1523 mg/day (152.3 μg/day) (FDA 1982a). Pennington et al. (1989) estimated the daily dietary intake of selenium by age group and by sex between 1982 and 1986, based on the FDA's Total Diet Studies for those years, to be between 0.020 mg/day (20 μg/day) for infants and 0.120 mg/day (120 μg/day) for adult males between 25 and 30 years of age. These values are sufficient to meet the RDA for selenium of 0.055 mg/day for men and women (NAS 2000).

Both inorganic selenium and selenomethionine are also found in selenium supplements. The amounts in these supplements generally range from 10 to 25 μg/tablet (Goodman et al. 1990). A guide to vitamin and minerals recommends that not more than 200 μg selenium/day should be taken in any form (Hender 1990).

The mean whole blood selenium concentration of residents from 19 U.S. cities ranged from 0.10 to 0.34 mg/L with a mean value of 0.21 mg/L (Barceloux 1999). A synopsis of selenium concentrations in human tissues has been summarized in Table 3-6.

The National Occupation Hazard Survey (NOHS), conducted by the National Institute for Occupational Safety and Health (NIOSH), estimated that 108,682 workers in 15,127 plants were potentially exposed to selenium in the workplace in 1970 (NOHS 1976). These estimates were derived from observations of the actual use of selenium (1% of total estimate), the use of trade name products known to contain selenium (4%), and the use of generic products suspected of containing the selenium compounds (95%). The

largest numbers of exposed workers were heavy equipment mechanics, painters, mechanics in service stations, and special trade contractors.

Preliminary data from a second workplace survey, the National Occupational Exposure Survey (NOES), conducted by NIOSH from 1980 to 1983, indicated that 27,461 workers, including 9,330 women, in 904 plants were potentially exposed to selenium in the workplace in 1980 (NIOSH 1989). The majority of these workers were employed in the health services (e.g., nursing), as janitors and cleaners, as machine operators, in the metals industry, or in work involving food and kindred products. These estimates were derived from observations of the actual use of selenium (87% of the total estimate) and the use of trade name products known to contain the selenium compounds (13%).

Neither the NOHS database nor the NOES database contain information on the frequency, level, or duration of the exposure of workers to any of the chemicals listed therein. They are surveys that provide estimates of workers potentially exposed to the chemicals.

The average selenium concentration in the blood of 20 workers employed in a rubber tire repair shop located in Mexico was 148 μg/L, while the average concentration in a control group of 18 healthy volunteers was 100 μg/L (Sánchez-Ocampo et al. 1996). Selenium was measured in the blood of 222 coal miners at concentrations ranging from 34.9–99.5 μg/L (Orszczyn et al. 1996). Selenium content in the blood decreased with age and unexpectedly, smokers had slightly lower blood plasma concentrations than nonsmokers. Furthermore, the most exposed miners (miners exposed to coal dust for more than 10 years) had lower selenium plasma levels than recently hired miners. Although the precise mechanism explaining the decrease in selenium concentration with dust exposure and smoking is unknown, the authors speculated that the decreased selenium levels might reflect its use by the increased demand in antioxidant protection, involving glutathione-peroxidase. Concentrations of selenium in the plasma and urine of copper refinery workers was studied (Rajotte et al. 1996). The levels of selenium in the urine and plasma of the 20 workers ranged from 34.02–189.95 μg/L and 113.93–173.57 μg/L, respectively. The selenium levels in a control group that were not occupationally exposed was 26.71–118.39 μg/L (urine) and 119.51–187.35 μg/L (plasma).

6.6 EXPOSURES OF CHILDREN

This section focuses on exposures from conception to maturity at 18 years in humans. Differences from adults in susceptibility to hazardous substances are discussed in 3.7 Children's Susceptibility.

Children are not small adults. A child's exposure may differ from an adult's exposure in many ways. Children drink more fluids, eat more food, breathe more air per kilogram of body weight, and have a larger skin surface in proportion to their body volume. A child's diet often differs from that of adults. The developing human's source of nutrition changes with age: from placental nourishment to breast milk or formula to the diet of older children who eat more of certain types of foods than adults. A child's behavior and lifestyle also influence exposure. Children crawl on the floor, put things in their mouths, sometimes eat inappropriate things (such as dirt or paint chips), and spend more time outdoors. Children also are closer to the ground, and they do not use the judgment of adults to avoid hazards (NRC 1993).

Children are exposed to selenium by the same pathways that affect adults. The primary route of exposure for children is through the ingestion of food sources. Selenium has been identified in pasteurized milk and milk-based infant formulas in the United States at mean concentrations in the range of 0.011–0.070 mg/kg (Table 6-4). Children may also be exposed to selenium by breast feeding mothers. Selenium was identified in the postpartum breast milk of women at different lactation stages at concentrations of 6.1–53.4 µg/L (Li et al. 1999). Using these concentrations, the daily intake of selenium for fully breast fed infants was estimated to range from 5.2 to 17.9 µg/day. Others have reported the estimated daily dietary intake of selenium for infants as 20 µg/day, while the daily intake for adult males was estimated as 120 µg/day (Pennington 1989). Selenium was detected in the umbilical blood of 350 subjects in the Czech Republic at concentrations of 4.0–82.6 µg/L („erná et al. 1997). The concentration of selenium in the blood of 388 children (196 males, 192 females) ranged from 5.0 to 98.2 µg/L („erná et al. 1995). Selenium was detected in fetal tissues at a mean concentration of 2.8 µg/g (Robkin et al. 1973). The concentration of selenium in various tissues of infants has been reported by Dickson and Tomlinson (1967) and is summarized in Table 3-6. In areas containing low (0.42 mg/kg), medium (3.09 mg/kg), and high (9.54 mg/kg) seleniferous soils, the mean whole blood selenium levels of school children (7–14 years of age) were 0.13, 0.37, and 1.57 mg/L, respectively (Yang et al. 1989b). Selenium was detected in postmortem liver, lung, and spleen samples of infants in Glasgow, Scotland at mean concentrations of 2.24, 0.76, and 0.099 ppm, respectively (Raie 1996).

The tendency of young children to ingest soil, either intentionally through pica or unintentionally through hand-to-mouth activity, is well documented. This potential route of exposure is most likely in areas that naturally have a high selenium content in the soil. Since children often play in fields and soils, dermal exposure is possible and the inhalation of dust particles from soil surfaces is also possible. The soluble forms of selenium such as the inorganic alkali selenites and selenates are more likely to be bioavailable in soils than the relatively insoluble selenides. Children are not likely to be exposed to selenium from their parents' work clothes, skin, hair, tools or other objects removed from the workplace. Selenium is contained in some household products such as shampoos and preparations to treat dandruff and eczema (IARC 1975a). It is also contained in some dietary supplements (Goodman et al. 1990). Since it is unlikely that children would use these products without adult supervision, the potential for overexposure to selenium from these products is low, except for the possibility of accidental poisoning.

6.7 POPULATIONS WITH POTENTIALLY HIGH EXPOSURES

Because selenium is a naturally occurring element found in rocks, soils, plants, and animals, the general population is commonly exposed to selenium through diet and drinking water. As a result of the uneven distribution of selenium in the earth's crust, populations living in certain areas of the United States are exposed to greater than average levels of selenium. Areas of the United States with highly seleniferous soils and plants include South Dakota, Wyoming, Montana, North Dakota, Nebraska, Kansas, Colorado, Utah, Arizona, and New Mexico (Valentine et al. 1978). Hawaii also has high levels of selenium in the soil, but not in plants (Smith et al. 1936; Valentine et al. 1978). Human exposure to selenium occurs through the ingestion of food (including meat, milk, eggs, and vegetables) and drinking water from these areas (Smith et al. 1936). Selenium was found at elevated levels in fish from rivers, creeks, and lakes in California, North Carolina, Texas, and Utah (RTI 1993). Farmers and fishermen living in these regions may be at higher risk of selenium exposure than people living in urban areas because farmers tend to consume a larger proportion of locally grown foods, and fishermen tend to consume seafood, whereas people in urban areas tend to consume foods grown over a wider geographic area. In addition, people who irrigate their home gardens with groundwater containing high selenium levels may grow and consume plants that contain high levels of selenium because this element accumulates in some plants. Fishermen and hunters of waterfowl who regularly consume fish and game from waterways with elevated selenium levels may increase their selenium body burden, but no reports of selenosis attributable to this practice have appeared in the literature.

People living in the vicinity of hazardous waste sites or coal burning plants may also be exposed to high levels of selenium. Selenosis has been reported in residents of the Wudang Mountains, China where food was grown in highly seleniferous soil (Yang et al. 1989a, 1989b). Selenium blood levels of five patients with long persisting, distinct clinical signs of selenosis ranged from 1.054 to 1.854 mg/L (Yang et al. 1989b). To attain selenium blood levels of this magnitude, it was estimated that the daily intake must be at least 910 µg/day. The mean selenium concentration in hair samples obtained from residents of a highly seleniferous region of Glasgow, Scotland was 18.92 ppm (Raie 1996). By comparison, the mean levels for adults from Iran and Iceland were only 5.72 and 1.81 ppm, respectively.

6.8 ADEQUACY OF THE DATABASE

Section 104(i)(5) of CERCLA, as amended, directs the Administrator of ATSDR (in consultation with the Administrator of EPA and agencies and programs of the Public Health Service) to assess whether adequate information on the health effects of selenium is available. Where adequate information is not available, ATSDR, in conjunction with the National Toxicology Program (NTP), is required to assure the initiation of a program of research designed to determine the health effects (and techniques for developing methods to determine such health effects) of selenium.

The following categories of possible data needs have been identified by a joint team of scientists from ATSDR, NTP, and EPA. They are defined as substance-specific informational needs that if met would reduce the uncertainties of human health assessment. This definition should not be interpreted to mean that all data needs discussed in this section must be filled. In the future, the identified data needs will be evaluated and prioritized, and a substance-specific research agenda will be proposed.

6.8.1 Identification of Data Needs

Physical and Chemical Properties. The physical and chemical properties of elemental selenium and most of the common environmental forms of selenium have been characterized (Budavari et al. 1996; Lide 2000) and no further data are needed (see Chapter 4).

Production, Import/Export, Use, Release, and Disposal. Knowledge of a chemical's production volume is important because it often correlates with possible environmental contamination and human exposure. Current data regarding the import (USGS 2001), export (USGS 2001), and use (Hoffmann and King 1997) of selenium are available. No statistics regarding the U.S. production of

selenium have been reported since 1996 (USGS 2001). Current information on the U.S. production of selenium would assist in identifying potential exposures, particularly in regions of the country where environmental exposure to selenium through food and drinking water is already relatively high.

According to the Emergency Planning and Community Right-to-Know Act of 1986, 42 U.S.C. Section 11023, industries are required to submit substance release and off-site transfer information to the EPA. The Toxics Release Inventory (TRI), which contains this information for 1999, became available in May of 2001. This database will be updated yearly and should provide a list of industrial production facilities and emissions.

Environmental Fate. Information is available to permit assessment of the environmental fate and transport of selenium in air (NAS 1976a), water (Chau and Riley 1965; NAS 1980b; Ohlendorf et al. 1986a; Rudd and Turner 1983a; Saiki and Lowe 1987), and soil (Kabatas-Pendias and Pendias 1984, NAS 1976b). Selenium released to the air will be removed by wet and dry deposition. The forms of selenium expected to be found in surface water and the water contained in soils are the salts of selenic and selenious acids. Selenic acid (H_2SeO_4) is a strong acid. The soluble selenate salts of this acid are expected to occur in alkaline waters. Sodium selenate is one of the most mobile selenium compounds in the environment because of its high solubility and inability to adsorb onto soil particles (NAS 1976a). Selenious acid (H_2SeO_3) is a weak acid, and the diselenite ion predominates in waters between pH 3.5 and 9. Most selenites are less soluble in water than the corresponding selenates (NAS 1980b).

It has been suggested that a biological cycle exists for selenium (Shrift 1964), but certain components of the cycle remain uncharacterized. The biological transformation of selenide to elemental selenium has not been well described in the literature (see Maier et al. 1988). Further research on the biological selenium cycle might help to identify "hot spots" of selenium in the environment. For example, further investigation of parameters that influence the tendency of selenium to move from one medium to another (e.g., from soil to water) would improve fate and transport modeling efforts.

Bioavailability from Environmental Media. The available monitoring data indicate that selenium is present in samples of air (Dams et al. 1970; Harrison et al. 1971; John et al. 1973; Peirson et al. 1973; Pillay et al. 1971), water (Besser et al. 1996; CRWQCB 1988; Cutter 1989; Glover et al. 1979; Lakin and Davidson 1967; Lewis 1988; Maier et al. 1988; NCDNR 1986; Oster et al. 1988a; Schutz and Turekiam 1965; Smith and Westfall 1937), soil/sediment (Glover et al. 1979; Lemly 1997; Sindeeva 1964), human tissues (Li et al. 1999; Orszczyn et al. 1996; Yang et al. 1989a, 1989b), fish (Besser et al. 1996; Lowe et

al. 1985; May and McKinney 1981; Ohlendorf et al. 1986b), and food (Beale et al. 1990; FDA 2000; Schubert et al.1987). Thus, it can be concluded that selenium is bioavailable from the environmental media.

Food Chain Bioaccumulation. Selenium in food contributed to the highest proportion of the daily selenium intake for human populations in the United States. Fruits, vegetables, milk, meat, and grains contain very low levels of selenium. However, selenium is bioaccumulated by aquatic organisms (Chau and Riley 1965; Ohlendorf et al. 1986a; Rudd and Turner 1983a). Based on reported BCFs and BAFs (Lemly 1982, 1985), selenium is expected to bioaccumulate in fish. Some evidence indicates that under natural conditions selenium might also biomagnify in aquatic organisms (Lemly 1985; Maier et al. 1988; NCDNR 1986; Sandholm et al. 1973).

Exposure Levels in Environmental Media. Selenium has been detected in air (Dams et al. 1970; Harrison et al. 1971; John et al. 1973; Peirson et al. 1973; Pillay et al. 1971), water (CRWQCB 1988; Cutter 1989; Glover et al. 1979; Lakin and Davidson 1967; Lewis 1988; Maier et al. 1988; NCDNR 1986; Oster et al. 1988a; Schutz and Turekiam 1965; Smith and Westfall 1937), soil and sediment (Beath et al. 1946; Coleman and Delevaux 1957; Glooschenko and Arafat 1988; Glover et al. 1979; Lemly 1997; NAS 1976a; Rosenfeld and Beath 1964; Shamberger 1981; Sindeeva 1964), coal and oil (Hashimoto et al. 1970; Pillay et al. 1969), plants (Arthur et al. 1992; Cappon 1981; Horne 1991; Rosenfeld and Beath 1964; Shane et al. 1988), and food (Beale et al. 1990; FDA 2000; Schubert et al. 1987). Continued monitoring data of selenium levels in the environment are necessary to understand current exposure levels.

Reliable monitoring data for the levels of selenium and selenium compounds in contaminated media at hazardous waste sites are needed so that the information obtained on levels of selenium and selenium compounds in the environment can be used in combination with the known body burden of selenium and selenium compounds to assess the potential risk of adverse health effects in populations living in the vicinity of hazardous waste sites.

Exposure Levels in Humans. Selenium has been detected in the blood (Barceloux 1999; Orszczyn et al. 1996), urine (Gromadzinska et al. 1996), hair (Raie 1996; Yang et al. 1989a, 1989b), and nails (Yang et al. 1989a, 1989b) of exposed individuals. Various estimates of selenium intake for the U.S. populations have been reported (FDA 1982a; Levander 1987; Pennington et al. 1989; Schrauzer and White 1978; Schubert et al. 1987; Welsh et al. 1981). The largest numbers of exposed workers were

heavy equipment mechanics, painters, mechanics, and special trade contractors (NOHS 1976). Preliminary data from another workplace study indicate that workplace exposure decreased from 1976 to 1984 (NIOSH 1989). Continued monitoring data are necessary to understand and evaluate human exposures to selenium in both occupational and nonoccupational settings.

Exposures of Children. Data are available regarding the exposure and body burdens of children to selenium. Children, like adults, are primarily exposed to selenium through diet. In areas containing low (0.42 mg/kg), medium (3.09 mg/kg), and high (9.54 mg/kg) seleniferous soils, the mean whole blood selenium levels of school children (7–14 years of age) were 0.13, 0.37, and 1.57 mg/L, respectively (Yang et al. 1989b). Selenium was detected in postmortem liver, lung, and spleen samples of infants in Glasgow, Scotland at mean concentrations of 2.24, 0.76, and 0.099 ppm, respectively (Raie 1996). Children can be exposed to selenium from breast feeding mothers. Selenium was identified in the postpartum breast milk of women at different lactation stages at concentrations of 6.1–53.4 μg/L (Li et al. 1999). Using these concentrations, the daily intake of selenium for fully breast fed infants was estimated to range from 5.2 to 17.9 μg/day. Others have reported the estimated daily dietary intake of selenium for infants as 20 μg/day, while the daily intake for adult males was estimated as 120 μg/day (Pennington 1989). Since selenium is found in soil surfaces and children ingest soil either intentionally through pica or unintentionally through hand-to-mouth activity, pica is a unique exposure pathway for children. While selenium is found in some home products like shampoos (IARC 1975a) and dietary supplements (Goodman et al. 1990), this exposure route should be low and will not disproportionally affect children. Continued monitoring data are necessary to understand potentially dangerous routes of childhood exposure.

Child health data needs relating to susceptibility are discussed in Section 3.12.2 Identification of Data Needs: Children's Susceptibility.

Exposure Registries. No exposure registries for selenium or selenium compounds were located. This substance is not currently one of the compounds for which a subregistry has been established in the National Exposure Registry. The substance will be considered in the future when chemical selection is made for subregistries to be established. The information that is amassed in the National Exposure Registry facilitates the epidemiological research needed to assess adverse health outcomes that may be related to exposure to this substance.

The development of a registry of exposures would provide a useful reference tool in assessing exposure levels and frequencies. In addition, a registry developed on the basis of exposure sources would allow an assessment of the variations in exposure levels from one source to another and of the effect of geographical, seasonal, or regulatory actions on the level of exposure from a certain source. These assessments, in turn, would provide a better understanding of the needs for research or data acquisition based on the current exposure levels.

6.8.2 Ongoing Studies

A summary of some pertinent ongoing research related to selenium is reported. Federally sponsored research that was reported in the CRIS/USDA (1999), CRISP (1999), and the Federal Research in Progress (FEDRIP 2001) databases are shown in Table 6-5.

As part of the Third National Health and Nutrition Evaluation Survey (NHANES III), the Environmental Health Laboratory Sciences Division of the National Center for Environmental Health, Centers for Disease Control and Prevention, will be analyzing human blood samples for selenium and selenium compounds. These data will give an indication of the frequency of occurrence and background levels of these compounds in the general population.

Table 6-5. Ongoing Studies on Selenium

Investigator	Institute	Research Area	Reference
Odom, JW	Auburn University	Monitoring selenium concentration in Alabama soils	CRIS/USDA 2001
Suarez, DL; Amrhein C	University of California Riverside	Selenium and arsenic speciation and mobilization in irrigated soils and drainage waters	CRIS/USDA 2001
Doner, H; Amundson, R	University of California, Berkeley	Quantify selenium levels in the Tulare Drainage District, San Joaquin Valley, California.	CRIS/USDA 2001
Tanji, KK	University of California, Davis	Remove selenium from irrigation drainage waters in the San Joaquin Valley	CRIS/USDA 2001
Jury, WA	University of California, Riverside	Characterize the flow and volatilization of selenium from soils	CRIS/USDA 2001
Schulthess, CP	University of Connecticut	Study the parameters that affect the adsorption and desorption of SeO_3, SeO_4 in soils.	CRIS/USDA 2001
McDowell, LR; Ramos-Santana, R	University of Florida	Compare bioavailability and tolerance of two Se sources (organic and inorganic).	CRIS/USDA 2001
Gottschall, EB	National Jewish Medical and Research Center	Study selenium and lung cancer risks in asbestos workers	CRISP 2001
Miller, WP	University of Georgia	Study waste streams from the poultry production (manure/litter), pulp and paper (sludges and ashes), and electric utility (fly ash/bottom ash) industries in Georgia	CRIS/USDA 2001

7. ANALYTICAL METHODS

The purpose of this chapter is to describe the analytical methods that are available for detecting, measuring, and/or monitoring selenium, its metabolites, and other biomarkers of exposure and effect to selenium. The intent is not to provide an exhaustive list of analytical methods. Rather, the intention is to identify well-established methods that are used as the standard methods of analysis. Many of the analytical methods used for environmental samples are the methods approved by federal agencies and organizations such as EPA and the National Institute for Occupational Safety and Health (NIOSH). Other methods presented in this chapter are those that are approved by groups such as the Association of Official Analytical Chemists (AOAC) and the American Public Health Association (APHA). Additionally, analytical methods are included that modify previously used methods to obtain lower detection limits and/or to improve accuracy and precision.

The analytical methods used to quantify selenium in biological and environmental samples are summarized below. Table 7-1 lists the applicable analytical methods used for determining selenium and selenium compounds in biological fluids and tissues, and Table 7-2 lists the methods used for determining selenium in environmental samples.

7.1 BIOLOGICAL MATERIALS

Sampling of biological material for determination of total selenium concentrations does not usually pose a problem unless specific selenium compounds are to be identified (Bem 1981). One exception is the collection and storage of urine samples without loss of volatile selenium compounds (Bem 1981). Unless special precautions are taken, most analyses of biological materials probably underestimate the concentration of these compounds. Ideally, selenium should be measured in 24-hour urine samples that have been stored in polyethylene containers in acid medium (Sanz Alaejos and Diaz Romero 1993). Blood samples should be separated into plasma or serum and cell fractions prior to freezing if the selenium levels in these components are to be measured separately. Freezing of biological samples immediately following collection is recommended to reduce enzymatic formation of volatile selenium compounds.

A variety of analytical methods can be used to determine trace concentrations (ng/g) of selenium in biological tissues. These include fluorometry, neutron activation analysis (NAA), atomic absorption

Table 7-1. Analytical Methods for Determining Selenium in Biological Materials

Sample matrix	Preparation method	Analytical method	Sample detection limit	Percent recovery	Reference
Air (breath)	Calibrate personal sampling pump; sample at a known flow rate for a total sample size of 5–2,000 L; analyze at 190.6 nm	ICP/AES	21 ng/mL	97–105%	NIOSH 1994 (method 7300)
Blood	Mineralize using HNO_3–$HClO_4$ mixture, generate hydride, and atomize	HGAAS	$1x10^{-8}$ g/g	No data	Clinton 1977
	Digest blood sample with a nitric/perchloric acid mixture; fume mixture at 200 EC and measure 2,3-diaminonaphthalene	Fluorometric	$1.2x10^{-9}$ g/g	98%	Rongpu et al. 1986
Blood, plasma, or tissue homogenate	Digest with $Mg(NO_3)_2$ or HNO_3 at a solution temperature of 100 EC for 60–90 minutes; add HCl; and add hydroxylamine sulfate, EDTA, and urea	GC/ECD	$1x10^{-8}$ g/g	95–105%	McCarthy et al. 1981
Serum	Dilute sample with matrix modifier containing $Mg(NO_3)_2$ and $Ni(NO_3)_2$ to thermally stabilize Se; heat, dry, atomize; use Zeeman background correction	ZAAS	No data	6.2% relative standard deviation	Lewis et al. 1986b
	Dilute sample with matrix modifier containing $NiCl_2$; heat, dry, and atomize	GFAAS	No data	84–116%	Oster and Prellwitz 1982
	Nitric-perchloric acid digestion; HCl reduction; sodium borohydride reduction; measure selenium hydride	HGAAS	No data	33–73%	Oster and Prellwitz 1982
	On-line acid ashing of sample followed by hydrive generation	ICP/AES	5.5 μg/L	98–106%	Recknagel et al. 1993

Table 7-1. Analytical Methods for Determining Selenium in Biological Materials (*continued*)

Sample matrix	Preparation method	Analytical method	Sample detection limit	Percent recovery	Reference
Urine	24-hour samples analyzed to measure CT and selenium concentration in urine	Folin-Wu method for CT measurement; fluorimetric method to measure Se	No data	No data	Hojo 1981b, 1982
	Digest sample with HNO_3 and $HClO_4$	Fluorometric	No data	100±22%	Koh and Benson 1983
	Add nitric acid, platinum, and nickel	EAAS	No data	4–8% relative deviation	Saeed 1986
Human spermatozoa and protasomes	Digest with 25% tetramethylammonium hydroxide in methanol	GFAAS	$1x10^{-8}$ g/g	95.1±5.2%	Suistomaa et al. 1987
Biological samples	Decompose sample with nitric acid; use 1,2-dibromobenzene as a reagent to measure piazselenol	GC/ECD	$1x10^{-9}$ g/g	No data	Shimoishi 1977
	Spike sample with ^{82}Se; digest; acidify with HCl; react with 4-nitro-o-phenylene-diamine; measure nitro-piazselenol	IDGC/MS	$5x10^{-11}$ g/g	No data	Lewis 1988
Liver	Lyophilize sample; irradiate the sample; digest with HNO_3, $HClO_4$, and the carrier source; distill sample, and use distillate for analysis	Radiochemical NAA	$2.2x10^{-10}$ g/g	No data	Lievens et al. 1977
Protein (human liver)		INAA and gel filtration	No data	No data	Norheim and Steinnes 1975

CT = creatinine; EAAS = electrothermal atomic absorption spectroscopy; EDTA = ethylenediaminetetraacetic acid; GC/ECD = gas chromatography/electron capture detection; GFAAS = graphite furnace atomic absorption spectroscopy; HCl = hydrochloric acid; HGAAS = hydride generation atomic absorption spectroscopy; HNO_3 = nitric acid; $HClO_4$ = perchloric acid; ICP/AES = inductively coupled plasma/atomic emission spectroscopy; IDGC/MS = isotope dilution gas chromatography/mass spectrometry; INAA = instrumental neutron activation analysis; $Mg(NO_3)_2$ = magnesium nitrate; NAA = neutron activation analysis; $NiCl_2$ = nickel chloride; Se = selenium; ZAAS = graphite furnace atomic absorption spectroscopy with Zeeman background correction

Table 7-2. Analytical Methods for Determining Selenium in Environmental Samples

Sample matrix	Preparation method	Analytical method	Sample detection limit	Percent recovery	Reference
Food	Reduce selenium in sample from SeVI to SeIV; add zinc to the acidified sample; pass gaseous selenium hydride to AA	AA, gaseous hydride	$2x10^{-9}$ g/mL	100%	EPA 1979a (method 270.3)
Water	Acidify sample with HCl, degas solution with N_2 bubbling	HGGC with photo-ionization detection	$1x10^{-12}$ g/mL (0.001 ppb)	No data	Vien and Fry 1988
	Reduce selenium to SeIV with HCl and KBr; coprecipitate with lanthanum hydroxide; centrifuge	ICP/AES	0.06 µg/L	100% selenite; 88% selenate	Adkins et al. 1995
Water and waste water	Acid digestion	AA, furnace	$2x10^{-9}$ g/mL	94–112%	EPA 1979a (method 270.2)
Water and wastes	Acid digestion	AA, furnace	$5x10^{-9}$ g/mL	No data	EPA 1984b (method 200.7 CLP-M)
Solid/solid waste/sludge	Aqueous samples subject to acid digestion	AA, furnace	$2x10^{-9}$ g/mL	No data	EPA 1984c (method 7740)
	Acid digestion; measure at 196 mm	ICP and GFAAS	$7.5x10^{-8}$ g/mL	94–112%	EPA 1986c (methods 3050, 6010)
	Acid digestion with HNO_3/sulfuric acid; convert SeIV to volatile hydride	AA, gaseous hydride	$2x10^{-9}$ g/mL	100%	EPA 1997a (method 7741a)
	Acid digestion	AA, furnace	$2x10^{-9}$ g/mL	No data	EPA 1984b (method 270.2 CLP-M)

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Table 7-2. Analytical Methods for Determining Selenium in Environmental Samples (*continued*)

Sample matrix	Preparation method	Analytical method	Sample detection limit	Percent recovery	Reference
Wastes/soil/ groundwater	Nitric acid digestion or nitric/peroxide/hydrochloric acid digestion	AA, furnace	$3x10^{-9}$ g/mL	100.5 %	EPA 1997b (method 7742)
Organic waste	Oxidize organic samples, absorb combustion products in NaOH; separate on an ion exchange column	Cathodic stripping	$5x10^{-9}$ g/mL	No data	DOE 1987
	Digest aqueous sample with HNO_3 and perchloric acid				
Marine biological tissues	Decompose tissue sample with HNO_3 under pressure; add sulfuric and perchloric acids; heat at 310 EC to evaporate excess acid; add HCl	HGAAS	$2x10^{-7}$ g/g	No data	Welz and Melcher 1985
Marine samples	Digest sample with concentrated HNO_3 at room temperature; add HNO_3, perchloric, and sulfuric acids to complete digestion; evaporate extra acids; dissolve residue in HCl	HGAES-ICP	$5x10^{-9}$ g/mL	No data	DOE 1987
Avian eggs and liver	Digest sample with HNO_3; and hydrogen peroxide to increase solubility	GFAAS	$4x10^{-7}$ g/g	No data	Krynitsky 1987
Fat materials (butter)	Melt butter under an infrared lamp; digest with HNO_3, sulfuric, and perchloric acids	HGAAS	10 ppb	No data	Narasaski 1985

Table 7-2. Analytical Methods for Determining Selenium in Environmental Samples (*continued*)

Sample matrix	Preparation method	Analytical method	Sample detection limit	Percent recovery	Reference
Plants		Gravimetric method	$2x10^{-6}$ g/g	No data	AOAC 1984 (method 3.101)
	Grind air-dried or fresh samples; acid digest with HNO_3 and HCl add EDTA; neutralize with NH_4OH; add HCl; shake with decalin; centrifuge decalin layer; read decalin solution with fluorometer at 525 nm within 5 minutes	Fluorometric method	$<4x10^{-6}$ g/g	No data	AOAC 1984 (method 3.102 to 3.107)
Food	Digest sample with HNO_3, perchloric, and sulfuric acids; heat; add H_2O_2; mix with EDTA, NH_4OH, and DAN; boil; add cyclohexane and shake; read cyclohexane layer at 525 nm	Titrimetric method	No data	No data	AOAC 1984 (methods 25.154 and 25.158)
Air (particulate)	Fileter particulate matter from air; irradiate and count sample	NAA, nondestructive	$1x10^{-10}$ g/m^3	No data	Dams et al. 1984
Air	Calibrate sampling pump; sample at a known flow rate for a total sample size of 13–2,000 L; analyze at 190.6 nm	ICP/AES	21 ng/mL	97–105%	NIOSH 1994a (method 7300)

AA = atomic absorption; AES = atomic emission spectrometry; DAN = 2,3-diaminonaphthalene; EDTA = ethylenediamine tetraacetate; CFAAS = graphite furnace atomic absorption spectroscopy; HCl = hydrochloric acid; HGAAS = hydride generation atomic absorption spectroscopy; HGAES=hydride generation atomic emission spectroscopy; HGGC = hydride generation gas chromatography; HNO_3 = nitric acid; ICP = inductively coupled plasma; KBr = potassium bromide; N_2 = nitrogen; NAA = neutron activation analysis; NaOH = sodium hydroxide; NH_4OH = ammonium hydroxide

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spectroscopy (AAS), gas chromatography (GC), spectrophotometry, x-ray fluorescence analysis, and others.

NAA techniques are preferred because of the lower detection limits for selenium (between 10^{-8} and 10^{-9} g selenium per gram of sample), but there are few reactors at which NAA facilities and expertise are available (Koirtyohann and Morris 1986). The most common NAA procedure for selenium determination is to produce the long-lived ^{75}Se radionucleide (half-life of 119 days) and count the samples after a 50–100 hour irradiation period and a 2–10 week cooling period. A faster NAA technique utilizes metastable ^{77m}Se, which has a much shorter half-life (17.4 seconds), so that counting can be initiated after an irradiation and cooling period of less than 1 minute (Koirtyohann and Morris 1986). The most common standard reference sample for NAA techniques is bovine liver tissue (Bem 1981). Biological tissues that can be analyzed for selenium using the NAA technique include bone, hair, liver, kidney, lung, serum, blood, feces, urine, brain, stomach, skin, aorta, heart, testis, pituitary gland, tooth enamel, tongue, muscle, spleen, and thyroid (Yukawa et al. 1980). For many NAA techniques, destructive sample pretreatment (involving radiochemical separation) is required to avoid interference from the biological matrix (Koirtyohann and Morris 1986). The advantages of NAA are its low detection limits and multielement capability (Molokhia et al. 1979). Because facilities at which NAA can be performed are extremely limited, NAA's most useful application is as a reference method against which other less expensive and more common methods can be compared for accuracy.

Classical flame AAS techniques do not have sufficiently low detection limits for selenium to be useful for determining its presence in biological samples (Koirtyohann and Morris 1986). Hydride generation atomic absorption spectroscopy (HGAAS) has been used instead for determination of selenium in biological samples such as blood and blood constituents and meat, fruits, and vegetables (Bem 1981).

Graphite furnace atomic absorption spectroscopy (GFAAS) offers high detectability (5×10^{-11} g selenium/g sample), but interference from the matrix can cause significant difficulties (Lewis 1988). GFAAS methods rely on the fact that numerous metal compounds react with selenium compounds to form relatively refractory metal selenides (Oster and Prellwitz 1982). Nickel, molybdenum, and platinum are commonly added to the sample to thermally stabilize the selenium. Organic materials are then destroyed by high temperature in the furnace prior to atomization of the sample at very high temperatures (e.g., 2,700 EC) (Oster and Prellwitz 1982). One advantage of GFAAS techniques is that the material in the graphite sample cell can be chemically treated *in situ* to reduce chemical interference. GFAAS techniques require correction for background absorption. Correction techniques include the deuterium

continuum light source method (Hoenig and Van Hoeyweghen 1986) and the Zeeman splitting of the absorption line (Koirtyohann and Morris 1986). A Zeeman-effect system, which applies a magnetic field to the atomizer, allows the background correction to be performed at the exact analyte wavelength without the use of auxiliary light sources (Fernandez and Giddings 1982). The Zeeman-effect background correction is necessary for the determination of selenium in blood and blood products when GFAAS is used because a spectral interference from iron occurs at the selenium wavelength that cannot be corrected by a deuterium continuum source.

A modification of the GFAAS method for determining selenium levels in human urine was described by Saeed (1986). In this electrothermal atomic absorption spectrometry (EAAS) method, nitric acid, nickel, and platinum are added to the graphite cell. The addition of nickel helps to mask the spectral interference from phosphates in urine. EAAS has been used to determine selenium levels in human spermatozoa (Suistomaa et al. 1987).

HGAAS offers reduced chemical interference but requires larger sample volumes than GFAAS techniques (Koirtyohann and Morris 1986). HGAAS techniques have been used to measure selenium concentrations in food (Fiorino et al. 1976). These techniques use wet-sample digestion (e.g., nitric-perchloric acid) to destroy organic matter. Sample reduction to convert Se(VI) (+6 valence state) to Se(IV) (+4 valence state) is necessary prior to using sodium borohydride to reduce all selenium present to selenium hydride (Macpherson et al. 1988). The selenium hydride is thermally decomposed and atomized in the sample beam of the AA spectrophotometer. Nitric-perchloric acid is commonly used for the digestion step. Because perchloric acid is potentially explosive, use of phosphoric acid instead is also common. Following the International Union of Pure and Applied Chemists (IUPAC) interlaboratory trial for the determination of selenium in human body fluids, Welz and Verlinden (1986) reported that it was important to use a temperature of at least 200 EC for sample decomposition when using HGAAS. They attributed the severe imprecision and systematic errors in measuring selenium in multiple samples to improper sample decomposition. Norheim and Haugen (1986) demonstrated that a combined system of a wet digestion and an automated hydride generator could analyze approximately 80 samples per day.

Application of gas-liquid chromatography (GLC) to determine selenium in biological samples allows for the elimination of interference from the biological matrix. GLC requires prior decomposition of organic matter with nitric acid. GLC techniques are based on measurement of the amount of piazselenol formed by the reaction of selenium(IV) with appropriate reagents in acidic media (Bem 1981). For gas chromatographic determination of selenium with an electron capture detector, 1,2-diaminoarenes can be

used as reagents to produce piazselenols (McCarthy et al. 1981; Poole et al. 1977; Shimoishi 1977; Young and Christian 1973). Using 1,2-diamino-3,5-dibromobenzene as a reagent, Shimoishi (1977) obtained a detection limit of $1x10^{-9}$ g selenium per gram of sample.

Isotope dilution gas chromatography/mass spectrometry (IDGC/MS) is a highly accurate technique that is more accessible than NAA techniques. IDGC/MS has been used to determine selenium in foods, plasma and serum, red blood cells, feces, urine, and human breast milk (Lewis 1988). The minimum sample size per determination is 0.5–10 g (0.5–10 mL). In the IDGC/MS method, a stable selenium isotope is added to the sample prior to digestion. This procedure eliminates the need for quantitative sample preparation and external standardization (Lewis 1988). With the recent development of benchtop capillary GC/MS systems, the use of IDGC/MS for determining selenium in biological matrices should increase (Lewis 1988).

Spectrophotometric, fluorometric, voltammetric, and x-ray fluorescence analysis methods have also been successfully employed to determine selenium levels in blood, tissue, and human hair. Of these, fluorometric methods are most commonly used (Koh and Benson 1983). The reaction of selenium(VI) with 2,3-diaminonaphthalene (DAN) or with 3,3-diaminobenzidine (DAB) to form a fluorescent Se-DAN or Se-DAB heterocyclic compound is the basis of the fluorometric method of selenium determination (Allaway and Cary 1964; Chen et al. 1982; Lewis 1988). The piazselenol formed with DAN as the reagent has a greater fluorescence sensitivity than the piazselenol formed with DAB as the reagent and is also extractable into organic solvents from acid solution (Chen et al. 1982). Fluorometric techniques require sample digestion to destroy organic matter and sample reduction to convert the selenium to the selenium(IV) oxidation state (Macpherson et al. 1988). Loss of volatile selenium compounds is possible during sample digestion and manipulation because several steps are required. Chen et al. (1982), Hasunuma et al. (1982), and Koh and Benson (1983) developed modifications of the digestion and treatment steps for selenium determination by fluorometric methods. Their methods allow small sample sizes, can be performed in a single flask, and measure submicrogram amounts of selenium.

Some of the methods for determining selenium in biological materials have been compared within the same laboratory for accuracy and precision. Macpherson et al. (1988) compared the accuracy of three methods for the determination of selenium in biological fluid samples from biological materials with certified selenium levels. Acid decomposition fluorometry, HGAAS, and EAAS gave equally accurate results. Lewis et al. (1986) compared the graphite furnace atomic absorption spectrometry with the Zeeman-effect background correction (ZAAS) to isotope dilution mass spectrometry (IDMS) for

determination of selenium in plasma and concluded that the ZAAS method compared favorably (correlation coefficient 0.987) but was half as precise as the IDMS method. Oster and Prellwitz (1982) compared HGAAS and GFAAS for the determination of selenium in serum. They concluded that the two techniques exhibited approximately equal detection limits in their laboratory.

In three studies that compared analytical methods for the detection of selenium in biological samples, all found that fluorometry gave both accurate and reliable results (Burguera et al. 1990; Heydorn and Griepink 1990; Macpherson et al. 1988). Burguera et al. (1990) indicated the acceptance of HGAAS as yielding reliable results, whereas Heydorn and Griepink (1990) reported HGAAS had a high relative standard deviation of 11.4%.

Decomposition procedures have been improved and analytical methods modified in recent years to increase the accuracy and speed of determination of selenium concentrations in plasma, serum, and urine. Reamer and Veillon (1983) used phosphoric acid along with nitric acid and hydrogen peroxide in digestion of biological fluids instead of perchloric acid to prepare samples for fluorometry. They concluded that phosphoric acid digestion increases the safety and convenience of the determination. Krynitsky (1987) used a modified wet digestion method for determination of selenium in biological samples such as eggs and liver of avian species. This method uses hydrogen peroxide to enhance the solubility of the sample. Digestion with HNO_3 and $HClO_4$ is essential for accurate analysis of the total selenium in urine to ensure complete oxidation of the trimethyl-selenonium ion (Koh and Benson 1983).

7.2 ENVIRONMENTAL SAMPLES

Many of the basic analytical methods used for determining selenium in biological media are also used for determining selenium levels in soil, water, and air. Precautions in the collection and storage of environmental samples, however, are necessary to prevent loss of the volatile selenium compounds to the air. The destruction of organic matter before selenium measurement is also often necessary. Acidification of water samples to a pH of 1.5 is recommended to preserve selenium compounds (Muñoz Olivas et al. 1994). Nitric acid can be used, although it interferes with the hydride generation method of analysis. The best storage method for selenium compounds in water is in glass containers at 4 EC (Wiedmeyer and May 1993).

The analytic methods generally fall into two groups: (1) those that do not require the destruction of organic materials in the sample and (2) those that require the elimination of interfering matter before the

selenium content can be measured. X-ray fluorescence and some of the neutron activation analysis techniques do not require sample destruction, whereas spectrophotometry, GC, atomic absorption spectrometry, polarography, titration, spark source, MS, fluorometry, and other neutron activation analysis techniques require some degree of sample destruction. Fluorometry, atomic absorption spectrometry, and neutron activation analysis are the most frequently used methods.

AAS techniques are commonly used for determination of selenium in environmental samples. Hydride generation AAS is more sensitive than flame or graphite furnace AAS for determination of selenium in materials of variable composition. Water samples, including freshwater, river water, sea water, and surface waters, and industrial wastes, muds, sediments, and soil samples have been analyzed by AAS techniques to detect selenium at parts-per-trillion levels (Bem 1981). Selenium(VI) and selenium(IV) can be distinguished in water samples with GFAAS by selective extraction procedures. HGAAS can also be used to distinguish between selenium(VI) and selenium(IV) in environmental samples because selenium(VI) does not readily form the hydride without reduction (Koirtyohann and Morris 1986). selenium(VI) is calculated on the basis of the total selenium minus selenium(IV) (Bem 1981).

NAA has been used to determine selenium levels in environmental samples. Dams et al. (1970) reported a detection limit of $1x10^{-10}$ g/m^3 selenium using nondestructive NAA for determining selenium in air particulate matter. For determining selenium levels in soil, radiochemical variants of NAA have been commonly employed (Bem 1981). Instrumental neutron activation analysis (INAA) is frequently used to determine selenium concentrations in water and can also be used to distinguish between selenium(IV) and selenium(VI) oxidation states (Bem 1981). INAA is also used to determine selenium concentrations in air (Bem 1981).

Inductively coupled plasma (ICP) emission techniques can be used to measure selenium concentrations in air. ICP techniques offer multielement capabilities, but instrumentation is costly and background interference can be a problem (Koirtyohann and Morris 1986). The NIOSH-recommended method for determining selenium in air is inductively coupled argon plasma atomic emission spectroscopy (NIOSH 1994). An ICP method with a detection limit of 0.06 μg selenium/L has also been described for water samples (Adkins et al. 1995).

Gas liquid chromatography allows for elimination of interference from the matrix when analyzing environmental samples. When analyzing biological samples, a variety of reagents can be used to convert selenium to piazselenols for measurement with an electron capture detector. Spectrophotometric

determinations of selenium are performed using organic reagents, whereas fluorometric analysis relies on piazselenol fluorescence to measure submicrogram levels of the element.

The hydride generation GC with photoionization detection (HGGC-PD) method for selenium determination was developed by Vien and Fry (1988). The combined usage of photoionization detector and a cold trap provided at least two orders of magnitude improvement in detectability over the existing GC systems. The detection limit for the HGGC-PD method was $1x10^{-12}$ g selenium/mL (0.001 ppb) for 28 mL samples. An advantage of the HGGC-PD technique is the ability to perform simultaneous determinations of at least four different hydride forming elements (Vien and Fry 1988).

EPA's Contract Laboratory Program (EPA 1984b) requires the participating laboratories to meet the Contract Required Detection Level (CRDL) for selenium of $5x10^{-9}$ g selenium/mL (5 µg selenium/L) using proven instruments and approved analytical techniques, including ICP and AA methods.

7.3 ADEQUACY OF THE DATABASE

Section 104(i)(5) of CERCLA, as amended, directs the Administrator of ATSDR (in consultation with the Administrator of EPA and agencies and programs of the Public Health Service) to assess whether adequate information on the health effects of selenium is available. Where adequate information is not available, ATSDR, in conjunction with the National Toxicology Program (NTP), is required to assure the initiation of a program of research designed to determine the health effects (and techniques for developing methods to determine such health effects) of selenium.

The following categories of possible data needs have been identified by a joint team of scientists from ATSDR, NTP, and EPA. They are defined as substance-specific informational needs that if met would reduce the uncertainties of human health assessment. This definition should not be interpreted to mean that all data needs discussed in this section must be filled. In the future, the identified data needs will be evaluated and prioritized, and a substance-specific research agenda will be proposed.

7.3.1 Identification of Data Needs

Methods for Determining Biomarkers of Exposure and Effect.

Exposure. Methods that distinguish among the various selenium compounds are not commonly used to estimate human exposure to selenium but have been used in specialized metabolic studies. Analytical methods currently used to measure concentrations of selenium in biological fluids or human tissue samples as an indication of human exposure are described in Table 7-1. Attempts to use measures of whole blood glutathione peroxidase (GSH-Px) activity levels as indicators of human exposure to selenium have not been successful. Errors can result if the selenium-dependent GSH-Px activity is not distinguished from the nonselenium-dependent GSH-Px activity (Edwards and Blackburn 1986). In addition, whole blood selenium concentrations and GSH-Px activity appear to correlate with one another only at low blood selenium levels (<0.010 mg selenium/L) (Allaway et al. 1968; Valentine et al. 1980). GSH-Px activity levels measured in platelets have provided an indication of selenium exposure levels at low blood selenium levels (Nève et al. 1988). Whether platelet GSH-Px activity levels would provide an indication of selenium status in populations with plasma selenium levels above 0.012 mg selenium/L is not known. There is great variability in the exposure data available for humans. Therefore, until larger databases of selenium concentrations in biological materials from affected and unaffected populations are available, no recommendations for analytical methods can be made.

Effect. There are no known sensitive and specific biomarkers of effect for selenium. Therefore, no analytical methods recommendations can be made for biomarkers of effect for selenium, at the present time.

Methods for Determining Parent Compounds and Degradation Products in Environmental Media. Numerous analytical methods are available for the determination of selenium levels in environmental media (AOAC 1984; Bem 1981; Dams et al. 1970; EPA 1984b, 1986c; Koirtyohann and Morris 1986; NIOSH 1994; DOE 1987; Vien and Fry 1988). However, most of these do not distinguish among the various selenium compounds. Many of the available methods can be used to detect selenium at subnanogram levels. For the determination of selenium only, fluorometry, chromatography, or spectrometry are the preferred techniques. When conducting a multielemental analysis or when analyzing a complex matrix, more sophisticated methods are required.

It is possible to detect selenium levels as low as 1 ng per cubic meter of air using neutron activation analysis. Standardized methods for selenium determination in different environmental samples such as water, soil, sludge, and industrial waste are available in the above-mentioned literature.

There are fewer methods available for distinguishing among the inorganic forms of selenium in the environment. HGAAS, INAA, and GFAAS with selective extraction procedures can be used to distinguish between selenium(VI) and selenium(IV) in samples of soil and water. Methods for determining selenium sulfide levels in the environment are lacking, but would be useful for the identification and measurement of this potentially carcinogenic selenium compound.

Very limited information is available regarding the sensitivity, reliability, and specificity of the existing methods. Further studies to determine these factors would be useful.

7.3.2 Ongoing Studies

D.L. Rabenstein (University of California, Riverside) is investigating the biological chemistry of selenium compounds in intact erythrocytes to characterize transport across the erythrocyte membrane and their intracellular metabolism. Reactions will be characterized in erythrocytes, plasma, and aqueous solution by nuclear magnetic resonance (NMR) spectroscopy. Methods will be developed for the determination of selenols, diselenides, and selenosulfides in biological fluids and erythrocytes by high performance liquid chromatography (HPLC) and electrochemical detection. The HPLC methods will be used to determine metabolites formed at trace levels.

8. REGULATIONS AND ADVISORIES

Because of its potential to cause adverse health effects in exposed people, a number of regulations and guidelines have been established for selenium by various national and state agencies. These values are summarized in Table 8-1.

The NRC RDAs (NAS 2000) are as follows:

Men: 0.055 mg/day
Women: 0.055 mg/day
Pregnant women: 0.060 mg/day
Lactating women: 0.070 mg/day
Infants (0–6 months): 0.015 mg/day
Infants (7–12 months): 0.020 mg/day
Children (1–3 years): 0.020 mg/day
Children (4–8 years): 0.030 mg/day
Children (9–18 years): 0.040 mg/day

A chronic oral MRL of 0.005 mg/kg/day has been derived for selenium based on a NOAEL of 0.015 mg/kg/day for dermal effects identified by Yang and Zhou (1994). The NOAEL was divided by an uncertainty factor of three to account for sensitive individuals. The EPA (EPA 1993) has derived a chronic oral reference dose (RfD) of 0.005 mg/kg/day for selenium based on studies by Yang et al. (1989a, 1989b). This value is equal to the average daily intake of selenium in the diet (0.853 mg/day) considered adequate and safe. The RfD was derived assuming an average body weight of 55 kg. It was based on a NOAEL of 0.015 mg selenium/kg/day. An uncertainty factor of 10 was not applied since similar NOAELs were identified in two human populations exposed to selenium levels in excess of the RDA throughout a lifetime without apparent clinical signs of selenosis. An uncertainty factor of 3 was applied to the NOAEL to account for sensitive individuals.

Table 8-1. Regulations and Guidelines Applicable to Selenium

Agency	Description	Information	References
INTERNATIONAL			
Guidelines:			
IARC	Carcinogenicity classification	Group 3[a]	IARC 2001
WHO	Guideline for drinking water	0.01 mg/L	WHO 2001
	Recommended daily intake—adults	0.9 μg/kg body weight	
NATIONAL			
Regulations and Guidelines:			
a. Air			
ACGIH	TLV (8-hour TWA) Selenium and compounds Selenium hexafluoride	 0.2 mg/m^3 0.16 mg/m^3	ACGIH 2000
EPA	Hazard rank under Section 112(g) of the Clean Air Act Amendments	42 out of 1–100, with 100 being the most toxic	EPA 2001a
	Reference air concentration	3.0 μg/m^3	EPA 2001b 40CFR 266, Appendix IV
NIOSH	REL (TWA) Selenium and compounds, except selenium hexafluoride	 0.2 mg/m^3	NIOSH 2001
	IDLH Selenium and compounds	 1.0 mg/m^3	
OSHA	General industry PEL (TWA) Selenium and compounds Selenium hexafluoride Hydrogen selenide	 0.2 mg/m^3 0.4 mg/m^3 0.2 mg/m^3	OSHA 2001 29CFR1910.1000, Table Z
	Construction industry PEL (TWA) Selenium and compounds Selenium hexafluoride	 0.2 mg/m^3 0.16 mg/m^3	OSHA 2001
b. Water			
EPA	MCLG	0.05 mg/L	EPA 2001c 40CFR141.51
	MCL	0.05 mg/L	EPA 2001d 40CFR141.62

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
NATIONAL *(cont.)*			
EPA	DWEL	0.2 mg/L	EPA 2000
	Health advisory—lifetime	0.05 mg/L	
	Groundwater monitoring (PQL)	750 µg/L	EPA 2001e 40CFR264, Appendix IX
	Groundwater monitoring—concentration limits	0.01 mg/L	EPA 2001f 40CFR264.94
	Water quality standards Freshwater Maximum concentration Continuous concentration Saltwater Maximum concentration Continuous concentration	 20 µg/L 5.0 µg/L 290 µg/L 71 µg/L	EPA 2001g 40CFR131.36
c. Food			
FDA	Approved use of selenium as a food additive in animal feeds—added to feed for chickens, swine, turkeys, sheep, cattle, and ducks	#0.3 ppm	FDA 2001a 21CFR573.920
	Bottled water—allowable level	0.05 mg/L	FDA 2001b 21CFR165.110
	RDA (mg/day) Men Women Pregnant women Lactating women Infants (0–6 months) Infants (7–12 months) Children (1–3 years) Children (4–8 years) Children (9–18 years)	 0.055 0.055 0.060 0.070 0.015 0.020 0.020 0.030 0.040	NAS 2000
d. Other			
EPA	Carcinogenicity classification Selenium and compounds Selenium sulfide	 Group D[b] Group B2[c]	IRIS 2001
	Designation of hazardous substances Selenium oxide Sodium selenite		EPA 2001h 40CFR116.4

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
NATIONAL *(cont.)*			
EPA	Determination of reportable quantities		EPA 2001i 40CFR117.3
	Selenium oxide	10 pounds	
	Sodium selenite	100 pounds	
	Extremely hazardous substance		EPA 2001j 40CFR355, Appendix B
	Reportable quantity		
	Hydrogen selenide	10 pounds	
	Selenious acid	10 pounds	
	Selenium oxychloride	500 pounds	
	Threshold planning quantity		
	Hydrogen selenide	10 pounds	
	Selenious acid	1,000/10,000 pounds	
	Selenium oxychloride	500 pounds	
	Identification and listing of hazardous waste		EPA 2001k 40CFR261, Appendix VIII
	Selenium		
	Selenium and compounds		
	Selenium dioxide		
	Selenium sulfide		
	Selenium tetrakis, (dimethyl-dithiocarbamate)		
	Selenious acid		
	Selenourea		
	Thallium selenite		
	Protection standards at inactive uranium processing sites—listed constituents		EPA 2001l 40CFR192, Appendix I
	Selenium and compounds		
	Selenium dioxide		
	Selenium sulfide		
	Recommended daily allowances		EPA 2001m
	Selenium and compounds		
	Men	0.7×10^{-1} mg/kg/day	
	Women	0.55×10^{-1} mg/kg/day	
	Infants	8.7×10^{-4} mg/kg/day	
	Reportable quantity		EPA 2001n 40CFR302.4, Appendix A
	Selenium and compounds	1 pound	
	Selenium dioxide	1,000 pounds	
	Selenium sulfide	1 pound	
	Selenious acid	1 pound	
	Selenourea	1 pound	
	Sodium selenite	1,000 pounds	
	Thallium selenite	1 pound	

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
NATIONAL *(cont.)*			
EPA	Reportable quantity Selenium oxide	 10 pounds	EPA 2001o 40CFR117.3
	Sewer sludge—disposal or use standards Ceiling concentration Cumulative pollutant loading rate Pollutant concentration[d] Annual pollutant loading rate	 100 mg/kg 100 kg/hectare 100 mg/kg 5.0 kg/hectare per 365-day period	EPA 2001p 40CFR503.13
	Toxic chemical release reporting; Community Right-to-Know—effective date	01/01/87	EPA 2001q 40CFR372.65
STATE			
Regulations and Guidelines:			
a. Air			
Hawaii	HAP		BNA 2001
Illinois	Toxic air contaminant		BNA 2001
Kansas	HAP		BNA 2001
Kentucky	HAP		BNA 2001
Maryland	Toxic air pollutant Selenium sulfide		BNA 2001
Minnesota	HAP threshold—de minimis level Selenium compounds Selenium sulfide	 0.1 ton/year 0.1 ton/year	BNA 2001
Nebraska	HAP—effective date	12/15/98	BNA 2001
New Hampshire	Regulated toxic air pollutant		BNA 2001
New Mexico	Toxic air pollutant OEL Emissions	 0.2 mg/m^3 0.0133 mg/m^3	BNA 2001
New York	HAP—selenium compounds		BNA 2001
Rhode Island	HAP		BNA 2001
South Carolina	Toxic air emissions—maximum allowable concentration Selenium compounds	 1 μg/m^3	BNA 2001

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
STATE *(cont.)*			
Vermont	Hazardous ambient air standards		BNA 2001
	Annual average	4.80 μg/m^3	
	Action level	0.40 pounds/8-hours	
Washington	HAP—threshold levels		BNA 2001
	Selenium and compounds	0.5 tons/year	
	Selenium hexafluoride	0.5 tons/year	
	Selenium sulfides	0.5 tons/year	
b. Water			
Alabama	Aquatic life criteria		BNA 2001
	Freshwater		
	Acute	20 μg/L	
	Chronic	5.0 μg/L	
	Marine		
	Acute	300 μg/L	
	Chronic	71 μg/L	
	MCL	0.05 mg/L	BNA 2001
	Primary drinking water standard	0.01 mg/L	BNA 2001
Alaska	Groundwater cleanup level	0.05 mg/L	BNA 2001
	MCL	0.05 mg/L	BNA 2001
Arizona	Aquifer water quality standards	0.05 mg/L	BNA 2001
	Drinking water guideline	45 μg/L	HSDB 2001
	MCL	0.05 mg/L	BNA 2001
	Water quality standards		EPA 2001r
	Conversion factor[e] for saltwater—acute criteria	0.998	40CFR131.38
	Conversion factor[e] for saltwater—chronic criteria	0.998	
Colorado	Groundwater protection—MCL	0.01 mg/L	BNA 2001
	MCL	0.05 mg/L	BNA 2001
	Primary drinking water standard	0.01 mg/L	BNA 2001
Connecticut	MCL	0.05 mg/L	BNA 2001
Delaware	Groundwater protection—MCL	0.01 mg/L	BNA 2001
	Primary drinking water standard	0.01 mg/L	BNA 2001
Florida	MCL	0.05 mg/L	BNA 2001
Georgia	MCL	0.05 mg/L	BNA 2001
Hawaii	MCL	0.05 mg/L	BNA 2001

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
STATE *(cont.)*			
Hawaii	Water quality criteria applicable to all waters		BNA 2001
	Freshwater		
	Acute	20 µg/L	
	Chronic	5.0 µg/L	
	Saltwater		
	Acute	300 µg/L	
	Chronic	71 µg/L	
Illinois	Concentration shall not be exceeded in water	1.0 mg/L	BNA 2001
	Groundwater quality standard	0.01 mg/L	BNA 2001
	MCL	0.05 mg/L	BNA 2001
Indiana	MCLG	0.05 mg/L	BNA 2001
	MCL	0.05 mg/L	
Iowa	MCL	0.05 mg/L	BNA 2001
Kansas	Surface water quality standard		BNA 2001
	Aquatic life		
	Acute	20 µg/L	
	Chronic	5.0 µg/L	
	Agriculture		
	Livestock	50 µg/L	
	Irrigation	20 µg/L	
	Public health food		
	Procurement	6,800 µg/L	
	Domestic water supply	50 µg/L	
Kentucky	Domestic water supply use—maximum allowable instream concentration	0.05 mg/L	BNA 2001
	MCL	0.05 mg/L	BNA 2001
	Maximum groundwater contaminant level	0.01 mg/L	BNA 2001
	Primary drinking water standard	0.01 mg/L	BNA 2001
	Warm water aquatic habitat criteria		BNA 2001
	Acute	20 µg/L	
	Chronic	5.0 µg/L	
Louisiana	Groundwater protection—MCL	0.01 mg/L	BNA 2001
Maine	Drinking water guideline	10 µg/L	HSDB 2001

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
STATE *(cont.)*			
Maryland	Criteria for toxic substances in surface waters		BNA 2001
	Freshwater		
	Acute	20 µg/L	
	Chronic	5.0 µg/L	
	Saltwater		
	Acute	300 µg/L	
	Chronic	71 µg/L	
	Drinking water	50 µg/L	
	MCL	0.05 mg/L	BNA 2001
	Primary drinking water standard	0.01 mg/L	BNA 2001
Massachusetts	Environmental toxicity values		BNA 2001
	Freshwater	20 µg/L	
	Acute	5.0 µg/L	
	Chronic		
	Marine		
	Acute	300 µg/L	
	Chronic	71 µg/L	
	Groundwater protection—MCL	0.01 mg/L	BNA 2001
	MCL	0.05 mg/L	BNA 2001
Michigan	MCL	0.05 mg/L	BNA 2001
	Effective date	07/30/92	
Minnesota	Drinking water guideline	30 µg/L	HSDB 2001
Mississippi	Groundwater standard	50 ppb	BNA 2001
	Water quality criteria—concentration shall not exceed	0.01 mg/L	BNA 2001
Montana	MCL	0.05 mg/L	BNA 2001
North Carolina	Fresh surface water quality standard for Class C waters	5.0 ug/L	BNA 2001
	Groundwater quality standard	0.05 mg/L	BNA 2001
Nebraska	Aquatic life		BNA 2001
	Acute	20 µg/L	
	Chronic	5.0 µg/L	
	Water supply	0.05 mg/L	BNA 2001
	MCL	0.05 mg/L	BNA 2001
New Hampshire	Groundwater quality standard	0.05 mg/L	

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
STATE *(cont.)*			
New Hampshire	MCLG MCL	0.05 mg/L 0.05 mg/L	BNA 2001
	Water quality criteria Protection of aquatic life Fresh acute chronic Marine acute chronic Protection of human health Water and fish ingestion	 5.0 µg/L 290µg/L 71 µg/L 170 µg/L 11,000 µg/L	BNA 2001
New Mexico	MCL	0.05 mg/L	BNA 2001
Nevada	Domestic water supply Dissolved selenium	 0.05 mg/L	BNA 2001
New York	MCL	0.05 mg/L	BNA 2001
North Dakota	MCL	0.05 mg/L	BNA 2001
Ohio	Groundwater concentration limit	0.01 mg/L	BNA 2001
Oklahoma	Public and private water supplies	0.01 mg/L	BNA 2001
	Fish and wildlife propagation Acute Chronic	 20 µg/L 5.0 µg/L	BNA 2001
Rhode Island	Groundwater quality standard Preventive action limit	0.05 mg/L 0.025 mg/L	BNA 2001
South Carolina	MCL	0.05 mgL	BNA 2001
South Dakota	Groundwater maximum allowable concentration	0.05 mg/L	BNA 2001
	Aquatic life value Acute Chronic	 20 µg/L 5.0 µg/L	BNA 2001
Tennessee	Groundwater criteria concentration	0.05 mg/L	BNA 2001
	MCL	0.05 mg/L	BNA 2001
Texas	MCL	0.05 mg/L	BNA 2001
Utah	MCL	0.05 mg/L	BNA 2001
	Water quality Domestic Agriculture	 0.01 mg/L 0.05 mg/L	BNA 2001

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
STATE *(cont.)*			
Vermont	Groundwater quality standards		BNA 2001
	Enforcement standard	50 µg/L	
	Preventive action level	25 µg/L	
	MCLG	0.05 mg/L	BNA 2001
	MCL	0.05 mg/L	
	Water quality criteria for protection of aquatic organisms		BNA 2001
	Maximum allowable concentration		
	Acute	20 µg/L	
	Average allowable concentration		
	Chronic	5.0 µg/L	
Virginia	Groundwater protection levels		BNA 2001
	Protection level	10 µg/L	
	Monitoring level	5.0 µg/L	
	MCL	0.01 mg/L	BNA 2001
	Surface water criteria		BNA 2001
	Freshwater		
	Acute	20 µg/L	
	Chronic	5.0 µg/L	
	Saltwater		
	Acute	300 µg/L	
	Chronic	71 µg/L	
	Human health		
	Public water supplies	170 µg/L	
	All other surface waters	11,000 µg/L	
Washington	MCL	0.05 mg/L	BNA 2001
Wisconsin	Groundwater quality standards		BNA 2001
	Enforcement standard	50 µg/L	
	Preventive action limit	10 µg/L	
	MCL	0.05 mg/L	BNA 2001
Wyoming	Water quality		BNA 2001
	Aquatic life		
	Acute	20 µg/L	
	Chronic	5.0 µg/L	
	Human health	10 µg/L	
c. Food			
New York	Bottled water sampling requirements—MCL	0.01 mg/L	BNA 2001

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
STATE *(cont.)*			
d. Other			
Alabama	Identification and listing of hazardous waste		BNA 2001
Arizona	Soil remediation levels		BNA 2001
	Residential	380 mg/kg	
	Non-residential	8,500 mg/kg	
California	Hazardous waste injection restrictions—waste specific prohibitions		EPA 2001s 40CFR148.12(b)(2)
	Selenium and/or compounds	100 mg/L	
	Known to cause cancer or reproductive toxicity—initial appearance of chemical on list		BNA 2001
	Selenium sulfide	10/01/89	
	Total threshold limit concentration	10,000 mg/kg	BNA 2001
Delaware	Regulated toxic substance—sufficient quantity		BNA 2001
	Selenium hexafluoride	900 pounds/hour	
Florida	Toxic substance in the workplace		BNA 2001
	Hydrogen selenium		
	Selenium		
	Selenium hexafluoride		
	Selenium oxychloride		
	Selenium sulfide		
Hawaii	Restricted use pesticides		BNA 2001
	Selenium compounds	All concentrations	
Kentucky	Threshold planning quantity	10 pounds	BNA 2001
	Hydrogen selenide	1,000/10,000 pounds	
	Selenious acid	500 pounds	
	Selenium oxychloride		
Massachusetts	Oil and hazardous material		BNA 2001
	Selenious acid		
	Selenium and compounds		
	Selenium dioxide		
	Selenium disulfide		
	Selenium oxide		
	Selenium oxychloride		
	Selenium sulfide		
	Selenourea		
Minnesota	RfD	0.005 mg/kg/day	BNA 2001
	Health risk limit	30 μg/L	
New Hampshire	Restricted use pesticide	All concentrations	BNA 2001

Table 8-1. Regulations and Guidelines Applicable to Selenium *(continued)*

Agency	Description	Information	References
STATE *(cont.)*			
New Jersey	Extraordinary hazardous substance—threshold quantity Selenium hexafluoride	 700 pounds	BNA 2001
Oregon	Toxic substance—de minimis concentration	1.0 percent	BNA 2001
Vermont	Restricted use pesticide Selenium and compounds	All concentrations	BNA 2001

[a]Group 3: not classifiable as to its carcinogenicity to humans
[b]Group D: not classifiable as to its carcinogenicity to humans
[c]Group B2: probable human carcinogen
[d]Monthly average concentrations
[e]Conversion factors are based on a hardness of 100 mg/L as calcium carbonate

ACGIH = American Conference of Governmental Industrial Hygienists; BNA = Bureau of National Affairs; CFR = Code of Federal Regulations; DWEL = drinking water equivalent level; EPA = Environmental Protection Agency; FDA = Food and Drug Administration; HAP = hazardous air pollutant; IARC = International Agency for Research on Cancer; IDLH = immediately dangerous to life and health; IRIS = Integrated Risk Information System; MCL = maximum contaminant level; MCLG = maximum contaminant level goal; NIOSH = National Institute for Occupational Safety and Health; OEL = occupational exposure limit; OSHA = Occupational Safety and Health Administration; PEL = permissible exposure limit; PQL = practical quantitation limit; RDA = recommended daily allowance; REL = recommended exposure limit; RfD = reference dose; TLV = threshold limit value; TWA = time-weighted average; WHO = World Health Organization

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10. GLOSSARY

Absorption—The taking up of liquids by solids, or of gases by solids or liquids.

Acute Exposure—Exposure to a chemical for a duration of 14 days or less, as specified in the Toxicological Profiles.

Adsorption—The adhesion in an extremely thin layer of molecules (as of gases, solutes, or liquids) to the surfaces of solid bodies or liquids with which they are in contact.

Adsorption Coefficient (K_{oc})—The ratio of the amount of a chemical adsorbed per unit weight of organic carbon in the soil or sediment to the concentration of the chemical in solution at equilibrium.

Adsorption Ratio (Kd)—The amount of a chemical adsorbed by a sediment or soil (i.e., the solid phase) divided by the amount of chemical in the solution phase, which is in equilibrium with the solid phase, at a fixed solid/solution ratio. It is generally expressed in micrograms of chemical sorbed per gram of soil or sediment.

Benchmark Dose (BMD)—Usually defined as the lower confidence limit on the dose that produces a specified magnitude of changes in a specified adverse response. For example, a BMD_{10} would be the dose at the 95% lower confidence limit on a 10% response, and the benchmark response (BMR) would be 10%. The BMD is determined by modeling the dose response curve in the region of the dose response relationship where biologically observable data are feasible.

Benchmark Dose Model—A statistical dose-response model applied to either experimental toxicological or epidemiological data to calculate a BMD.

Bioconcentration Factor (BCF)—The quotient of the concentration of a chemical in aquatic organisms at a specific time or during a discrete time period of exposure divided by the concentration in the surrounding water at the same time or during the same period.

Biomarkers—Broadly defined as indicators signaling events in biologic systems or samples. They have been classified as markers of exposure, markers of effect, and markers of susceptibility.

Cancer Effect Level (CEL)—The lowest dose of chemical in a study, or group of studies, that produces significant increases in the incidence of cancer (or tumors) between the exposed population and its appropriate control.

Carcinogen—A chemical capable of inducing cancer.

Case-Control Study—A type of epidemiological study which examines the relationship between a particular outcome (disease or condition) and a variety of potential causative agents (such as toxic chemicals). In a case-controlled study, a group of people with a specified and well-defined outcome is identified and compared to a similar group of people without outcome.

Case Report—Describes a single individual with a particular disease or exposure. These may suggest some potential topics for scientific research but are not actual research studies.

Case Series—Describes the experience of a small number of individuals with the same disease or exposure. These may suggest potential topics for scientific research but are not actual research studies.

Ceiling Value—A concentration of a substance that should not be exceeded, even instantaneously.

Chronic Exposure—Exposure to a chemical for 365 days or more, as specified in the Toxicological Profiles.

Cohort Study—A type of epidemiological study of a specific group or groups of people who have had a common insult (e.g., exposure to an agent suspected of causing disease or a common disease) and are followed forward from exposure to outcome. At least one exposed group is compared to one unexposed group.

Cross-sectional Study—A type of epidemiological study of a group or groups which examines the relationship between exposure and outcome to a chemical or to chemicals at one point in time.

Data Needs—Substance-specific informational needs that if met would reduce the uncertainties of human health assessment.

Developmental Toxicity—The occurrence of adverse effects on the developing organism that may result from exposure to a chemical prior to conception (either parent), during prenatal development, or postnatally to the time of sexual maturation. Adverse developmental effects may be detected at any point in the life span of the organism.

Dose-Response Relationship—The quantitative relationship between the amount of exposure to a toxicant and the incidence of the adverse effects.

Embryotoxicity and Fetotoxicity—Any toxic effect on the conceptus as a result of prenatal exposure to a chemical; the distinguishing feature between the two terms is the stage of development during which the insult occurs. The terms, as used here, include malformations and variations, altered growth, and *in utero* death.

Environmental Protection Agency (EPA) Health Advisory—An estimate of acceptable drinking water levels for a chemical substance based on health effects information. A health advisory is not a legally enforceable federal standard, but serves as technical guidance to assist federal, state, and local officials.

Epidemiology—Refers to the investigation of factors that determine the frequency and distribution of disease or other health-related conditions within a defined human population during a specified period.

Genotoxicity—A specific adverse effect on the genome of living cells that, upon the duplication of affected cells, can be expressed as a mutagenic, clastogenic or carcinogenic event because of specific alteration of the molecular structure of the genome.

Half-life—A measure of rate for the time required to eliminate one half of a quantity of a chemical from the body or environmental media.

Immediately Dangerous to Life or Health (IDLH)—The maximum environmental concentration of a contaminant from which one could escape within 30 minutes without any escape-impairing symptoms or irreversible health effects.

Incidence—The ratio of individuals in a population who develop a specified condition to the total number of individuals in that population who could have developed that condition in a specified time period.

Intermediate Exposure—Exposure to a chemical for a duration of 15–364 days, as specified in the Toxicological Profiles.

Immunologic Toxicity—The occurrence of adverse effects on the immune system that may result from exposure to environmental agents such as chemicals.

Immunological Effects—Functional changes in the immune response.

In Vitro—Isolated from the living organism and artificially maintained, as in a test tube.

In Vivo—Occurring within the living organism.

Lethal Concentration$_{(LO)}$ (LC_{LO})—The lowest concentration of a chemical in air which has been reported to have caused death in humans or animals.

Lethal Concentration$_{(50)}$ (LC_{50})—A calculated concentration of a chemical in air to which exposure for a specific length of time is expected to cause death in 50% of a defined experimental animal population.

Lethal Dose$_{(LO)}$ (LD_{LO})—The lowest dose of a chemical introduced by a route other than inhalation that has been reported to have caused death in humans or animals.

Lethal Dose$_{(50)}$ (LD_{50})—The dose of a chemical which has been calculated to cause death in 50% of a defined experimental animal population.

Lethal Time$_{(50)}$ (LT_{50})—A calculated period of time within which a specific concentration of a chemical is expected to cause death in 50% of a defined experimental animal population.

Lowest-Observed-Adverse-Effect Level (LOAEL)—The lowest exposure level of chemical in a study, or group of studies, that produces statistically or biologically significant increases in frequency or severity of adverse effects between the exposed population and its appropriate control.

Lymphoreticular Effects—Represent morphological effects involving lymphatic tissues such as the lymph nodes, spleen, and thymus.

Malformations—Permanent structural changes that may adversely affect survival, development, or function.

Minimal Risk Level (MRL)—An estimate of daily human exposure to a hazardous substance that is likely to be without an appreciable risk of adverse noncancer health effects over a specified route and duration of exposure.

Modifying Factor (MF)—A value (greater than zero) that is applied to the derivation of a minimal risk level (MRL) to reflect additional concerns about the database that are not covered by the uncertainty factors. The default value for a MF is 1.

Morbidity—State of being diseased; morbidity rate is the incidence or prevalence of disease in a specific population.

Mortality—Death; mortality rate is a measure of the number of deaths in a population during a specified interval of time.

Mutagen—A substance that causes mutations. A mutation is a change in the DNA sequence of a cell's DNA. Mutations can lead to birth defects, miscarriages, or cancer.

Necropsy—The gross examination of the organs and tissues of a dead body to determine the cause of death or pathological conditions.

Neurotoxicity—The occurrence of adverse effects on the nervous system following exposure to a chemical.

No-Observed-Adverse-Effect Level (NOAEL)—The dose of a chemical at which there were no statistically or biologically significant increases in frequency or severity of adverse effects seen between the exposed population and its appropriate control. Effects may be produced at this dose, but they are not considered to be adverse.

Octanol-Water Partition Coefficient (K_{ow})—The equilibrium ratio of the concentrations of a chemical in *n*-octanol and water, in dilute solution.

Odds Ratio (OR)—A means of measuring the association between an exposure (such as toxic substances and a disease or condition) which represents the best estimate of relative risk (risk as a ratio of the incidence among subjects exposed to a particular risk factor divided by the incidence among subjects who were not exposed to the risk factor). An odds ratio of greater than 1 is considered to indicate greater risk of disease in the exposed group compared to the unexposed.

Organophosphate or Organophosphorus Compound—A phosphorus containing organic compound and especially a pesticide that acts by inhibiting cholinesterase.

Permissible Exposure Limit (PEL)—An Occupational Safety and Health Administration (OSHA) allowable exposure level in workplace air averaged over an 8-hour shift of a 40-hour workweek.

Pesticide—General classification of chemicals specifically developed and produced for use in the control of agricultural and public health pests.

Pharmacokinetics—The science of quantitatively predicting the fate (disposition) of an exogenous substance in an organism. Utilizing computational techniques, it provides the means of studying the absorption, distribution, metabolism and excretion of chemicals by the body.

Pharmacokinetic Model—A set of equations that can be used to describe the time course of a parent chemical or metabolite in an animal system. There are two types of pharmacokinetic models: data-based and physiologically-based. A data-based model divides the animal system into a series of compartments which, in general, do not represent real, identifiable anatomic regions of the body whereby the physiologically-based model compartments represent real anatomic regions of the body.

Physiologically Based Pharmacodynamic (PBPD) Model—A type of physiologically-based dose-response model which quantitatively describes the relationship between target tissue dose and toxic end points. These models advance the importance of physiologically based models in that they clearly describe the biological effect (response) produced by the system following exposure to an exogenous substance.

Physiologically Based Pharmacokinetic (PBPK) Model—Comprised of a series of compartments representing organs or tissue groups with realistic weights and blood flows. These models require a variety of physiological information: tissue volumes, blood flow rates to tissues, cardiac output, alveolar ventilation rates and, possibly membrane permeabilities. The models also utilize biochemical information such as air/blood partition coefficients, and metabolic parameters. PBPK models are also called biologically based tissue dosimetry models.

Prevalence—The number of cases of a disease or condition in a population at one point in time.

Prospective Study—A type of cohort study in which the pertinent observations are made on events occurring after the start of the study. A group is followed over time.

q_1*—The upper-bound estimate of the low-dose slope of the dose-response curve as determined by the multistage procedure. The q_1* can be used to calculate an estimate of carcinogenic potency, the incremental excess cancer risk per unit of exposure (usually μg/L for water, mg/kg/day for food, and μg/m^3 for air).

Recommended Exposure Limit (REL)—A National Institute for Occupational Safety and Health (NIOSH) time-weighted average (TWA) concentrations for up to a 10-hour workday during a 40-hour workweek.

Reference Concentration (RfC)—An estimate (with uncertainty spanning perhaps an order of magnitude) of a continuous inhalation exposure to the human population (including sensitive subgroups) that is likely to be without an appreciable risk of deleterious noncancer health effects during a lifetime. The inhalation reference concentration is for continuous inhalation exposures and is appropriately expressed in units of mg/m^3 or ppm.

Reference Dose (RfD)—An estimate (with uncertainty spanning perhaps an order of magnitude) of the daily exposure of the human population to a potential hazard that is likely to be without risk of deleterious effects during a lifetime. The RfD is operationally derived from the no-observed-adverse-effect level (NOAEL-from animal and human studies) by a consistent application of uncertainty factors that reflect various types of data used to estimate RfDs and an additional modifying factor, which is based on a professional judgment of the entire database on the chemical. The RfDs are not applicable to nonthreshold effects such as cancer.

Reportable Quantity (RQ)—The quantity of a hazardous substance that is considered reportable under the Comprehensive Environmental Response, Compensation, and Liability Act (CERCLA). Reportable quantities are (1) 1 pound or greater or (2) for selected substances, an amount established by regulation either under CERCLA or under Section 311 of the Clean Water Act. Quantities are measured over a 24-hour period.

Reproductive Toxicity—The occurrence of adverse effects on the reproductive system that may result from exposure to a chemical. The toxicity may be directed to the reproductive organs and/or the related endocrine system. The manifestation of such toxicity may be noted as alterations in sexual behavior, fertility, pregnancy outcomes, or modifications in other functions that are dependent on the integrity of this system.

Retrospective Study—A type of cohort study based on a group of persons known to have been exposed at some time in the past. Data are collected from routinely recorded events, up to the time the study is undertaken. Retrospective studies are limited to causal factors that can be ascertained from existing records and/or examining survivors of the cohort.

Risk—The possibility or chance that some adverse effect will result from a given exposure to a chemical.

Risk Factor—An aspect of personal behavior or lifestyle, an environmental exposure, or an inborn or inherited characteristic, that is associated with an increased occurrence of disease or other health-related event or condition.

Risk Ratio—The ratio of the risk among persons with specific risk factors compared to the risk among persons without risk factors. A risk ratio greater than 1 indicates greater risk of disease in the exposed group compared to the unexposed.

Short-Term Exposure Limit (STEL)—The American Conference of Governmental Industrial Hygienists (ACGIH) maximum concentration to which workers can be exposed for up to 15 min continually. No more than four excursions are allowed per day, and there must be at least 60 min between exposure periods. The daily Threshold Limit Value - Time Weighted Average (TLV-TWA) may not be exceeded.

Target Organ Toxicity—This term covers a broad range of adverse effects on target organs or physiological systems (e.g., renal, cardiovascular) extending from those arising through a single limited exposure to those assumed over a lifetime of exposure to a chemical.

Teratogen—A chemical that causes structural defects that affect the development of an organism.

Threshold Limit Value (TLV)—An American Conference of Governmental Industrial Hygienists (ACGIH) concentration of a substance to which most workers can be exposed without adverse effect. The TLV may be expressed as a Time Weighted Average (TWA), as a Short-Term Exposure Limit (STEL), or as a ceiling limit (CL).

Time-Weighted Average (TWA)—An allowable exposure concentration averaged over a normal 8-hour workday or 40-hour workweek.

Toxic Dose$_{(50)}$ (TD$_{50}$)—A calculated dose of a chemical, introduced by a route other than inhalation, which is expected to cause a specific toxic effect in 50% of a defined experimental animal population.

Toxicokinetic—The study of the absorption, distribution and elimination of toxic compounds in the living organism.

Uncertainty Factor (UF)—A factor used in operationally deriving the Minimal Risk Level (MRL) or Reference Dose (RfD) or Reference Concentration (RfC) from experimental data. UFs are intended to account for (1) the variation in sensitivity among the members of the human population, (2) the uncertainty in extrapolating animal data to the case of human, (3) the uncertainty in extrapolating from data obtained in a study that is of less than lifetime exposure, and (4) the uncertainty in using lowest-observed-adverse-effect level (LOAEL) data rather than no-observed-adverse-effect level (NOAEL) data. A default for each individual UF is 10; if complete certainty in data exists, a value of one can be used; however a reduced UF of three may be used on a case-by-case basis, three being the approximate logarithmic average of 10 and 1.

Xenobiotic—Any chemical that is foreign to the biological system.

APPENDIX A
ATSDR MINIMAL RISK LEVEL AND WORKSHEETS

The Comprehensive Environmental Response, Compensation, and Liability Act (CERCLA) [42 U.S.C. 9601 et seq.], as amended by the Superfund Amendments and Reauthorization Act (SARA) [Pub. L. 99–499], requires that the Agency for Toxic Substances and Disease Registry (ATSDR) develop jointly with the U.S. Environmental Protection Agency (EPA), in order of priority, a list of hazardous substances most commonly found at facilities on the CERCLA National Priorities List (NPL); prepare toxicological profiles for each substance included on the priority list of hazardous substances; and assure the initiation of a research program to fill identified data needs associated with the substances.

The toxicological profiles include an examination, summary, and interpretation of available toxicological information and epidemiologic evaluations of a hazardous substance. During the development of toxicological profiles, Minimal Risk Levels (MRLs) are derived when reliable and sufficient data exist to identify the target organ(s) of effect or the most sensitive health effect(s) for a specific duration for a given route of exposure. An MRL is an estimate of the daily human exposure to a hazardous substance that is likely to be without appreciable risk of adverse noncancer health effects over a specified duration of exposure. MRLs are based on noncancer health effects only and are not based on a consideration of cancer effects. These substance-specific estimates, which are intended to serve as screening levels, are used by ATSDR health assessors to identify contaminants and potential health effects that may be of concern at hazardous waste sites. It is important to note that MRLs are not intended to define clean-up or action levels.

MRLs are derived for hazardous substances using the no-observed-adverse-effect level/uncertainty factor approach. They are below levels that might cause adverse health effects in the people most sensitive to such chemical-induced effects. MRLs are derived for acute (1–14 days), intermediate (15–364 days), and chronic (365 days and longer) durations and for the oral and inhalation routes of exposure. Currently, MRLs for the dermal route of exposure are not derived because ATSDR has not yet identified a method suitable for this route of exposure. MRLs are generally based on the most sensitive chemical-induced end point considered to be of relevance to humans. Serious health effects (such as irreparable damage to the liver or kidneys, or birth defects) are not used as a basis for establishing MRLs. Exposure to a level above the MRL does not mean that adverse health effects will occur.

MRLs are intended only to serve as a screening tool to help public health professionals decide where to look more closely. They may also be viewed as a mechanism to identify those hazardous waste sites that are not expected to cause adverse health effects. Most MRLs contain a degree of uncertainty because of the lack of precise toxicological information on the people who might be most sensitive (e.g., infants, elderly, nutritionally or immunologically compromised) to the effects of hazardous substances. ATSDR uses a conservative (i.e., protective) approach to address this uncertainty consistent with the public health principle of prevention. Although human data are preferred, MRLs often must be based on animal studies because relevant human studies are lacking. In the absence of evidence to the contrary, ATSDR assumes that humans are more sensitive to the effects of hazardous substance than animals and that certain persons may be particularly sensitive. Thus, the resulting MRL may be as much as a hundredfold below levels that have been shown to be nontoxic in laboratory animals.

Proposed MRLs undergo a rigorous review process: Health Effects/MRL Workgroup reviews within the Division of Toxicology, expert panel peer reviews, and agencywide MRL Workgroup reviews, with participation from other federal agencies and comments from the public. They are subject to change as new information becomes available concomitant with updating the toxicological profiles. Thus, MRLs in the most recent toxicological profiles supersede previously published levels. For additional information regarding MRLs, please contact the Division of Toxicology, Agency for Toxic Substances and Disease Registry, 1600 Clifton Road, Mailstop E-29, Atlanta, Georgia 30333.

APPENDIX A

MINIMAL RISK LEVEL (MRL) WORKSHEET

Chemical Name: Selenium
CAS Number: 7782-49-2 (elemental)
Date: March 30, 2001
Profile Status: Final Post Public Comments
Route: [ ] Inhalation [X] Oral
Duration: [ ] Acute [ ] Intermediate [X] Chronic
Graph Key: 98
Species: Human

Minimal Risk Level: 0.005 [X] mg/kg/day [ ] ppm

Reference: Yang G, Zhou R. 1994. Further observations on the human maximum safe dietary selenium intake in a seleniferous area of China. J Trace Elem Electrolytes Health Dis 8:159-165.

Experimental design: This study was an examination of a group of five individuals recovering from selenosis who had been drawn from a larger population studied by the same authors (Yang et al. 1989a, 1989b). Yang et al. (1989a, 1989b) examined a population in an area of China where selenosis occurred. Data were collected on selenium levels in the diet, blood, nails, hair, urine, and milk of residents, and the incidence of clinical symptoms of selenosis (morphological changes in finger nails) was compared with dietary intake of selenium and selenium levels in blood. They found that selenium levels in blood corresponded to the dietary intake of selenium, and that symptoms of selenosis were found at or above a selenium intake level of 910 µg/day (0.016 mg/kg/day) (Yang et al 1989a). In 1992, Yang and Zhou (1994) reexamined five individuals from the high selenium site who had been suffering from symptoms of selenosis (loss of fingernails and hair), but were recovering (nails were regrowing). Since their earlier report, the living conditions of the population had improved; they had been cautioned against consuming high selenium foods and parts of their locally produced corn had been replaced with rice or cereals. Yang and Zhou (1994) found that the mean concentration of selenium in the blood of these selenosis patients had fallen from 1,346 µg/L (measured in 1986) to 968 µg/L (measured in 1992). Using a regression equation derived from the data in their earlier report (Yang et al. 1989b) and average body weights of 55 kg, Yang and Zhou (1994) calculated that the mean dietary intake of selenium associated with selenosis in these individuals was 1,270 µg/day (LOAEL of 0.023 mg/kg/day), while a mean intake of 819 µg selenium/day (NOAEL of 0.015 mg/kg/day) was associated with recovery.

Effects noted in study and corresponding doses: A NOAEL of 0.015 mg/kg/day for nail disease based on recovery from symptoms of selenosis, and a LOAEL of 0.023 mg/kg/day based on nail damage were calculated from selenium concentrations in blood using average body weights of 55 kg and the regression equation: $Y_{\text{blood-Se (mg/L)}} = 8230 \,.\, 10^{-4}\, X_{\text{se-intake (µg)}} + 0.176$ derived in Yang et al. (1989b).

Dose and end point used for MRL derivation: 0.015 mg/kg/day; nail disease (selenosis)

[X] NOAEL [ ] LOAEL

Uncertainty Factors used in MRL derivation:

[ ] 10 for use of a LOAEL
[ ] 10 for extrapolation from animals to humans
[X] 3 for human variability

A factor of 3 was considered appropriate because the individuals in this report were sensitive individuals drawn from the larger population in the Yang et al. (1989a, 1989b) studies and because of the supporting studies described below.

Was a conversion used from ppm in food or water to a mg/body weight dose? No.
If so, explain:

If an inhalation study in animals, list the conversion factors used in determining human equivalent dose:
NA

Other additional studies or pertinent information which lend support to this MRL:

Yang et al. (1989a, 1989b) examined a population of 349 individuals in an area of China where selenosis occurred. They collected data on selenium levels in the diet, blood, nails, hair, urine, and milk of residents at three sites with low, medium, and high selenium, and compared the incidence of clinical symptoms of selenosis (morphological changes in finger nails) with dietary intake of selenium and selenium levels in blood. They found that selenium levels in blood corresponded to the dietary intake of selenium, and that symptoms of selenosis were found at or above a selenium intake level of 910 µg/day (0.016 mg/kg/day) (Yang et al 1989a). The population included adult men and women, teenagers, children, and infants. High selenium levels were found in individuals of all ages, but symptoms of selenosis were generally confined to adults (97% of cases) and were never observed in children younger than 12 years of age (Yang et al. 1989b). The manifestation of symptoms of selenosis was not solely dependent on selenium intake, but was subject to individual variability, as individuals who exhibited selenosis did not necessarily have the highest blood selenium levels.

Longnecker et al. (1991) examined two groups of adults (142 individuals) in areas of Wyoming and South Dakota with elevated selenium intake. The average daily intake of selenium in this population was 239 µg/day (0.003 mg/kg/day) and some individuals consumed as much as 724 µg/day (0.01 mg/kg/day). The highest blood concentration of selenium noted in this population was 0.67 mg/kg, a concentration lower than the 1.05 mg/L concentration associated with effects in China. No symptoms of selenosis or any other significant health effects associated with selenium exposure were reported for individuals in this study. This study suggests that the estimates of dietary intake of selenium produced by the regression equation in Yang et al. (1989b) may be conservative. Longnecker et al. (1991) reported doses of 68–724 µg/day associated with blood concentrations of 0.18–0.67 mg/kg. If the doses from the Longnecker et al. (1991) study are placed in the regression equation from Yang et al. (1989b), blood concentrations of 0.14 and 0.88 mg/L are calculated. If it is assumed that a liter of blood weighs approximately 1 kg, then this regression equation over predicts blood levels of selenium at the higher doses in the population from North Dakota. This provides support for additional exposure (e.g., inhalation exposure) in the Chinese population that was not accounted for in the regression equation.

Selenium is a component of all three members of the deiodinase enzyme family, the enzymes responsible for deiodination of the thyroid hormones (St. Germain and Galton 1997). Two human studies were located that describe significant decreases in triiodothyronine levels in response to elevated selenium; however, the hormone levels observed in these studies were within the normal human range and the biological significance of the effect is not clear. Brätter and Negretti De Brätter (1996) examined a Venezuelan population with high selenium intake. Serum, erythrocyte, toenail, and breast milk selenium concentrations were determined for 65 women living in three seleniferous regions of Venezuela. Selenium dietary intakes were determined from the selenium concentration of breast milk by regression (Bratter et al. 1991), and free thyroxine (T4), free triiodothyronine (T3), and human thyroid stimulating hormone (TSH) levels were measured. Selenium intake ranged from 170 to 980 µg/day. Free T4, free T3, and TSH levels were found to be within normal range, but a significant inverse correlation was found

between free T3 and selenium levels (Spearman R test). The influence of selenium on the free T3 concentration became significant at a dietary selenium intake in the range of 350–450 µg/day. No symptoms of selenosis were found in the women included in this study. In a second human study, increased selenium intake produced a significant decrease (11%) in T3 levels in five men involved in a dietary study in a metabolic unit for 120 days (Hawkes and Keim 1995). The men consumed 356 µg selenium/day for 14 weeks and their T3 levels were significantly reduced compared with baseline measurements taken while they consumed 80 µg selenium/day (Hawkes and Keim 1995).

Chemical Manager: John Risher, Ph.D.

APPENDIX B

USER'S GUIDE

Chapter 1

Public Health Statement

This chapter of the profile is a health effects summary written in non-technical language. Its intended audience is the general public especially people living in the vicinity of a hazardous waste site or chemical release. If the Public Health Statement were removed from the rest of the document, it would still communicate to the lay public essential information about the chemical.

The major headings in the Public Health Statement are useful to find specific topics of concern. The topics are written in a question and answer format. The answer to each question includes a sentence that will direct the reader to chapters in the profile that will provide more information on the given topic.

Chapter 2

Relevance to Public Health

This chapter provides a health effects summary based on evaluations of existing toxicologic, epidemiologic, and toxicokinetic information. This summary is designed to present interpretive, weight-of-evidence discussions for human health end points by addressing the following questions.

1. What effects are known to occur in humans?

2. What effects observed in animals are likely to be of concern to humans?

3. What exposure conditions are likely to be of concern to humans, especially around hazardous waste sites?

The chapter covers end points in the same order they appear within the Discussion of Health Effects by Route of Exposure section, by route (inhalation, oral, dermal) and within route by effect. Human data are presented first, then animal data. Both are organized by duration (acute, intermediate, chronic). *In vitro* data and data from parenteral routes (intramuscular, intravenous, subcutaneous, etc.) are also considered in this chapter. If data are located in the scientific literature, a table of genotoxicity information is included.

The carcinogenic potential of the profiled substance is qualitatively evaluated, when appropriate, using existing toxicokinetic, genotoxic, and carcinogenic data. ATSDR does not currently assess cancer potency or perform cancer risk assessments. Minimal risk levels (MRLs) for noncancer end points (if derived) and the end points from which they were derived are indicated and discussed.

Limitations to existing scientific literature that prevent a satisfactory evaluation of the relevance to public health are identified in the Chapter 3 Data Needs section.

Interpretation of Minimal Risk Levels

Where sufficient toxicologic information is available, we have derived minimal risk levels (MRLs) for inhalation and oral routes of entry at each duration of exposure (acute, intermediate, and chronic). These MRLs are not meant to support regulatory action; but to acquaint health professionals with exposure levels at which adverse health effects are not expected to occur in humans. They should help physicians and public health officials determine the safety of a community living near a chemical emission, given the concentration of a contaminant in air or the estimated daily dose in water. MRLs are based largely on toxicological studies in animals and on reports of human occupational exposure.

MRL users should be familiar with the toxicologic information on which the number is based. Chapter 2, "Relevance to Public Health," contains basic information known about the substance. Other sections such as Chapter 3 Section 3.9, "Interactions with Other Substances," and Section 3.10, "Populations that are Unusually Susceptible" provide important supplemental information.

MRL users should also understand the MRL derivation methodology. MRLs are derived using a modified version of the risk assessment methodology the Environmental Protection Agency (EPA) provides (Barnes and Dourson 1988) to determine reference doses for lifetime exposure (RfDs).

To derive an MRL, ATSDR generally selects the most sensitive end point which, in its best judgement, represents the most sensitive human health effect for a given exposure route and duration. ATSDR cannot make this judgement or derive an MRL unless information (quantitative or qualitative) is available for all potential systemic, neurological, and developmental effects. If this information and reliable quantitative data on the chosen end point are available, ATSDR derives an MRL using the most sensitive species (when information from multiple species is available) with the highest NOAEL that does not exceed any adverse effect levels. When a NOAEL is not available, a lowest-observed-adverse-effect level (LOAEL) can be used to derive an MRL, and an uncertainty factor (UF) of 10 must be employed. Additional uncertainty factors of 10 must be used both for human variability to protect sensitive subpopulations (people who are most susceptible to the health effects caused by the substance) and for interspecies variability (extrapolation from animals to humans). In deriving an MRL, these individual uncertainty factors are multiplied together. The product is then divided into the inhalation concentration or oral dosage selected from the study. Uncertainty factors used in developing a substance-specific MRL are provided in the footnotes of the LSE Tables.

Chapter 3

Health Effects

Tables and Figures for Levels of Significant Exposure (LSE)

Tables (3-1, 3-2, and 3-3) and figures (3-1 and 3-2) are used to summarize health effects and illustrate graphically levels of exposure associated with those effects. These levels cover health effects observed at increasing dose concentrations and durations, differences in response by species, minimal risk levels (MRLs) to humans for noncancer end points, and EPA's estimated range associated with an upper- bound individual lifetime cancer risk of 1 in 10,000 to 1 in 10,000,000. Use the LSE tables and figures for a quick review of the health effects and to locate data for a specific exposure scenario. The LSE tables and figures should always be used in conjunction with the text. All entries in these tables and figures represent studies that provide reliable, quantitative estimates of No-Observed-Adverse-Effect Levels (NOAELs), Lowest-Observed-Adverse-Effect Levels (LOAELs), or Cancer Effect Levels (CELs).

The legends presented below demonstrate the application of these tables and figures. Representative examples of LSE Table 3-1 and Figure 3-1 are shown. The numbers in the left column of the legends correspond to the numbers in the example table and figure.

LEGEND

See LSE Table 3-1

(1) Route of Exposure One of the first considerations when reviewing the toxicity of a substance using these tables and figures should be the relevant and appropriate route of exposure. When sufficient data exists, three LSE tables and two LSE figures are presented in the document. The three LSE tables present data on the three principal routes of exposure, i.e., inhalation, oral, and dermal (LSE Table 3-1, 3-2, and 3-3, respectively). LSE figures are limited to the inhalation (LSE Figure 3-1) and oral (LSE Figure 3-2) routes. Not all substances will have data on each route of exposure and will not therefore have all five of the tables and figures.

(2) Exposure Period Three exposure periods - acute (less than 15 days), intermediate (15–364 days), and chronic (365 days or more) are presented within each relevant route of exposure. In this example, an inhalation study of intermediate exposure duration is reported. For quick reference to health effects occurring from a known length of exposure, locate the applicable exposure period within the LSE table and figure.

(3) Health Effect The major categories of health effects included in LSE tables and figures are death, systemic, immunological, neurological, developmental, reproductive, and cancer. NOAELs and LOAELs can be reported in the tables and figures for all effects but cancer. Systemic effects are further defined in the "System" column of the LSE table (see key number 18).

(4) Key to Figure Each key number in the LSE table links study information to one or more data points using the same key number in the corresponding LSE figure. In this example, the study represented by key number 18 has been used to derive a NOAEL and a Less Serious LOAEL (also see the 2 "18r" data points in Figure 3-1).

(5) Species The test species, whether animal or human, are identified in this column. Chapter 2, "Relevance to Public Health," covers the relevance of animal data to human toxicity and Section 3.4, "Toxicokinetics," contains any available information on comparative toxicokinetics. Although NOAELs and LOAELs are species specific, the levels are extrapolated to equivalent human doses to derive an MRL.

(6) Exposure Frequency/Duration The duration of the study and the weekly and daily exposure regimen are provided in this column. This permits comparison of NOAELs and LOAELs from different studies. In this case (key number 18), rats were exposed to 1,1,2,2-tetrachloroethane via inhalation for 6 hours per day, 5 days per week, for 3 weeks. For a more complete review of the dosing regimen refer to the appropriate sections of the text or the original reference paper, i.e., Nitschke et al. 1981.

(7) System This column further defines the systemic effects. These systems include: respiratory, cardiovascular, gastrointestinal, hematological, musculoskeletal, hepatic, renal, and dermal/ocular. "Other" refers to any systemic effect (e.g., a decrease in body weight) not covered in these systems. In the example of key number 18, 1 systemic effect (respiratory) was investigated.

(8) NOAEL A No-Observed-Adverse-Effect Level (NOAEL) is the highest exposure level at which no harmful effects were seen in the organ system studied. Key number 18 reports a NOAEL of 3 ppm for the respiratory system which was used to derive an intermediate exposure, inhalation MRL of 0.005 ppm (see footnote "b").

(9) LOAEL A Lowest-Observed-Adverse-Effect Level (LOAEL) is the lowest dose used in the study that caused a harmful health effect. LOAELs have been classified into "Less Serious" and "Serious" effects. These distinctions help readers identify the levels of exposure at which adverse health effects first appear and the gradation of effects with increasing dose. A brief description of the specific end point used to quantify the adverse effect accompanies the LOAEL. The respiratory effect reported in key number 18 (hyperplasia) is a Less serious LOAEL of 10 ppm. MRLs are not derived from Serious LOAELs.

(10) Reference The complete reference citation is given in Chapter 9 of the profile.

(11) CEL A Cancer Effect Level (CEL) is the lowest exposure level associated with the onset of carcinogenesis in experimental or epidemiologic studies. CELs are always considered serious effects. The LSE tables and figures do not contain NOAELs for cancer, but the text may report doses not causing measurable cancer increases.

(12) Footnotes Explanations of abbreviations or reference notes for data in the LSE tables are found in the footnotes. Footnote "b" indicates the NOAEL of 3 ppm in key number 18 was used to derive an MRL of 0.005 ppm.

LEGEND

See Figure 3-1

LSE figures graphically illustrate the data presented in the corresponding LSE tables. Figures help the reader quickly compare health effects according to exposure concentrations for particular exposure periods.

(13) Exposure Period The same exposure periods appear as in the LSE table. In this example, health effects observed within the intermediate and chronic exposure periods are illustrated.

(14) Health Effect These are the categories of health effects for which reliable quantitative data exists. The same health effects appear in the LSE table.

(15) Levels of Exposure concentrations or doses for each health effect in the LSE tables are graphically displayed in the LSE figures. Exposure concentration or dose is measured on the log scale "y" axis. Inhalation exposure is reported in mg/m^3 or ppm and oral exposure is reported in mg/kg/day.

(16) NOAEL In this example, 18r NOAEL is the critical end point for which an intermediate inhalation exposure MRL is based. As you can see from the LSE figure key, the open-circle symbol indicates to a NOAEL for the test species-rat. The key number 18 corresponds to the entry in the LSE table. The dashed descending arrow indicates the extrapolation from the exposure level of 3 ppm (see entry 18 in the Table) to the MRL of 0.005 ppm (see footnote "b" in the LSE table).

(17) CEL Key number 38r is 1 of 3 studies for which Cancer Effect Levels were derived. The diamond symbol refers to a Cancer Effect Level for the test species-mouse. The number 38 corresponds to the entry in the LSE table.

(18) Estimated Upper-Bound Human Cancer Risk Levels This is the range associated with the upper-bound for lifetime cancer risk of 1 in 10,000 to 1 in 10,000,000. These risk levels are derived from the EPA's Human Health Assessment Group's upper-bound estimates of the slope of the cancer dose response curve at low dose levels (q_1*).

(19) Key to LSE Figure The Key explains the abbreviations and symbols used in the figure.

SAMPLE

1 6 **Table 3-1. Levels of Significant Exposure to [Chemical x] – Inhalation**

Key to figure[a]	Species	Exposure frequency/ duration	System	NOAEL (ppm)	LOAEL (effect) Less serious (ppm)	LOAEL (effect) Serious (ppm)	Reference
2 6 INTERMEDIATE EXPOSURE							
	5	6	7	8	9		10
3 6 Systemic	9	9	9	9	9		9
4 6 18	Rat	13 wk 5 d/wk 6 hr/d	Resp	3[b]	10 (hyperplasia)		Nitschke et al. 1981
CHRONIC EXPOSURE							
						11	
Cancer						9	
38	Rat	18 mo 5 d/wk 7 hr/d				20 (CEL, multiple organs)	Wong et al. 1982
39	Rat	89–104 wk 5 d/wk 6 hr/d				10 (CEL, lung tumors, nasal tumors)	NTP 1982
40	Mouse	79–103 wk 5 d/wk 6 hr/d				10 (CEL, lung tumors, hemangiosarcomas)	NTP 1982

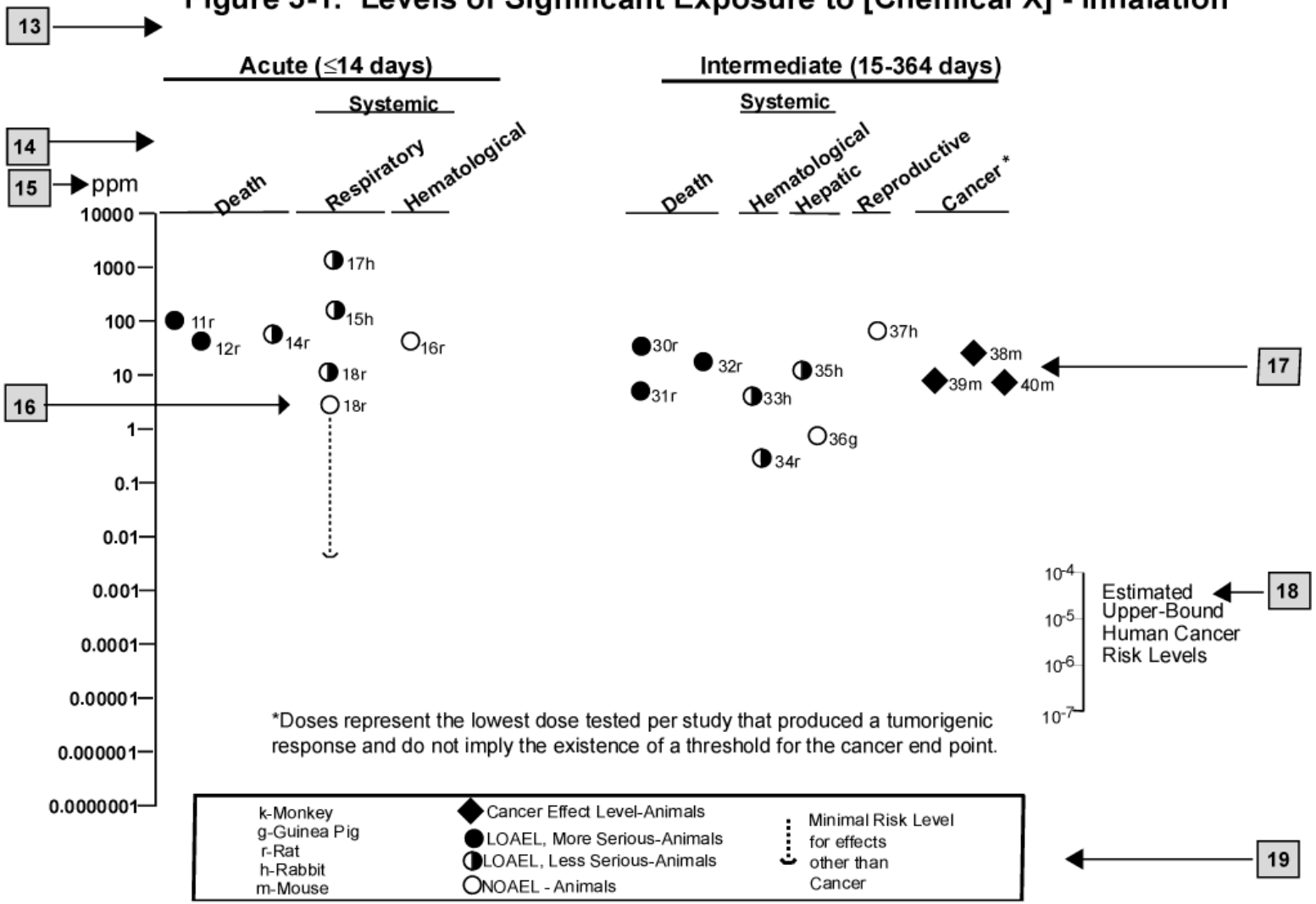
[a] The number corresponds to entries in Figure 3-1.

12 6 [b] Used to derive an intermediate inhalation Minimal Risk Level (MRL) of 5 x 10^{-3} ppm; dose adjusted for intermittent exposure and divided by an uncertainty factor of 100 (10 for extrapolation from animal to humans, 10 for human variability).

DRAFT FOR PUBLIC COMMENT

SAMPLE

Figure 3-1. Levels of Significant Exposure to [Chemical X] - Inhalation



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APPENDIX C

ACRONYMS, ABBREVIATIONS, AND SYMBOLS

ACGIH	American Conference of Governmental Industrial Hygienists
ADI	Acceptable Daily Intake
ADME	Absorption, Distribution, Metabolism, and Excretion
AFID	alkali flame ionization detector
AFOSH	Air Force Office of Safety and Health
AML	acute myeloid leukemia
AOAC	Association of Official Analytical Chemists
atm	atmosphere
ATSDR	Agency for Toxic Substances and Disease Registry
AWQC	Ambient Water Quality Criteria
BAT	Best Available Technology
BCF	bioconcentration factor
BEI	Biological Exposure Index
BSC	Board of Scientific Counselors
C	Centigrade
CAA	Clean Air Act
CAG	Cancer Assessment Group of the U.S. Environmental Protection Agency
CAS	Chemical Abstract Services
CDC	Centers for Disease Control and Prevention
CEL	Cancer Effect Level
CELDS	Computer-Environmental Legislative Data System
CERCLA	Comprehensive Environmental Response, Compensation, and Liability Act
CFR	Code of Federal Regulations
Ci	curie
CL	ceiling limit value
CLP	Contract Laboratory Program
cm	centimeter
CML	chronic myeloid leukemia
CNS	central nervous system
CPSC	Consumer Products Safety Commission
CWA	Clean Water Act
d	day
Derm	dermal
DHEW	Department of Health, Education, and Welfare
DHHS	Department of Health and Human Services
DNA	deoxyribonucleic acid
DOD	Department of Defense
DOE	Department of Energy
DOL	Department of Labor
DOT	Department of Transportation
DOT/UN/ NA/IMCO	Department of Transportation/United Nations/ North America/International Maritime Dangerous Goods Code
DWEL	Drinking Water Exposure Level
ECD	electron capture detection
ECG/EKG	electrocardiogram

EEG	electroencephalogram
EEGL	Emergency Exposure Guidance Level
EPA	Environmental Protection Agency
F	Fahrenheit
F_1	first-filial generation
FAO	Food and Agricultural Organization of the United Nations
FDA	Food and Drug Administration
FEMA	Federal Emergency Management Agency
FIFRA	Federal Insecticide, Fungicide, and Rodenticide Act
FPD	flame photometric detection
fpm	feet per minute
ft	foot
FR	*Federal Register*
g	gram
GC	gas chromatography
Gd	gestational day
gen	generation
GLC	gas liquid chromatography
GPC	gel permeation chromatography
HPLC	high-performance liquid chromatography
hr	hour
HRGC	high resolution gas chromatography
HSDB	Hazardous Substance Data Bank
IDLH	Immediately Dangerous to Life and Health
IARC	International Agency for Research on Cancer
ILO	International Labor Organization
in	inch
IRIS	Integrated Risk Information System
Kd	adsorption ratio
kg	kilogram
kkg	metric ton
K_{oc}	organic carbon partition coefficient
K_{ow}	octanol-water partition coefficient
L	liter
LC	liquid chromatography
LC_{Lo}	lethal concentration, low
LC_{50}	lethal concentration, 50% kill
LD_{Lo}	lethal dose, low
LD_{50}	lethal dose, 50% kill
LT_{50}	lethal time, 50% kill
LOAEL	lowest-observed-adverse-effect level
LSE	Levels of Significant Exposure
m	meter
MA	*trans,trans*-muconic acid
MAL	Maximum Allowable Level
mCi	millicurie
MCL	Maximum Contaminant Level
MCLG	Maximum Contaminant Level Goal
mg	milligram
min	minute
mL	milliliter

mm	millimeter
mm Hg	millimeters of mercury
mmol	millimole
mo	month
mppcf	millions of particles per cubic foot
MRL	Minimal Risk Level
MS	mass spectrometry
NAAQS	National Ambient Air Quality Standard
NAS	National Academy of Science
NATICH	National Air Toxics Information Clearinghouse
NATO	North Atlantic Treaty Organization
NCE	normochromatic erythrocytes
NCI	National Cancer Institute
NIEHS	National Institute of Environmental Health Sciences
NIOSH	National Institute for Occupational Safety and Health
NIOSHTIC	NIOSH's Computerized Information Retrieval System
NFPA	National Fire Protection Association
ng	nanogram
NLM	National Library of Medicine
nm	nanometer
NHANES	National Health and Nutrition Examination Survey
nmol	nanomole
NOAEL	no-observed-adverse-effect level
NOES	National Occupational Exposure Survey
NOHS	National Occupational Hazard Survey
NPD	nitrogen phosphorus detection
NPDES	National Pollutant Discharge Elimination System
NPL	National Priorities List
NR	not reported
NRC	National Research Council
NS	not specified
NSPS	New Source Performance Standards
NTIS	National Technical Information Service
NTP	National Toxicology Program
ODW	Office of Drinking Water, EPA
OERR	Office of Emergency and Remedial Response, EPA
OHM/TADS	Oil and Hazardous Materials/Technical Assistance Data System
OPP	Office of Pesticide Programs, EPA
OPPTS	Office of Prevention, Pesticides and Toxic Substances, EPA
OPPT	Office of Pollution Prevention and Toxics, EPA
OSHA	Occupational Safety and Health Administration
OSW	Office of Solid Waste, EPA
OTS	Office of Toxic Substances
OW	Office of Water
OWRS	Office of Water Regulations and Standards, EPA
PAH	Polycyclic Aromatic Hydrocarbon
PBPD	Physiologically Based Pharmacodynamic
PBPK	Physiologically Based Pharmacokinetic
PCE	polychromatic erythrocytes
PEL	permissible exposure limit
PID	photo ionization detector

pg	picogram
pmol	picomole
PHS	Public Health Service
PMR	proportionate mortality ratio
ppb	parts per billion
ppm	parts per million
ppt	parts per trillion
PSNS	Pretreatment Standards for New Sources
REL	recommended exposure level/limit
RfC	Reference Concentration
RfD	Reference Dose
RNA	ribonucleic acid
RTECS	Registry of Toxic Effects of Chemical Substances
RQ	Reportable Quantity
SARA	Superfund Amendments and Reauthorization Act
SCE	sister chromatid exchange
sec	second
SIC	Standard Industrial Classification
SIM	selected ion monitoring
SMCL	Secondary Maximum Contaminant Level
SMR	standard mortality ratio
SNARL	Suggested No Adverse Response Level
SPEGL	Short-Term Public Emergency Guidance Level
STEL	short term exposure limit
STORET	Storage and Retrieval
TD_{50}	toxic dose, 50% specific toxic effect
TLV	threshold limit value
TOC	Total Organic Compound
TPQ	Threshold Planning Quantity
TRI	Toxics Release Inventory
TSCA	Toxic Substances Control Act
TRI	Toxics Release Inventory
TWA	time-weighted average
U.S.	United States
UF	uncertainty factor
VOC	Volatile Organic Compound
yr	year
WHO	World Health Organization
wk	week
$>$	greater than
$\geq$	greater than or equal to
$=$	equal to
$<$	less than
$\leq$	less than or equal to
%	percent
α	alpha
β	beta
γ	gamma
δ	delta
μm	micrometer

µg	microgram
q_1^*	cancer slope factor
–	negative
+	positive
(+)	weakly positive result
(–)	weakly negative result

APPENDIX D

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